



Cambridge
IGCSE

Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CO-ORDINATED SCIENCES (DOUBLE) (US)

0442/13

Paper 1 Multiple Choice

May/June 2015

45 minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Center number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

This document consists of **16** printed pages.

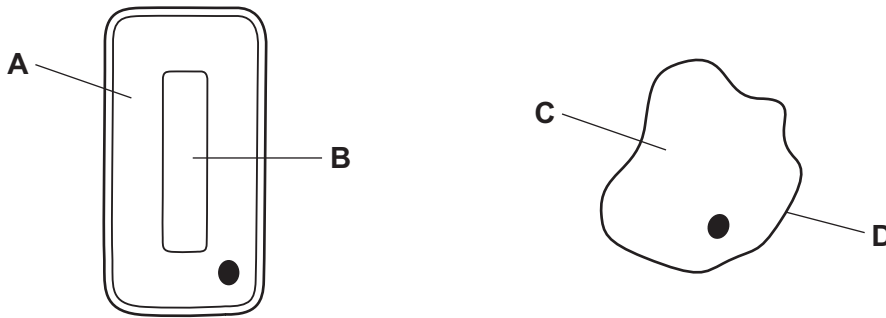


1 What are all living things capable of?

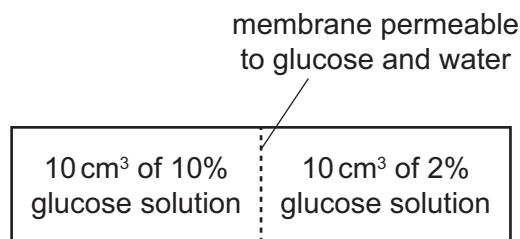
- A excretion
- B digestion
- C photosynthesis
- D sexual reproduction

2 The diagram shows two cells.

Which labeled part might contain chloroplasts?



3 Diffusion occurs between the two solutions shown.

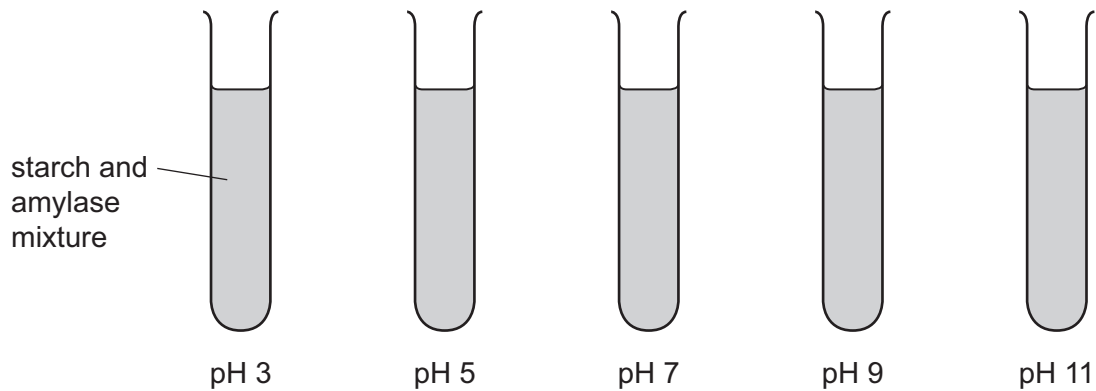


What is the final concentration of glucose solution in each region?

- A 2%
- B 6%
- C 8%
- D 12%

- 4 A student carried out an experiment to investigate the effect of pH on the activity of the enzyme amylase.

She set up five test-tubes of starch and amylase mixture, each at a different pH.



