



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/12

Paper 1 Multiple Choice

May/June 2011

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

You may use a calculator.

This document consists of **16** printed pages.



- 4 An aqueous solution is coloured.

Which method of separation would show that the solution contains ions of different colours?

- A chromatography
- B crystallisation
- C distillation
- D filtration

- 5 The table gives the solubility of four substances in ethanol and in water.

A mixture containing all four substances is added to ethanol, stirred and filtered.

The solid residue is added to water, stirred and filtered.

The filtrate is evaporated to dryness, leaving a white solid.

Which is the white solid?

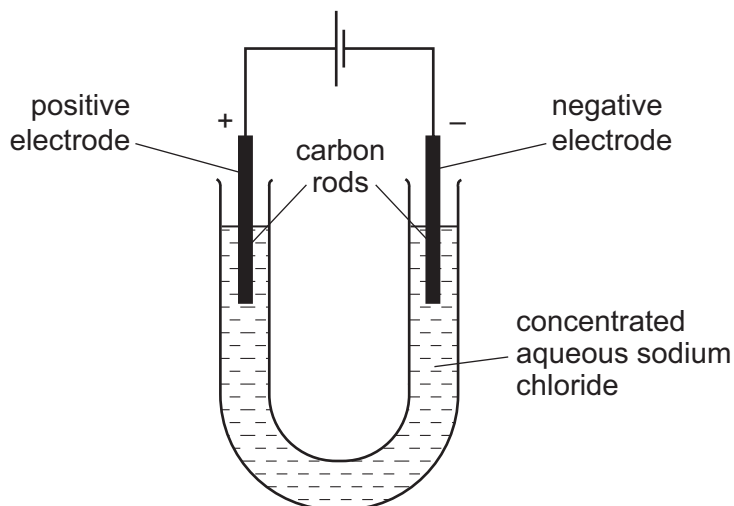
	solubility in	
	ethanol	water
A	insoluble	insoluble
B	insoluble	soluble
C	soluble	insoluble
D	soluble	soluble

- 6 Which two elements react together to form an ionic compound?

element	electronic structure
W	2,4
X	2,8
Y	2,8,1
Z	2,8,7

- A** W and X
- B** X and Y
- C** Y and Z
- D** Z and W

- 7 Electricity is passed through concentrated aqueous sodium chloride, as shown.



What is the test for the gas formed at the positive electrode?

- A** bleaches damp litmus paper
B 'pops' with a lighted splint
C relights a glowing splint
D turns damp red litmus paper blue
- 8 Electricity from a power station passes through overhead cables to a substation and then to a school where it is used to electrolyse concentrated hydrochloric acid using inert electrodes.

Which substances are used for the overhead cables and for the electrodes?

	overhead cables	electrodes
A	aluminium	copper
B	aluminium	platinum
C	copper	platinum
D	platinum	aluminium

- 9 The nucleon number and proton number of the lithium atom are shown by the symbol ${}^7_3\text{Li}$.

What is the correct symbol for the lithium ion in lithium chloride?

- A** ${}^6_2\text{Li}^-$ **B** ${}^6_3\text{Li}^+$ **C** ${}^7_3\text{Li}^+$ **D** ${}^7_3\text{Li}^-$

10 Three processes are listed.

burning methane in air

radioactive decay of ^{235}U

reacting hydrogen with oxygen.

Which statements about these processes are correct?

- 1 Hydrogen and methane are being used as fuels.
- 2 All the processes involve oxidation.
- 3 All the processes are used to produce energy.

A 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3

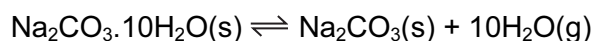
11 Which statement about the electrolysis of molten lead(II) bromide is correct?

- A** A colourless gas is seen at the cathode.
- B** A grey metal is seen at the anode.
- C** A red/brown gas is seen at the anode.
- D** A red/brown metal is seen at the cathode.

