



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
General Certificate of Education Ordinary Level

CANDIDATE  
NAME

CENTRE  
NUMBER

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CANDIDATE  
NUMBER

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**CHEMISTRY**

**5070/21**

Paper 2 Theory

**May/June 2011**

**1 hour 30 minutes**

Candidates answer on the Question Paper.

No additional materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**DO NOT WRITE IN ANY BARCODES.**

**Section A**

Answer **all** questions.

Write your answers in the spaces provided in the Question Paper.

**Section B**

Answer any **three** questions.

Write your answers in the spaces provided in the Question Paper.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
<b>Section A</b>	
<b>B6</b>	
<b>B7</b>	
<b>B8</b>	
<b>B9</b>	
<b>Total</b>	

This document consists of **17** printed pages and **3** blank pages.



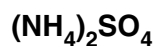
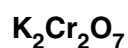
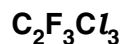
## Section A

Answer **all** the questions in this section in the spaces provided.

The total mark for this section is 45.

For  
Examiner's  
Use

**A1** Choose from the following formulae to answer the questions below.



Each formula can be used once, more than once, or not at all.

Which is the formula of a compound that

(a) is a catalyst in the Contact process,

.....[1]

(b) in aqueous solution reacts with aqueous sodium hydroxide to give a white precipitate that redissolves in excess sodium hydroxide,

.....[1]

(c) is an insoluble salt,

.....[1]

(d) is involved in ozone depletion in the upper atmosphere,

.....[1]

(e) in aqueous solution will react with aqueous barium chloride to make a white precipitate,

.....[1]

(f) is an alkane,

.....[1]

(g) is used as a fertiliser?

.....[1]

[Total: 7]

**A2** Small pieces of copper were added to excess concentrated sulfuric acid and the mixture heated for 30 minutes. A colourless gas **Z** was formed. When **Z** was tested with filter paper dipped into acidified potassium dichromate(VI), there was a colour change from orange to green.

For  
Examiner's  
Use

The reaction mixture was cooled and then diluted with water. A blue solution, **Y**, was formed. Aqueous sodium hydroxide was added drop by drop to the blue solution. Eventually a blue precipitate, **X**, was formed. On heating the blue precipitate turned black to form compound **V**. Analysis of **V** showed that it contained 79.9 % copper and 20.1 % oxygen by mass.

**(a)** Name gas **Z**.

.....[1]

**(b)** Name the blue solution **Y**.

.....[1]

**(c)** When aqueous sodium hydroxide was added to the cooled reaction mixture, it initially reacted with excess sulfuric acid.

Write the ionic equation for this reaction.

[1]

**(d) (i)** Name the blue precipitate **X**.

.....[1]

**(ii)** Write an ionic equation, including state symbols, to show the formation of this blue precipitate.

[2]

**(e)** Calculate the empirical formula of the black solid **V**.

empirical formula of **V** is ..... [2]

[Total: 8]

**A3** Uranium is a radioactive metal. It has two main isotopes, uranium-235 with a nucleon number of 235 and uranium-238 with a nucleon number of 238.

For  
Examiner's  
Use

- (a) (i)** State one similarity, in terms of sub-atomic particles, between uranium-235 and uranium-238.

.....  
.....[1]

- (ii)** State one difference, in terms of sub-atomic particles, between uranium-235 and uranium-238.

.....  
.....[1]

- (b)** Uranium is manufactured from uranium(IV) oxide,  $\text{UO}_2$ , in a two-step process.

**Step 1** – uranium(IV) oxide is heated with hydrogen fluoride to make uranium(IV) fluoride,  $\text{UF}_4$ , and water.

**Step 2** – uranium(IV) fluoride is reduced by magnesium to give uranium and one other product.

- (i)** Construct the equation for step 1.

[1]

- (ii)** Construct the equation for step 2.

[1]

- (iii)** Step 2 involves a reduction.  
Explain the meaning of the term *reduction*?

.....  
.....  
.....[1]

- (iv) Calculate the mass of uranium that can be made from 1.00 tonne of uranium(IV) oxide.