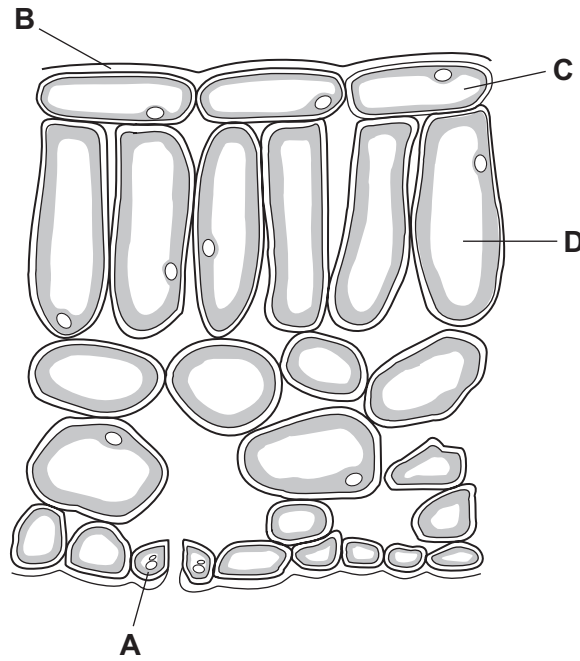
At which temperature(s) should the test-tubes be kept during this experiment?

- A all at 37 °C
 - B all at 100 °C
 - C at a range of temperatures between 0 °C and 50 °C
 - D at a range of temperatures between 20 °C and 70 °C
- 5 In a balanced diet, which constituents provide most energy?
- A carbohydrate and protein
 - B fat and carbohydrate
 - C fat and fiber
 - D vitamins and protein

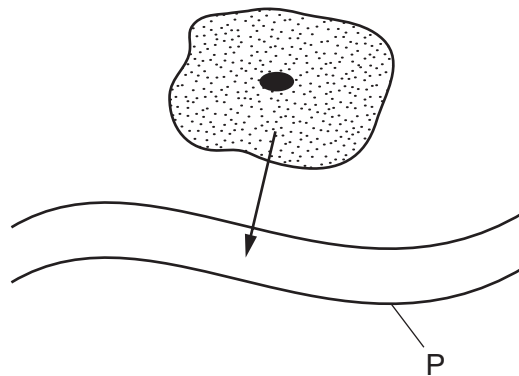
4

6 The diagram shows a section through a leaf.

Which is a cell with **no** chloroplasts?



7 The arrow shows urea leaving a cell and passing into structure P.



What is P?

- A a capillary
- B an artery
- C a vein
- D the small intestine

- 8 What describes respiration?
- A absorption of oxygen in the alveoli
 - B carbohydrate production in plant cells
 - C the break down of nutrient molecules to release energy
 - D the inspiration of gases in an animal
- 9 A person touches a hot object which triggers a reflex action.
In which order does the signal travel in the reflex arc?
- A relay neurone → spinal cord → sensory neurone
 - B sensory neurone → spinal cord → motor neurone
 - C spinal cord → sensory neurone → stimulus
 - D stimulus → motor neurone → spinal cord

- 10 Which are target organs for adrenaline?

