

Location Entry Codes

As part of CIE's continual commitment to maintaining best practice in assessment, CIE uses different variants of some question papers for our most popular assessments with large and widespread candidature. The question papers are closely related and the relationships between them have been thoroughly established using our assessment expertise. All versions of the paper give assessment of equal standard.

The content assessed by the examination papers and the type of questions is unchanged.

This change means that for this component there are now two variant Question Papers, Mark Schemes and Principal Examiner's Reports where previously there was only one. For any individual country, it is intended that only one variant is used. This document contains both variants which will give all Centres access to even more past examination material than is usually the case.

The diagram shows the relationship between the Question Papers, Mark Schemes and Principal Examiners' Reports that are available.

Question Paper	Mark Scheme	Principal Examiner's Report
Introduction	Introduction	Introduction
First variant Question Paper	First variant Mark Scheme	First variant Principal Examiner's Report
Second variant Question Paper	Second variant Mark Scheme	Second variant Principal Examiner's Report

Who can I contact for further information on these changes?

Please direct any questions about this to CIE's Customer Services team at:

international@cie.org.uk

The titles for the variant items should correspond with the table above, so that at the top of the first page of the relevant part of the document and on the header, it has the words:

- First variant Question Paper / Mark Scheme / Principal Examiner's Report

or

- Second variant Question Paper / Mark Scheme / Principal Examiner's Report

as appropriate.



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE NAME

CENTRE NUMBER

CANDIDATE NUMBER

* 7 1 9 0 2 5 0 8 2 4 *

CHEMISTRY **0620/31**
Paper 3 (Extended) **May/June 2008**
1 hour 15 minutes

Candidates answer on the Question Paper.
No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use a pencil for any diagrams, graphs or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.
DO NOT WRITE IN ANY BARCODES

Answer **all** questions.
A copy of the Periodic Table is printed on page 12.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [] at the end of each question or part questions.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
8	
Total	

This document consists of 11 printed pages and 1 blank page.

1 For each of the following select an element from Period 4, potassium to krypton that matches the description.

(a) It is a brown liquid at room temperature.

(b) It forms a compound with hydrogen having the formula XH_4

(c) A metal that reacts violently with cold water.

(d) It has a complete outer energy level.

(e) It has oxidation states of 2 and 3 only.

(f) It can form an ion of the type X^-

(g) One of its oxides is the catalyst in the Contact Process.

[Total: 7]

- 2 (a) Complete the table which gives the names, symbols, relative masses and relative charges of the three subatomic particles.

name	symbol	relative mass	relative charge
electron	e^-		
proton		1	
	n		0

[3]

- (b) Use the information in the table to explain the following.

- (i) Atoms contain charged particles but they are electrically neutral because they have no overall charge.

.....
 [2]

- (ii) Atoms can form positive ions.

.....
 [2]

- (iii) Atoms of the same element can have different masses.

.....
 [2]

- (iv) Scientists are certain that there are no undiscovered elements missing from the Periodic Table from hydrogen to lawrencium.

..... [1]

[Total: 10]

3 Copper is purified by electrolysis.

(a) Complete the following.

The positive electrode (anode) is made from

The negative electrode (cathode) is made from

The electrolyte is aqueous [3]

(b) Write an ionic equation for the reaction at the positive electrode (anode).

..... [2]

(c) (i) Give **two** reasons why copper is used,

in electric wiring,

..... [2]

in cooking utensils.

..... [2]

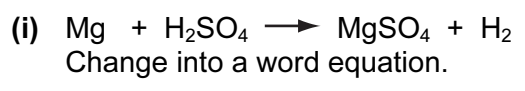
(ii) Give another use of copper.

..... [1]

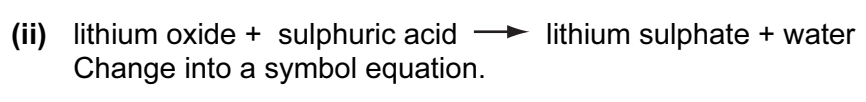
[Total: 10]

4 Sulphuric acid is a typical strong acid.

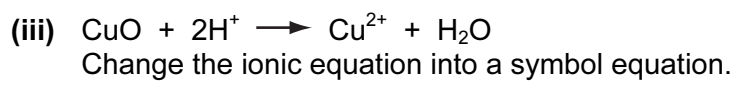
(a) Change the equations given into a different format.



