

Centre Number	Candidate Number	Name
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CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/02

Paper 2

October/November 2003

1 hour

Candidates answer on the Question Paper.
No Additional Materials required

READ THESE INSTRUCTIONS FIRST

Write your name, centre number and candidate number in the spaces at the top of this page.
Write in dark blue or black pen in the spaces provided on the Question Paper.
You may use a pencil for any diagrams, graphs, or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.

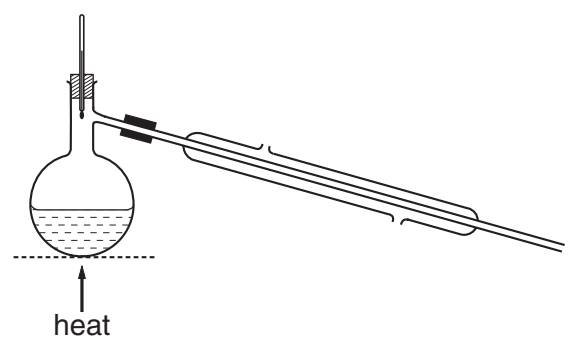
Answer **all** questions.
The number of marks is given in brackets [] at the end of each question or part question.
A copy of the Periodic Table is provided on page 20.

For Examiner's Use	
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TOTAL	

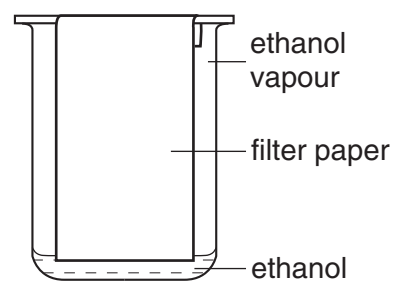
If you have been given a label, look at the details. If any details are incorrect or missing, please fill in your correct details in the space given at the top of this page.

Stick your personal label here, if provided.

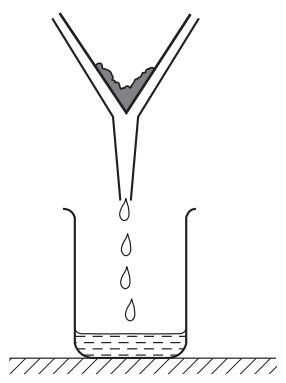
1 The diagrams show four methods of purifying substances.



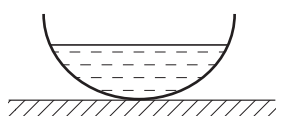
A



B



C



D

(a) Which of these methods, **A**, **B**, **C** or **D**, is best used for

(i) separating the different colours in a sample of ink?

(ii) separating two liquids with different boiling points?

(iii) separating mud from water?

(iv) making crystals of copper sulphate from copper sulphate solution?

[4]

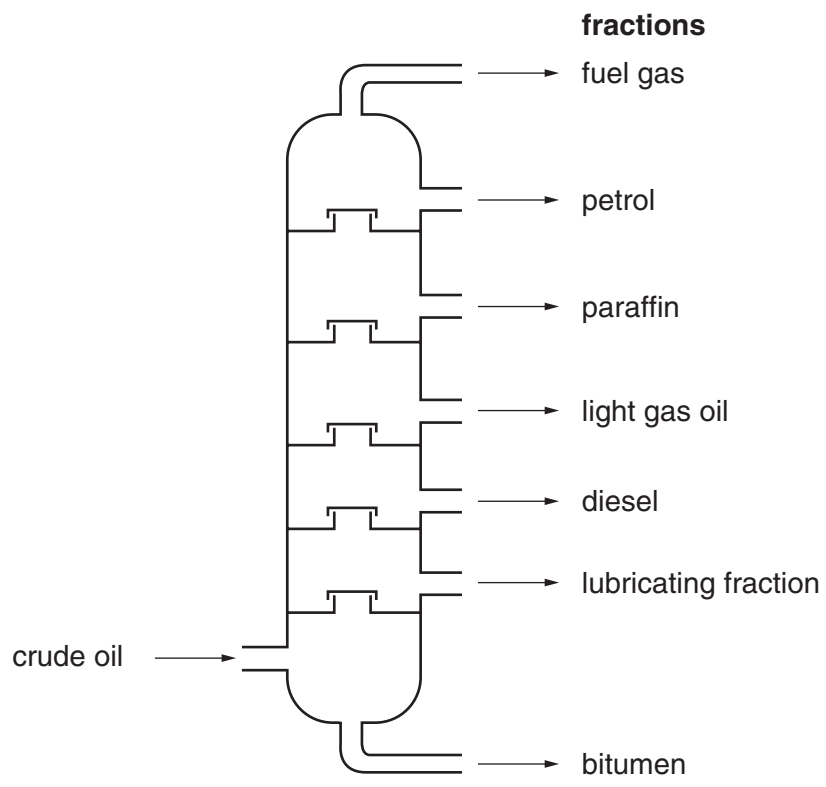
(b) State the name given to the method of separation shown in

