



## Cambridge IGCSE™

CANDIDATE  
NAME
CENTRE  
NUMBER

--	--	--	--	--

CANDIDATE  
NUMBER

--	--	--	--

**CHEMISTRY****0620/33**

Paper 3 Theory (Core)

**October/November 2022****1 hour 15 minutes**

You must answer on the question paper.

No additional materials are needed.

**INSTRUCTIONS**

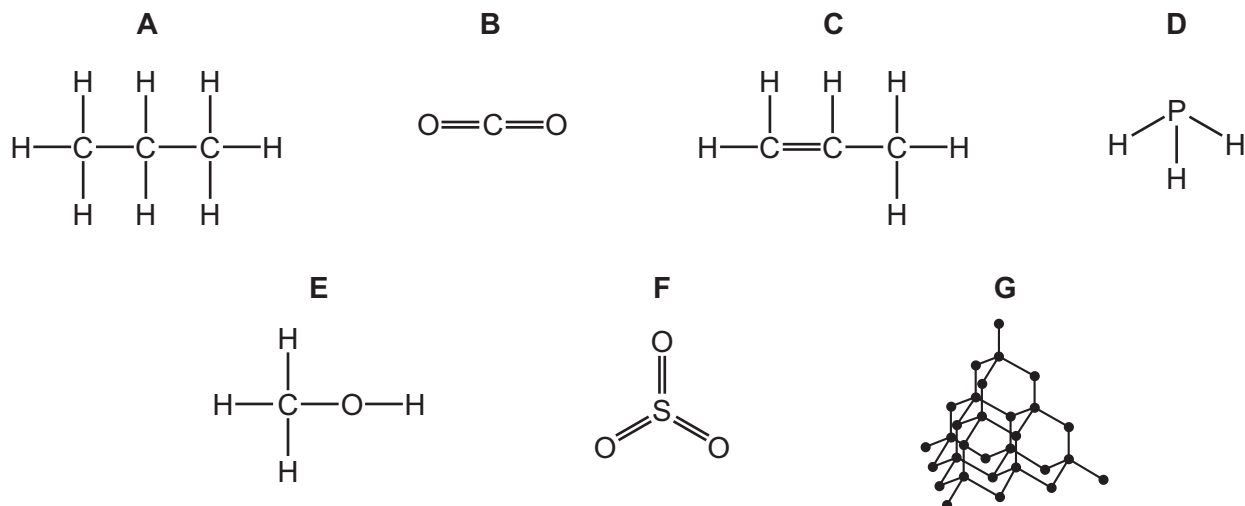
- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

**INFORMATION**

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.

1 (a) The structures of seven compounds or elements, **A**, **B**, **C**, **D**, **E**, **F** and **G**, are shown.



Answer the following questions about these structures.  
Each structure may be used once, more than once or not at all.

State which structure, **A**, **B**, **C**, **D**, **E**, **F** or **G**, represents:

(i) a compound that contains an atom of an element in Group V of the Periodic Table

..... [1]

(ii) an element

..... [1]

(iii) a substance that is used for cutting tools

..... [1]

(iv) a compound that is a major contributor to climate change

..... [1]

(v) a saturated hydrocarbon.

..... [1]

(b) Describe how aqueous bromine can be used to distinguish between a saturated hydrocarbon and an unsaturated hydrocarbon.

observations with saturated hydrocarbon .....

.....

observations with unsaturated hydrocarbon .....

.....

[2]

(c) When a hydrocarbon undergoes incomplete combustion, carbon monoxide is formed.

(i) State the meaning of the term *incomplete combustion*.

..... [1]

(ii) Give **one** adverse effect of carbon monoxide on health.

..... [1]

[Total: 9]

- 2 (a) The table compares the percentage by mass of the elements in the whole Earth and in the oceans.

element	percentage by mass in the whole Earth	percentage by mass in the oceans
calcium	1.1	0.1
chlorine	less than 0.01	1.0
hydrogen	less than 0.01	11.0
iron	34.6	0.0
magnesium	12.7	1.0
oxygen	29.5	85.0
silicon	15.2	0.0
sodium	0.6	1.0
other elements	6.3	
total	100.00	100.00

Answer these questions using only the information in the table.

- (i) Deduce the percentage by mass of the other elements in the oceans.

..... [1]

- (ii) State which non-metallic element is present in the whole Earth in the greatest percentage by mass.

..... [1]

- (iii) Give **two** major differences in the composition of the whole Earth and of the oceans.

1 .....

