



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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CANDIDATE  
NUMBER

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**CHEMISTRY**

**0620/22**

Paper 2

**May/June 2011**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may need to use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 16.

At the end of the examination, fasten all your work securely together.

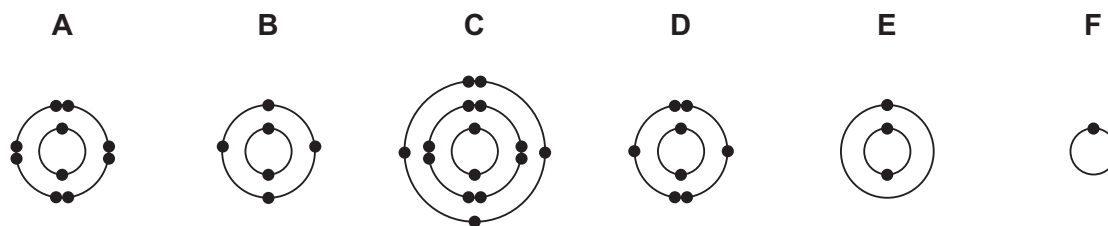
The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
<b>Total</b>	

This document consists of **14** printed pages and **2** blank pages.



1 The diagram shows the electronic structures of six atoms.



(a) Answer the following questions by choosing from the structures **A**, **B**, **C**, **D**, **E** or **F**. You can use each structure once, more than once or not at all.

Which structure represents

- (i) an atom in Period 3 of the Periodic Table,
- (ii) an atom containing six protons,
- (iii) an atom of a Group I metal,
- (iv) an atom of phosphorus,
- (v) an atom with six electrons in its outer shell,
- (vi) an atom with a complete outer shell of electrons?

[6]

(b) Atoms **D** and **F** can combine to form a covalent molecule.

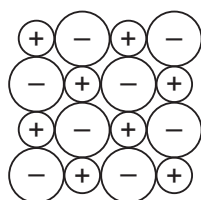
(i) Complete the following sentence using words from the list.

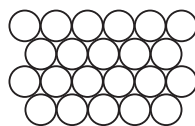
**atoms      electrons      ions      neutrons      solids**

A covalent bond is formed when a pair of ..... is shared between two .....

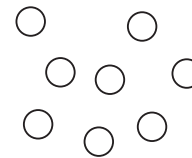
[2]

(ii) Which one of the following structures represents an ionic compound? Tick **one** box.










[1]

[Total: 9]

- 2 Many metal oxides can be reduced with carbon. The table shows the temperatures required to produce different metals from their oxides by heating with carbon.

reaction	temperature/°C
aluminium oxide → aluminium	2100
iron oxide → iron	425
nickel oxide → nickel	475
zinc oxide → zinc	925

- (a) (i) Use the information in the table to arrange aluminium, iron, nickel and zinc in order of their reactivity.

least reactive  $\xrightarrow{\hspace{15em}}$  most reactive

[1]

- (ii) Suggest why aluminium is extracted by electrolysis rather than by heating with carbon.

.....  
 ..... [1]

- (iii) State the name of the main ore of aluminium.

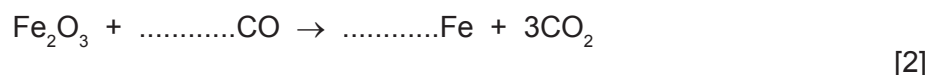
..... [1]

- (b) Iron is extracted by heating iron ore with carbon in a blast furnace.

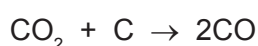
- (i) Apart from iron ore and carbon (coke), state the names of **two** other raw materials used in the blast furnace for the extraction of iron.

..... and ..... [2]

- (ii) Complete this equation for the reaction of iron(III) oxide with carbon.



- (iii) In the blast furnace, carbon dioxide reacts with red hot carbon to form carbon monoxide.



Which substance gets reduced during this reaction? Explain your answer.

.....  
 ..... [2]

(iv) State **one** adverse effect of carbon monoxide on health.

..... [1]

(v) The reaction between carbon dioxide and red hot carbon is endothermic.  
What do you understand by the term *endothermic*?

..... [1]

(c) Iron is usually made into steel alloys.

(i) What do you understand by the term *alloy*?