(i) diagram A,

(ii) diagram B.

[2]

(c) Method A can be modified to separate petroleum into useful fractions.
The diagram below shows the different fractions obtained from a fractionating column.



(i) Which of these fractions has the lowest boiling point?
.....

(ii) State **one** use for each of the following fractions.
paraffin

bitumen

[3]

(d) Petroleum is a mixture of organic compounds.
Which **one** of the following best describes the compounds found in petroleum?
Put a ring around the correct answer.

- acids**
- alcohols**
- carbohydrates**
- hydrocarbons**

[1]

- (e) Before petroleum is fractionated, it is often heated to remove dissolved natural gas. Most of this natural gas is methane, CH_4 . Draw a diagram to show how the electrons are arranged in methane.

show hydrogen electrons as

•

show carbon electrons as

×

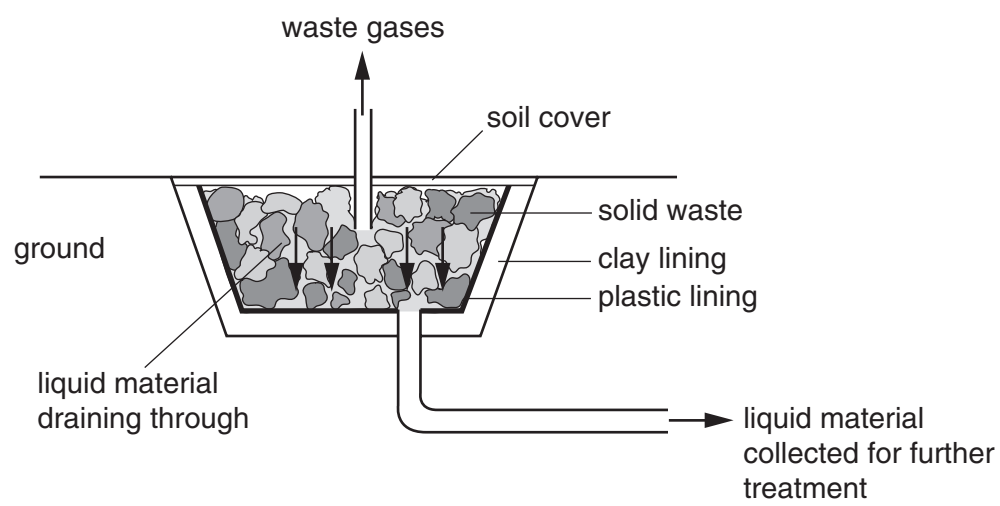
[2]

- (f) Methane, ethane and propane belong to a particular homologous series of compounds. State the name of the homologous series to which these three compounds belong.

.....

[1]

2 The diagram below shows a modern landfill site for the disposal of waste materials.



The waste materials are broken down naturally in several stages.

(a) In the first stage, micro-organisms (mainly bacteria) break down some of the organic material in the waste to carbon dioxide.
 What is the name given to the process by which organisms use food to produce carbon dioxide?

.....[1]

(b) In the second stage, the micro-organisms break down organic substances to produce ammonia, hydrogen and more carbon dioxide.

(i) Describe a test for hydrogen.

test

result

(ii) The large volumes of hydrogen produced may be hazardous.
 Explain why hydrogen may be hazardous when mixed with air.

.....

(iii) Ammonia is a base.
 Describe a test for ammonia.

test

result

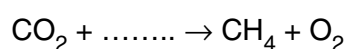
[5]

- (c) In the third stage, ethanoic acid is produced.
Draw the structure of ethanoic acid showing all atoms and bonds.

