

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
GCE Ordinary Level

**MARK SCHEME for the May/June 2010 question paper
for the guidance of teachers**

5070 CHEMISTRY

5070/22

Paper 2 (Theory), maximum raw mark 75

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A1 (a) CF_3Cl [1]

(b) $\text{CH}_4 / \text{CO}_2$ [1]

(c) CaCO_3 [1]

(d) $\text{BaSO}_4 / \text{CaCO}_3$ [1]

(e) $\text{K}_2\text{Cr}_2\text{O}_7$ [1]

(f) C_2H_4 [1]

[Total: 6]

A2 (a) 1 / one [1]

(b) proton (atomic) number = 87
 number of protons = 87
 number of electrons = 87
 number of neutrons = 136

All correct = 2 marks [2]
 Any 3 correct = 1 mark

(c) Any two of:

- thermal conductor /
- electrical conductor /
- soft **or** cuts easily /
- low melting point **or** low boiling point /
- (relatively) low density **or** lightweight IGNORE: light
- malleable /
- ductile /
- shiny **or** silvery ALLOW: grey IGNORE: white /

IGNORE: floats on water / sonorous.
 IGNORE: chemical properties
 IGNORE: comparisons e.g. heavier than lithium [2]

(d) $2\text{Fr} + 2\text{H}_2\text{O} \rightarrow 2\text{FrOH} + \text{H}_2$ [1]
 ALLOW: multiples
 ALLOW: $\text{Fr} + \text{H}_2\text{O} \rightarrow \text{FrOH} + \frac{1}{2}\text{H}_2$
 IGNORE: state symbols

[Total: 6]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
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A3 (a) $\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{ZnCl}_2\text{(aq)} + \text{H}_2\text{(g)}$ [2]
 1 mark for correctly balanced equation ;
 1 mark for correct state symbols (dependent on all formulae being correct)

(b) (i) gas escapes / hydrogen escapes / gas given off / hydrogen given off / gas released / hydrogen released / gas produced / gas evolved / hydrogen is a gas ; [1]
 NOT: hydrogen produced without qualification. ALLOW: ecf from wrong gas in part **(a)**

(ii) downwards curve starting at the same point as the original curve but displayed to the left (at least at first) ; [1]

Line ends at the same mass as the original ; [1]
 NOT: curve dipping markedly below the horizontal section and then going upwards to meet it

(c) (acid) particles in dilute acid are less crowded / there are fewer particles (of acid) in a given volume / the particles (of acid) are further apart ; [1]
 ALLOW: concentration of HCl particles is lower
 ALLOW: molecules / ions in place of particles
 ALLOW: reverse argument e.g. particles in concentrated acid are more crowded / there are more particles (of acid) in a given volume etc
 IGNORE: there are fewer molecules unqualified / there is more water there are more moles in a given volume.

fewer collisions (in dilute acid) / less chance of collisions (in dilute acid) / frequency of collisions lower (in dilute acid) ; [1]
 ALLOW: reverse argument e.g. more collisions (in concentrated acid) / more chance of collisions (in concentrated acid) ;
 IGNORE: effective (collisions)

(d) more particles exposed / large(r) surface area ; [1]
 ALLOW: atoms / ions in place of particles

more collisions / greater chance of collisions / particles collide more often / greater frequency of collisions ; [1]
 IGNORE: effective (collisions)

(e) white precipitate / ppt **or** white solid ; [1]
 IGNORE: bubbles / colourless ppt / incorrectly named ppt

precipitate redissolves (in excess) / precipitate goes to (colourless) solution (in excess) ; [1]
 ALLOW: this mark if wrong colour precipitate
 NOTE: second mark dependent on ppt or solid stated for first mark

[Total: 11]

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
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- A4 (a)** graphite has electron(s) that can move / are mobile / are delocalised ; [1]
ALLOW: graphite has free electron(s) / graphite has a sea of electrons
REJECT: implications of layers moving / ions have free electrons

diamond has all its electrons involved in bonding / has electron(s) that cannot move / are not mobile / no delocalised electrons ; [1]
ALLOW: diamond has no free electron(s)
REJECT: mention of ions

- (b)** solid sodium chloride has ions fixed in position / ions cannot move ; [1]
IGNORE: electrons cannot move / ions can't carry electricity / references to intermolecular forces
ALLOW: ions are not free
REJECT: no ions to move
aqueous sodium chloride has ions that can move / are mobile ; [1]
ALLOW: ions are free
REJECT: reference to moving electrons as well as ions
IGNORE: ions carry electric charge / ions dislocated / ions delocalised /