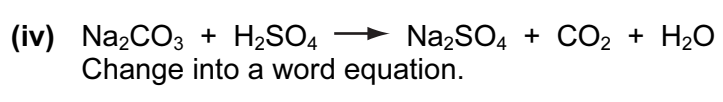
..... [1]



..... [2]

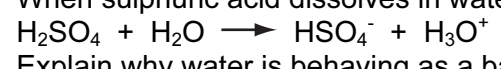


..... [2]



..... [1]

(b) When sulphuric acid dissolves in water, the following reaction occurs.



Explain why water is behaving as a base in this reaction.

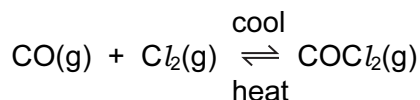
..... [2]

(c) Sulphuric acid is a strong acid, ethanoic acid is a weak acid.
Explain the difference between a strong acid and a weak acid.

.....
..... [2]

[Total: 10]

- 5 Carbonyl chloride, COCl_2 , is a colourless gas. It is made by the following reaction.



- (a) When the pressure on the equilibrium mixture is decreased, the position of equilibrium moves to left.

- (i) How does the concentration of each of the three chemicals change?

.....
 [2]

- (ii) Explain why the position of equilibrium moves to left.

.....
 [2]

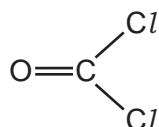
- (b) Using the information given with the equation, is the forward reaction exothermic or endothermic? Give a reason for your choice.

.....
 [2]

- (c) Carbonyl chloride reacts with water to form two acidic compounds. Suggest which acidic compounds are formed.

1.
 2. [2]

- (d) The structural formula of carbonyl chloride is given below.



Draw a diagram that shows the arrangement of the valency electrons in one molecule of this covalent compound.

Use x for an electron from a chlorine atom.

Use o for an electron from a carbon atom.

Use • for an electron from an oxygen atom.

[4]
 [Total: 12]

6 Three of the factors that can influence the rate of a chemical reaction are:

- physical state of the reactants
- light
- the presence of a catalyst

(a) The first recorded dust explosion was in a flour mill in Italy in 1785. Flour contains carbohydrates. Explosions are very fast exothermic reactions.

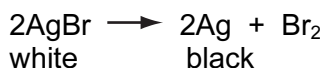
(i) Use the collision theory to explain why the reaction between the particles of flour and the oxygen in the air is very fast.

.....
 [2]

(ii) Write a word equation for this exothermic reaction.

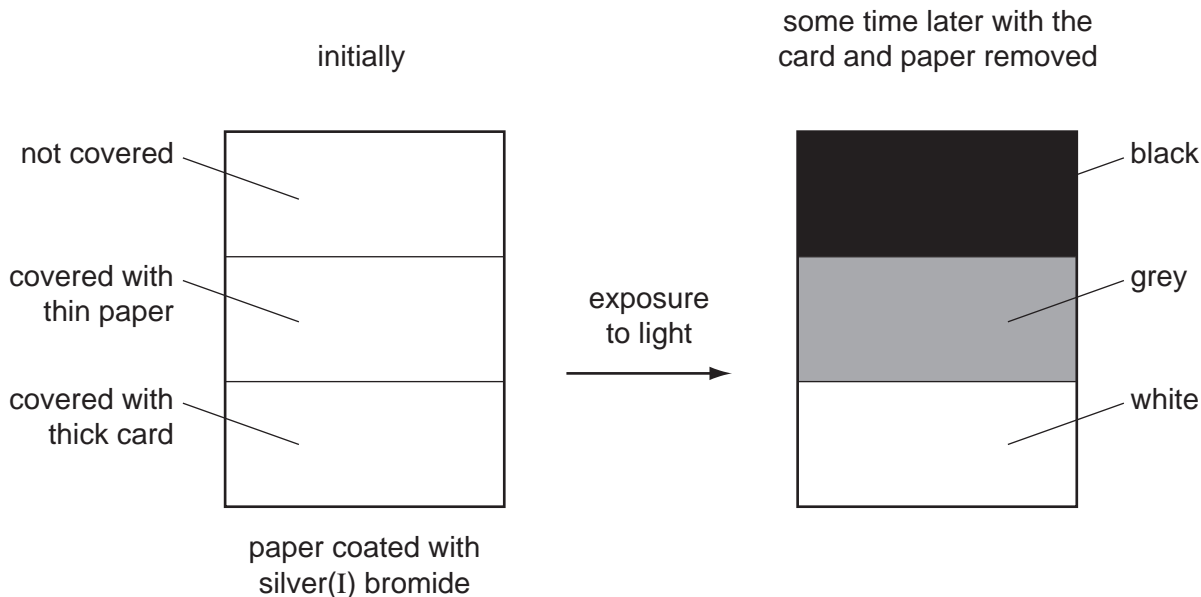
..... [1]

