



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/12

Paper 1 Multiple Choice (Core)

May/June 2018

45 minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

* 8 0 5 3 9 3 3 2 7 5 1 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **15** printed pages and **1** blank page.

2

1 When iodine is heated it turns from a solid to a gas.

When liquid ammonia is cooled it turns into a solid.

When ice is heated it turns into water.

Which terms describe these changes of state?

	when iodine is heated	when liquid ammonia is cooled	when ice is heated
A	boiling	freezing	melting
B	freezing	sublimation	boiling
C	sublimation	condensation	freezing
D	sublimation	freezing	melting

2 Which piece of apparatus **cannot** be used to collect and measure the volume of gas produced in an experiment?

A burette

B gas syringe

C measuring cylinder

D pipette

3 Pure ethanol has a melting point of $-114\text{ }^{\circ}\text{C}$ and a boiling point of $78\text{ }^{\circ}\text{C}$.

What are the melting and boiling points of a sample of ethanol with glucose dissolved in it?

	melting point/ $^{\circ}\text{C}$	boiling point/ $^{\circ}\text{C}$
A	-116	77
B	-116	79
C	-112	77
D	-112	79

4 Which atom has an equal number of protons, neutrons and electrons?

A ^{40}Ar

B ^1H

C ^{23}Na

D ^{14}N

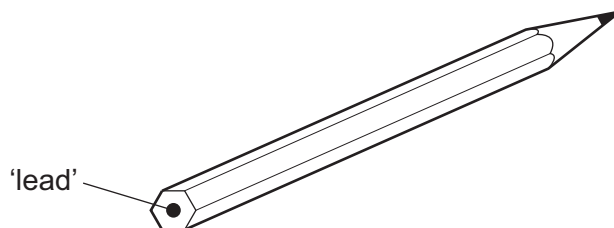
- 5 Which description of brass is correct?
- A alloy
 - B compound
 - C element
 - D non-metal
- 6 The bonding between elements X and Y in compound XY_2 is shown.



Which row shows the type of bond in XY_2 and the type of element X?

	type of bond	type of element X
A	covalent	metal
B	covalent	non-metal
C	ionic	metal
D	ionic	non-metal

- 7 The 'lead' in a pencil is made of a mixture of graphite and clay.



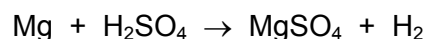
When the percentage of graphite is increased, the pencil slides across the paper more easily.

Which statement explains this observation?

- A Graphite has a high melting point.
- B Graphite is a form of carbon.
- C Graphite is a lubricant.
- D Graphite is a non-metal.

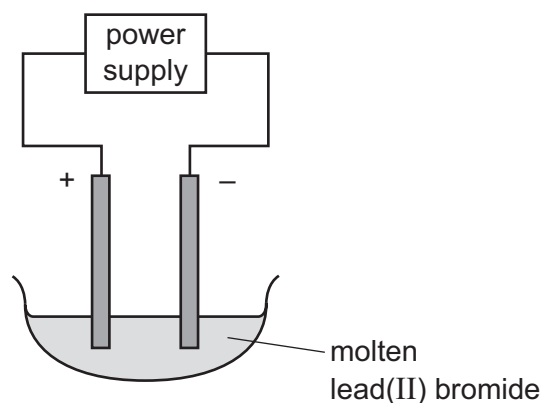
- 8 The equation for the reaction between magnesium and dilute sulfuric acid is shown.

The M_r of $MgSO_4$ is 120.



Which mass of magnesium sulfate is formed when 12g of magnesium completely reacts with dilute sulfuric acid?

- A** 5g **B** 10g **C** 60g **D** 120g
- 9 The electrolysis of molten lead(II) bromide is shown.



