

Cambridge International AS & A Level

CHEMISTRY

9701/31

Paper 3 Advanced Practical Skills 1

May/June 2025

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the May/June 2025 series for most Cambridge IGCSE, Cambridge International A and AS Level components, and some Cambridge O Level components.

This document consists of **12** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptions for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.

2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.

3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).

4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.

5 'List rule' guidance

For questions that require *n* responses (e.g. State **two** reasons ...):

- The response should be read as continuous prose, even when numbered answer spaces are provided.
- Any response marked *ignore* in the mark scheme should not count towards *n*.
- Incorrect responses should not be awarded credit but will still count towards *n*.
- Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
- Non-contradictory responses after the first *n* responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.












Annotations guidance for centres

Examiners use a system of annotations as a shorthand for communicating their marking decisions to one another. Examiners are trained during the standardisation process on how and when to use annotations. The purpose of annotations is to inform the standardisation and monitoring processes and guide the supervising examiners when they are checking the work of examiners within their team. The meaning of annotations and how they are used is specific to each component and is understood by all examiners who mark the component.

We publish annotations in our mark schemes to help centres understand the annotations they may see on copies of scripts. Note that there may not be a direct correlation between the number of annotations on a script and the mark awarded. Similarly, the use of an annotation may not be an indication of the quality of the response.

The annotations listed below were available to examiners marking this component in this series.

Annotations

Annotation	Meaning
	Correct point or mark awarded
	Incorrect point or mark not awarded
	Information missing or insufficient for credit
	Benefit of the doubt given
	Contradiction in response otherwise markworthy, mark not given
	Error in number of decimal places
	Error carried forward applied
	Incorrect or insufficient point ignored while marking the rest of the response
	Benefit of the doubt not applied in this instance
	Rounding error
	Repeat error

Annotation	Meaning
SEEN or /	Blank page or part of script seen
SF	Error in number of significant figures
TE	Transcription error

Question	Answer	Marks
1(a)	<p>I Unambiguous headings in space for results AND correct units in the space for results for:</p> <ul style="list-style-type: none"> • (mass of) container with FA 1 / solid • (mass of) container (+ residual FA 1 / solid) • (mass of) FA 1 • (volume of) gas / CO₂ (produced) 	1
	<p>II Both weighings to a consistent number of decimal places (either to 2 dp or to 3 dp) AND mass of FA 1 correctly calculated AND volume of gas (in cm³) produced recorded as an integer value (No headings are required for this mark.)</p>	1
	<p>III Volume of gas produced Award in range 90–210 cm³ (0.090–0.210 dm³)</p>	1
1(b)(i)	<p>Correctly calculates (amount of CO₂ =) volume of gas collected / 24000 mol AND amount of CaCO₃ = amount of CO₂ AND both answers to 2–4 sf</p>	1
1(b)(ii)	<p>Correct display % purity = $[(b)(i) \times 100.1 / \text{mass of FA 1}] \times 100$ AND answer to 2–4 sf</p>	1
1(c)	<p>M1: the bottom box is ticked AND the volume of gas collected is less</p> <p>M2: more gas escapes (before the bung is replaced) AND the (initial) rate is greater</p>	2

Question	Answer	Marks
1(d)(i)	<i>One of:</i> <ul style="list-style-type: none">• unreacted solid / FA 1 / CaCO₃ remains in the (conical) flask (after bubbling stops)• not all the solid / FA 1 has reacted / disappeared (at end of reaction)	1
1(d)(ii)	% purity is less AND M1: volume of gas collected is less % purity is less AND M2: so amount / moles / mass of calcium carbonate (calculated) is lower	2

Question	Answer	Marks
2(a)	<p>I The following data must be shown</p> <ul style="list-style-type: none"> • two burette readings AND titre for rough titration • initial and final burette readings for two (or more) accurate titrations <p>II Titre values recorded for accurate titrations AND correct headings and units in the accurate titration table</p> <ul style="list-style-type: none"> • initial / start AND (burette) reading / volume • final / end AND (burette) reading / volume • titre OR volume / FA 5 AND used / added <p>unit: / cm³ OR (cm³) OR in cm³ (for each heading) OR cm³ unit given for each volume recorded</p> <p>III All accurate burette readings are to nearest 0.05 cm³</p> <p>IV The final accurate titre recorded must be within 0.10 cm³ of any other accurate titre.</p> <p>Accuracy (Q) marks: marks V, VI and VII will be determined by comparison of candidate and supervisor</p> <p>Round burette readings to the nearest 0.05 cm³. Check and correct titre subtractions where necessary. Select the best mean titre, using the following hierarchy:</p> <ul style="list-style-type: none"> • 2 identical accurate titres (ignoring any that are labelled 'rough'), <i>then</i> • accurate titres within 0.05 cm³, <i>then</i> • accurate titres within 0.10 cm³, <i>etc.</i> <p>Calculate the Supervisor's mean titre to two decimal places. Calculate the candidate's mean titre to two decimal places. Calculate the difference (δ) between the supervisor's and candidate's mean titre.</p> <p>V, VI, VII Award V if $\delta \leq 0.80 \text{ cm}^3$ Award VI if $\delta \leq 0.50 \text{ cm}^3$ Award VII if $\delta \leq 0.30 \text{ cm}^3$</p>	7
2(b)	<p>Correct calculation of the mean titre</p> <ul style="list-style-type: none"> • candidate must take the average of two (or more) titres that are within a total spread of not more than 0.20 cm³. • working/ explanation must be shown OR ticks must be put next to the two(or more) accurate readings selected. • the mean must be quoted to 2 decimal place and be rounded to the nearest 0.01 cm³, e.g., 26.675 must be rounded to 26.68. 	1

