

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

COMBINED SCIENCE

0653/01

Paper 1 Multiple Choice

October/November 2006

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

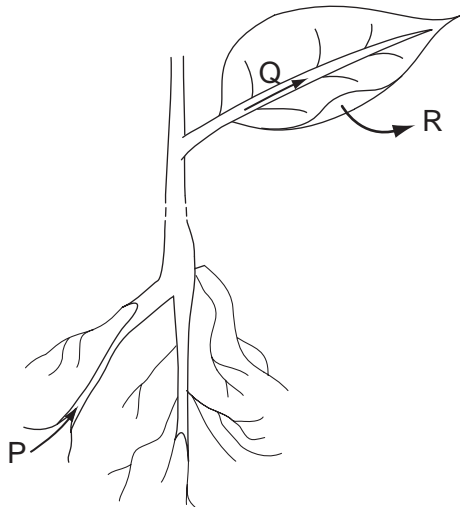
This document consists of **16** printed pages.



- 1 Living and dead plant cells are placed in a dilute solution of red dye. After a few minutes they are observed using a microscope. Only the dead cells are stained red.

Which part of the living cells stops the uptake of the red dye?

- A cell membrane
 B cell wall
 C cytoplasm
 D nucleus
- 2 Which type of chemical is the enzyme catalase?
- A fat
 B protein
 C starch
 D sugar
- 3 The diagram shows the pathway taken by water as it passes through a plant.



In which state is the water at positions P, Q and R?

	P	Q	R
A	liquid	liquid	liquid
B	liquid	liquid	vapour
C	liquid	vapour	vapour
D	vapour	vapour	vapour

- 4 Four foods are each tested separately with Benedict's, biuret and iodine solutions.

Which food contains starch and reducing sugar?

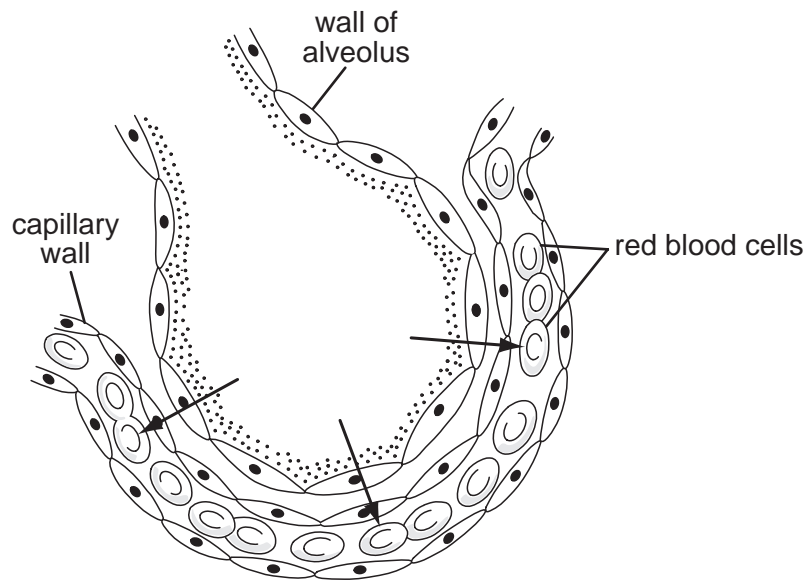
food	Benedict's test	biuret test	iodine test
A	✓	✓	x
B	✓	x	✓
C	x	x	✓
D	x	✓	x

key

✓ = positive result

x = negative result

- 5 The diagram shows an alveolus and one of its capillaries.

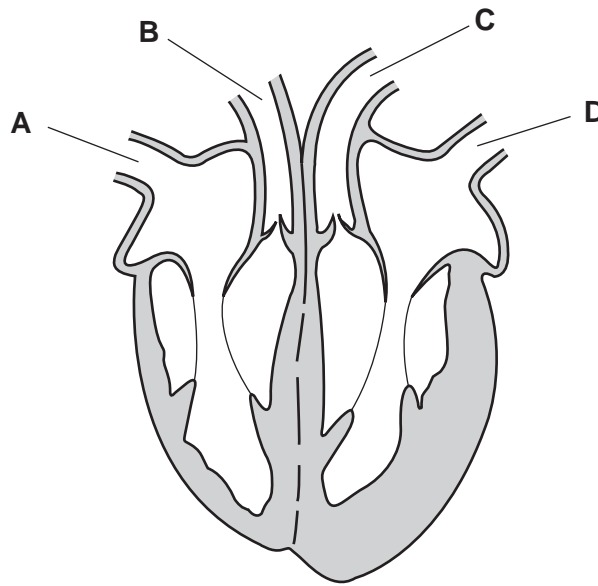


What moves in the direction shown by the arrows?