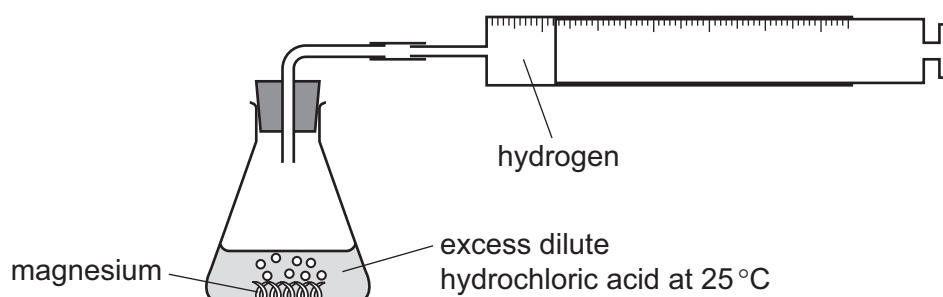
Which statement describes what happens at the negative electrode?

- A** Bromide ions gain electrons to form bromine molecules.
B Bromine molecules gain electrons to form bromide ions.
C Lead atoms lose electrons to form lead ions.
D Lead ions gain electrons to form lead atoms.
- 10 Which statement about the combustion of fuels is correct?
- A** It always produces carbon dioxide.
B It always produces carbon monoxide.
C It is always endothermic.
D It is always exothermic.

11 Which statement about chemical reactions is correct?

- A Endothermic reactions show a temperature decrease because energy is absorbed from the surroundings.
- B Endothermic reactions show a temperature increase because energy is released into the surroundings.
- C Exothermic reactions show a temperature increase because energy is absorbed from the surroundings.
- D Exothermic reactions show a temperature decrease because energy is released into the surroundings.

12 The diagram shows a rate of reaction experiment.

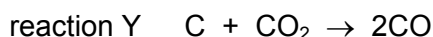
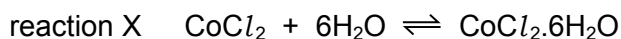


Increasing the concentration of the acid and increasing the temperature both affect the rate of reaction.

Which row is correct?

	increase the concentration of acid	increase the temperature
A	decrease rate of reaction	decrease rate of reaction
B	decrease rate of reaction	increase rate of reaction
C	increase rate of reaction	decrease rate of reaction
D	increase rate of reaction	increase rate of reaction

- 13 Reaction X shows a test for water. Reaction Y occurs in the blast furnace for extracting iron.

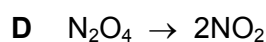
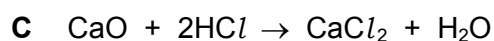
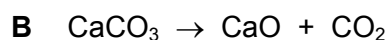
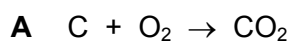


Reaction X is1..... . In reaction Y, the oxide CO_2 is2..... .

Which words correctly complete gaps 1 and 2?

	1	2
A	irreversible	oxidised
B	irreversible	reduced
C	reversible	oxidised
D	reversible	reduced

- 14 Which equation shows an oxidation reaction?



- 15 Which two gases each give the same result for the test shown?

	test	gas 1	gas 2
A	damp blue litmus paper	ammonia	chlorine
B	damp blue litmus paper	ammonia	oxygen
C	lighted splint	hydrogen	chlorine
D	lighted splint	hydrogen	oxygen

- 16 Which statement about oxides is correct?

A A solution of magnesium oxide has a pH less than pH 7.

B A solution of sulfur dioxide has a pH greater than pH 7.

C Magnesium oxide reacts with nitric acid to make a salt.

D Sulfur dioxide reacts with hydrochloric acid to make a salt.

17 Which methods are suitable for preparing **both** zinc sulfate and copper(II) sulfate?

- 1 reacting the metal oxide with warm dilute aqueous sulfuric acid
- 2 reacting the metal with dilute aqueous sulfuric acid
- 3 reacting the metal carbonate with dilute aqueous sulfuric acid

A 1, 2 and 3 **B** 1 and 2 only **C** 1 and 3 only **D** 2 and 3 only

18 Two separate tests are done on separate solutions of compound X.

- 1 Addition of aqueous sodium hydroxide forms a green precipitate that dissolves in an excess of aqueous sodium hydroxide.
- 2 Addition of dilute nitric acid and aqueous silver nitrate forms a white precipitate.

