



General Certificate of Secondary Education
2015–2016

Centre Number

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Candidate Number

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Double Award Science: Chemistry

Unit C1
Higher Tier



[GSD22]

THURSDAY 12 NOVEMBER 2015, MORNING

TIME

1 hour, plus your additional time allowance.

INSTRUCTIONS TO CANDIDATES

Write your Centre Number and Candidate Number in the spaces provided at the top of this page.
Write your answers in the spaces provided in this question paper.
Answer **all eight** questions.

INFORMATION FOR CANDIDATES

The total mark for this paper is 70.
Figures in brackets printed down the right-hand side of pages indicate the marks awarded to each question or part question.
Quality of written communication will be assessed in Question **2(b)**.
A Data Leaflet, which includes a Periodic Table of the Elements, is included in this question paper.

For Examiner's use only	
Question Number	Marks
1	
2	
3	
4	
5	
6	
7	
8	

Total Marks	
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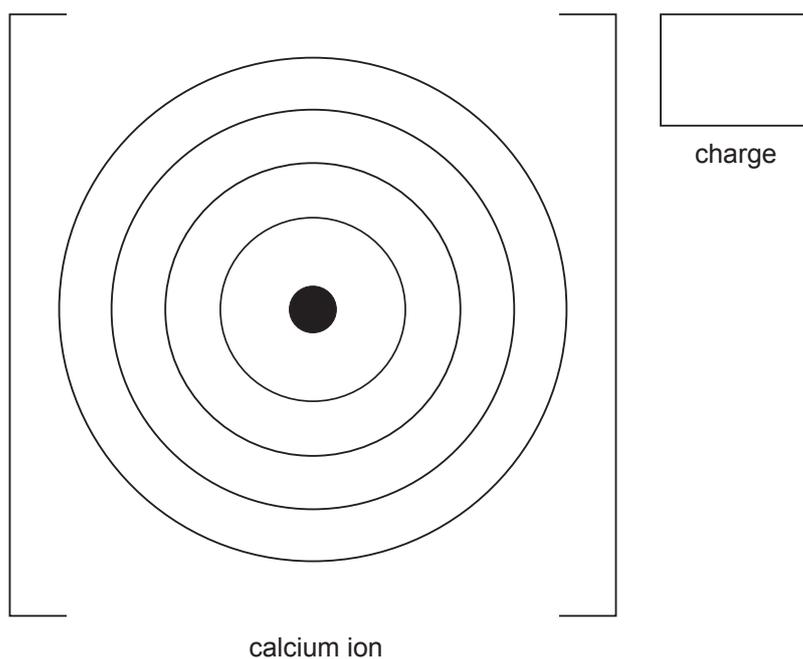
1 Some tap water contains dissolved calcium chloride (CaCl_2) and dissolved calcium hydrogen carbonate, $\text{Ca}(\text{HCO}_3)_2$.

(a) Write two reasons why calcium hydrogen carbonate and calcium chloride could be **ionic** compounds.

1. _____

2. _____ [2]

(b) Draw the electronic configuration of the calcium **ion** and write in the charge.



[2]

(c) How many oxygen atoms are there in the formula $\text{Ca}(\text{HCO}_3)_2$?

_____ [1]

Some tap water can contain dissolved magnesium sulfate or dissolved potassium carbonate.

(d) Write the formulae for magnesium sulfate and potassium carbonate.

magnesium sulfate _____

potassium carbonate _____ [2]

(e) What colour are solid magnesium sulfate and solid potassium carbonate?

_____ [1]

Examiner Only	
Marks	Remark
○	○

2 Chemists use the Periodic Table to help them understand the reactions of the elements. John Newlands was one of the first chemists to see repeating patterns in the properties and reactions of the elements.

(a) What name did Newlands give to the repeating pattern he saw in the properties of the elements?

_____ [1]

Examiner Only	
Marks	Remark
○	○

3 Metal oxides are bases. They react with strong acids to form salts.

(a) What is the pH range of a strong acid?

_____ [1]

Copper oxide reacts with sulfuric acid.

(b) (i) Complete the **word** equation below for this reaction.

copper oxide + sulfuric acid → _____ + _____ [2]

(ii) Why can this reaction be described as a neutralisation reaction?

 _____ [2]

(iii) What colour change is seen **in the solution** as this reaction is happening?

from _____ to _____ [2]

Sodium oxide reacts with hydrochloric acid.

(c) (i) Complete a balanced symbol equation for the reaction between sodium oxide and hydrochloric acid.