- A** carbon dioxide
- B** hydrogen
- C** oxygen
- D** water

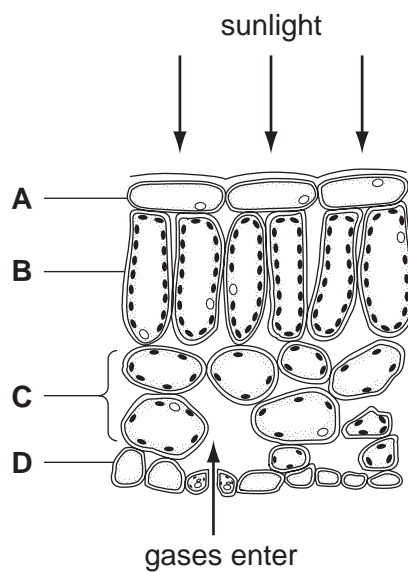
6 The diagram shows a section through the human heart.

Which vessel is a vein containing oxygenated blood?



7 The diagram shows some cells in a leaf of a green plant.

In which layer of cells does most photosynthesis occur?



- 8 In a healthy person, which shows the correct relationship between blood sugar level, insulin level, and liver activity?

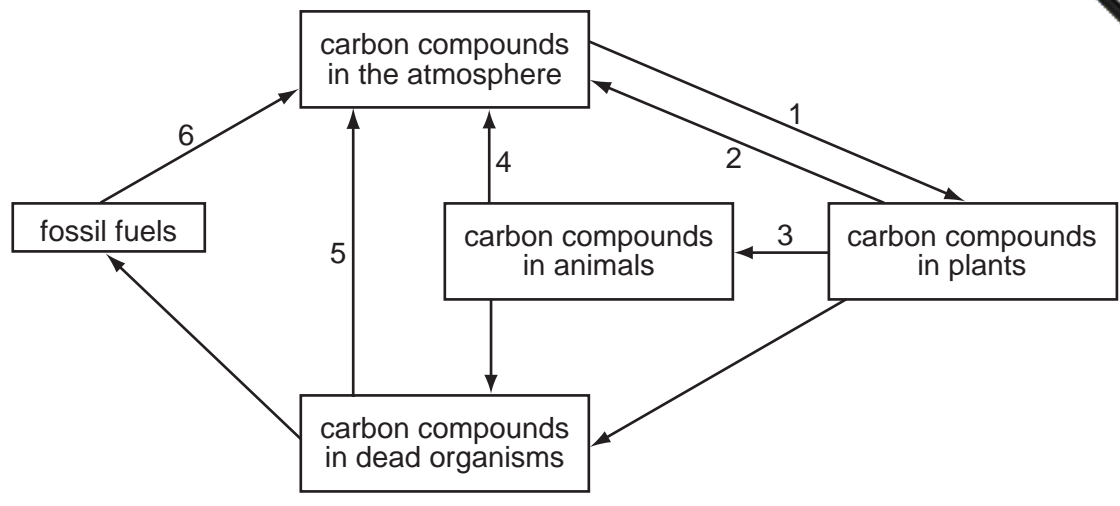
	blood sugar level	insulin level	liver activity
A	high	high	removes glucose from the blood
B	high	low	releases glucose into the blood
C	low	high	releases glucose into the blood
D	low	low	removes glucose from the blood

- 9 A variety of potato plant produces red tubers ('potatoes') that grow into new potato plants which then produce red 'potatoes' the following year.

Why is this?

- A** Asexual reproduction produces identical potato plants.
- B** Asexual reproduction results in different coloured 'potatoes'.
- C** Sexual reproduction requires the potato plant to produce flowers.
- D** Sexual reproduction produces only red coloured 'potatoes'.
- 10 After it has been fertilised, which part of a flower develops into a seed?
- A** egg
- B** ovary
- C** ovule
- D** pollen
- 11 Which pairs of human features are inherited and **not** affected by the environment?
- A** blood group and body mass
- B** blood group and sex
- C** hair colour and height
- D** sex and body mass

12 The diagram shows the carbon cycle.



Which of the numbered processes represent respiration and photosynthesis?

	respiration	photosynthesis
A	3	1
B	4	1
C	5	2
D	6	3

13 Which statement describes species diversity?

- A** the number of different types of habitat in which species are found
- B** the total number of habitats in which a species is found
- C** the number of species in a community
- D** the number of variations within a species

14 Which substance is an element?

- A** air
- B** brass
- C** iron
- D** water

15 Atoms of four different elements are shown.

Which atom contains six neutrons?