Question	Answer	Marks
2(c)(i)	Significant figures All quoted answers in (c)(ii)–(c)(v) are expressed to 3 or 4 sf.	1
2(c)(ii)	Correctly calculates amount of NaOH = $0.09 \times (\text{vol (b)}/1000)$ mol	1
2(c)(iii)	Correctly calculates amount of HCl in 250 cm ³ FA 4 = (c)(ii) × 10 mol AND amount of HCl in 250 cm ³ FA 3 = [(c)(ii) × 10] × 10 mol	1
2(c)(iv)	Correctly calculates amount of HCl used to prepare FA 3 = $(250/1000) \times 2 = 0.500$ mol	1
2(c)(v)	Correctly uses (c)(iv) – (c)(iii)	1
2(c)(vi)	Correctly uses amount of CaCO ₃ present = $(\text{c)(v)}/2$ mol	1
2(c)(vii)	Correctly uses % purity = [(c)(vi) × 100.1 / 19.0] × 100	1
2(d)(i)	Any two of: <ul style="list-style-type: none"> • some gas / carbon dioxide dissolves in the water (in trough / measuring cylinder) • some gas escapes before the bung is replaced on the conical flask • the impurity may react to produce gas 	2
2(d)(ii)	Any one of: <ul style="list-style-type: none"> • description of both CaCO₃ and acid being separated within the stoppered flask until reaction starts, e.g. use of small container for CaCO₃ which can be tipped over, package of CaCO₃ held by a thread, etc. • saturate the water with CO₂ prior to experiment / use of warmer water • use of gas syringe 	1

Question	Answer	Marks																		
FA 7 is NH₄Cl and CuCl₂; FA 8 is NaHCO₃; FA 9 is Na₂CO₃																				
3(a)(i)	<p>2 * = 1 mark (round down)</p> <table border="1" data-bbox="347 347 1561 1182"> <tbody> <tr> <td data-bbox="347 347 564 448">Test 1 NH₃(aq)</td> <td data-bbox="564 347 1561 448">(pale) blue ppt * soluble in excess * (to give a) dark / deep blue solution *</td> </tr> <tr> <td data-bbox="347 448 564 549">Test 2 NaOH(aq)</td> <td data-bbox="564 448 1561 549">(pale) blue ppt AND insoluble in excess *</td> </tr> <tr> <td data-bbox="347 549 564 616">warm</td> <td data-bbox="564 549 1561 616">gas / bubbles / NH₃ turns (red) litmus blue * black solid / ppt / residue *</td> </tr> <tr> <td data-bbox="347 616 564 716">Test 3 AgNO₃</td> <td data-bbox="564 616 1561 716">white ppt *</td> </tr> <tr> <td data-bbox="347 716 564 817">Test 4 Ba²⁺(aq)</td> <td data-bbox="564 716 1561 817">no (visible) change / remains pale blue solution / no (visible) reaction / no ppt *</td> </tr> <tr> <td data-bbox="347 817 564 951">Test 5 Dilute HNO₃</td> <td data-bbox="564 817 1561 951">no (visible) change / remains pale blue solution / no (visible) reaction / no effervescence AND</td> </tr> <tr> <td data-bbox="347 951 564 1018">standing</td> <td data-bbox="564 951 1561 1018">no change / remains pale blue solution *</td> </tr> <tr> <td data-bbox="347 1018 564 1118">Test 6 KI(aq)</td> <td data-bbox="564 1018 1561 1118">Brown *</td> </tr> <tr> <td data-bbox="347 1118 564 1182">starch</td> <td data-bbox="564 1118 1561 1182">Blue-black / <u>dark</u> blue / black *</td> </tr> </tbody> </table>	Test 1 NH ₃ (aq)	(pale) blue ppt * soluble in excess * (to give a) dark / deep blue solution *	Test 2 NaOH(aq)	(pale) blue ppt AND insoluble in excess *	warm	gas / bubbles / NH ₃ turns (red) litmus blue * black solid / ppt / residue *	Test 3 AgNO ₃	white ppt *	Test 4 Ba ²⁺ (aq)	no (visible) change / remains pale blue solution / no (visible) reaction / no ppt *	Test 5 Dilute HNO ₃	no (visible) change / remains pale blue solution / no (visible) reaction / no effervescence AND	standing	no change / remains pale blue solution *	Test 6 KI(aq)	Brown *	starch	Blue-black / <u>dark</u> blue / black *	5
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Test 6 KI(aq)	Brown *																			
starch	Blue-black / <u>dark</u> blue / black *																			
3(a)(ii)	<p>Cu²⁺(aq) + 2OH⁻(aq) → Cu(OH)₂(s) OR Ag⁺(aq) + Cl⁻(aq) → AgCl(s) OR 2Cu²⁺(aq) + 4I⁻(aq) → 2CuI(s) + I₂(aq)</p>	1																		

Question	Answer	Marks
3(a)(iii)	Cations: Cu^{2+} , NH_4^+ Anions: Cl^- 2 correct = 1 mark. All 3 correct = 2 marks	2
3(b)(i)	condensation / water vapour / (colourless) droplets / steam	1
3(b)(ii)	M1: effervescence / bubbling / fizzing M2: EITHER gas / fizzing / CO_2 gives a white ppt with limewater OR (white solid / FA 9 reacts and) colourless solution formed	2
3(b)(iii)	FA 8 is NaHCO_3 from appropriate observations: condensation in (b)(i) AND evidence of CO_2 in (b)(ii)	1