$\text{Na}_2\text{O} + \text{HCl} \rightarrow$ _____ + _____ [3]

(ii) Sodium oxide is an alkali. Why could sodium oxide be described as an alkali?

 _____ [2]

Examiner Only	
Marks	Remark
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(Questions continue overleaf)

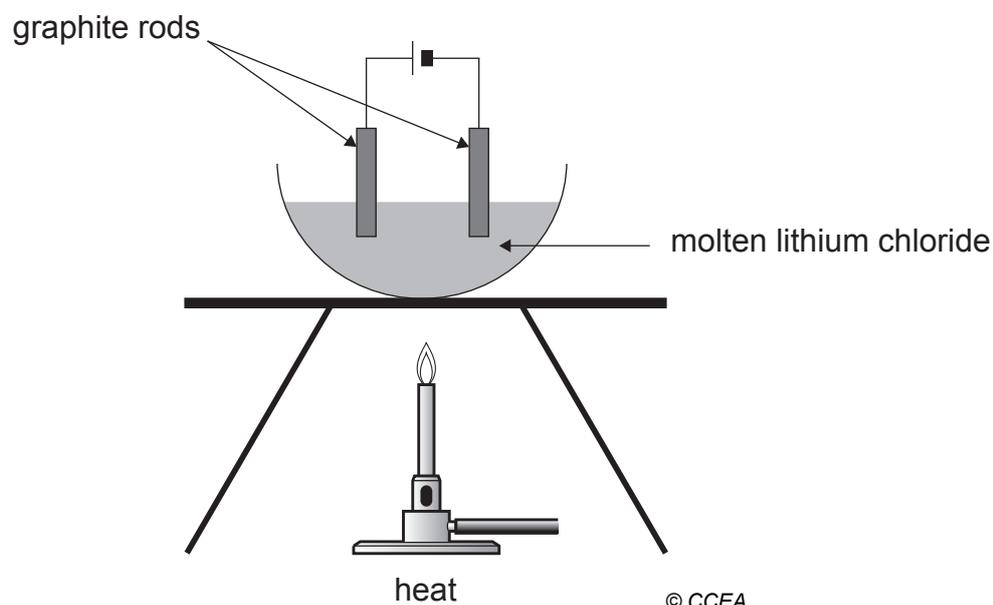
- 4 Metals, such as copper, are good conductors of electricity. Metal compounds, such as sodium chloride, will conduct electricity when they are molten or dissolved in water.

(a) Complete the table below to give information about how copper and sodium chloride solution conduct electricity.

Name of conductor	Name of the type of particle which moves and carries the charge	Effect of the passage of electricity on the conductor
		(A) No effect (B) Conductor breaks down (C) The conductor melts
copper		
sodium chloride solution		

[4]

The diagram below shows the apparatus used in the laboratory for the electrolysis of molten lithium chloride.



- (b) (i) Write down one reason why graphite is used to make these electrodes.

_____ [1]

- (ii) Write a half equation for the reaction that happens at the cathode when molten lithium chloride undergoes electrolysis.

_____ [2]

- 6 Look at the table below. It summarises the observations made when solutions of the halogens are added to solutions of the halide ions.

halogen \ halide solution	Examiner Only		
	potassium chloride solution	potassium bromide solution	potassium iodide solution
chlorine		colourless solution becomes orange-brown	colourless solution becomes dark brown
bromine	no reaction		colourless solution becomes dark brown
iodine	no reaction	no reaction	

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- (a) Explain why bromine solution does not react with potassium chloride solution.

_____ [2]

- (b) Name the **type** of reaction that takes place between potassium bromide and chlorine.

_____ [1]

- (c) Write a balanced **symbol** equation for the reaction between chlorine and potassium iodide.