- A** ${}^4_2\text{He}$
- B** ${}^6_3\text{Li}$
- C** ${}^{11}_5\text{B}$
- D** ${}^{14}_6\text{C}$

16 Which substance is an ionic compound?

	melting point	electrical conductivity when melted
A	high	high
B	high	low
C	low	high
D	low	low

17 Which property of an element **cannot** be predicted from its position in the Periodic Table?

- A** the charge on its ion
- B** the melting point of the element
- C** the metallic/ non-metallic character of the element
- D** the number of protons in its nucleus

18 The diagram shows some elements in Groups III, IV and V of the Periodic Table.

III	IV	V
Al	Si	P
Ga	Ge	As
In	Sn	Sb
Tl	Pb	Bi

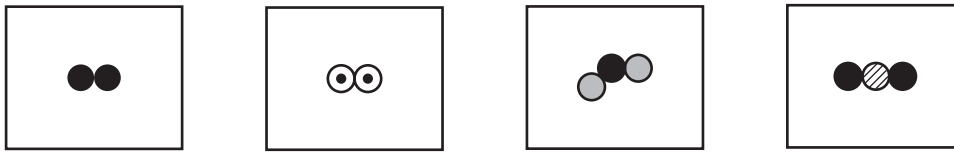
Which two elements would be expected to form an oxide of the type XO_2 ?

	In and Sn	Sn and Pb
A	✓	✓
B	✓	x
C	x	✓
D	x	x

19 Which element, present in fossil fuels, is responsible for causing 'acid rain'?

- A** carbon
- B** hydrogen
- C** oxygen
- D** sulphur

- 20 The diagrams show some molecules of substances present in air. Different circles represent atoms of different elements.



Which elements could be shown as \odot and \bullet ?

	\odot = nitrogen	\bullet = oxygen
A	✓	✓
B	✓	x
C	x	✓
D	x	x

- 21 Three metals are listed.

copper

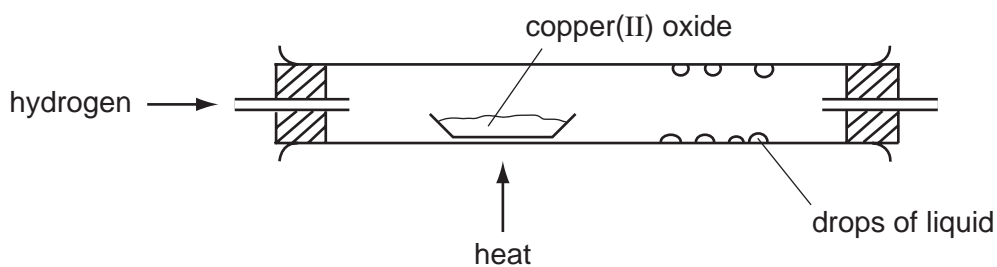
magnesium

zinc

Which of these metals react with dilute sulphuric acid?

- A** copper and magnesium only
B copper and zinc only
C magnesium and zinc only
D copper, magnesium and zinc

22 Hydrogen is passed over heated copper(II) oxide as shown.



The copper(II) oxide is reduced.

Which other statement also describes a change that occurs during the reaction?

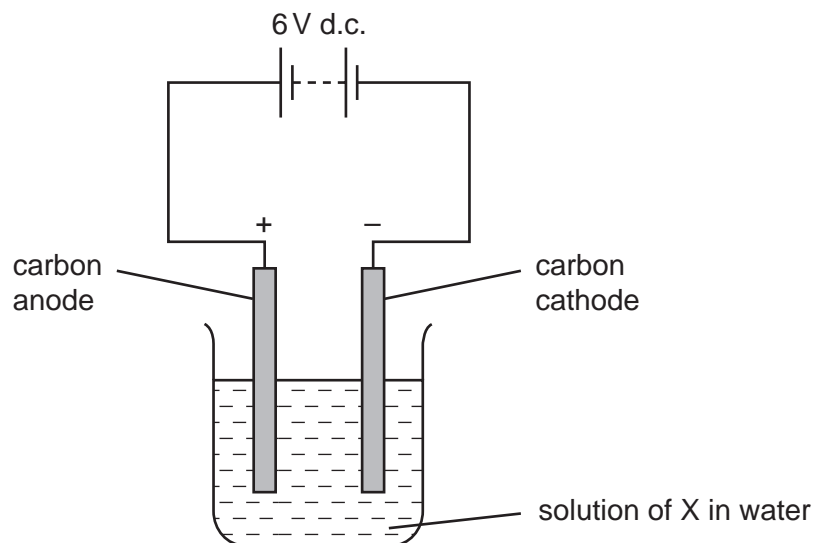
- A Copper is distilled.
- B Copper(II) oxide is thermally decomposed.
- C Hydrogen is condensed.
- D Hydrogen is oxidised.