- (c)** 1st row: lead at cathode **and** bromine at anode ; [1]
ALLOW: Pb at cathode / Br₂ at anode
REJECT: lead(II) / Pb²⁺ / Br⁻ / bromide
IGNORE: Br
2nd row: oxygen / O₂ ; [1]
REJECT: O²⁻
IGNORE: O
3rd row: hydrogen / H₂ ; [1]
REJECT: H⁺
IGNORE: H

- (d)** commercial use e.g. extraction of aluminium or any other element which is definitely extracted by electrolysis / purification of copper / (electro)plating ; [1]
ALLOW: coating metals / hair removal / production of sodium hydroxide
NOT: electrolysis **of** named substance unqualified / reference to electrochemical cells

correct electrolyte / correct formula of electrolyte:

This mark is dependent on the correct use BUT allow if it is feasible e.g. zinc sulphate (given incorrect use of zinc in the first part).

e.g. molten aluminium oxide dissolved in cryolite / (aqueous) copper sulfate or copper sulfate (solution) / for hair removal accept sweat or sodium chloride (solution). [1]

correct ionic equation: This mark is dependent on the electrolyte used; [1]

e.g. $Al^{3+} + 3e^{-} \rightarrow Al$ / $Cu^{2+} + 2e^{-} \rightarrow Cu$ / $2H^{+} + 2e^{-} \rightarrow H_2$

[Total: 10]

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
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A5 (a) cracking / thermal decomposition ; [1]

(b) (i) $C_2H_4 + H_2O \rightarrow C_2H_5OH$ [1]
ALLOW: C_2H_6O for the product

(ii) propanol; [1]
ALLOW: propan-1-ol / propan-2-ol
IGNORE: formulae

(c) (i) any **two** from: [2]

- temperature between 25°C to 40°C /
- REJECT: high temperature IGNORE: room temperature
- yeast / zymase / enzymes /
- IGNORE: catalyst alone
- absence of oxygen / anaerobic (conditions) / not exposed to air
- water REJECT: moisture / damp
- pH neutral / near neutral / pH 7
- IGNORE: pressure / presence of glucose

(ii) any one of: [1]
 renewable raw materials used or renewable fuel made NOT: renewable process /
 conserves valuable resources / lower energy costs / lower temperature required / lower
 pressure required / consumes less energy / atmospheric pressure required / specialised
 equipment not required / simple apparatus required;
 ALLOW: carbon neutral / carbon dioxide made (in this process) can be used for
 photosynthesis (to make more glucose) NOT: carbon dioxide can be used for
 photosynthesis alone
 IGNORE: not as complicated / references to pollution / consumes energy without
 qualification
 NOT: costs alone / faster / uses glucose without qualification

(d) (fractional) distillation / fractionation; [1]
 ALLOW: description of distillation e.g. evaporating then condensing the alcohol (first)
 IGNORE: using an anhydrous salt / named anhydrous salt

(e) lime water goes milky / cloudy / chalky / misty / white precipitate [1]

[Total: 8]

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
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- A6 (a) (i)** addition ; [1]
ALLOW: additional
IGNORE: specific names
- (ii)** minimum required is $C_2H_5CH=CH_2$ [1]
- (iii)** no (carbon-carbon) double bonds / only has (carbon-carbon) single bonds [1]
ALLOW: no hydrogen can be added / no addition reactions / carbons fully occupied by (hydrogen atoms)
NOT: occupied by wrong atoms e.g. Cl atoms
NOT: has carbon-carbon single bonds
- (b)** non-biodegradeable / can't be broken down by bacteria / insoluble in water / only soluble in organic solvents [1]
ALLOW: doesn't react with water / unreactive
IGNORE: it is a hydrocarbon / it is strongly bonded

[Total: 4]

Page 7	Mark Scheme: Teachers' version	Syllabus	Paper
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B7 (a) non-polluting gases formed / harmless gases formed / nitrogen and water are harmless / nitrogen and water are non-polluting / the products are non-polluting/the products are harmless ; [1]

ALLOW: nitrogen and water don't affect ozone / don't contribute (as much) to greenhouse effect / don't contribute to acid rain

NOT: nitrogen and water less harmful / nitrogen and water are formed (without qualification) / environmentally friendly products