What is compound X?

- A** chromium(III) carbonate
- B** chromium(III) chloride
- C** iron(II) carbonate
- D** iron(II) chloride

19 Which statement about the Periodic Table is correct?

- A** Elements in the same group have the same number of electron shells.
- B** It contains elements arranged in order of increasing proton number.
- C** Metals are on the right and non-metals are on the left.
- D** The most reactive elements are at the bottom of every group.

20 Chlorine, bromine and iodine are elements in Group VII of the Periodic Table.

Which statement about these elements is **not** correct?

- A** The colour gets darker down the group.
- B** The density increases down the group.
- C** They are all gases at room temperature and pressure.
- D** They are all non-metals.

21 Which row describes the properties of a transition element?

	property 1	property 2
A	forms colourless compounds	acts as a catalyst
B	forms colourless compounds	low electrical conductivity
C	high density	acts as a catalyst
D	high density	low electrical conductivity

22 Which statement about the elements in Group VIII is correct?

- A** They all form diatomic molecules.
- B** They all have eight electrons in their outer shells.
- C** They all react with oxygen to form oxides.
- D** They are all gases at room temperature.

23 Stainless steel is an alloy of iron, carbon and other metals.

Which row is correct?

	stainless steel is harder than pure iron	stainless steel resists corrosion better than pure iron
A	✓	✓
B	✓	x
C	x	✓
D	x	x

24 A student is given metal Z and its oxide.

The student does some experiments to find out the position of metal Z in the reactivity series.

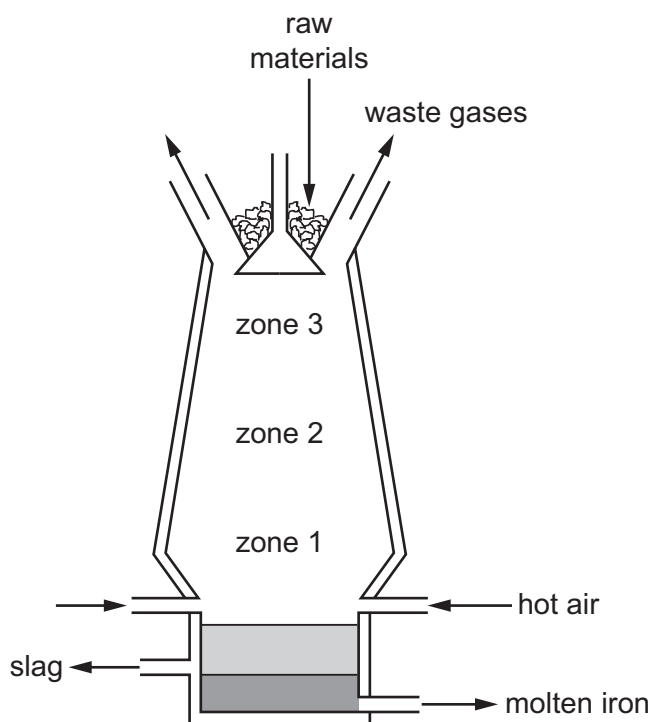
The results are shown.

- Metal Z reacted slowly with dilute hydrochloric acid.
- Metal Z reacted slowly with steam but not with water.
- The oxide of metal Z reacted when heated with carbon.

Which statement about the position of metal Z in the reactivity series is correct?

- A** It is between calcium and sodium.
- B** It is between copper and hydrogen.
- C** It is between hydrogen and iron.
- D** It is between magnesium and calcium.