23 Sodium chloride is an ionic salt that is used industrially as an electrolyte.

Under which conditions does sodium chloride behave as an electrolyte?

	solid	molten	in aqueous solution
A	no	no	yes
B	no	yes	yes
C	yes	no	no
D	yes	yes	no

24 A substance, X, is dissolved in water and electrolysed as shown.



A yellow-green gas is given off at the anode and the cathode becomes brown.

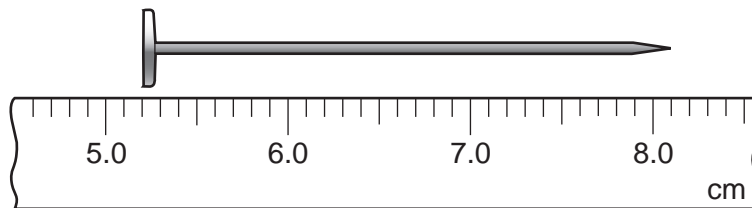
What is X?

- A copper(II) chloride
 - B lead(II) bromide
 - C sodium bromide
 - D sodium chloride
- 25 Which fuel burns to form only one product?
- A coal
 - B hydrogen
 - C methane
 - D petrol
- 26 Why is water often used to extinguish fires?
- A Water is a compound.
 - B Water is neutral.
 - C Water reacts with most fuels.
 - D Water removes heat from the fire.

- 27 Some man-made plastics are made from small molecules which join together by covalent bonds. What is the main source of these small molecules and what is the structure of the plastics?

	<i>source of small molecules</i>	<i>structure</i>
A	coal	$\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \text{H} \\ \\ \text{H} \end{array}$
B	coal	$\begin{array}{cccc} \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ \cdots - \text{C} & - \text{C} & - \text{C} & - \text{C} - \cdots \\ & & & \\ \text{H} & \text{H} & \text{H} & \text{H} \end{array}$
C	oil	$\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \text{H} \\ \\ \text{H} \end{array}$
D	oil	$\begin{array}{cccc} \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ \cdots - \text{C} & - \text{C} & - \text{C} & - \text{C} - \cdots \\ & & & \\ \text{H} & \text{H} & \text{H} & \text{H} \end{array}$

- 28 A ruler is used to measure the length of a nail.



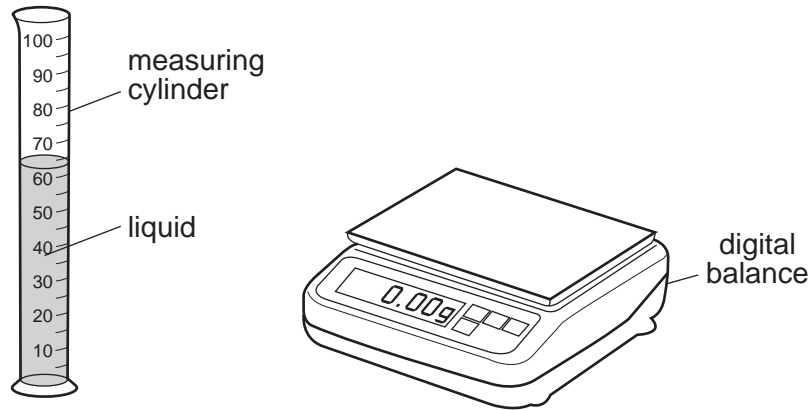
What is the length of the nail?

- A** 1.3 cm **B** 2.9 cm **C** 5.2 cm **D** 8.1 cm
- 29 A newton is a unit of force.

Which quantity is measured in newtons?