.....  
..... [1]

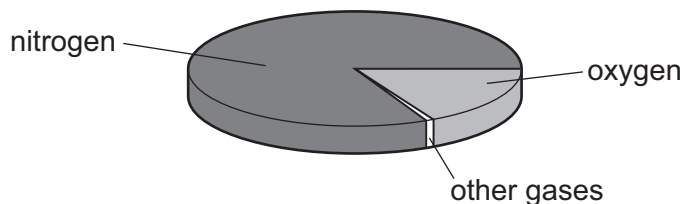
(ii) Mild steel is an alloy.  
State **one** use of mild steel.

..... [1]

[Total: 13]

5

- 3 The pie chart shows the composition of air.



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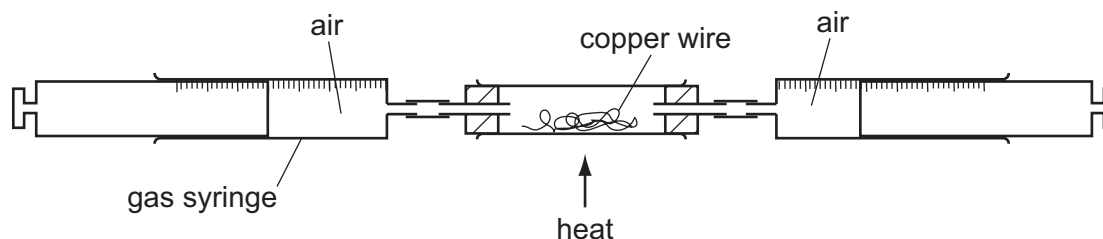
- (a) (i) What is the percentage of nitrogen in the air?

..... [1]

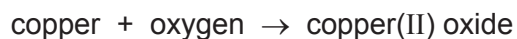
- (ii) Apart from nitrogen and oxygen, state the names of **two** gases present in unpolluted air.

..... and ..... [2]

- (b) The percentage of oxygen in air can be found using the apparatus shown below.



Air is passed backwards and forwards over the heated copper using the syringes. The copper reacts with oxygen in the air.



As the experiment proceeds, suggest what happens to

- (i) the total volume of air in the gas syringes,

..... [1]

- (ii) the mass of the wire in the tube.

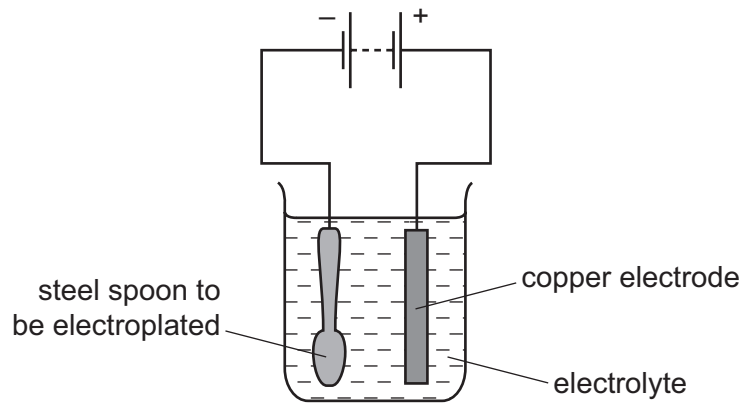
..... [1]

- (c) State **one** use of copper.

..... [1]

6

(d) A steel spoon can be electroplated using the apparatus shown.



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Use

Give a description of this electroplating.  
In your answer, refer to:

- a suitable electrolyte that can be used;
- the anode and the cathode;
- changes to the spoon.

.....

.....

.....

.....

.....

..... [3]

[Total: 9]