(b) bond breaking endothermic / requires energy / absorbs energy
AND bond making exothermic / releases energy / gives out energy ; [1]

more energy is released than absorbed (or similar wording) ; [1]

REJECT: implication that energy needed in bond formation

NOTE: energy released on forming bonds is greater than energy taken in to break bonds (or similar wording) = 2 marks

(c) (i) moles $\text{N}_2\text{H}_4 = 1\,000\,000 / 32 = 31\,250$; [1]

moles $\text{O}_2 = \text{moles } \text{N}_2\text{H}_4$ or implication of this in working ; [1]

ALLOW: ecf from wrong moles of N_2H_4

Volume of O_2 ($31\,250 \times 24$) = $750\,000 \text{ dm}^3 / 7.5 \times 10^5 \text{ dm}^3$; [1]

ALLOW: ecf from second mark.

Alternative for 1st two stages:

32 g $\text{N}_2\text{H}_4 \rightarrow 32\text{g } \text{O}_2$ (1 mark)

moles $\text{O}_2 = 1\,000\,000 / 32 = 31\,250$ (allow ecf) (1 mark)

(ii) it / liquid oxygen takes up less space / room ; [1]

ALLOW: able to store more in liquid form / gaseous volume too high / maximum storage capacity.

IGNORE: less easily spread out/no gas can escape / less possibility of an explosion / to prevent reaction with other substances

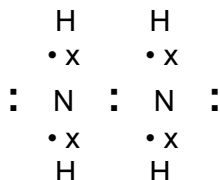
(d) (i) $\text{N}_2\text{H}_5\text{Cl} / \text{N}_2\text{H}_6\text{Cl}_2$ [1]

ALLOW: any order of atoms

ALLOW: correct displayed formulae or mixtures of displayed and molecular

REJECT: $\text{N}_2\text{H}_5\text{Cl}$ in equation if more than one product given

(ii) [2]



Structure completely correct = 2 marks

NOTE: (i) only outer shells need be shown

(ii) no distinction need be made between dots and crosses

IF: inner shells incorrect = 1 mark maximum.

IF: structure with a triple bond and no lone pairs = 1 mark

NOT: structures with separate nitrogen atoms / double bonds (= 0)

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B8 (a) (i) butanoic acid / methylpropanoic acid ; [1]

(ii) minimum is $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ / $(\text{CH}_3)_2\text{CHCOOH}$ [1]
ALLOW: correct displayed formulae or mixture of structural and displayed

(iii) $\text{C}_2\text{H}_4\text{O}$ [1]

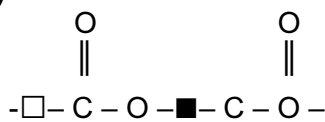
(b) molar ratio correct C = 4.35, H = 13.0, O = 2.18 ; [1]

$\text{C}_2\text{H}_6\text{O}$ [1]
ALLOW: correct error carried forward as long as there is not too much rounding up or down from the first stage
ALLOW: $\text{C}_2\text{H}_5\text{OH}$

(c) (i) ethyl ethanoate ; [1]

(ii) solvent / flavouring / perfume / aroma /
ALLOW: to make the taste in sweets / deodorants
IGNORE: food additive

(d) (i) [2]



correct structure of ester linkage showing ALL atoms and bonds (including bonds to the boxes) = 1 mark

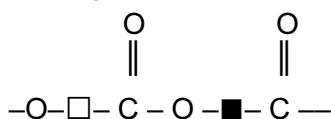
at least 2 units shown with continuation bonds = 1 mark

ALLOW: ester linkages reversed

ALLOW: boxes or part formulae between ester linkages the same

NOT: more than three type of 'boxes'

ALLOW:



ALLOW: single unit shown bracketed and continuation bonds

2nd mark dependent on ester linkage being shown correctly, or as $-\text{COO}-$ or $-\text{CO}_2-$ etc

(ii) fat / lipid / (tri)glyceride; [1]

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B9 (a) reaction in which there is electron transfer / one reactant loses electrons and the other gains electrons / both oxidation and reduction occur ; [1]
 ALLOW: a reaction involving changes in oxidation state
 IGNORE: gaining and losing oxygen / gaining and losing hydrogen

(b) (i) less iodine present / lower concentration of iodine ; [1]
 NOT: less reactants present / diluted in colour because more colourless HI present

(position of) equilibrium moves to the right / increased yield / reaction moves to the right ; [1]