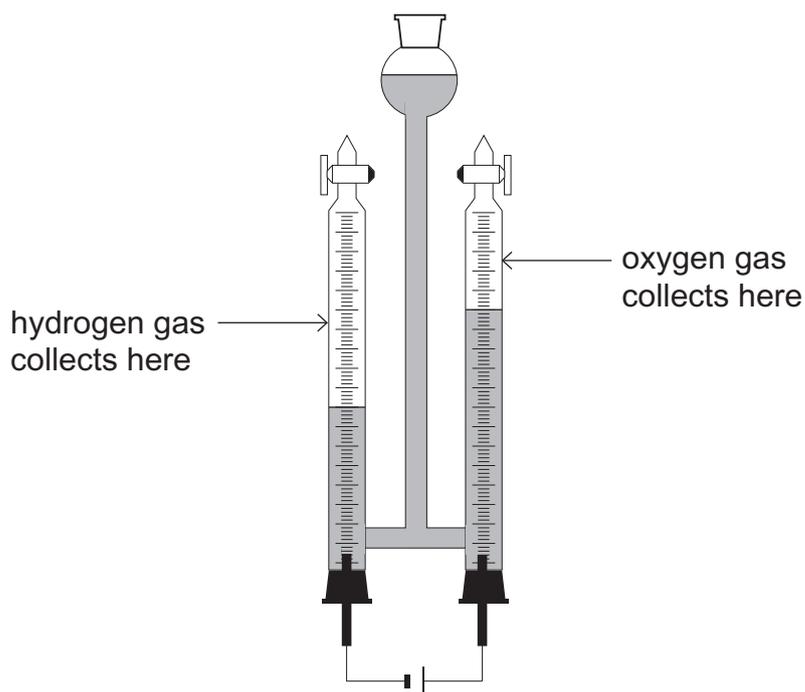
_____ [3]

- (d) Explain, in terms of their electronic structures, why the halogens have similar chemical properties.

_____ [2]

Examiner Only	
Marks	Remark
○	○

- 7 Look at the diagram below. This apparatus may be used in the laboratory to show that water (H_2O) is a compound made from the elements hydrogen and oxygen.



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- (a) Describe a **chemical** test for water.

[3]

In one demonstration the water decomposed to form 10 cm^3 of hydrogen gas and 5 cm^3 of oxygen gas.

- (b) Explain why the volume of hydrogen formed was exactly twice as much as the volume of oxygen formed.

[1]

- (c) Describe a test for hydrogen.

[2]

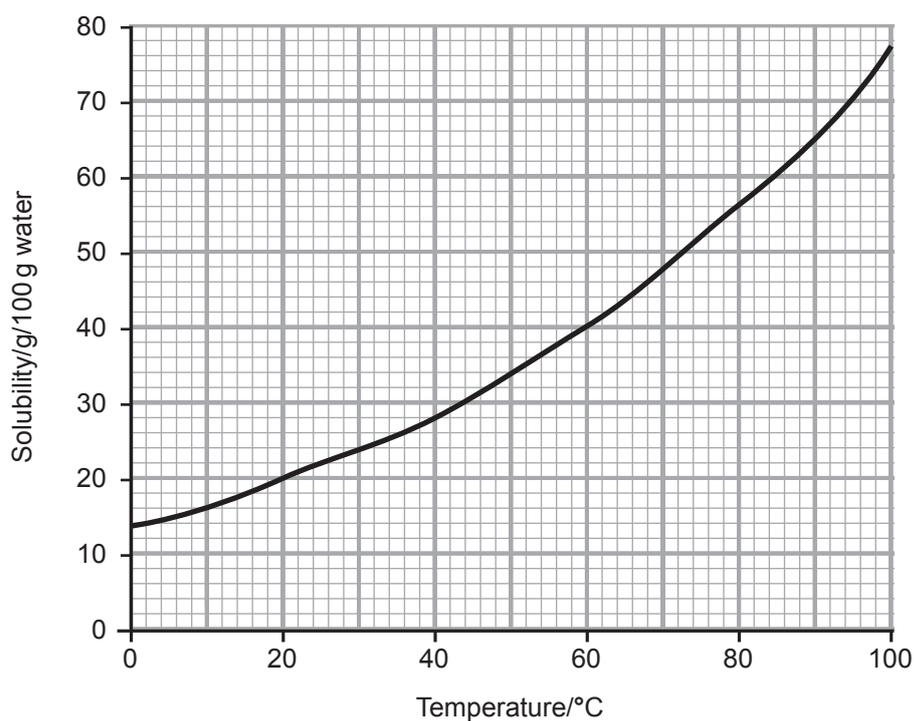
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Marks	Remark
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- 8 When a hot saturated solution of copper(II) sulfate cools down, crystals of hydrated copper(II) sulfate form.

(a) Explain why crystals form when the solution is cooled.

_____ [2]

The solubility curve for copper(II) sulfate is shown below.



- (b) Calculate the mass of copper(II) sulfate that will crystallise when a saturated solution, in 20 g water, is cooled from 82 °C to 30 °C.

_____ g [4]

THIS IS THE END OF THE QUESTION PAPER

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Marks	Remark
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