The decomposition of silver(I) bromide is the basis of film photography. The equation for this decomposition is:



This reaction is photochemical.

A piece of white paper was coated with silver(I) bromide and the following experiment was carried out.

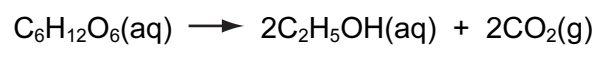


(b) Explain the results.

.....

 [3]

(c) The fermentation of glucose is catalysed by enzymes from yeast. Yeast is added to an aqueous glucose, the solution starts to bubble and becomes cloudy as more yeast cells are formed.



The reaction is exothermic.

Eventually the fermentation stops when the concentration of ethanol is about 12%.

(i) What is an enzyme?

..... [1]

(ii) Pasteur said that fermentation was respiration in the absence of air. Suggest a definition of *respiration*.

.....
..... [2]

(iii) On a large scale, the reaction mixture is cooled. Suggest a reason why this is necessary.

..... [1]

(iv) Why does the fermentation stop? Suggest **two** reasons.

.....
..... [2]

(v) When the fermentation stops, there is a mixture of dilute aqueous ethanol and yeast. Suggest a technique which could be used to remove the cloudiness due to the yeast.

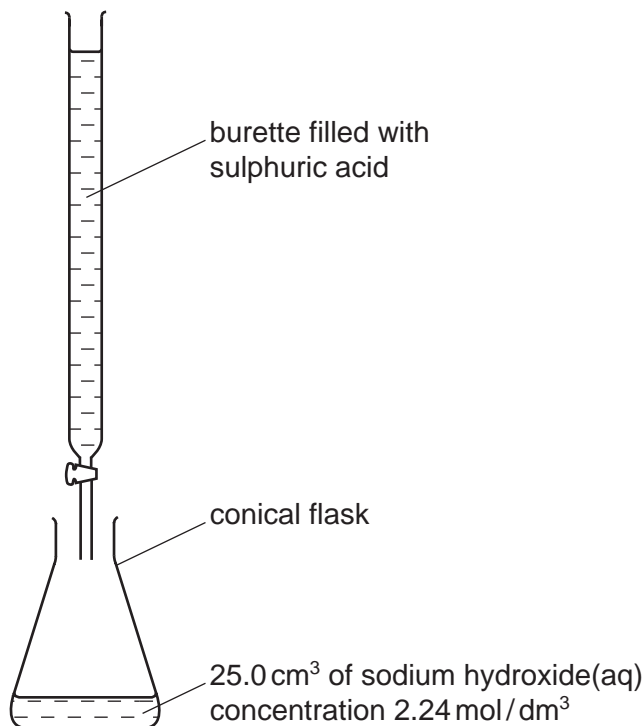
..... [1]

Name a technique which will separate the ethanol from the ethanol/water mixture.

..... [1]

[Total: 14]

7 Crystals of sodium sulphate-10-water, $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$, are prepared by titration.



- (a) 25.0 cm³ of aqueous sodium hydroxide is pipetted into a conical flask. A few drops of an indicator are added. Using a burette, dilute sulphuric acid is slowly added until the indicator just changes colour. The volume of acid needed to neutralise the alkali is noted.

Suggest how you would continue the experiment to obtain pure, dry crystals of sodium sulphate-10-water.