25 Iron is produced from iron ore in a blast furnace.



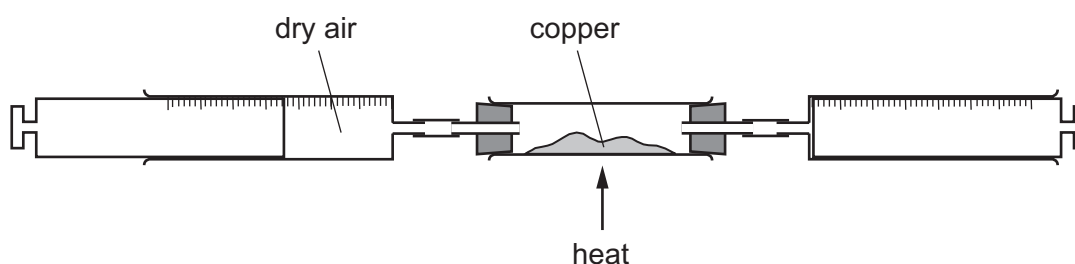
Which equation represents the main reaction that happens in zone 1?

- A** $C(s) + CO_2(g) \rightarrow 2CO(g)$
- B** $C(s) + O_2(g) \rightarrow CO_2(g)$
- C** $Fe_2O_3(s) + 3CO(g) \rightarrow 2Fe(l) + 3CO_2(g)$
- D** $Fe_3O_4(s) + CO(g) \rightarrow 3FeO(s) + CO_2(g)$

26 Which row describes the use of an alloy and the property upon which the use depends?

	alloy	use	property
A	mild steel	cutlery	resistant to corrosion
B	mild steel	machinery	strong
C	stainless steel	cutlery	low density
D	stainless steel	machinery	good conductor of electricity

27 Dry air is passed over hot copper until all the oxygen has reacted.



The volume of gas at the end of the reaction is 120 cm^3 .

What is the starting volume of dry air?

- A** 132 cm^3 **B** 152 cm^3 **C** 180 cm^3 **D** 570 cm^3

28 A steel bicycle which had been left outdoors for several months was starting to rust.

What would **not** reduce the rate of corrosion?

- A** Remove the rust and paint the bicycle.
B Remove the rust and store the bicycle in a dry shed.
C Remove the rust and wipe the bicycle with a clean, damp cloth.
D Remove the rust and wipe the bicycle with an oily cloth.

29 Which statements about water are correct?

- 1 Household water contains dissolved salts.
- 2 Water for household use is filtered to remove soluble impurities.
- 3 Water is treated with chlorine to kill bacteria.
- 4 Water is used in industry for cooling.

- A** 1, 2, 3 and 4
B 1, 2 and 3 only
C 1, 3 and 4 only
D 2, 3 and 4 only

30 Ammonium nitrate is a common fertiliser used by farmers to increase the yield of their crops.

Which compound reacts with ammonium nitrate to form ammonia?

- A** calcium hydroxide
B potassium nitrate
C sodium chloride
D sodium phosphate

31 Which process does **not** release a greenhouse gas?

- A** digestion of food in cows
B reaction between zinc and hydrochloric acid
C respiration by animals
D thermal decomposition of calcium carbonate

32 Which row describes the uses of sulfur and sulfur dioxide?

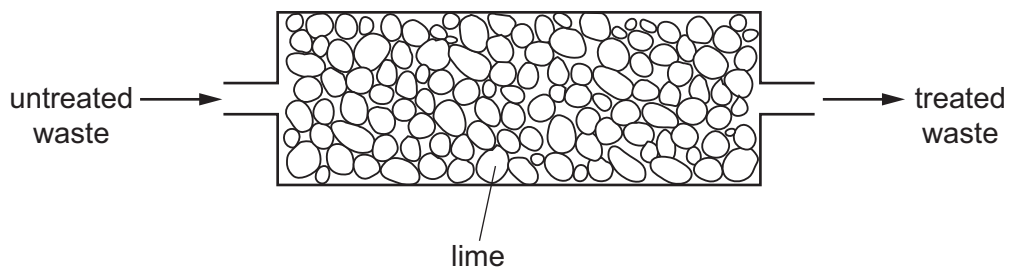
	sulfur	sulfur dioxide
A	extraction of aluminium	food preservative
B	extraction of aluminium	water treatment
C	manufacture of sulfuric acid	food preservative
D	manufacture of sulfuric acid	water treatment

33 Limestone is used in many industrial processes.

In which process is it **not** used?