- A** acceleration
B density
C mass
D weight

30 A student pours liquid into a measuring cylinder.



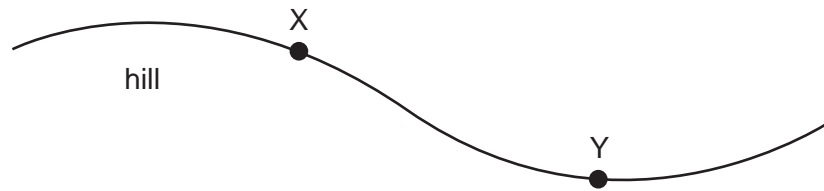
The student records the volume of the liquid from the scale on the measuring cylinder. He then puts the measuring cylinder containing the liquid on a balance and records the mass.

What else needs to be measured before the density of the liquid can be calculated?

- A the depth of the liquid in the measuring cylinder
 - B the mass of the empty measuring cylinder
 - C the temperature of the liquid in the measuring cylinder
 - D the volume of the empty measuring cylinder
- 31 Which source of energy uses the production of steam to generate electricity?
- A hydroelectric
 - B nuclear
 - C tides
 - D waves

32 A cyclist travels down a hill from rest at point X without pedalling.

The cyclist applies his brakes and the cycle stops at point Y.



Which energy changes have taken place between X and Y?

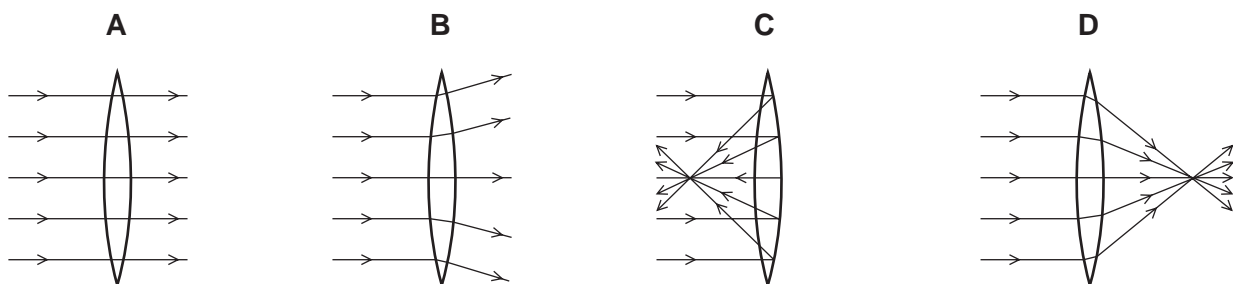
- A kinetic \rightarrow heat \rightarrow potential
- B kinetic \rightarrow potential \rightarrow heat
- C potential \rightarrow heat \rightarrow kinetic
- D potential \rightarrow kinetic \rightarrow heat

33 Which line in the table is correct about conduction and convection?

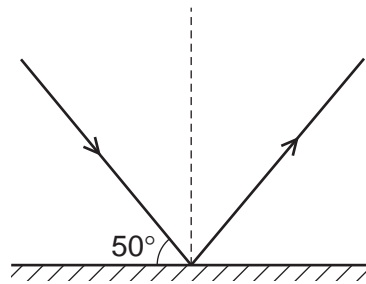
	conduction	convection
A	can happen in a solid	can happen in a solid
B	can happen in a solid	only happens in fluids
C	only happens in fluids	can happen in a solid
D	only happens in fluids	only happens in fluids

34 A parallel beam of light falls on a converging lens.

Which diagram shows what happens to the beam of light?

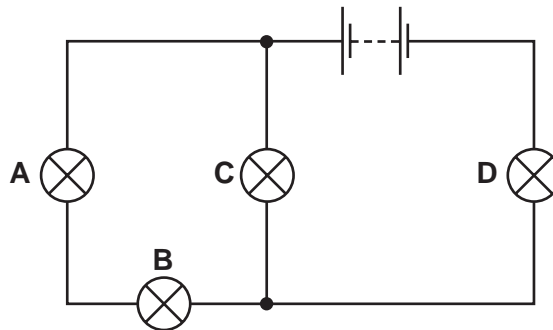


- 35 The diagram shows a ray of light striking a plane mirror.



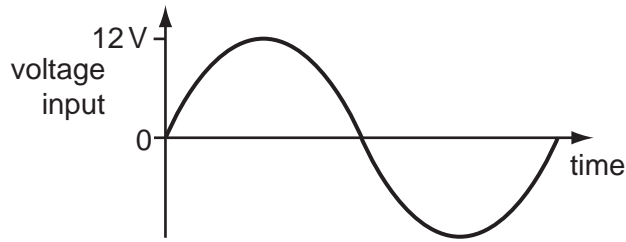
What is the angle of reflection?