	heart	liver
A	x	x
B	x	✓
C	✓	x
D	✓	✓

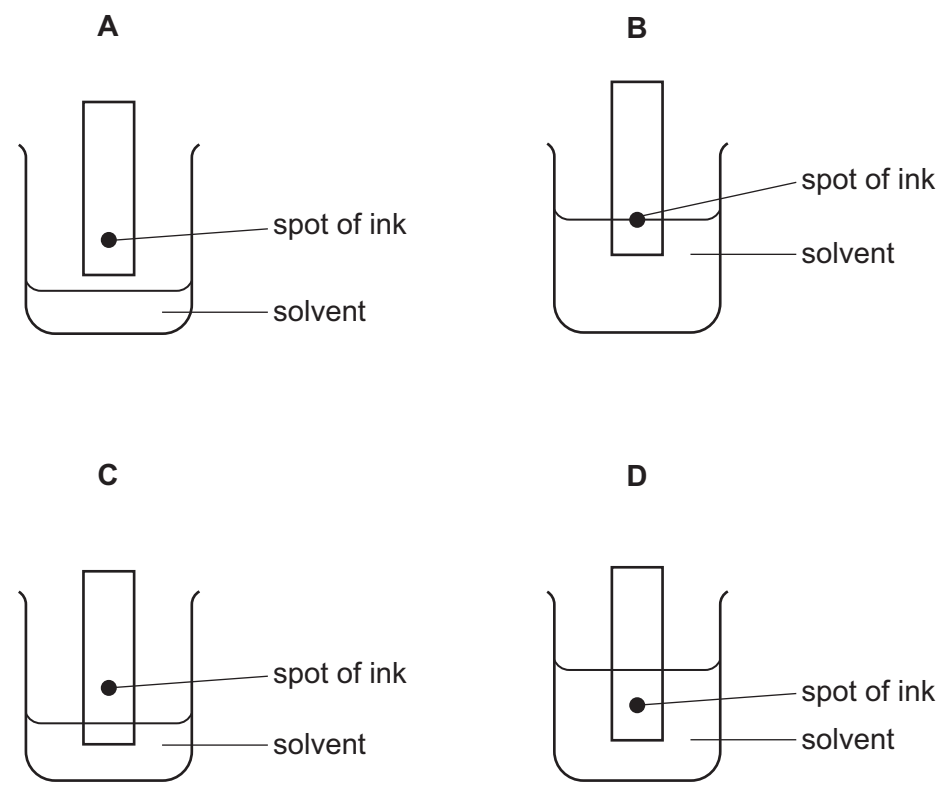
- 11 What is fertilization?
- A a pollen tube nucleus reaching an ovule
 - B a sperm reaching an ovum
 - C a zygote being formed
 - D pollen grains reaching a stigma
- 12 Which process is responsible for the flow of energy along a food chain?
- A excretion
 - B feeding
 - C respiration
 - D seed dispersal

13 Which two gases are considered to be air-polluting gases and contribute to global warming?

- A carbon dioxide and methane
- B carbon dioxide and nitrogen
- C oxygen and methane
- D oxygen and nitrogen

14 The colors in an ink can be separated by chromatography.

Which diagram shows the correct way to set up the apparatus?



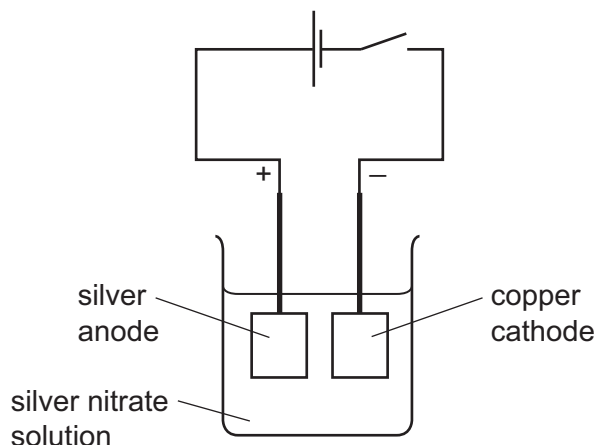
15 The positions of four elements are shown on the outline of part of the Periodic Table.

Which element forms an ion with a charge of 2+?

A	B																		

7

16 The diagram shows an electroplating experiment.



Which row shows the change in mass of each electrode?

	anode	cathode
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

17 Which statement about the energetics of a reaction is correct?

- A** In an endothermic reaction heat is given out and the temperature decreases.
- B** In an endothermic reaction heat is taken in and the temperature increases.
- C** In an exothermic reaction heat is given out and the temperature increases.
- D** In an exothermic reaction heat is taken in and the temperature decreases.

18 Dilute sulfuric acid reacts with a piece of zinc.

Which change does **not** speed up the reaction?

- A** Use a catalyst.
- B** Use a larger volume of dilute sulfuric acid.
- C** Use an equal volume of more concentrated sulfuric acid.
- D** Use the same mass of powdered zinc.

19 Hydrogen and oxygen react explosively to form water.

Which words describe this reaction?

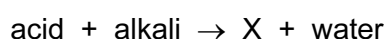
	combustion	oxidation
A	✓	✓
B	✓	x
C	x	✓
D	x	x

key

✓ = yes

x = no

20 The equation shown is incomplete.



What is X?

- A** base
- B** carbon dioxide
- C** hydrogen
- D** salt

21 Sodium hydroxide solution and aluminum powder are added to a salt solution and warmed.

A gas is produced that turns moist red litmus paper blue.

Which anion is present in the salt?