[One tonne is one million grams.]

For  
Examiner's  
Use

mass of uranium = ..... tonnes [3]

- (c) Uranium reacts with dilute hydrochloric acid to form hydrogen.  
Using this information and your knowledge of the reactivity of metals, suggest where in the following reactivity series you would place uranium.

**most reactive**

**potassium  
sodium  
calcium  
magnesium  
copper  
silver**

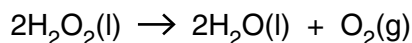
**least reactive**

.....[1]

[Total: 9]

- A4** Hydrogen peroxide,  $\text{H}_2\text{O}_2$ , is a covalent compound. Hydrogen peroxide decomposes to form water and oxygen.

For  
Examiner's  
Use



- (a) Draw a 'dot-and-cross' diagram for a molecule of hydrogen peroxide.

[2]

- (b) The decomposition of hydrogen peroxide involves a change from the liquid state to the gaseous state. Describe the difference in both the movement and arrangement of particles in a liquid and in a gas.

.....  
 .....  
 .....  
 .....[2]

- (c) At room temperature pure hydrogen peroxide decomposes much faster than dilute aqueous hydrogen peroxide. Explain why in terms of collision theory.

.....  
 .....  
 .....  
 .....[2]

- (d) When aqueous iron(II) ions are warmed with aqueous hydrogen peroxide, iron(III) ions are formed.

- (i) Construct an ionic equation for the oxidation of iron(II) ions to iron(III) ions.

[1]

- (ii) Describe a chemical test that can be used to confirm that iron(II) ions have been oxidised to form iron(III) ions.

For  
Examiner's  
Use

.....  
.....  
.....  
.....[2]

- (e) Aqueous hydrogen peroxide was added to acidified aqueous potassium manganate(VII). The purple solution turned colourless.

Aqueous hydrogen peroxide was added to acidified aqueous potassium iodide. The colourless solution turned brown.

What deductions can you make about hydrogen peroxide from these two observations? Explain your answer.

.....  
.....  
.....  
.....[2]

[Total: 11]

**A5** Clean, dry air contains a mixture of gases including oxygen, nitrogen, carbon dioxide and the noble gases.

*For  
Examiner's  
Use*

**(a)** Give the percentage by volume of nitrogen in clean, dry air.

.....[1]

**(b)** State and explain how oxygen is extracted from air.

.....  
.....  
.....  
.....  
.....  
.....[3]

**(c)** Explain how the carbon cycle helps to keep the composition of air relatively constant.

.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....[4]

**(d)** Many electricity power stations burn fossil fuels. Sulfur dioxide is a pollutant produced during the burning of fossil fuels. Sulfur dioxide causes acid rain.

Describe **two** ways in which calcium carbonate can be used to reduce the effects of burning fossil fuels.

1 .....

.....

2 .....

.....[2]

[Total: 10]



## Section B

Answer **three** questions from this section in the spaces provided.

The total mark for this section is 30.

For  
Examiner's  
Use

**B6** Electrolysis involves the chemical decomposition of a compound, either when molten or in aqueous solution, by the passage of an electric current.

(a) Explain why aqueous calcium nitrate can be electrolysed but liquid pentane cannot.

.....  
.....  
.....  
.....[2]

(b) State the products of the electrolysis of molten sodium chloride.

.....[1]

(c) State the products of the electrolysis of concentrated aqueous sodium chloride.