12 What is the relative molecular mass (M_r) of HNO_3 ?

A 5 **B** 31 **C** 32 **D** 63

13 The equation for the effect of heat on hydrated sodium carbonate is as shown.



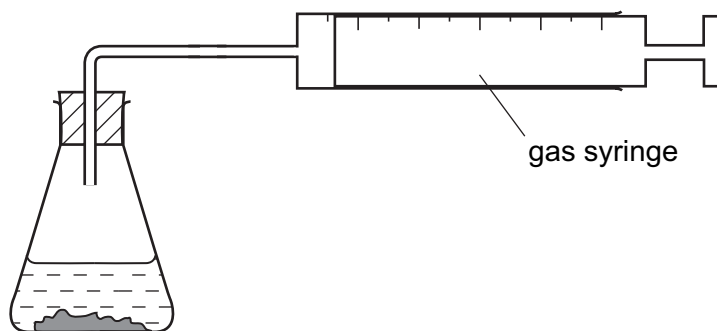
Statements made by four students about the reaction are given.

- P** Anhydrous sodium carbonate is formed.
- Q** Steam is formed.
- R** There is a colour change from blue to white.
- S** The reaction is reversible.

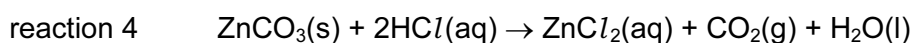
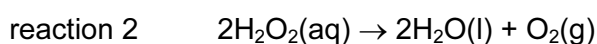
Which students' statements are correct?

- A** P, Q and R only
- B** P, Q and S only
- C** Q, R and S only
- D** P, Q, R and S

14 The apparatus shown can be used to measure the rate of some chemical reactions.



For which two reactions would the apparatus be suitable?



A 1 and 2

B 1 and 3

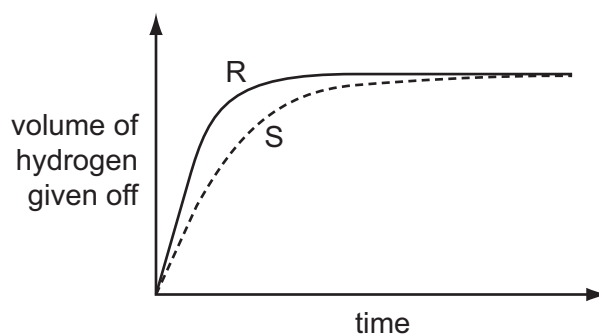
C 2 and 4

D 3 and 4

15 A student investigates the rate of reaction between magnesium and excess sulfuric acid.

The volume of hydrogen given off in the reaction is measured over time.

The graph shows the results of two experiments, R and S.



Which change in conditions would cause the difference between R and S?

A A catalyst is added in S.

B The acid is more concentrated in R than in S.

C The magnesium is less finely powdered in R than in S.

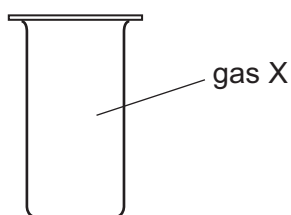
D The temperature in R is lower than in S.

16 Butane, ethanol and hydrogen are fuels.

Which substances produce **both** carbon dioxide and water when used as a fuel?

	butane	ethanol	hydrogen
A	✓	✓	✓
B	✓	✓	x
C	✓	x	✓
D	x	✓	x

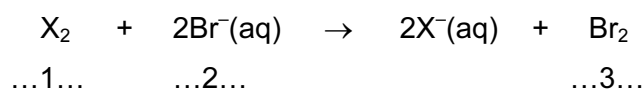
17 X is a monatomic gas.



Which statement about X is correct?

- A** X burns in air.
- B** X is coloured.
- C** X is unreactive.
- D** X will displace iodine from potassium iodide.

18 The equation shows the reaction between a halogen and aqueous bromide ions.



Which words correctly complete gaps 1, 2 and 3?