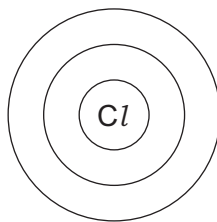
.....

2 .....

.....

[2]

(b) Complete the diagram to show the electron arrangement in a chlorine atom.



[2]

(c) Iron is extracted from iron(III) oxide in a blast furnace.

(i) Explain why air is blown into the blast furnace.

.....  
 ..... [1]

(ii) In the blast furnace, carbon dioxide reacts with carbon to produce carbon monoxide.

Complete the chemical equation for this reaction.



(iii) Carbon monoxide reduces iron(III) oxide to iron.

State the meaning of the term *reduction*.

..... [1]

(iv) When 80 g of iron(III) oxide reacts with excess carbon monoxide, 56 g of iron is produced.

Calculate the minimum mass of iron(III) oxide needed to produce 14 g of iron.

..... g [1]

(d) Steel is an alloy of iron.

(i) Choose **two** substances from the list that are used in the conversion of iron from the blast furnace into steel.

- calcium oxide
- carbon dioxide
- hydrogen
- nitrogen
- oxygen
- silicon(IV) oxide
- sulfur dioxide

1 .....

2 ..... [2]

(ii) State the meaning of the term *alloy*.

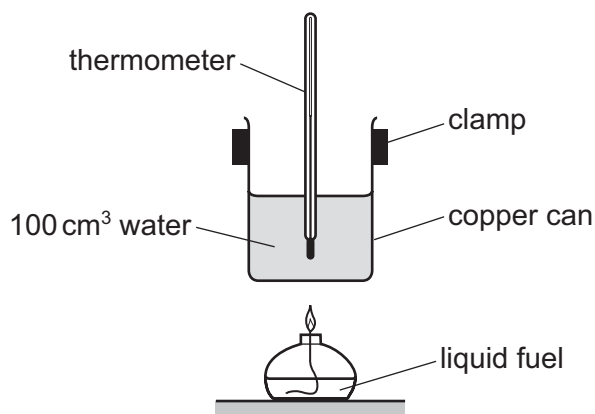
.....

..... [1]

[Total: 13]

3 This question is about fuels and energy production.

- (a) The diagram shows the apparatus used to compare the energy released when  $100\text{ cm}^3$  of water is heated by burning different liquid fuels, **J**, **K**, **L** and **M**.



All conditions are kept the same apart from the type of fuel and mass of fuel burned.

The results are shown.

fuel	mass of fuel burned /g	increase in temperature /°C
<b>J</b>	2	4
<b>K</b>	4	8
<b>L</b>	1	3
<b>M</b>	2	5

Deduce which fuel, **J**, **K**, **L** or **M**, releases the most energy per gram.

..... [1]

- (b) The fractional distillation of petroleum produces fractions, such as gasoline and diesel, which are used as fuels.

(i) Name one **other** petroleum fraction that is used as a fuel.

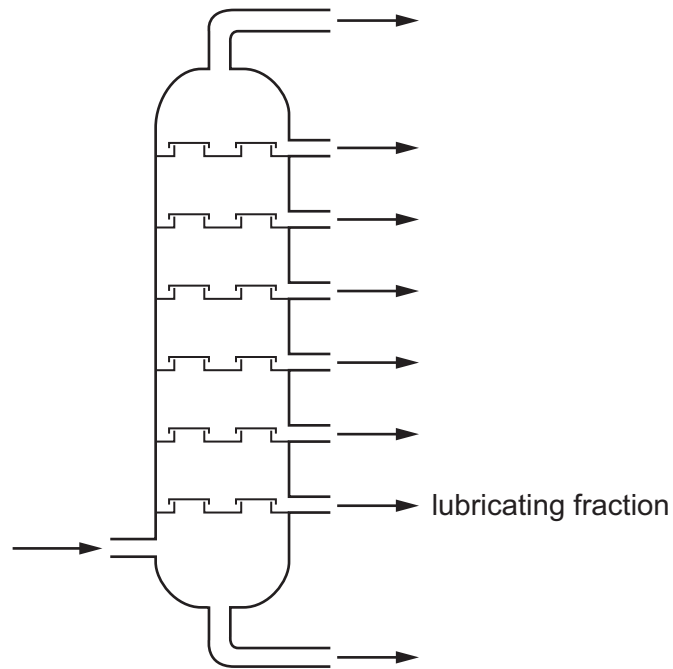
..... [1]

(ii) State the physical property on which the fractional distillation of petroleum depends.

..... [1]

(iii) Bitumen is a fraction of petroleum.