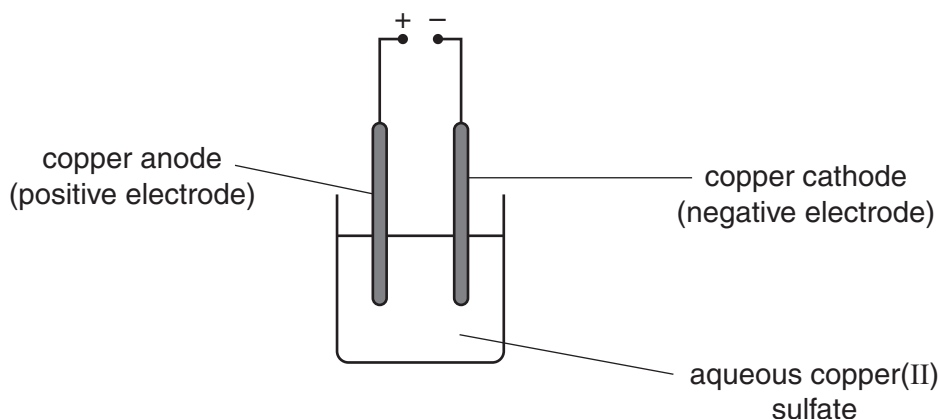
.....[1]

(d) Describe the essential details of the manufacture of aluminium by electrolysis.

.....  
.....  
.....  
.....[2]

- (e) A student investigates the electrolysis of aqueous copper(II) sulfate using the apparatus shown below.

For  
Examiner's  
Use



The student weighs the copper cathode before and after the electrolysis.

experiment number	current used / A	time taken / s	mass of cathode	
			before starting / g	after electrolysis / g
1	2.0	180	1.24	1.36
2	4.0	180	1.20	1.44
3	2.0	360	1.34	1.58

- (i) Explain, with the aid of an equation, why the cathode increases in mass.

.....  
 .....  
 .....[2]

- (ii) In experiment 2 the student measures the mass of the anode both before and after the electrolysis.

At the start the anode has a mass of 1.45 g.

Determine the mass of the anode at the end of the electrolysis.

mass of anode at end = ..... g [1]

- (iii) The student does a fourth experiment, this time using a current of 8.0A for 90 seconds. At the start the cathode has a mass of 1.51 g. Predict the mass of the cathode at the end of the electrolysis.

*For  
Examiner's  
Use*

mass of cathode at end = ..... g [1]

[Total: 10]

- B7** Alcohols are a homologous series of organic compounds.  
The table shows some information about the first five alcohols.

For  
Examiner's  
Use

name	molecular formula
methanol	$\text{CH}_4\text{O}$
ethanol	$\text{C}_2\text{H}_6\text{O}$
	$\text{C}_3\text{H}_8\text{O}$
butanol	$\text{C}_4\text{H}_{10}\text{O}$
pentanol	$\text{C}_5\text{H}_{12}\text{O}$

- (a) Suggest the name of the alcohol with the molecular formula  $\text{C}_3\text{H}_8\text{O}$ .  
.....[1]
- (b) Draw the structure of an alcohol with the molecular formula  $\text{C}_4\text{H}_{10}\text{O}$  and explain why this alcohol is saturated.  
.....  
.....[2]
- (c) Deduce the molecular formula of an alcohol that contains seven carbon atoms.  
.....[1]
- (d) Ethanol reacts with ethanoic acid to form ethyl ethanoate.
- (i) Draw the structure of ethyl ethanoate.  
.....[1]
- (ii) Suggest a use for ethyl ethanoate.  
.....[1]

(e) Describe, with the aid of an equation, how ethanol is manufactured by fermentation.

For  
Examiner's  
Use

.....  
.....  
.....  
.....  
.....[3]

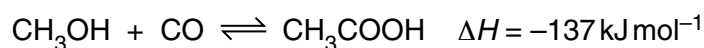
(f) When ethanol is heated with concentrated sulfuric acid a colourless gas, **A**, is produced. Gas **A** will decolourise aqueous bromine.

Identify gas **A**.

.....[1]

[Total: 10]

**B8** Ethanoic acid is manufactured by a reaction between methanol,  $\text{CH}_3\text{OH}$ , and carbon monoxide.



For  
Examiner's  
Use

This reaction is exothermic.

**(a)** The reaction is carried out at a pressure of 30 atmospheres and a temperature of  $180^\circ\text{C}$ .

**(i)** Predict and explain the effect on the position of equilibrium if the reaction is carried out at 30 atmospheres pressure and  $20^\circ\text{C}$  rather than  $180^\circ\text{C}$ .