4 Ethanol can be produced by fermentation or by the catalytic addition of steam to ethene.

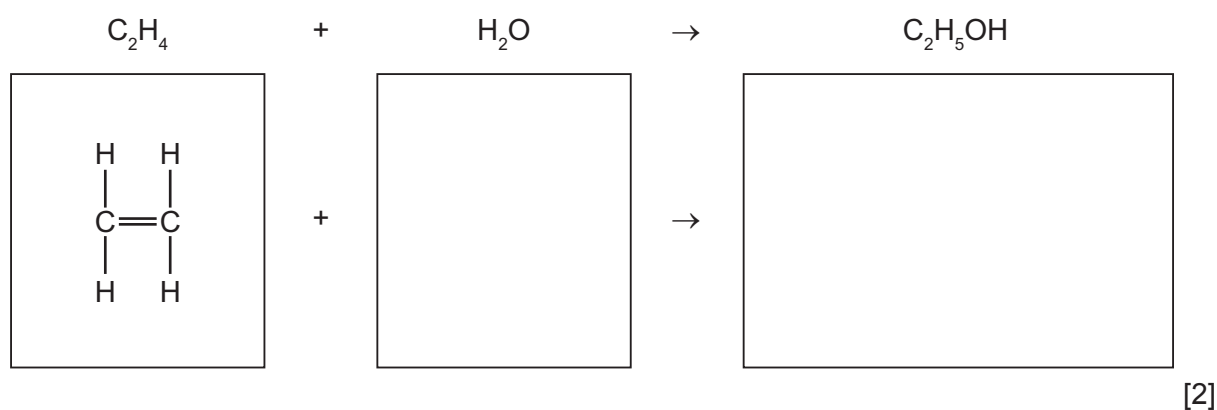
(a) (i) Complete the word equation for fermentation.

glucose  $\rightarrow$  ..... + ethanol [1]

(ii) State the conditions needed for fermentation.

.....  
..... [2]

(b) (i) Complete the equation for the catalytic addition of steam to ethene by drawing the structures of water and ethanol in the boxes.



(ii) Ethene is an unsaturated hydrocarbon.  
Describe a test for an unsaturated hydrocarbon.

test .....

result ..... [2]

(c) Ethanol can be used as a fuel.

State the names of the products formed when ethanol undergoes complete combustion.

..... and ..... [2]

(d) Complete these sentences using words from the list.

**different   functional   homologous   similar   unreactive   unsaturated**

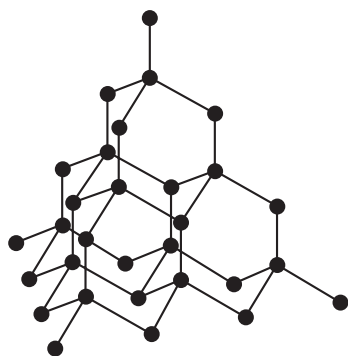
Ethanol is a member of the alcohol ..... series.

All alcohols have ..... chemical properties because they contain the same  
..... group. [3]

[Total: 12]

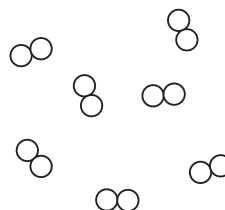
- 5 The structures of diamond and chlorine are shown below.

For  
Examiner's  
Use



diamond

● = carbon atom



chlorine

○ = chlorine atom

- (a) Describe the structure of these two substances.  
Use the list of words to help you.

**covalent    diatomic    giant    macromolecule    molecule    structure**

diamond .....

.....

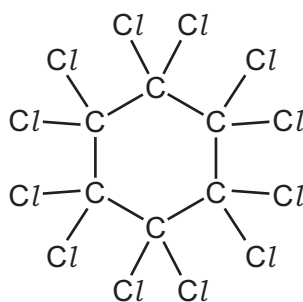
.....

chlorine .....

.....

..... [4]

- (b) The structure of a compound containing carbon and chlorine is shown below.



What is the molecular formula of this compound?

..... [1]



(c) Chlorine is a halogen.

(i) State the colour of chlorine.

..... [1]

The table shows some properties of the halogens.

element	boiling point / °C	density in liquid state / g per cm <sup>3</sup>	colour
fluorine	-188	1.51	yellow
chlorine	-35	1.56	
bromine	-7		red-brown
iodine	+114	4.93	grey-black

Use the information in the table to answer the following questions.

(ii) Predict the density of liquid bromine.

..... [1]

(iii) Describe the trend in boiling point of the halogens down the group.

..... [1]

(d) (i) Complete the word equation for the reaction of bromine with aqueous potassium iodide.

bromine + potassium iodide → ..... + .....

..... [2]

(ii) Explain why bromine does **not** react with aqueous potassium chloride.

..... [1]

(e) Potassium chloride is an ionic substance but iodine is a molecular substance. How do most ionic and molecular substances differ in their

solubility in water, .....

.....

electrical conductivity? .....

..... [2]

[Total: 13]

6 Iron(II) sulfate is a light green salt.

(a) Describe how you can prepare pure dry crystals of iron(II) sulfate from iron powder and dilute sulfuric acid.

.....

.....

.....

.....

..... [3]

(b) The formula for iron(II) sulfate is  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ .

