



Cambridge International Examinations
Cambridge Ordinary Level

CHEMISTRY**5070/31**

Paper 3 Practical Test

October/November 2016

MARK SCHEME

Maximum Mark: 40

Published

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Page 2	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks	Guidance
1(a)	Titration Measurements (1) Both readings, i.e. initial and final are present for each titration and readings are recorded to 1dp. Titres (1) All the titres are calculated correctly, i.e. no subtraction errors Accuracy (6) For the two best titres give: 3 marks for a titre within 0.2 cm^3 of the Supervisor's value. 2 marks for a titre within 0.3 cm^3 of the Supervisor's value. 1 mark for a titre within 0.4 cm^3 of the Supervisor's value. Concordance (3) Give 3 marks if all the ticked values are within 0.2 cm^3. Give 2 marks if all the ticked values are within 0.3 cm^3. Give 1 marks if all the ticked values are within 0.4 cm^3. Average (1) Give 1 mark for calculating the correct average of selected titres.	12	Reject final readings in excess of 50 Reject initial readings of 50 Accuracy marks are awarded using the candidate's correct values. Concordance marks are awarded using the uncorrected titres.
1(b)	Assuming a pipette volume of 25 cm^3 and the average volume of P used = 19.8 cm^3: Mole of iron(II) sulfate in the average volume $= (25.0 \times 0.0800) / 1000$ $= 0.002$	1	
1(c)	Answer from (b)/5 $= 0.002 / 5$ $= 0.0004$	1	

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Question	Answer	Marks	Guidance
1(d)	Answer from (c) $\times 250$ /average volume of P $= 0.0004 \times 250 / 19.8$ $= 0.00505$	1	
1(e)	Answer from (d) $\times 55$ $= 0.00505 \times 55$ $= 0.278 \text{ g}$	1	
1(f)	Answer from (e) $\times 100/2.12$ $= 0.278 \times 100/2.12$ $= 13.1\%$	1	

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Question	Answer	Marks	Guidance
Question 2 General points R is nitric acid S is calcium carbonate For ppt: accept solid / suspension / powder but ignore substance / particles / deposit / residue / sediment / gelatinous / insoluble Ignore cloudy / milky / white / gelatinous solution for ppt forms but accept cloudy / milky / white / gelatinous solution for ppt remains Ignore solution / ppt turns colourless for ppt dissolves but accept clears for ppt dissolves For gases: to gain credit for the name of the gas produced, the test must be at least partially correct. For the evolution of a gas in a liquid accept the observation effervescence / bubbles / fizz / gas vigorously evolved but ignore gas evolved. Solutions: colourless is not equivalent to clear and clear is not equivalent to colourless Marks awarded for conclusions are dependent on correct evidence.			
2 (test 1)	(a) solution turns red or pink (1) (b) solution turns blue (1)	19	
2 (test 2)	gas turns damp red litmus blue (1) ammonia (1)		To score ammonia mark there must be an indication of a test, i.e. a smell of ammonia, alkaline gas, tested with litmus.
2 (test 3)	(a) solution turns yellow (1) (b) red-brown ppt (1) insoluble in excess (1)		
2 (test 4)	bubbles (1) gas pops with a lighted splint (1) hydrogen (1) piece of metal disappears or dissolves (1)		To score hydrogen mark there must be an indication of a test, i.e. pops (with a splint)
2 (test 5)	bubbles (1) gas turns limewater milky (1) carbon dioxide (1) solid disappears or dissolves (allow to score 1 if mark missed in test 4)		To score carbon dioxide mark there must be an indication of a test, i.e. tested with limewater.

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Question	Answer	Marks	Guidance
2 (test 6)	white ppt (1) insoluble in excess (1)		
2 (test 7)	no reaction (1)		Accept very slight white ppt
2 (test 8)	gas turns damp red litmus blue (1) ammonia (1) Allow the test and identification of carbon dioxide marks if not awarded in test 5		To score ammonia mark there must be an indication of a test – see test 2.
Conclusions	Cation in R is H^+ (1) Anion in R is NO_3^- (1) Cation in S is Ca^{2+} (1) Anion in S is CO_3^{2-} (1)	4	Evidence: Test 1(a) red or pink with litmus Test 2 alkaline gas / ammonia In test 6 white ppt remains and in 7 no reaction Carbon dioxide identified in test 5 or 8