.....

.....

.....

[4]

- (b) Using 25.0 cm³ of aqueous sodium hydroxide, 2.24 mol / dm³, 3.86 g of crystals were obtained. Calculate the percentage yield.



Number of moles of NaOH used =

Maximum number of moles of $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ that could be formed =

Mass of one mole of $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ = 322 g

Maximum yield of sodium sulphate-10-water = g

Percentage yield = % [4]

[Total: 8]

DATA SHEET
The Periodic Table of the Elements

		Group																																																																																																								
I	II	III	IV	V	VI	VII	0					0																																																																																														
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10	23 Na Sodium 11	24 Mg Magnesium 12	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulphur 16	35.5 Cl Chlorine 17	40 Ar Argon 18	39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36	85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	131 Xe Xenon 54	133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	226 Fr Francium 87	226 Ra Radium 88	227 Ac Actinium 89	227 Th Thorium 90	232 Pa Protactinium 91	238 U Uranium 92	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103	238 Lu Lutetium 71	238 Yb Ytterbium 70	238 Er Erbium 68	238 Ho Holmium 67	238 Dy Dysprosium 66	238 Tb Terbium 65	238 Gd Gadolinium 64	238 Eu Europium 63	238 Sm Samarium 62	238 Pm Promethium 61	238 Ce Cerium 58	238 Pr Praseodymium 59	238 Nd Neodymium 60	238 Os Osmium 76	238 Ir Iridium 77	238 Pt Platinum 78	238 Au Gold 79	238 Hg Mercury 80	238 Tl Thallium 81	238 Pb Lead 82	238 Bi Bismuth 83	238 Po Polonium 84	238 At Astatine 85	238 Rn Radon 86

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X	a = relative atomic mass
	X	X = atomic symbol
b		b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).



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NAME

CENTRE
NUMBER

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CHEMISTRY

0620/32

Paper 3 (Extended)

May/June 2008

1 hour 15 minutes

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(b) It forms a covalent compound with hydrogen having the formula H_2X .

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(c) A metal that reacts violently with cold water.

.....

(d) It has a complete outer energy level.

.....

(e) It has oxidation states of 2 and 3 only.

.....

(f) It can form an ion of the type X^+ .

.....

(g) This metal is the catalyst in the Haber Process.

.....

[Total: 7]

2 (a) Complete the table which gives the names, symbols, relative masses and charges of the three subatomic particles.

name	symbol	relative mass	relative charge
electron	e ⁻		
proton		1	
neutron	n		

[3]

(b) Use the information in the table to explain the following.

(i) Atoms contain charged particles but they are electrically neutral - they have no overall charge.

.....
..... [2]

(ii) Atoms can form negative ions.

.....
..... [2]

(iii) Different atoms of the element chlorine are $^{35}_{17}\text{Cl}$ and $^{37}_{17}\text{Cl}$.

How are they different?

How are they the same? [2]

(iv) Scientists are certain that there are no undiscovered elements missing from the Periodic Table from hydrogen to lawrencium.

..... [1]

[Total: 10]

3 Copper is purified by electrolysis.

(a) Complete the following.

The positive electrode (anode) is made from

The negative electrode (cathode) is made from

The electrolyte is aqueous [3]

(b) Write an ionic equation for the reaction at the positive electrode (anode).

..... [2]

(c) (i) Give **two** reasons why copper is used,

in electric wiring, [2]
.....

in cooking utensils. [2]
.....

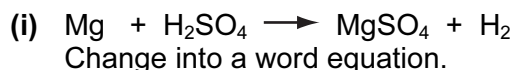
(ii) Give another use of copper.

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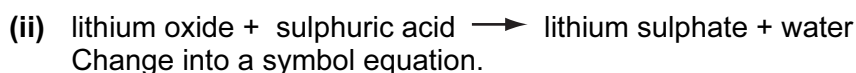
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4 Sulphuric acid is a typical strong acid.