	1	2	3
A	chlorine	brown	colourless
B	chlorine	colourless	brown
C	iodine	brown	colourless
D	iodine	colourless	brown

19 Carbon dioxide is an acidic oxide that reacts with aqueous calcium hydroxide.

Which type of reaction takes place?

- A decomposition
- B fermentation
- C neutralisation
- D oxidation

20 A solution contains barium ions and silver ions.

What could the anion be?

- A chloride only
- B nitrate only
- C sulfate only
- D chloride or nitrate or sulfate

21 A mixture containing two anions was tested and the results are shown below.

test	result
dilute nitric acid added	effervescence of a gas which turned limewater milky
dilute nitric acid added, followed by aqueous silver nitrate	yellow precipitate formed

Which anions were present?

- A carbonate and chloride
- B carbonate and iodide
- C sulfate and chloride
- D sulfate and iodide

22 Which is **not** a typical property of an acid?

- A They react with alkalis producing water.
- B They react with all metals producing hydrogen.
- C They react with carbonates producing carbon dioxide.
- D They turn litmus paper red.

23 The diagram shows a section of the Periodic Table.

	I	II	III	IV	V	VI	VII	0
	V			W			X	
		Y				Z		

Which elements will conduct electricity at room temperature?

- A** V, W and X **B** V, Y and W **C** W, X and Z **D** Y and Z

24 Water from a reservoir flows to the water works where purification processes 1 takes place followed by process 2.

What are purification processes 1 and 2?

	purification process 1	purification process 2
A	chlorination	filtration
B	filtration	chlorination
C	fractional distillation	filtration
D	filtration	fractional distillation

25 The properties of a metal are important in deciding its use.

Which row lists a property that is **not** correct for the use given?

	use of the metal	metal property needed
A	aluminium in aircraft wings	low density
B	aluminium in food containers	resists corrosion
C	mild steel in car bodies	high density
D	stainless steel in cutlery	does not rust

26 Brass is an alloy of copper and zinc.

Which statement is correct?

- A Brass can be represented by a chemical formula.
- B Brass is formed by a chemical reaction between copper and zinc.
- C The alloy will dissolve completely in dilute hydrochloric acid.
- D The zinc in the alloy will dissolve in dilute hydrochloric acid.

27 Which statement is correct for the element of proton number 19?

- A It is a gas that dissolves in water.
- B It is a hard metal that is not very reactive with water.
- C It is a non-metal that burns quickly in air.
- D It is a soft metal that is highly reactive with water.

28 Which row describes the conditions used to make steel from the iron produced by a blast furnace?

	calcium oxide (lime)	oxygen	heat
A	✓	✓	✓
B	✓	✓	x
C	x	✓	✓
D	x	✓	x

- 29 The table shows the results of adding three metals, P, Q and R, to dilute hydrochloric acid and to water.

metal	dilute hydrochloric acid	water
P	hydrogen produced	hydrogen produced
Q	no reaction	no reaction
R	hydrogen produced	no reaction

What is the order of reactivity of the metals?

	most reactive	→	least reactive
A	P		Q
B	P		R
C	R		P
D	R		Q

- 30 Which substance is a metal?

	electrical conductivity (solid)	electrical conductivity (molten)
A	high	high
B	high	low
C	low	high
D	low	low

- 31 Greenhouse gases may contribute to climate change.

Two of these gases are emitted into the atmosphere as a result of processes within animals.

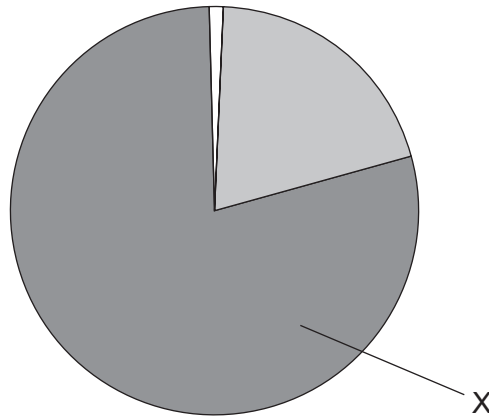
Gas1..... is produced by process3..... .