(i) What is the meaning of the symbol (II) in this formula?

..... [1]

(ii) Describe a test for iron(II) ions.

test .....

result ..... [3]

(iii) A student warmed a sample of iron(II) sulfate,  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ , in a test-tube. Drops of a colourless liquid appeared at the mouth of the test-tube and the iron(II) sulfate turned white.

Explain these observations.

..... [1]

(iv) The reaction in part (iii) can be reversed.

Write down the symbol which shows that a reaction is reversible.

[1]

(c) Many centuries ago, sulfuric acid was made from iron(II) sulfate. Sulfuric acid is a typical acid.

(i) Describe what you would observe when sulfuric acid is added to

blue litmus paper, .....

iron powder. .... [2]

- (ii) Farmers can add a solution of dilute sulfuric acid to their fields if the soil is too alkaline.

Explain why farmers do **not** want their fields to become too alkaline.

..... [1]

- (iii) Which **one** of the following pH values best describes a solution which is slightly alkaline?

Put a ring around the correct answer.

pH 1

pH 6

pH 7

pH 8

pH 14

[1]

- (iv) Give the name of a compound which farmers add to the soil to make it less acidic.

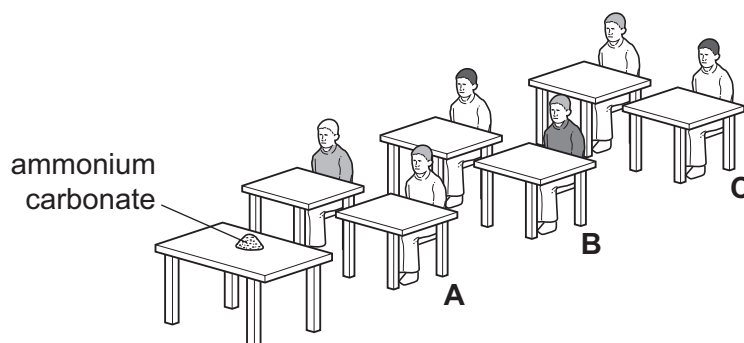
..... [1]

[Total: 14]

- 7 Smelling salts are sometimes used to 'wake up' people who feel faint. The main chemical in smelling salts is ammonium carbonate. Ammonium carbonate breaks down slowly at room temperature to release ammonia gas.



- (a) A few crystals of ammonium carbonate were put on a bench at the front of a classroom.



The students in row **A** could smell the ammonia 10 seconds after the smelling salts had been put on the bench. The students in row **C** could smell the ammonia after 40 seconds.

- (i) Suggest how long it took the students in row **B** to smell the ammonia.

..... [1]

- (ii) Explain these results using ideas about moving particles.

.....

.....

..... [3]

- (b) Ammonium carbonate has the formula  $(\text{NH}_4)_2\text{CO}_3$ . Calculate the relative formula mass of ammonium carbonate.

[1]

- (c) Ammonia is used to make fertilisers such as ammonium sulfate.

- (i) State the names of the **three** chemical elements, essential for plant growth, that are present in fertilisers.

..... [3]

- (ii) Which one of the following statements about ammonium sulfate is correct?  
Tick **one** box.

When ammonium sulfate is heated with sodium hydroxide, carbon dioxide is given off.

When ammonium sulfate is added to a solution of barium chloride, a blue precipitate is formed.

Ammonium sulfate can be made by adding sulfuric acid to ammonia.

Farmers add ammonium sulfate to the soil to make it alkaline.

[1]

- (d) Ammonium carbonate is made by heating ammonium sulfate with excess calcium carbonate.

ammonium sulfate + calcium carbonate  $\rightarrow$  ammonium carbonate + calcium sulfate

When 132 g of ammonium sulfate is reacted with calcium carbonate, 96 g of ammonium carbonate is formed.

Calculate the mass of ammonium sulfate needed to make 240 g of ammonium carbonate.