- A** 40° **B** 50° **C** 80° **D** 130°
- 36 When electricity is transmitted over long distances energy is wasted.
How can the wasted energy be kept as small as possible?
- A** Keep the current in the transmission lines as large as possible.
B Keep the power supplied to the transmission lines as large as possible.
C Keep the resistance of the transmission lines as large as possible.
D Keep the voltage supplied to the transmission lines as large as possible.
- 37 In the circuit below, one of the lamps breaks, causing all the other lamps to go out.
Which lamp breaks?

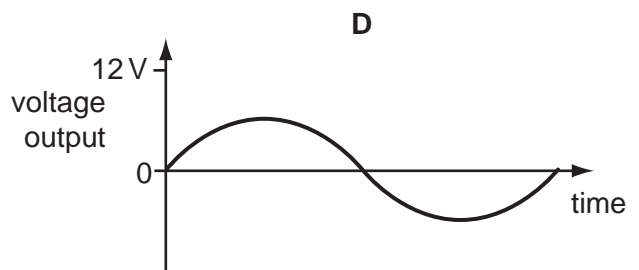
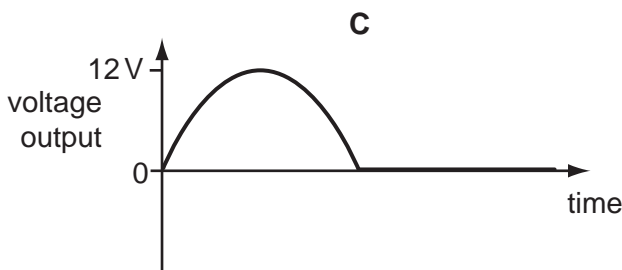
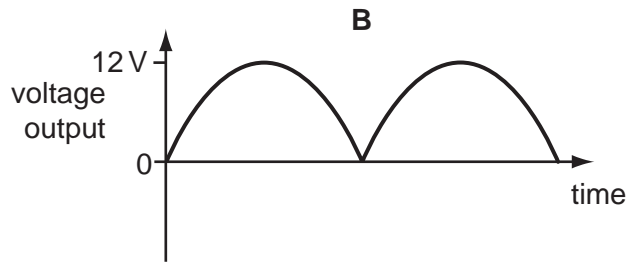
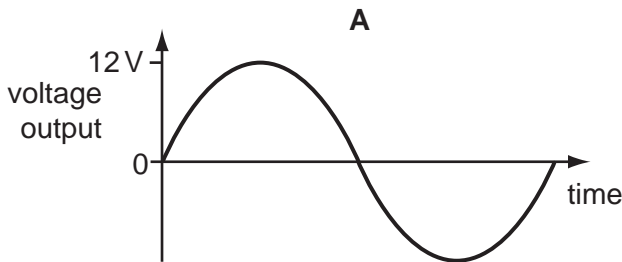


- 38 An electric heater is connected to the mains using insulated copper wires. The wires become very warm.
What can be done to prevent so much heat being produced in the connecting wires?
- A** Use thicker copper wires.
B Use thinner copper wires.
C Use thicker insulation.
D Use thinner insulation.

39 The graph shows the voltage input to a step-down transformer.



Which diagram shows the voltage output from the transformer?



40 Which line in the table describes the nature of an alpha-particle and of a gamma-ray?

	alpha-particle	gamma-ray
A	helium nucleus	electromagnetic radiation
B	helium nucleus	electron
C	proton	electromagnetic radiation
D	proton	electron

DATA SHEET
The Periodic Table of the Elements

		Group																						
	I	II	III	IV	V	VI	VII	0																
	1 H Hydrogen 1																							
7 Li Lithium 3	9 Be Beryllium 4																							
23 Na Sodium 11	24 Mg Magnesium 12																							
39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36							
85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54									
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	190 Os Osmium 76	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86									
226 Ra Radium 88	227 Ac Actinium 89																							
											140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	146 Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	
											232 Th Thorium 90	232 Pa Protactinium 91	238 U Uranium 92	238 Np Neptunium 93	244 Pu Plutonium 94	244 Am Americium 95	244 Cm Curium 96	244 Bk Berkelium 97	244 Cf Californium 98	244 Es Einsteinium 99	244 Fm Fermium 100	244 Md Mendelevium 101	244 No Nobelium 102	244 Lr Lawrencium 103

*58-71 Lanthanoid series
†90-103 Actinoid series

a = relative atomic mass

X = atomic symbol

b = proton (atomic) number

Key

a	X
b	

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).