Gas2..... is produced by process4..... .

Which words correctly complete gaps 1, 2, 3 and 4?

	1	2	3	4
A	CO	C ₂ H ₆	digestion	respiration
B	CO	C ₂ H ₆	respiration	digestion
C	CO ₂	CH ₄	digestion	respiration
D	CO ₂	CH ₄	respiration	digestion

32 The diagram shows the composition by volume of air.



What is X?

- A argon
- B carbon dioxide
- C nitrogen
- D oxygen

33 The table gives the composition of the atmosphere of four newly discovered planets.

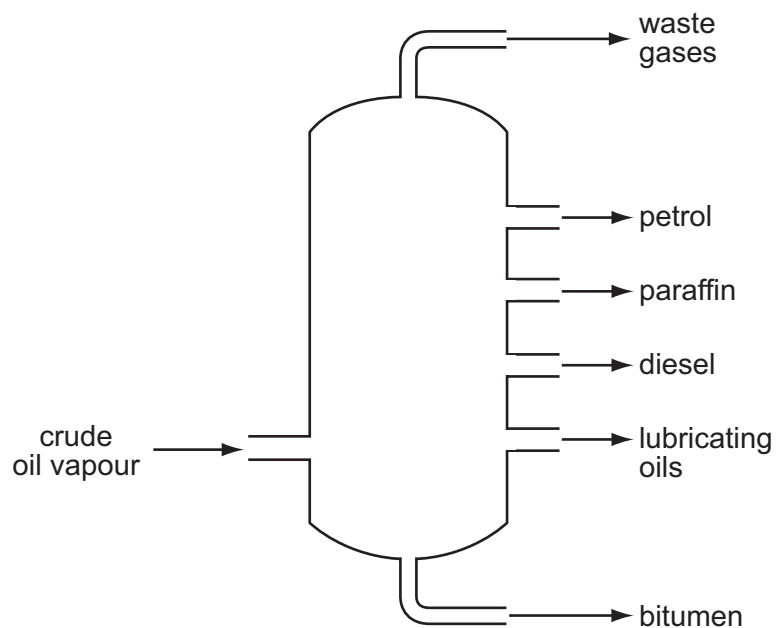
planet	composition of atmosphere
W	argon, carbon dioxide and oxygen
X	argon, nitrogen and oxygen
Y	argon, carbon dioxide and methane
Z	methane, nitrogen and oxygen

On which planets is the greenhouse effect likely to occur?

- A W only
- B W, X and Z
- C W and Y only
- D W, Y and Z

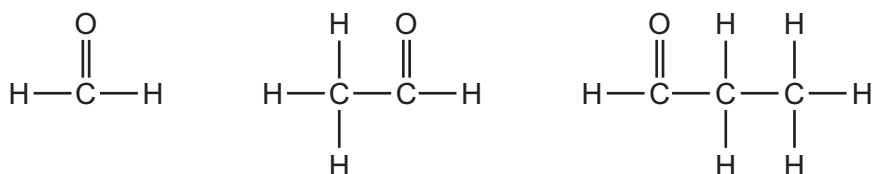
- 34 Which two substances, when reacted together, would form a salt that contains two of the essential elements provided by fertilisers?
- A potassium hydroxide and nitric acid
 - B potassium hydroxide and sulfuric acid
 - C sodium hydroxide and nitric acid
 - D sodium hydroxide and sulfuric acid
- 35 Statement 1: Alloying iron with other materials to form stainless steel prevents iron from rusting by excluding oxygen.
- Statement 2: Painting, oiling and electroplating are all methods of preventing iron from rusting.
- Which is correct?
- A Both statements are correct and statement 2 explains statement 1.
 - B Both statements are correct but statement 2 does not explain statement 1.
 - C Statement 1 is correct but statement 2 is incorrect.
 - D Statement 2 is correct but statement 1 is incorrect.
- 36 What is the main constituent of natural gas?
- A carbon dioxide
 - B ethane
 - C hydrogen
 - D methane
- 37 What is **not** essential for the formation of ethanol by fermentation?
- A light
 - B sugar
 - C yeast
 - D water