[1]

- (d) In the fourth stage, carbon dioxide reacts with hydrogen to form methane and oxygen.

- (i) Complete the equation for this reaction.



- (ii) State one use of methane.

.....

- (iii) Methane is a gas.
Which **two** of the following statements about gas molecules are true?
Tick **two** boxes.

The molecules are far apart.

The molecules are not moving.

The molecules are randomly arranged.

The molecules are arranged in a regular manner.

[4]

(e) The list below shows some of the substances which are found in the liquid which
through the waste.

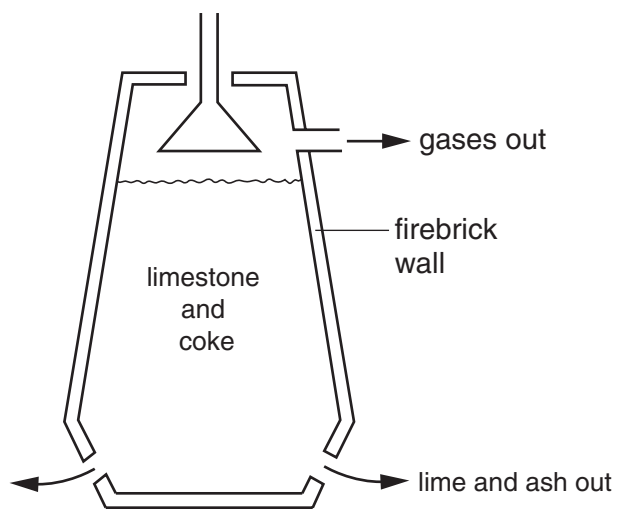
- aluminium
- calcium carbonate
- iron
- lead
- magnesium
- nickel
- sodium sulphate
- zinc

From this list choose

- (i) a metal used to galvanise iron.
- (ii) a transition metal.
- (iii) a metal which is in Group IV of the periodic table.
- (iv) a substance which will release carbon dioxide when an acid is added.
.....
- (v) a metal which is used to make aircraft bodies.

[5]

3 One way of making lime from limestone (calcium carbonate) is shown in the diagram.



The limestone is mixed with coke and dropped into the limekiln. The coke is burnt and releases heat.

(a) State **one** use of limestone, other than in making lime.

.....[1]

(b) Coke is mainly carbon.
Write a symbol equation for the burning of carbon.

[2]

(c) State the name of the type of reaction which releases heat energy.

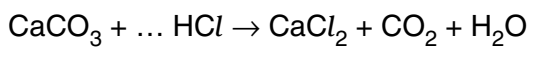
.....[1]

(d) The heat produced by the burning coke causes thermal decomposition of the limestone. Complete the word equation for the thermal decomposition of calcium carbonate.

calcium carbonate → +
.....

[2]

- (e) (i) Complete the following equation for the reaction of calcium carbonate and hydrochloric acid.



- (ii) Describe how you would test for the gas given off in this reaction.

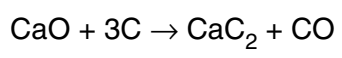
test

result

[3]

- (f) Quicklime, CaO, is a product of the thermal decomposition of calcium carbonate.

When quicklime is heated strongly with coke, calcium carbide is formed.



- (i) What type of reaction is the conversion of C to CO?
Explain your answer.

.....
.....

- (ii) When water is added to calcium carbide, CaC₂, acetylene is formed.
State a use of acetylene.

.....
[3]

4 Bromine is an element in Group VII of the Periodic Table.

(a) State the name given to the Group VII elements.

.....[1]

(b) Bromine has two isotopes.
The nucleon (mass) number of bromine-79 is 79 and of bromine-81 is 81.

(i) What is the meaning of the term *isotopes*?

.....
.....

(ii) Complete the table to show the numbers of electrons, neutrons and protons in one atom of bromine-79 and bromine-81. A copy of the Periodic Table is printed on page 20.

number of	bromine-79	bromine-81
electrons		
neutrons		
protons		

[5]

(c) Bromine is extracted from seawater by treatment with chlorine.
When chlorine is bubbled through a solution of potassium bromide, the solution turns orange-red.