Write an X **on the diagram** to show where bitumen is obtained.



[1]

(c) (i) Name a radioactive element that is used as a source of energy in nuclear power stations.

..... [1]

(ii) State one **other** industrial use of radioactive isotopes.

..... [1]

[Total: 6]



4 (a) The table shows some properties of four Group I elements.

element	melting point /°C	boiling point /°C	hardness /MPa
lithium	181	1342	5.00
sodium	98	.....	0.70
potassium	63	760	0.36
rubidium	39	686	.....

(i) Complete the table by predicting:

- the boiling point of sodium
- the hardness of rubidium.

[2]

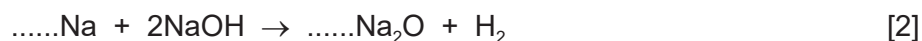
(ii) Predict the physical state of potassium at 50 °C.  
Give a reason for your answer.

.....

..... [2]

(b) When a mixture of sodium and sodium hydroxide is heated, sodium oxide and hydrogen are formed.

(i) Complete the chemical equation for this reaction.



[2]

(ii) Describe a test for hydrogen.

test .....

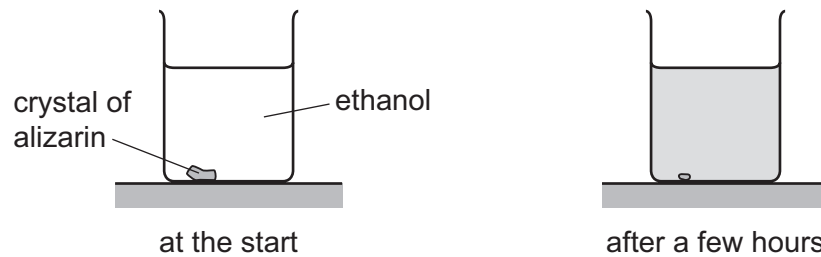
observations .....

[2]

- (c) Sodium hydroxide is used to make the red dye alizarin.  
Alizarin is soluble in ethanol.

A crystal of alizarin is placed in a beaker of ethanol.

After a few hours, the red colour has spread throughout the beaker.



Explain these observations using the kinetic particle model.

.....

.....

.....

..... [3]

[Total: 11]

5 This question is about compounds of nitrogen and fertilisers.

(a) Ammonium chloride is heated with sodium hydroxide.

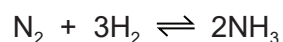


Choose from the list the word that describes this reaction.

Draw a circle around your answer.

**addition**      **displacement**      **oxidation**      **reduction**      [1]

(b) Ammonia is manufactured from nitrogen and hydrogen.



(i) Give the meaning of the symbol  $\rightleftharpoons$ .

..... [1]

(ii) The nitrogen for this process is obtained from the air.

State the percentage of nitrogen in clean, dry air.

..... [1]

(c) Fertilisers contain nitrogen.

Name two **other** elements found in most fertilisers that are essential for plant growth.

..... and ..... [2]

(d) Bacteria in the soil can convert ammonium ions into oxides of nitrogen.

Oxides of nitrogen contribute to acid rain.

(i) Name one **other** pollutant in the air that contributes to acid rain.

..... [1]

(ii) State **one** adverse effect of acid rain on buildings.

..... [1]

[Total: 7]

6 This question is about acids, bases and salts.

(a) Describe the reaction of excess dilute hydrochloric acid with magnesium and with magnesium carbonate. Give the names of the products and any observations.

reaction with magnesium

- products

.....

- observations

.....

.....

reaction with magnesium carbonate

- products

.....

- observations

.....

.....

[4]

(b) (i) Different sized pieces of magnesium react with excess dilute hydrochloric acid. The time taken for each reaction to finish is recorded.

The sizes of the pieces of magnesium are:

- large
- small
- very small.

Equal masses of magnesium are used in each reaction.

All other conditions stay the same.

Complete the table by writing the size of the magnesium pieces in the first column.

size of magnesium pieces	time taken for the reaction to finish /s
	30
	200
	90

[1]

- (ii) Describe the effect on the time taken for small pieces of magnesium to react with hydrochloric acid of a lower concentration.

All other conditions stay the same.

..... [1]

- (c) (i) Sodium hydroxide is an alkali.

State the colour change when excess aqueous sodium hydroxide is added to a solution of methyl orange in acid.

from ..... to ..... [2]

- (ii) Choose the pH value of an alkali.

Draw a circle around your answer.

**pH 3**                      **pH 5**                      **pH 7**                      **pH 14**                      [1]

- (d) The salt magnesium chloride can be prepared by reacting hydrochloric acid with magnesium oxide.