.....  
 .....  
 .....[2]

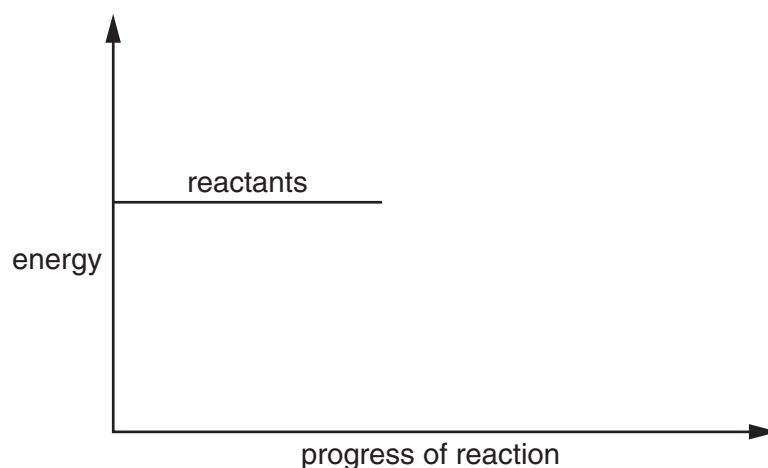
**(ii)** Suggest one reason why the reaction is carried out at  $180^\circ\text{C}$  rather than  $20^\circ\text{C}$ .

.....  
 .....[1]

**(b)** Complete the energy profile diagram for the reaction between methanol and carbon monoxide.

On your diagram label the

- product,
- activation energy,  $E_a$ ,
- enthalpy change for the reaction,  $\Delta H$ .



[3]

- (c) The manufacture of ethanoic acid from methanol also uses a catalyst to increase the speed of reaction.

Explain how a catalyst increases the speed of reaction.

.....  
.....[1]

- (d) In an investigation 10.0 moles of methanol are mixed with 20.0 moles of carbon monoxide.

At the end of the reaction 9.8 moles of ethanoic acid are formed.

Calculate the percentage yield of ethanoic acid.

percentage yield = ..... % [2]

- (e) Ethanoic acid reacts with ammonia to form a salt.

Give the formula of this salt.

.....[1]

[Total: 10]

For  
Examiner's  
Use

**B9** Sulfamic acid,  $\text{SO}_3\text{NH}_2$ , is a weak acid used to remove limescale from kettles.

For  
Examiner's  
Use

(a) Explain the meaning of the term *weak acid*?

.....  
.....[1]

(b) The pH of an aqueous solution of sulfamic acid can be determined using a pH meter. Describe another way of estimating the pH of a solution of sulfamic acid.

.....  
.....  
.....[2]

(c) A 0.105 g sample of sulfamic acid is dissolved in  $25.0\text{cm}^3$  of water. The sulfamic acid solution requires  $10.8\text{cm}^3$  of  $0.100\text{mol dm}^{-3}$  potassium hydroxide for complete neutralisation.

Calculate the number of moles of sulfamic acid that react with one mole of potassium hydroxide.

number of moles of sulfamic acid = ..... [3]

(d) Aqueous sulfamic acid reacts with magnesium to form magnesium sulfamate,  $\text{Mg}(\text{SO}_3\text{NH}_2)_2$ .

(i) Write an equation for this reaction.

[1]

(ii) Limescale contains calcium carbonate. Describe, with the aid of an equation, how aqueous sulfamic acid reacts with calcium carbonate.

.....[2]

(e) Sulfamic acid reacts with sodium nitrite,  $\text{NaNO}_2$ , to form water, sodium hydrogensulfate,  $\text{NaHSO}_4$ , and a colourless gas. Suggest the identity of the colourless gas.