- A manufacture of alkenes
- B manufacture of cement
- C manufacture of iron
- D manufacture of lime

34 Lime is used to treat an industrial waste.



Which change occurs in the treatment?

	untreated waste		treated waste
A	acidic	→	neutral
B	alkaline	→	acidic
C	alkaline	→	neutral
D	neutral	→	acidic

35 What is **not** the correct use of the fraction named?

	name of fraction	use
A	fuel oil	making waxes
B	gas oil	fuel in diesel engines
C	kerosene	jet fuel
D	naphtha	making chemicals

36 Four organic compounds are listed.

ethane

ethanoic acid

ethanol

ethene

Which bond do all four compounds contain?

- A C–C B C–H C C–O D O–H

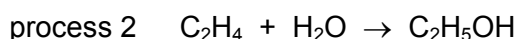
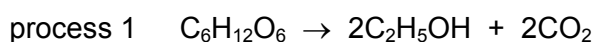
37 Which compounds belong to the same homologous series?

- A ethane and propane
B ethanoic acid and ethanol
C methane and ethene
D propene and ethanoic acid

38 Which substances can be obtained by cracking hydrocarbons?

- A ethanol and ethene
B ethanol and hydrogen
C ethene and hydrogen
D ethene and poly(ethene)

39 The equations for two important processes used to manufacture ethanol are shown.



Which statement is **not** correct?

- A Both processes require a catalyst.
B Both processes use a starting material obtained from petroleum.
C Process 1 shows the production of a renewable fuel.
D Process 2 is an addition reaction.

40 Part of the label on the packet of a potato product is shown.

This potato product contains:

starch
ethanoic acid
sodium chloride
sugar

Which constituent is a natural polymer?

- A ethanoic acid
- B sodium chloride
- C starch
- D sugar

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The Periodic Table of Elements