[1]

[Total: 10]





**DATA SHEET**  
**The Periodic Table of the Elements**

Group																						
I	II	III	IV	V	VI	VII	0															
1 <b>H</b> Hydrogen 1								2 <b>He</b> Helium 2														
3 <b>Li</b> Lithium 3	4 <b>Be</b> Beryllium 4	5 <b>B</b> Boron 5	6 <b>C</b> Carbon 6	7 <b>N</b> Nitrogen 7	8 <b>O</b> Oxygen 8	9 <b>F</b> Fluorine 9	10 <b>Ne</b> Neon 10															
11 <b>Na</b> Sodium 11	12 <b>Mg</b> Magnesium 12	13 <b>Al</b> Aluminium 13	14 <b>Si</b> Silicon 14	15 <b>P</b> Phosphorus 15	16 <b>S</b> Sulfur 16	17 <b>Cl</b> Chlorine 17	18 <b>Ar</b> Argon 18															
19 <b>K</b> Potassium 19	20 <b>Ca</b> Calcium 20	21 <b>Sc</b> Scandium 21	22 <b>Ti</b> Titanium 22	23 <b>V</b> Vanadium 23	24 <b>Cr</b> Chromium 24	25 <b>Mn</b> Manganese 25	26 <b>Fe</b> Iron 26	27 <b>Co</b> Cobalt 27	28 <b>Ni</b> Nickel 28	29 <b>Cu</b> Copper 29	30 <b>Zn</b> Zinc 30	31 <b>Ga</b> Gallium 31	32 <b>Ge</b> Germanium 32	33 <b>As</b> Arsenic 33	34 <b>Se</b> Selenium 34	35 <b>Br</b> Bromine 35	36 <b>Kr</b> Krypton 36					
37 <b>Rb</b> Rubidium 37	38 <b>Sr</b> Strontium 38	39 <b>Y</b> Yttrium 39	40 <b>Zr</b> Zirconium 40	41 <b>Nb</b> Niobium 41	42 <b>Mo</b> Molybdenum 42	43 <b>Tc</b> Technetium 43	44 <b>Ru</b> Ruthenium 44	45 <b>Rh</b> Rhodium 45	46 <b>Pd</b> Palladium 46	47 <b>Ag</b> Silver 47	48 <b>Cd</b> Cadmium 48	49 <b>In</b> Indium 49	50 <b>Sn</b> Tin 50	51 <b>Sb</b> Antimony 51	52 <b>Te</b> Tellurium 52	53 <b>I</b> Iodine 53	54 <b>Xe</b> Xenon 54					
55 <b>Cs</b> Caesium 55	56 <b>Ba</b> Barium 56	57 <b>La</b> Lanthanum 57	72 <b>Hf</b> Hafnium 72	73 <b>Ta</b> Tantalum 73	74 <b>W</b> Tungsten 74	75 <b>Re</b> Rhenium 75	76 <b>Os</b> Osmium 76	77 <b>Ir</b> Iridium 77	78 <b>Pt</b> Platinum 78	79 <b>Au</b> Gold 79	80 <b>Hg</b> Mercury 80	81 <b>Tl</b> Thallium 81	82 <b>Pb</b> Lead 82	83 <b>Bi</b> Bismuth 83	84 <b>Po</b> Polonium 84	85 <b>At</b> Astatine 85	86 <b>Rn</b> Radon 86					
87 <b>Fr</b> Francium 87	88 <b>Ra</b> Radium 88	89 <b>Ac</b> Actinium 89 †																				

140 <b>Ce</b> Cerium 58	141 <b>Pr</b> Praseodymium 59	144 <b>Nd</b> Neodymium 60	150 <b>Sm</b> Samarium 62	152 <b>Eu</b> Europium 63	157 <b>Gd</b> Gadolinium 64	162 <b>Dy</b> Dysprosium 66	165 <b>Ho</b> Holmium 67	167 <b>Er</b> Erbium 68	169 <b>Tm</b> Thulium 69	173 <b>Yb</b> Ytterbium 70	175 <b>Lu</b> Lutetium 71
232 <b>Th</b> Thorium 90	238 <b>U</b> Uranium 92	238 <b>Pa</b> Protactinium 91	94 <b>Pu</b> Plutonium 94	95 <b>Am</b> Americium 95	96 <b>Cm</b> Curium 96	98 <b>Cf</b> Californium 98	99 <b>Es</b> Einsteinium 99	100 <b>Fm</b> Fermium 100	101 <b>Md</b> Mendelevium 101	102 <b>No</b> Nobelium 102	103 <b>Lr</b> Lawrencium 103

**\*58-71 Lanthanoid series**  
**†90-103 Actinoid series**

**Key**

a	<b>X</b>
b	

a = relative atomic mass  
X = atomic symbol  
b = proton (atomic) number

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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