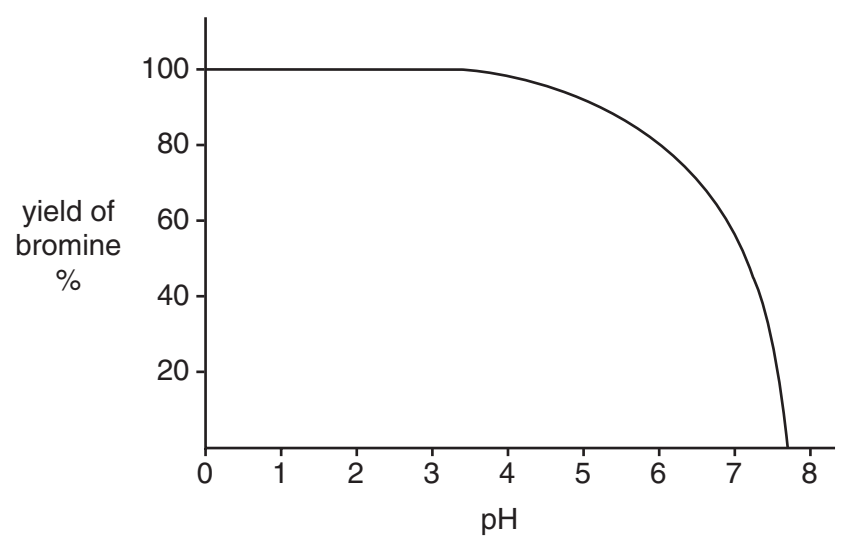
(i) What does this tell you about the reactivity of chlorine compared with bromine?

.....

(ii) Write a word equation for this reaction.

[2]

(d) In order to get the maximum yield of bromine from seawater, acid is added during the extraction procedure.
The graph shows how the yield of bromine changes with pH.



(i) What is the highest pH at which the yield of bromine is 100%?

.....

(ii) The pH scale is used to measure acidity.
Some pH values are given below.

- pH 3
- pH 5
- pH 7
- pH 9
- pH 11

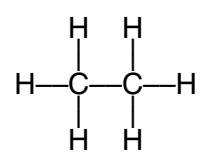
From this list of pH values choose

the pH which is most acidic.

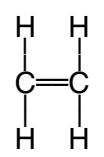
the pH of a neutral solution.

[3]

(e) Bromine water can be used to distinguish between ethane and ethene.



ethane



ethene

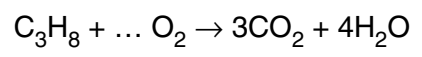
Describe what you would observe when bromine water is added to ethene.

.....

[1]

5 When fuels are burnt, carbon dioxide and water are formed.

(a) Complete the equation for the burning of propane.



[1]

(b) Describe a chemical test for water.

test

result

[2]

(c) In which **two** of the following is carbon dioxide produced.
Tick **two** boxes.

a car driven by a petrol engine

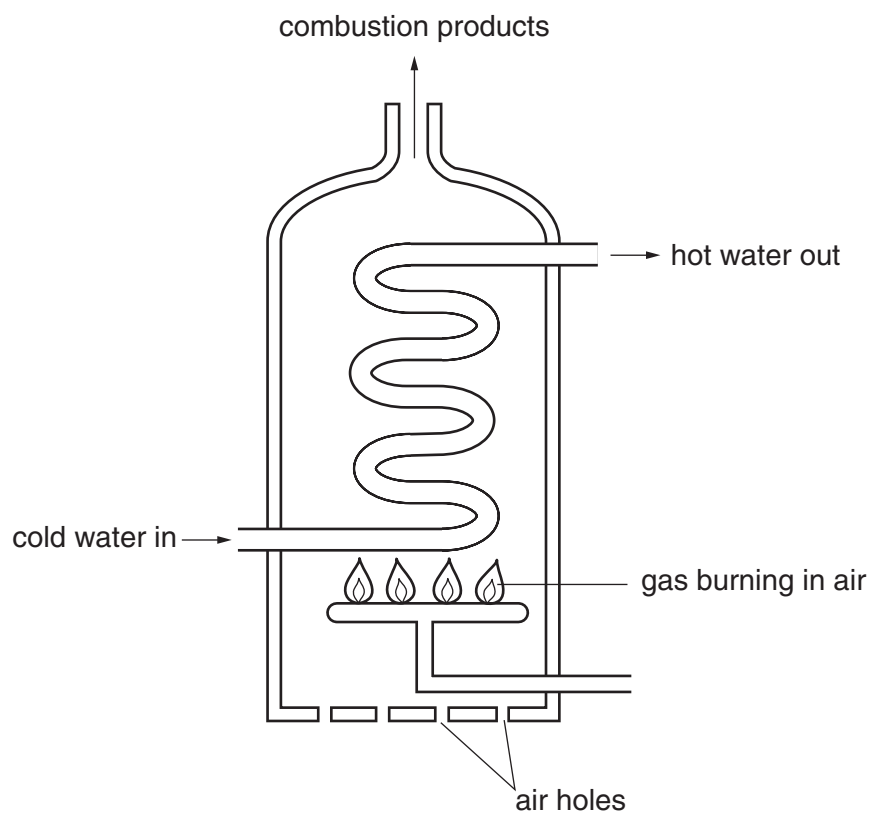
magnesium carbonate reacting with an acid