		Group											
I	II	III	IV	V	VI	VII	VIII						
3 Li lithium 7	4 Be beryllium 9	11 Na sodium 23	12 Mg magnesium 24	19 K potassium 39	20 Ca calcium 40	37 Rb rubidium 85	55 Cs caesium 133	87 Fr francium —	1 H hydrogen 1	2 He helium 4			
57 La lanthanum 139	89 Ac actinium —	72 Hf hafnium 178	74 W tungsten 184	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —
58 Ce cerium 140	90 Th thorium 232	73 Ta tantalum 181	75 Re rhenium 186	79 Au gold 197	80 Cu copper 64	81 Zn zinc 65	82 Ga gallium 70	83 Ge germanium 73	84 As arsenic 75	85 Se selenium 79	86 Br bromine 80	87 Kr krypton 84	88 Xe xenon 131
59 Pr praseodymium 141	91 Pa protactinium 231	74 Zr zirconium 91	76 Mo molybdenum 96	78 Pd palladium 106	80 Ni nickel 59	81 Co cobalt 59	82 Fe iron 56	83 Mn manganese 55	84 Cr chromium 52	85 V vanadium 51	86 Ti titanium 48	87 Sc scandium 45	88 Y yttrium 89
60 Nd neodymium 144	92 U uranium 238	75 Nb niobium 93	77 Tc technetium —	79 Ru ruthenium 101	80 Rh rhodium 103	81 Pd palladium 106	82 Ag silver 108	83 Cd cadmium 112	84 In indium 115	85 Sn tin 119	86 Sb antimony 122	87 Te tellurium 128	88 I iodine 127
61 Pm promethium —	93 Np neptunium —	76 Zr zirconium 91	78 Y yttrium 89	80 Zn zinc 65	81 Cu copper 64	82 Ni nickel 59	83 Co cobalt 59	84 Mn manganese 55	85 Cr chromium 52	86 V vanadium 51	87 Ti titanium 48	88 Sc scandium 45	89 Y yttrium 89
62 Sm samarium 150	94 Pu plutonium —	77 Hf hafnium 178	79 Re rhenium 186	81 Ru ruthenium 101	82 Rh rhodium 103	83 Pd palladium 106	84 Ag silver 108	85 Cd cadmium 112	86 In indium 115	87 Sn tin 119	88 Sb antimony 122	89 Te tellurium 128	90 I iodine 127
63 Eu europium 152	95 Am americium —	78 Ta tantalum 181	80 Os osmium 190	82 Pd palladium 106	83 Co cobalt 59	84 Ni nickel 59	85 Fe iron 56	86 Mn manganese 55	87 Cr chromium 52	88 V vanadium 51	89 Ti titanium 48	90 Sc scandium 45	91 Y yttrium 89
64 Gd gadolinium 157	96 Cm curium —	79 W tungsten 184	81 Ir iridium 192	83 Ru ruthenium 101	84 Rh rhodium 103	85 Pd palladium 106	86 Ag silver 108	87 Cd cadmium 112	88 In indium 115	89 Sn tin 119	90 Sb antimony 122	91 Te tellurium 128	92 I iodine 127
65 Tb terbium 159	97 Bk berkelium —	80 Hf hafnium 178	82 Pt platinum 195	84 Rh rhodium 103	85 Co cobalt 59	86 Ni nickel 59	87 Fe iron 56	88 Mn manganese 55	89 Cr chromium 52	90 V vanadium 51	91 Ti titanium 48	92 Sc scandium 45	93 Y yttrium 89
66 Dy dysprosium 163	98 Cf californium —	81 Ta tantalum 181	83 Au gold 197	85 Pd palladium 106	86 Co cobalt 59	87 Ni nickel 59	88 Fe iron 56	89 Mn manganese 55	90 Cr chromium 52	91 V vanadium 51	92 Ti titanium 48	93 Sc scandium 45	94 Y yttrium 89
67 Ho holmium 165	99 Es einsteinium —	82 W tungsten 184	84 Pt platinum 195	86 Ru ruthenium 101	87 Rh rhodium 103	88 Pd palladium 106	89 Ag silver 108	90 Cd cadmium 112	91 In indium 115	92 Sn tin 119	93 Sb antimony 122	94 Te tellurium 128	95 I iodine 127
68 Er erbium 167	100 Fm fermium —	83 Re rhenium 186	85 Au gold 197	87 Rh rhodium 103	88 Co cobalt 59	89 Ni nickel 59	90 Fe iron 56	91 Mn manganese 55	92 Cr chromium 52	93 V vanadium 51	94 Ti titanium 48	95 Sc scandium 45	96 Y yttrium 89
69 Tm thulium 169	101 Md mendelevium —	84 Os osmium 190	86 Pt platinum 195	88 Ru ruthenium 101	89 Rh rhodium 103	90 Pd palladium 106	91 Ag silver 108	92 Cd cadmium 112	93 In indium 115	94 Sn tin 119	95 Sb antimony 122	96 Te tellurium 128	97 I iodine 127
70 Yb ytterbium 173	102 No nobelium —	85 Ir iridium 192	87 Pt platinum 195	89 Rh rhodium 103	90 Co cobalt 59	91 Ni nickel 59	92 Fe iron 56	93 Mn manganese 55	94 Cr chromium 52	95 V vanadium 51	96 Ti titanium 48	97 Sc scandium 45	98 Y yttrium 89
71 Lu lutetium 175	103 Lr lawrencium —	86 Au gold 197	88 Pt platinum 195	90 Rh rhodium 103	91 Co cobalt 59	92 Ni nickel 59	93 Fe iron 56	94 Mn manganese 55	95 Cr chromium 52	96 V vanadium 51	97 Ti titanium 48	98 Sc scandium 45	99 Y yttrium 89

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).