

MARK SCHEME for the May/June 2013 series

9185 CHEMISTRY (US)

9185/23

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Cambridge is publishing the mark schemes for the May/June 2013 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

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	GCE AS/A LEVEL – May/June 2013	9185

1 (a) (i)



S atom has 6 **and** C atom has 4 electrons (1)

S=C double bonds (4 electrons) clearly shown (1)

(ii) linear **and** 180° (1) [3]

(b) (i) $\text{CS}_2 + 3\text{O}_2 \rightarrow \text{CO}_2 + 2\text{SO}_2$ (1)

(ii) enthalpy change when 1 mol of a substance (1)

is burnt in an excess of oxygen/air

or is completely combusted

under standard conditions (1) [3]

(c)

$$\begin{array}{rclcl}
 \text{CS}_2 & + & 3\text{O}_2 & \rightarrow & \text{CO}_2 & + & 2\text{SO}_2 & & \\
 \Delta H_f^\ominus / \text{kJ mol}^{-1} & & & & -395 & & 2(-298) & & (1) \\
 \Delta H_{\text{reaction}} = -395 + 2(-298) - x = -1110 \text{ kJ mol}^{-1} & & & & & & & & (1) \\
 \text{gives } x = -395 + (-596) + 1110 = +119 \text{ kJ mol}^{-1} & & & & & & & & (1) [3]
 \end{array}$$

(d) (i) $\text{CS}_2 + 2\text{NO} \rightarrow \text{CO}_2 + 2\text{S} + \text{N}_2$
or
 $\text{CS}_2 + 2\text{NO} \rightarrow \text{CO} + 2\text{S} + \text{N}_2\text{O}$

correct products (1)

correct equation (1)

(ii) from -2 to 0 **both** required (1) [3]

[Total: 12]

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2 (a) (i) if the conditions of a system in equilibrium are changed

the position of equilibrium moves so as to reduce that change (1)

(ii) lower temperature (1)

because the forward reaction is exothermic (1)

higher pressure (1)

because the forward reaction shows a reduction in volume

or

there are fewer molecules/moles on RHS of equilibrium (1) [4]

(b)	CO ₂	+	H ₂	⇌	CO	+	H ₂ O	
initial moles	0.70		0.70		0.30		0.30	
equil. moles	(0.70-x)		(0.70-x)		(0.30+x)		(0.30+x)	(1)
equil. concn.	$\frac{(0.70-x)}{1}$		$\frac{(0.70-x)}{1}$		$\frac{(0.30+x)}{1}$		$\frac{(0.30+x)}{1}$	

$$K_c = \frac{(0.30+x)^2}{(0.70-x)^2} = 1.44 \quad (1)$$

gives $x = 0.25$ (1)

at equilibrium,

$$n(\text{CO}_2) = n(\text{H}_2) = 0.70 - 0.25 = 0.45 \text{ moles}$$

and

$$n(\text{CO}) = n(\text{H}_2\text{O}) = 0.3 + 0.25 = 0.55 \text{ moles} \quad (1) \quad [4]$$

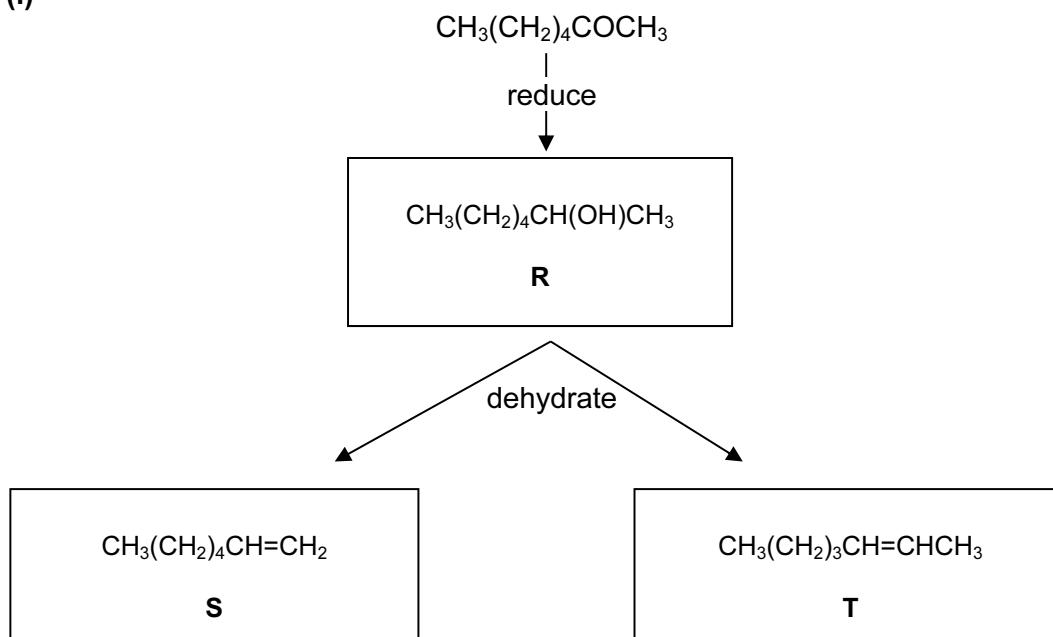
[Total: 10]