sodium reacting with water

zinc reacting with hydrochloric acid

[2]

(d) The diagram shows a water heater.



If some of the air holes become blocked, a poisonous gas is produced.

(i) State the name of this poisonous gas.

.....

(ii) Explain how this poisonous gas has been formed.

.....

.....

[2]

(e) The table below compares the amounts of carbon dioxide and sulphur dioxide when 1 kilogram of different fuels are burnt.

fuel	mass of carbon dioxide produced / g	mass of sulphur dioxide produced / g
oil	2900	5.0
gas	2500	0.1
coal	2500	11.0

(i) Which fuel is **least** polluting?

.....

(ii) Which fuel when burnt, contributes most to the formation of acid rain?

.....

(iii) State two harmful effects of acid rain.

.....

.....

(iv) When acid rain falls on the ground, it can react with insoluble aluminium compounds in the soil. A solution of aluminium ions is formed.

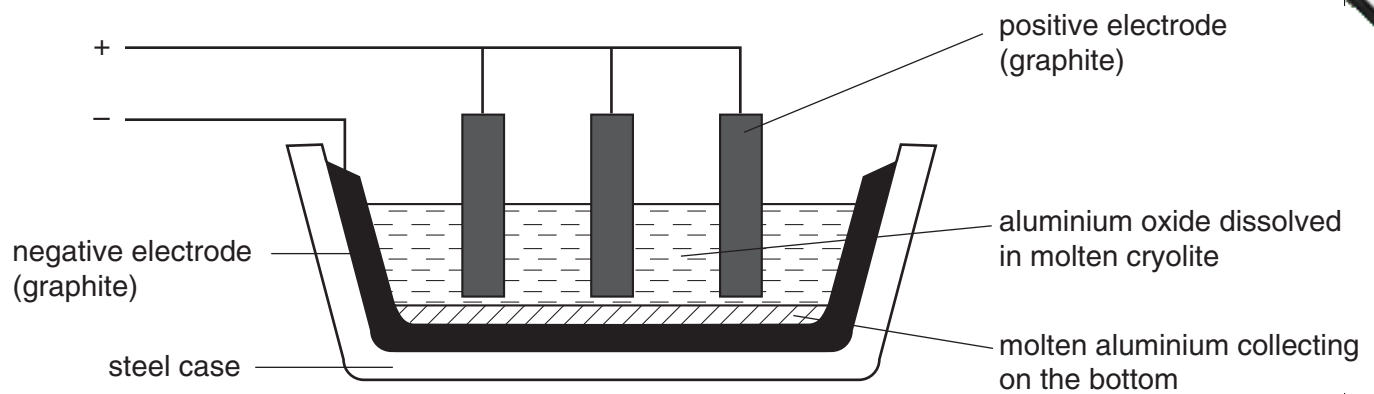
Describe what you would observe when aqueous sodium hydroxide is added to a solution containing aluminium ions.

.....

.....

[6]

6 Aluminium is extracted from its ore, bauxite.
 The bauxite is purified to give aluminium oxide.
 Electrolysis is then used to extract the aluminium from aluminium oxide dissolved in cryolite.



The melting point of pure aluminium oxide is 2070 °C.
 The melting point of the mixture of aluminium oxide and cryolite is about 1000 °C.