38 Which industrial process is shown in the diagram?



- A cracking
- B fermentation
- C fractional distillation
- D polymerisation

39 The diagram shows the structures of three compounds.



Why do these three compounds belong to the same homologous series?

- A They all contain carbon, hydrogen and oxygen.
- B They all contain the same functional group.
- C They are all carbon based molecules.
- D They are all flammable liquids.

40 Compounds containing five carbon atoms in a molecule may have names beginning with 'pent...'.
What is the name of the compound shown?



key

○ = carbon

● = oxygen

● = hydrogen

- A pentane
- B pentanoic acid
- C pentanol
- D pentene

DATA SHEET
The Periodic Table of the Elements

		Group															
		I	II	III	IV	V	VI	VII	VIII	IX	X						
		1 H Hydrogen 1															
		4 He Helium 2															
7	9	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
Li Lithium	Be Beryllium	B Boron	C Carbon	N Nitrogen	O Oxygen	F Fluorine	Ne Neon	Na Sodium	Mg Magnesium	Al Aluminium	Si Silicon	P Phosphorus	S Sulfur	Cl Chlorine	Ar Argon	K Potassium	Ca Calcium
11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28
Sc Scandium	Ti Titanium	V Vanadium	Cr Chromium	Mn Manganese	Fe Iron	Co Cobalt	Ni Nickel	Cu Copper	Zn Zinc	Ga Gallium	Ge Germanium	As Arsenic	Se Selenium	Br Bromine	Kr Krypton	Rb Rubidium	Sr Strontium
31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48
Y Yttrium	Zr Zirconium	Nb Niobium	Mo Molybdenum	Tc Technetium	Ru Ruthenium	Rh Rhodium	Pd Palladium	Ag Silver	Cd Cadmium	In Indium	Sn Tin	Sb Antimony	Te Tellurium	I Iodine	Xe Xenon	Cs Caesium	Ba Barium
49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66
La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium	Lu Lutetium	Rn Radon	Fr Francium	Ra Radium
71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88
Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium	Lr Lawrencium	Po Polonium	At Astatine	Rn Radon
89	90	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106
Fr Francium	Ra Radium	Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium	Lr Lawrencium	Lu Lutetium
107	108	109	110	111	112	113	114	115	116	117	118	119	120	121	122	123	124
Cs Caesium	Ba Barium	La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium	Lu Lutetium	Hf Hafnium
127	128	129	130	131	132	133	134	135	136	137	138	139	140	141	142	143	144
Fr Francium	Ra Radium	Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium	Lr Lawrencium	Lu Lutetium
151	152	153	154	155	156	157	158	159	160	161	162	163	164	165	166	167	168
Cs Caesium	Ba Barium	La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium	Lu Lutetium	Hf Hafnium
171	172	173	174	175	176	177	178	179	180	181	182	183	184	185	186	187	188
Fr Francium	Ra Radium	Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium	Lr Lawrencium	Lu Lutetium
191	192	193	194	195	196	197	198	199	200	201	202	203	204	205	206	207	208
Cs Caesium	Ba Barium	La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium	Lu Lutetium	Hf Hafnium
211	212	213	214	215	216	217	218	219	220	221	222	223	224	225	226	227	228
Fr Francium	Ra Radium	Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium	Lr Lawrencium	Lu Lutetium

*58-71 Lanthanoid series
†90-103 Actinoid series

Key

a	X
b	

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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