- (i) The method for preparing pure dry crystals of magnesium chloride is given.

Complete the missing steps 4 and 6.

1 Add excess magnesium oxide to dilute hydrochloric acid.

2 Warm the mixture to complete the reaction.

3 Filter off the excess magnesium oxide and collect the filtrate.

4 .....

5 Remove the crystals.

6 .....

[2]

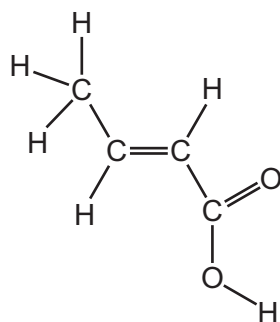
- (ii) Magnesium oxide is used as a catalyst in some reactions.

State the purpose of using a catalyst.

..... [1]

[Total: 12]

- 7 (a) The structure of an organic compound, **S**, is shown.



- (i) On the structure, draw a circle around the carboxylic acid functional group. [1]

- (ii) Deduce the formula of compound **S** to show the number of carbon, hydrogen and oxygen atoms.

..... [1]

- (iii) Compound **S** is a solid at room temperature.

Use the kinetic particle model to describe the arrangement of the particles in a solid.

..... [1]

- (b) Compound **S** reacts with ethanol.

- (i) Draw the structure of ethanol to show all of the atoms and all of the bonds.

[1]

- (ii) Ethanol can be manufactured by fermentation.

Describe one **other** method of manufacturing ethanol.

.....

..... [2]

(c) Compound **S** can be polymerised.

(i) State the general name given to the small units that join together to form a polymer.

..... [1]

(ii) Name **one** natural polymer.

..... [1]

(iii) Some plastics are non-biodegradable.

Describe **one** pollution problem caused by non-biodegradable plastics.

..... [1]

[Total: 9]

8 This question is about metals.

- (a) (i) Copper is a transition element. Sodium is an element in Group I of the Periodic Table. Copper is harder than sodium.

Give two **other** ways in which the physical properties of copper differ from the physical properties of sodium.

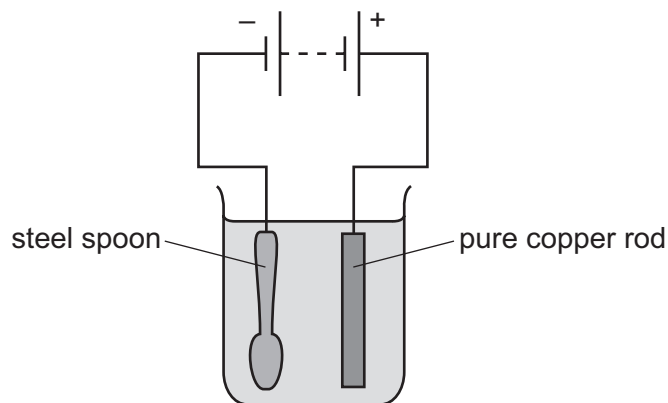
1 .....

2 ..... [2]

- (ii) Give **one** use of copper.

..... [1]

(b) The apparatus used to electroplate a steel spoon with copper is shown.



- (i) Label the diagram to show:

- the anode
- the electrolyte.

[2]

- (ii) Describe the observations made during this electroplating at the:

steel spoon .....

.....

copper rod. ....

..... [2]



- (c) Deduce the number of electrons and neutrons in one atom of the isotope of copper shown.



number of electrons .....

number of neutrons .....

[2]

- (d) A compound of copper has the formula  $\text{K}_2\text{CuF}_4$ .

Complete the table to calculate the relative molecular mass of  $\text{K}_2\text{CuF}_4$ .

atom	number of atoms	relative atomic mass	
potassium	2	39	$2 \times 39 = 78$
copper		64	
fluorine		19	

relative molecular mass = ..... [2]

- (e) The table shows the rates of reaction of four metals with steam.

metal	rate of reaction with steam
zinc	reacts quickly
gold	does not react
iron	reacts slowly
lanthanum	reacts very quickly

Put the four metals in order of their reactivity.

Put the least reactive metal first.

least reactive  $\longrightarrow$  most reactive

--	--	--	--

[2]

[Total: 13]

**BLANK PAGE**

**BLANK PAGE**

---

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge Assessment International Education Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at [www.cambridgeinternational.org](http://www.cambridgeinternational.org) after the live examination series.