(a) Suggest why electrolysis is used to extract aluminium from aluminium oxide rather than reduction using carbon.

.....
[1]

(b) How is the electrolyte of aluminium oxide and cryolite kept molten?

.....[1]

(c) What property of graphite makes it suitable for use as electrodes?

.....[1]

(d) State the name given to the negative electrode.

.....[1]

(e) The melting point of steel is about 1500 °C.
 Suggest **two** reasons why molten aluminium oxide is not used by itself in this electrolysis.

.....
[2]

(f) During the electrolysis, hot oxygen is formed at the positive electrodes.
 Suggest why the positive electrodes have to be replaced frequently.

.....
[2]

(g) Aluminium is formed at the negative electrode.
Complete the following equation for the reaction at the negative electrode.



(h) Why do aluminium ions move towards the negative electrode?
.....[1]

(i) A sample of bauxite ore had the following composition:

aluminium oxide	120g
iron(III) oxide	30g
silica	40g
titanium(IV) oxide	10g

Calculate the percentage of aluminium oxide in this sample of bauxite.

[1]

(j) Aluminium is a metal in Group III of the Periodic Table.
State three physical properties which are typical of most metals.

- 1
 - 2
 - 3
- [3]

18
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DATA SHEET
The Periodic Table of the Elements

		Group																																																																																																												
I	II	III	IV	V	VI	VII	0																																																																																																							
7 Li Lithium 4	9 Be Beryllium	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	13 Al Aluminium 13	14 N Nitrogen 7	15 O Oxygen 8	16 F Fluorine 9	17 Ne Neon 10	18 Ar Argon 18	19 K Potassium 19	20 Ca Calcium 20	21 Sc Scandium 21	22 Ti Titanium 22	23 V Vanadium 23	24 Cr Chromium 24	25 Mn Manganese 25	26 Fe Iron 26	27 Co Cobalt 27	28 Ni Nickel 28	29 Cu Copper 29	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36	37 Rb Rubidium 37	38 Sr Strontium 38	39 Y Yttrium 39	40 Zr Zirconium 40	41 Nb Niobium 41	42 Mo Molybdenum 42	43 Tc Technetium 43	44 Ru Ruthenium 44	45 Rh Rhodium 45	46 Pd Palladium 46	47 Ag Silver 47	48 Cd Cadmium 48	49 In Indium 49	50 Sn Tin 50	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54	55 Cs Caesium 55	56 Ba Barium 56	57 La Lanthanum 57	58 Ce Cerium 58	59 Pr Praseodymium 59	60 Nd Neodymium 60	61 Pm Promethium 61	62 Sm Samarium 62	63 Eu Europium 63	64 Gd Gadolinium 64	65 Tb Terbium 65	66 Dy Dysprosium 66	67 Ho Holmium 67	68 Er Erbium 68	69 Tm Thulium 69	70 Yb Ytterbium 70	71 Lu Lutetium 71	72 Hf Hafnium 72	73 Ta Tantalum 73	74 W Tungsten 74	75 Re Rhenium 75	76 Os Osmium 76	77 Ir Iridium 77	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	81 Tl Thallium 81	82 Pb Lead 82	83 Bi Bismuth 83	84 Po Polonium 84	85 At Astatine 85	86 Rn Radon 86	87 Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89	90 Th Thorium 90	91 Pa Protactinium 91	92 U Uranium 92	93 Np Neptunium 93	94 Pu Plutonium 94	95 Am Americium 95	96 Cm Curium 96	97 Bk Berkelium 97	98 Cf Californium 98	99 Es Einsteinium 99	100 Fm Fermium 100	101 Md Mendelevium 101	102 No Nobelium 102	103 Lr Lawrencium 103	104 Rf Rutherfordium 104	105 Db Dubnium 105	106 Sg Seaborgium 106	107 Bh Bohrium 107	108 Hs Hassium 108	109 Mt Meitnerium 109	110 Ds Darmstadtium 110	111 Rg Roentgenium 111	112 Cn Copernicium 112	113 Nh Nihonium 113	114 Fl Flerovium 114	115 Mc Moscovium 115	116 Lv Livermorium 116	117 Ts Tennessine 117	118 Og Oganesson 118

87-103 Lanthanoid series
103-108 Actinoid series

a = relative atomic mass
 X = atomic symbol
 b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).