- A** carbonate
- B** chloride
- C** nitrate
- D** sulfate

22 A gas is used in welding metals together at high temperatures.

The gas is used to provide an inert atmosphere.

What is the gas?

- A** argon
- B** carbon dioxide
- C** fluorine
- D** oxygen

23 The table shows information about some minerals.

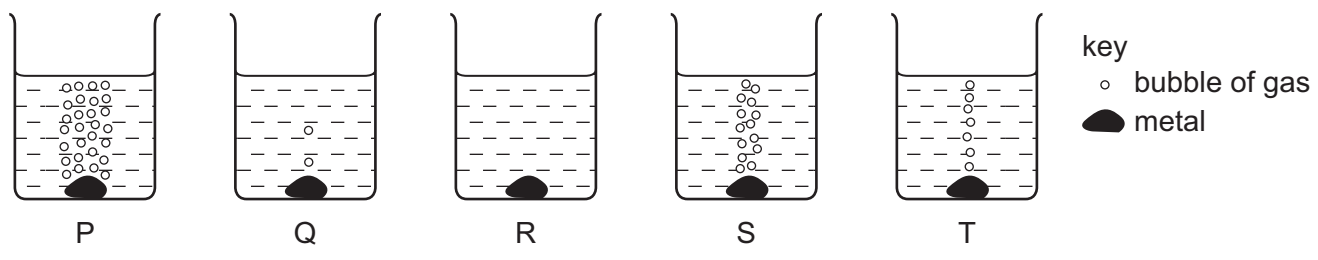
mineral	chemical formula
bauxite	Al_2O_3
galena	PbS
hematite	Fe_2O_3
rutile	TiO_2

Which minerals contain a transition element?

- A bauxite and galena
- B bauxite and hematite
- C galena and rutile
- D hematite and rutile

24 Similar sized pieces of five different metals P, Q, R, S and T are reacted with equal volumes of dilute hydrochloric acid of the same concentration.

The results are shown below.



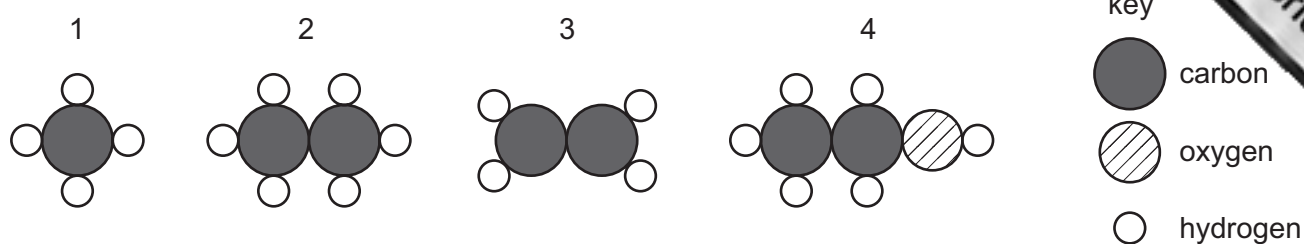
What is the order of reactivity?

	least reactive		→	most reactive	
A	P	S	T	Q	R
B	R	Q	T	S	P
C	R	T	Q	P	S
D	Q	R	S	T	P

25 Which gas forms 78% of the air?

- A argon
- B carbon dioxide
- C nitrogen
- D water vapor

26 The structures of four organic molecules are shown.



Which row correctly identifies these compounds?

	1	2	3	4
A	ethane	ethanol	methane	ethene
B	ethanol	ethene	ethane	methane
C	ethene	methane	ethanol	ethane
D	methane	ethane	ethene	ethanol

27 A fuel used for cooking food is the hydrocarbon ...1... that burns in an ...2... reaction.

Which words correctly complete gaps 1 and 2?

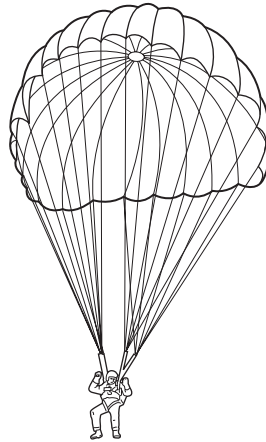
	1	2
A	coke	endothermic
B	coke	exothermic
C	methane	endothermic
D	methane	exothermic

28 The circuit of a motor racing track is 3.0 km in length. In a race, a car goes 25 times round the circuit in 30 minutes.