Cambridge Assessment International Education is part of Cambridge Assessment. Cambridge Assessment is the brand name of the University of Cambridge Local Examinations Syndicate (UCLES), which is a department of the University of Cambridge.

## The Periodic Table of Elements

		Group							
I	II	III	IV	V	VI	VII	VIII		
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	1 <b>H</b> hydrogen 1	5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20	2
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24	<b>Key</b> atomic number atomic symbol name relative atomic mass							
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	13 <b>Al</b> aluminium 27	14 <b>Si</b> silicon 28	15 <b>P</b> phosphorus 31	16 <b>S</b> sulfur 32	17 <b>Cl</b> chlorine 35.5	18 <b>Ar</b> argon 40	36 <b>Kr</b> krypton 84	37 <b>Rb</b> rubidium 85
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	54 <b>Xe</b> xenon 131	55 <b>Cs</b> caesium 133
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	86 <b>Rn</b> radon —	87 <b>Fr</b> francium —	88 <b>Ra</b> radium —
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	116 <b>Lv</b> livermorium —	117 <b>Ts</b> tennessine —	118 <b>Og</b> oganesson —
57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	29 <b>Cu</b> copper 64	28 <b>Ni</b> nickel 59	27 <b>Co</b> cobalt 59	26 <b>Fe</b> iron 56	25 <b>Mn</b> manganese 55	24 <b>Cr</b> chromium 52	23 <b>V</b> vanadium 51	22 <b>Ti</b> titanium 48
89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	47 <b>Ag</b> silver 108	46 <b>Pd</b> palladium 106	45 <b>Rh</b> rhodium 103	44 <b>Ru</b> ruthenium 101	43 <b>Tc</b> technetium —	42 <b>Mo</b> molybdenum 96	41 <b>Nb</b> niobium 93	40 <b>Zr</b> zirconium 91
89–103 actinoids	104 <b>Rf</b> rutherfordium —	79 <b>Au</b> gold 197	78 <b>Pt</b> platinum 195	77 <b>Ir</b> iridium 192	76 <b>Os</b> osmium 190	75 <b>Re</b> rhenium 186	74 <b>W</b> tungsten 184	73 <b>Ta</b> tantalum 181	72 <b>Hf</b> hafnium 178
105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	111 <b>Rg</b> roentgenium —	110 <b>Ds</b> darmstadtium —	109 <b>Mt</b> meitnerium —	108 <b>Hs</b> hassium —	107 <b>Bh</b> bohrium —	106 <b>Sg</b> seaborgium —	105 <b>Db</b> dubnium —	104 <b>Rf</b> rutherfordium —
109 <b>Me</b> meitnerium —	110 <b>Ds</b> darmstadtium —	112 <b>Cn</b> copernicium —	114 <b>Fl</b> flerovium —	115 <b>Mc</b> moscovium —	116 <b>Lv</b> livermorium —	117 <b>Ts</b> tennessine —	118 <b>Og</b> oganesson —	119 <b>Uue</b> unbinilium —	120 <b>Uub</b> unbinilium —
69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	68 <b>Er</b> erbium 167	67 <b>Ho</b> holmium 165	66 <b>Dy</b> dysprosium 163	65 <b>Tb</b> terbium 159	64 <b>Gd</b> gadolinium 157	63 <b>Eu</b> europium 152	62 <b>Sm</b> samarium 150	61 <b>Pm</b> promethium —
101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	100 <b>Fm</b> fermium —	99 <b>Es</b> einsteinium —	98 <b>Cf</b> californium —	97 <b>Bk</b> berkelium —	96 <b>Cm</b> curium —	95 <b>Am</b> americium —	94 <b>Pu</b> plutonium —	93 <b>Np</b> neptunium —
103 <b>Lr</b> lawrencium —	104 <b>Rf</b> rutherfordium —	103 <b>Nh</b> nihonium —	102 <b>Fl</b> flerovium —	101 <b>Mc</b> moscovium —	100 <b>Ds</b> darmstadtium —	99 <b>Cn</b> copernicium —	98 <b>Bh</b> bohrium —	97 <b>Hs</b> hassium —	96 <b>Mt</b> meitnerium —
105 <b>Uut</b> ununtrium —	106 <b>Uuq</b> ununquadium —	104 <b>Uuq</b> ununquadium —	103 <b>Uub</b> ununbium —	102 <b>Uuo</b> ununoctium —	101 <b>Uua</b> unununium —	100 <b>Uuh</b> ununhexium —	99 <b>Uuq</b> ununquadium —	98 <b>Uuh</b> ununhexium —	97 <b>Uuo</b> ununoctium —

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).