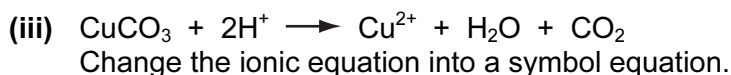
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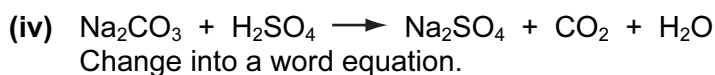
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..... [2]

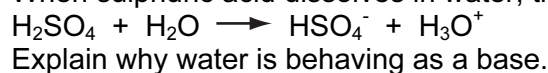


..... [2]



..... [1]

(b) When sulphuric acid dissolves in water, the following reaction occurs.



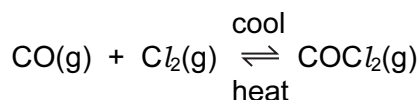
..... [2]

(c) Sulphuric acid is a strong acid, ethanoic acid is a weak acid. One way of distinguishing between them is to measure their pH. The weaker acid will have the higher pH. Describe another way by which they could be distinguished.

..... [2]
.....

[Total: 10]

- 5 Carbonyl chloride, COCl_2 , is a colourless gas. It is made by the following reaction.



- (a) When the pressure on the equilibrium mixture is increased, the position of equilibrium moves to right.

- (i) How does the concentration of each of the three chemicals change?

.....
 [2]

- (ii) Explain why the position of equilibrium moves to right.

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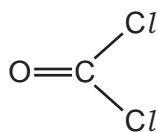
- (b) Using the information given with the equation, is the forward reaction exothermic or endothermic? Give a reason for your choice.

.....
 [2]

- (c) Carbonyl chloride reacts with water to form two acidic compounds. Name them.

..... [2]

- (d) The structural formula of carbonyl chloride is given below.



Draw a diagram that shows the arrangement of the valency electrons in one molecule of this covalent compound.

Use x for an electron from a chlorine atom.

Use o for an electron from a carbon atom.

Use ● for an electron from an oxygen atom.

[4]
 [Total: 12]

6 Three of the factors that can influence the rate of a chemical reaction are:

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(a) The first recorded dust explosion was in a flour mill in Italy in 1785. Flour contains carbohydrates. Explosions are very fast exothermic reactions.

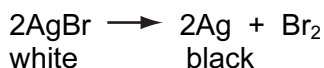
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(ii) Write a word equation for this exothermic reaction.

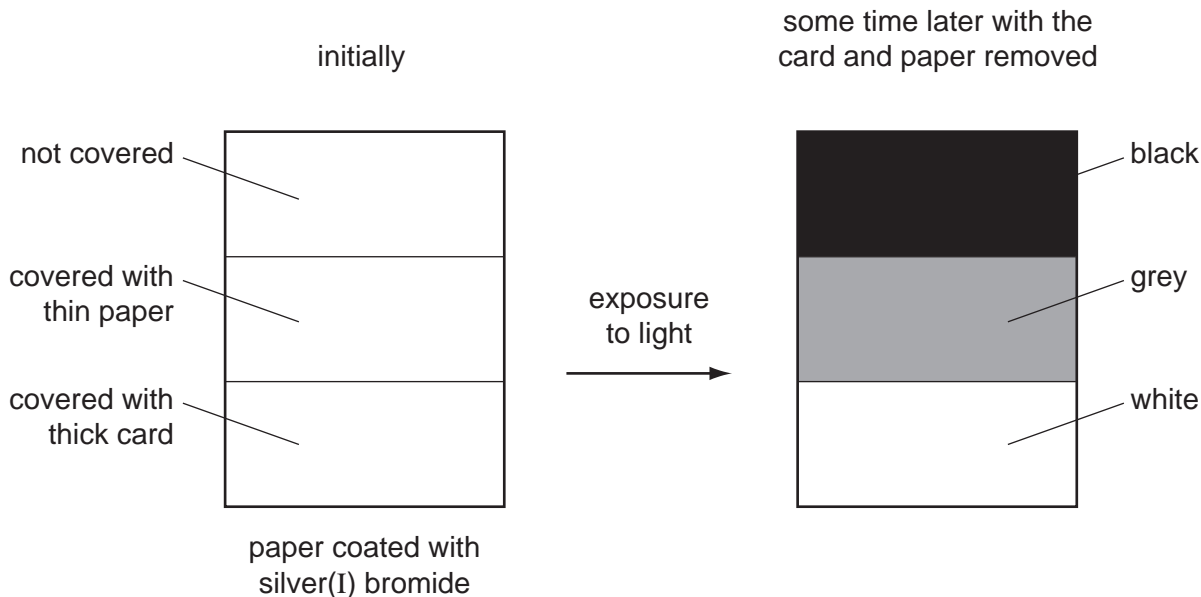
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(b) This reaction is photochemical.