What is the average speed of the car?

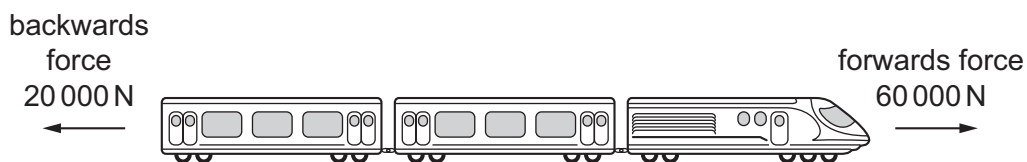
- A** 75 km/hour
- B** 90 km/hour
- C** 150 km/hour
- D** 750 km/hour

- 29 Inside an airplane, a parachutist has a mass of 70 kg.



What is his mass after he has jumped from the airplane?

- A 0 kg
B between 0 kg and 70 kg
C 70 kg
D greater than 70 kg
- 30 A train travels along a horizontal track at constant speed. Two of the forces acting on the train are shown in the diagram.



A force of air resistance is also acting on the train to give it a resultant force of zero.

What is this air resistance force?

- A 40 000 N backwards
B 80 000 N backwards
C 40 000 N forwards
D 80 000 N forwards
- 31 Which energy sources are both renewable?
- A oil and coal
B oil and tidal
C tidal and geothermal
D tidal and nuclear fission

32 Which row describes the molecules of a solid at 0 °C, a liquid at 0 °C and a gas at 0 °C?

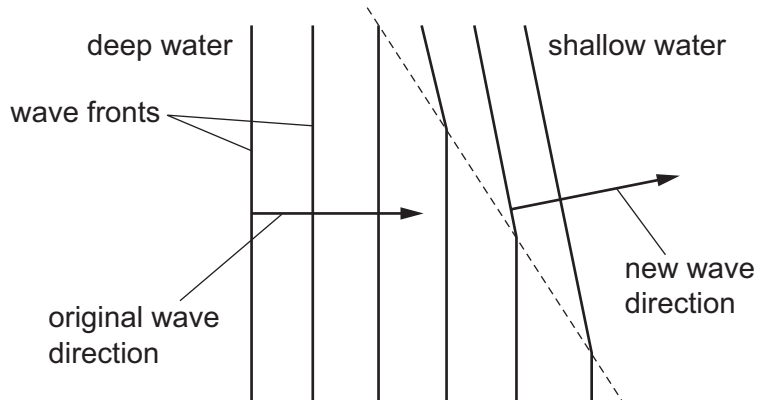
	solid	liquid	gas
A	stationary	stationary	stationary
B	stationary	stationary	moving
C	stationary	moving	moving
D	moving	moving	moving

33 There is a vacuum between the double walls of a vacuum flask.

Which types of heat transfer are reduced by the vacuum?

- A** conduction, convection and radiation
- B** conduction and convection only
- C** conduction and radiation only
- D** convection and radiation only

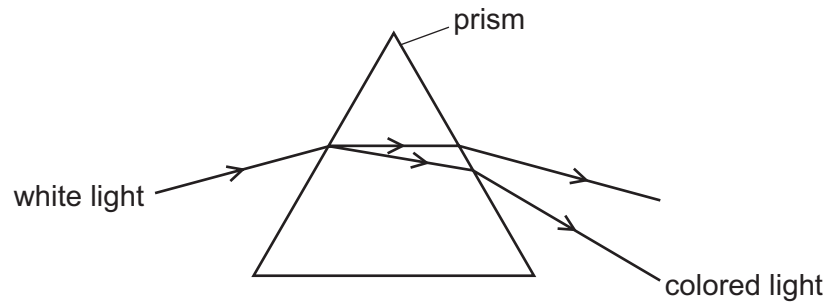
34 The diagram represents a water wave changing direction as it moves into a shallower region.