ALLOW: more hydrogen and iodine react to form hydrogen iodide

ALLOW: more hydrogen iodide formed / more product formed / rate of forward reaction increases (to achieve new equilibrium)

The reaction is endothermic / the reaction absorbs heat (or energy) / ΔH is positive; [1]

(c) moles of hydrogen = $45.3 / 2 = 22.65$ [1]
 answer only scores mark
 ALLOW: 22.7

moles of HI = 45.3; [1]
 ALLOW: ecf / indication that moles HI 2× moles of hydrogen i.e. use of 1:2 ratio

mass = $(45.3 \times 128) = 5798 \text{ g} / 5798.4 \text{ g}$; [1]
 ALLOW: ecf moles HI / 5800 g

Alternative method:

2 g hydrogen $\rightarrow 2 \times 128 = 256 \text{ g HI}$ (1 mark)

so 1 g hydrogen $\rightarrow 128 \text{ g HI}$ (1 mark)

45.3 g hydrogen $\rightarrow 45.3 \times 256 / 2 = 5798(.4) \text{ g}$ (1 mark)

(d) (i) $\text{Pb}^{2+}(\text{aq}) + 2\text{I}^{-}(\text{aq}) \rightarrow \text{PbI}_2(\text{s})$ [2]

balanced equation = 1 mark

correct state symbols = 1 mark (dependent on correct formulae above)

ALLOW: full ionic equation

NOT: X^{-} in place I^{-} and PbX_2 in place of PbI_2

(ii) it or X is a reducing agent / HI is a reducing agent / it or X can be oxidised / HI can be oxidised ; [1]

Page 10	Mark Scheme: Teachers' version	Syllabus	Paper
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- B10(a) (i)** $\text{KOH} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2\text{O}$ [1]
 ALLOW: $\text{K}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{KCl} + \text{H}_2\text{O} + \text{CO}_2$
 ALLOW: $\text{KHCO}_3 + \text{HCl} \rightarrow \text{KCl} + \text{H}_2\text{O} + \text{CO}_2$
 IGNORE: state symbols
 NOT: word equation
- (ii)** titrate (acid against alkali) / titration / description of titration e.g. add one solution to the other until neutralised / add one solution to another until (acid-base) indicator changes colour ; [1]
 IGNORE: lack of repeating the titration without indicator
- Evaporate the solution (from the titration flask to dryness) ; [1]
 ALLOW: evaporate / heat / boil
 ALLOW: ecf from wrongly named solution in first marking point
 ALLOW: evaporation etc from potassium chloride / salt solution without reference to titration
 REJECT: if method incorrect e.g. precipitation the mark for part **(ii)** is zero in total.
- (b) (i)** $(\text{NH}_4)_3\text{PO}_4$ [1]
 ALLOW: $\text{PO}_4(\text{NH}_4)_3$
- (ii)** molar mass $(\text{NH}_4)_3\text{PO}_4 = 149$; [1]
 ALLOW: ecf from wrong formula in part **(i)**
- % by mass = 28.2 [1]
 ALLOW: 28.19 / 28
 ALLOW: ecf from wrong molar mass
- (c) (i)** $\text{Ca}(\text{OH})_2 + 2\text{H}^+ \rightarrow \text{Ca}^{2+} + 2\text{H}_2\text{O}$ [1]
 ALLOW: $\text{Ca}^{2+} + 2\text{OH}^- + 2\text{H}^+ \rightarrow \text{Ca}^{2+} + 2\text{H}_2\text{O}$
 ALLOW: $\text{OH}^- + \text{H}^+ \rightarrow \text{H}_2\text{O}$ (or multiples)
- (ii)** ammonium phosphate (reacts with calcium hydroxide to) give ammonia / there is loss of nitrogen (content) with ammonium phosphate [1]
 ALLOW: reverse arguments
 IGNORE: ammonia poisonous / potassium nitrate is more soluble
 REJECT: loses nitrogen gas / potassium nitrate has a greater % of nitrogen
- (d)** add (excess) sodium hydroxide **and** aluminium (powder / foil and warm) ; [1]
 ALLOW: add sodium hydroxide **and** Devarda's alloy
- ammonia given off / gas (given off) turns red litmus blue; [1]
 NOTE: this mark is dependent on correct reagents $\text{Al} + \text{NaOH}$
- Alternative:
 add iron(II) sulfate then concentrated sulfuric acid (1 mark)
 brown ring forms at the interface (1 mark)