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- 3 (a) (i) He or Ne or Ar or Kr
- (ii) P or As (1)
- (iii) Br (1)
- (iv) Na allow Ar (1)
- (v) Si (1)
- (vi) P allow Si (1)
- (vii) Cl or F or Br (1) [7]
- (b) (i) any **two** from P_4O_6 , SO_2 and Cl_2O_7 (1+1)
- (ii) Al_2O_3 or SiO_2 (1)
- (iii) $MgSO_3$ (1) [4]
- (c) (i) Si is giant molecular/giant covalent or
P, S, and Cl are simple molecular (1)
- (ii) the molecules are S_8 , P_4 , Cl_2 (1)
- larger molecules have more electrons (1)
- and hence greater van der Waals' forces (1) [4]

[Total: 15]

4 (a) (i)



one mark for each correct compound, **R**, **S** and **T**

allow correct *cis* **and** *trans* versions of compound **T** for 2 marks (3 × 1)

(ii) reduction

NaBH_4 **or** LiAlH_4 **or** H_2/Ni **or** $\text{Na}/\text{C}_2\text{H}_5\text{OH}$ (1)

dehydration

$\text{P}_4\text{O}_{10}/\text{P}_2\text{O}_5$ **or** H_3PO_4 **or** conc. H_2SO_4 **or** Al_2O_3 (1) [5]

(b)

Tollens' reagent	NO REACTION
HCN	$\begin{array}{c} \text{CH}_3(\text{CH}_2)_4\text{C}(\text{OH})\text{CH}_3 \\ \\ \text{CN} \end{array}$
$\text{K}_2\text{Cr}_2\text{O}_7/\text{H}^+$	NO REACTION

one mark for each correct answer (3 × 1) [3]

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(c) Na_2CO_3 or NaHCO_3 effervescence/colourless gas

or

Na colourless gas

or

$\text{PCl}_3/\text{PCl}_5$ etc. steamy fumes

or

$\text{C}_2\text{H}_5\text{OH}/\text{conc. H}_2\text{SO}_4$ sweet smell of ester

or

$\text{K}_2\text{Cr}_2\text{O}_7/\text{H}^+$ orange solution becomes green

correct reagent

(1)

correct observation

(1) [2]

[Total: 10]

Page 7	Mark Scheme	Syllabus
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- 5 (a) (i) $\text{CH}_2=\text{CHCO}_2\text{H}$
- (ii) $\text{BrCH}_2\text{CHBrCH}_2\text{OH}$ (1)
- (iii) product is $\text{HOCH}_2\text{CH}(\text{OH})\text{CH}_2\text{OH}$
- correct addition across $>\text{C}=\text{C}<$ (1)
- original $-\text{CH}_2\text{OH}$ remains (1)
- (iv) $\text{HO}_2\text{CCO}_2\text{H}$ (1) [5]
- (b) (i) nucleophilic substitution (1)
- (ii) oxidation (1) [2]
- (c) (i) **step I**
- H_2 (1)
- heat with Ni catalyst (1)
- step II**
- acidified $\text{K}_2\text{Cr}_2\text{O}_7$ (1)
- heat **or** distil off product (1)
- (ii) structural isomerism
- or**
- functional group isomerism (1) [5]
- (d) **both** oxidation **and** reduction have occurred **or**
- disproportionation has taken place (1) [1]

[Total: 13]