What happens to the speed and what happens to the wavelength of the wave as it changes direction?

	speed	wavelength
A	changes	changes
B	changes	stays the same
C	stays the same	changes
D	stays the same	stays the same

- 35 One of the effects of passing a ray of white light through a prism is to split the light into



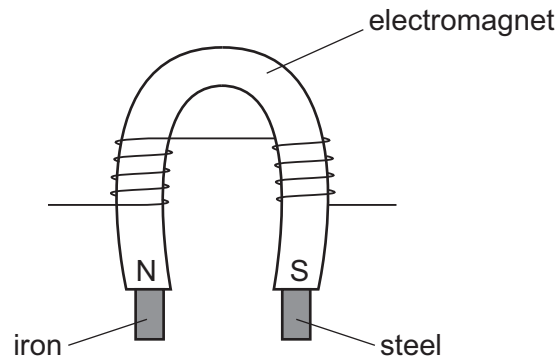
What is the name given to this splitting effect?

- A dispersion
 - B radiation
 - C reflection
 - D refraction
- 36 Sound waves may cause an echo.

What happens to sound waves to cause an echo and what is the nature of sound waves?

	what an echo is caused by	nature of sound waves
A	reflection	longitudinal
B	reflection	transverse
C	refraction	longitudinal
D	refraction	transverse

37 A piece of iron and a piece of steel are attracted to an electromagnet as shown.



The electromagnet is now switched off.

What happens?

- A Both the iron and the steel remain magnetized.
- B Neither the iron nor the steel remains magnetized.
- C Only the iron remains magnetized.
- D Only the steel remains magnetized.

38 The diagrams show two possible ways in which a fuse and a lamp can be connected in a circuit. The current in the lamp is 2.0 A.

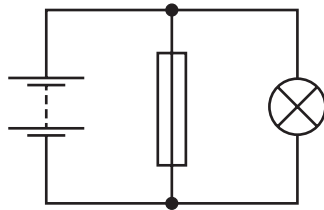


diagram 1

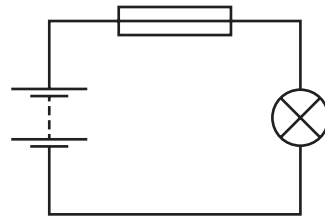


diagram 2

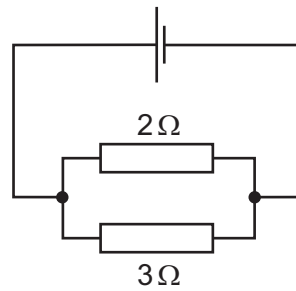
A fault develops. The current in the lamp increases, and the fuse blows.

Which row gives the correct position of the fuse, and the effect of the fuse when it blows?

	correct position	effect
A	as in diagram 1	reduces current to 0
B	as in diagram 1	reduces current to 2.0 A
C	as in diagram 2	reduces current to 0
D	as in diagram 2	reduces current to 2.0 A

15

39 A $2\ \Omega$ and a $3\ \Omega$ resistor are connected as shown.



What is the total resistance of the two resistors?

- A less than $2\ \Omega$
 - B between $2\ \Omega$ and $3\ \Omega$
 - C between $3\ \Omega$ and $5\ \Omega$
 - D exactly $5\ \Omega$
- 40 Which type of radiation has the greatest ionizing effect?
- A infra-red rays
 - B α -particles
 - C β -particles
 - D γ -rays

DATA SHEET
The Periodic Table of the Elements

		Group																																																																	
I	II	III	IV	V	VI	VII	0					0																																																							
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10	23 Na Sodium 11	24 Mg Magnesium 12	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18	39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36	85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	101 Ru Ruthenium 44	106 Pd Palladium 46	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54	133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86	226 Ra Radium 88	227 Ac Actinium 89	†
												140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	232 Th Thorium 90	238 U Uranium 92	238 Pa Protactinium 91	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103																													

*58-71 Lanthanoid series
†90-103 Actinoid series

Key

a	X	a = relative atomic mass
b	X	X = atomic symbol
		b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).