.....[1]

[Total: 10]



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## DATA SHEET The Periodic Table of the Elements

Group		Group											
		I	II	III	IV	V	VI	VII	0				
		1 <b>H</b> Hydrogen 1											
7 <b>Li</b> Lithium 3	9 <b>Be</b> Beryllium 4												
23 <b>Na</b> Sodium 11	24 <b>Mg</b> Magnesium 12												
39 <b>K</b> Potassium 19	40 <b>Ca</b> Calcium 20	55 <b>Mn</b> Manganese 25	59 <b>Co</b> Cobalt 27	65 <b>Zn</b> Zinc 30	73 <b>Ge</b> Germanium 32	75 <b>As</b> Arsenic 33	79 <b>Se</b> Selenium 34	80 <b>Br</b> Bromine 35	84 <b>Kr</b> Krypton 36	86 <b>Rn</b> Radon 86			
85 <b>Rb</b> Rubidium 37	88 <b>Sr</b> Strontium 38	91 <b>Zr</b> Zirconium 40	93 <b>Nb</b> Niobium 41	101 <b>Ru</b> Ruthenium 44	106 <b>Pd</b> Palladium 46	108 <b>Ag</b> Silver 47	112 <b>Cd</b> Cadmium 48	115 <b>In</b> Indium 49	119 <b>Sn</b> Tin 50	122 <b>Sb</b> Antimony 51	127 <b>I</b> Iodine 53	131 <b>Xe</b> Xenon 54	
133 <b>Cs</b> Caesium 55	137 <b>Ba</b> Barium 56	147 <b>Pm</b> Promethium 61	150 <b>Sm</b> Samarium 62	159 <b>Tb</b> Terbium 65	162 <b>Dy</b> Dysprosium 66	165 <b>Ho</b> Holmium 67	167 <b>Er</b> Erbium 68	173 <b>Yb</b> Ytterbium 70	175 <b>Lu</b> Lutetium 71	209 <b>Po</b> Polonium 84	210 <b>At</b> Astatine 85	222 <b>Rn</b> Radon 86	
223 <b>Fr</b> Francium 87	226 <b>Ra</b> Radium 88	144 <b>Nd</b> Neodymium 60	141 <b>Pr</b> Praseodymium 59	157 <b>Gd</b> Gadolinium 64	162 <b>Dy</b> Dysprosium 66	165 <b>Ho</b> Holmium 67	167 <b>Er</b> Erbium 68	173 <b>Yb</b> Ytterbium 70	175 <b>Lu</b> Lutetium 71	209 <b>Po</b> Polonium 84	210 <b>At</b> Astatine 85	222 <b>Rn</b> Radon 86	
		232 <b>Th</b> Thorium 90	231 <b>Pa</b> Protactinium 91	247 <b>Bk</b> Berkelium 97	251 <b>Cf</b> Californium 98	252 <b>Es</b> Einsteinium 99	257 <b>Fm</b> Fermium 100	259 <b>No</b> Nobelium 102	260 <b>Lr</b> Lawrencium 103				
		140 <b>Ce</b> Cerium 58	141 <b>Pr</b> Praseodymium 59	150 <b>Sm</b> Samarium 62	152 <b>Eu</b> Europium 63	157 <b>Gd</b> Gadolinium 64	159 <b>Tb</b> Terbium 65	162 <b>Dy</b> Dysprosium 66	165 <b>Ho</b> Holmium 67	167 <b>Er</b> Erbium 68	169 <b>Tm</b> Thulium 69	173 <b>Yb</b> Ytterbium 70	175 <b>Lu</b> Lutetium 71
		232 <b>Th</b> Thorium 90	231 <b>Pa</b> Protactinium 91	244 <b>Pu</b> Plutonium 94	243 <b>Am</b> Americium 95	247 <b>Cm</b> Curium 96	247 <b>Bk</b> Berkelium 97	251 <b>Cf</b> Californium 98	252 <b>Es</b> Einsteinium 99	257 <b>Fm</b> Fermium 100	258 <b>Md</b> Mendelevium 101	259 <b>No</b> Nobelium 102	260 <b>Lr</b> Lawrencium 103

\* 58–71 Lanthanoid series  
† 90–103 Actinoid series

**Key**  
a **X**  
b

a = relative atomic mass  
 X = atomic symbol  
 b = atomic (proton) number

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).