A piece of white paper was coated with silver(I) bromide and the following experiment was carried out.

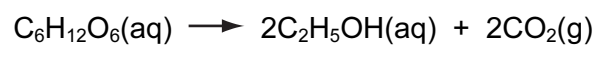


Explain the results.

.....

 [3]

(c) The fermentation of glucose is catalysed by enzymes from yeast. Yeast is added to an aqueous glucose, the solution starts to bubble and becomes cloudy as more yeast cells are formed.



The reaction is exothermic.

Eventually the fermentation stops when the concentration of ethanol is about 12%.

(i) What is an enzyme?

..... [1]

(ii) Pasteur said that fermentation was respiration in the absence of air. Define *respiration*.

.....
..... [2]

(iii) On a large scale, the reaction mixture is cooled. Suggest a reason why this is necessary.

..... [1]

(iv) Why does the fermentation stop? Suggest **two** reasons.

.....
..... [2]

(v) When the fermentation stops, there is a mixture of dilute aqueous ethanol and yeast. Suggest a technique which could be used to remove the cloudiness due to the yeast.

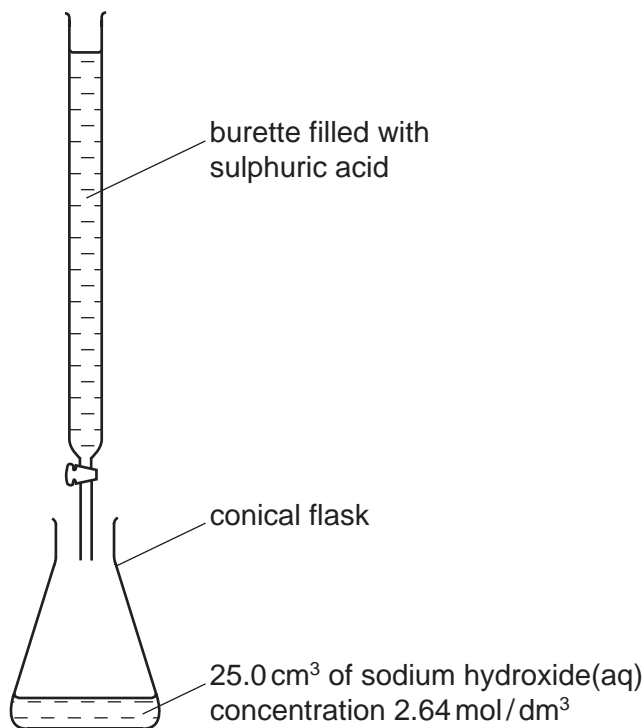
..... [1]

Name another technique which will separate the ethanol from the ethanol / water mixture.

..... [1]

[Total: 14]

7 Crystals of sodium sulphate-10-water, Na₂SO₄.10H₂O, are prepared by titration.



(a) 25.0 cm³ of aqueous sodium hydroxide is pipetted into a conical flask. A few drops of an indicator are added. Using a burette, dilute sulphuric acid is slowly added until the indicator just changes colour. The volume of acid needed to neutralise the alkali is noted.

Suggest how you would continue the experiment to obtain pure, dry crystals of sodium sulphate-10-water.

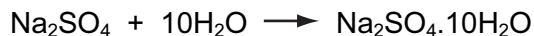
.....

.....

.....

[4]

(b) Using 25.0 cm³ of aqueous sodium hydroxide, 2.64 mol / dm³, 3.95 g of crystals were obtained. Calculate the percentage yield.



Number of moles of NaOH used =

Maximum number of moles of Na₂SO₄.10H₂O that could be formed =

Mass of one mole of Na₂SO₄.10H₂O = 322 g

Maximum yield of sodium sulphate-10-water = g

Percentage yield = % [4]

[Total: 8]

