



Rewarding Learning

General Certificate of Secondary Education
2012–2013

Centre Number

71

Candidate Number

Science: Single Award

Unit 2 (Chemistry)

Higher Tier

[GSS22]

ML

MONDAY 20 MAY 2013, AFTERNOON

TIME

1 hour 15 minutes, plus your additional time allowance.

INSTRUCTIONS TO CANDIDATES

Write your Centre Number and Candidate Number in the spaces provided at the top of this page.

Write your answers in the spaces provided in this question paper.
Answer **all eleven** questions.

INFORMATION FOR CANDIDATES

The total mark for this paper is 75.

Quality of written communication will be assessed in Questions **3** and **9(b)**.

Figures in brackets printed down the right-hand side of pages indicate the marks awarded to each question or part question. A Data Leaflet, which includes a Periodic Table of the Elements, is included in this question paper.

For Examiner's
use only

Question Number	Marks
1	
2	
3	
4	
5	
6	
7	
8	
9	
10	
11	

Total
Marks



1 The table below shows properties of some plastics.

Plastic	Melting point/°C	Resistance to alkali	Other properties	Cost per kg /£	Examiner Only	
					Marks	Remark
A	20	highly resistant	strong and flexible	1.1		
B	120	slowly reacts	strong and flexible	1.5		
C	200	highly resistant	strong and shatters easily	0.5		
D	160	highly resistant	strong and not very flexible	2.4		

Use the information in the table and your knowledge to answer the questions below.

(a) Why is plastic **C** **not** suitable to cover the copper wire in an electrical cable?

_____ [1]

(b) A shop owner has a large warehouse where he stores things. He wants to cover the things he stores with plastic sheets. Why is plastic **B** a better choice than plastic **D**?

_____ [1]

(c) The shop owner needs large plastic containers to hold a corrosive alkali. The containers will be loaded on and off lorries.

Which plastic (**A**, **B**, **C** or **D**) would be the best one to use for the containers? Explain your answer fully.

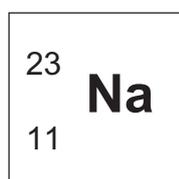
 _____ [3]

2 (a) Complete the table below about the particles in an atom.

Particle	Relative charge	Relative mass	Location in an atom
proton	+1		the nucleus
electron		$\frac{1}{1840}$	orbits the nucleus
neutron	0	1	

[3]

(b) Below is the atomic number and mass number of sodium.



(i) How many protons does an atom of sodium have?

_____ [1]

(ii) Calculate the number of neutrons in an atom of sodium.

_____ [1]

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Marks Remark

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(Questions continue overleaf)

- 4 Look at the table below. It shows the colour of four indicators at different pH values.

Indicator	pH value													
	1	2	3	4	5	6	7	8	9	10	11	12	13	14
Universal	R	R	O	O	Y	Y	G	B	B	I	I	I	V	V
Methyl Red	R	R	R	R	Y	Y	Y	Y	Y	Y	Y	Y	Y	Y
Thymol Blue	Y	Y	Y	Y	Y	Y	Y	Y	Y	B	B	B	B	B
Alizarin Yellow	Y	Y	Y	Y	Y	Y	Y	Y	Y	Y	Y	R	R	R

Key

R=Red O=Orange Y=Yellow G=Green B=Blue I=Indigo V=Violet

Use the information above to answer the following questions.

- (a) (i) What colour is Methyl Red indicator in a solution of pH 7?

_____ [1]

- (ii) What colour is Alizarin Yellow indicator in strong alkali?

_____ [1]

- (iii) What colour is Universal indicator in hydrochloric acid?

_____ [1]

- (b) A scientist has some acid and is going to add an alkali to it. He needs to stop adding the alkali when the pH value is 7.

- (i) What name is given to the reaction of an acid with an alkali?

_____ [1]

- (ii) Choose the most suitable indicator from the table above for the scientist's experiment. Explain why you chose this indicator.

 _____ [2]

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Marks Remark

- 5 Below is the ingredients label found on an indigestion tablet. Each tablet has a mass of 1500 mg.

Each tablet contains:

Calcium carbonate (700 mg)
 Magnesium carbonate (80 mg)
 Sucrose (250 mg)
 Glucose (250 mg)
 Peppermint flavour (20 mg)
Talc
 Saccharin sodium (40 mg)
 Magnesium stearate (60 mg)

- (a) Calculate the mass of talc in each tablet.
 (Show your working out.)

_____ mg [2]

- (b) The active ingredient in this indigestion tablet is calcium carbonate.

- (i) What is the chemical formula for calcium carbonate?

_____ [1]

- (ii) Write down the name of the gas produced when calcium carbonate is added to an acid.

_____ [1]

- (c) Explain fully how an indigestion tablet works to cure indigestion.

 _____ [3]

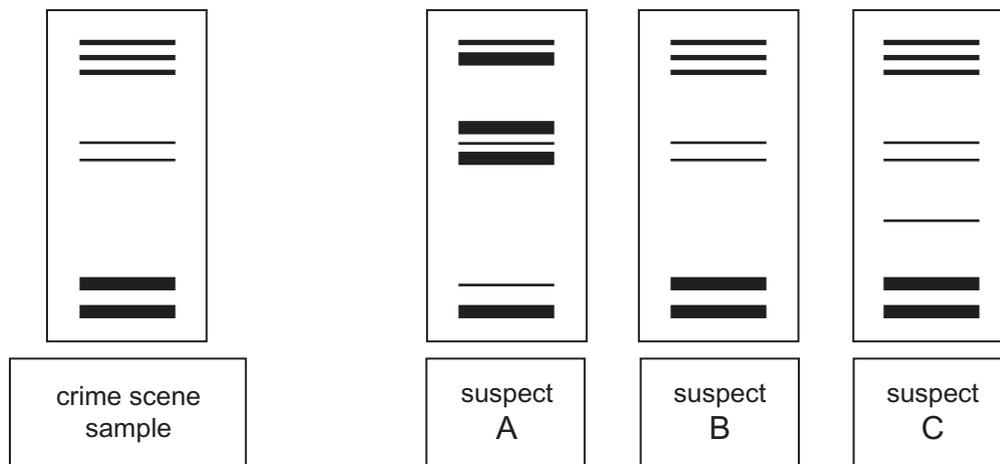
Examiner Only	
Marks	Remark

6 Genetic or DNA fingerprinting was discovered 30 years ago. It has really helped forensic science and crime solving.

(a) This technique only needs small amounts of DNA. Write down **two** pieces of evidence that could be collected at a crime scene that could be used for DNA analysis.

1. _____
2. _____ [2]

(b) Below are the genetic fingerprints taken from a crime scene and from three police suspects.



Which suspect does the crime scene sample belong to? Explain your answer.

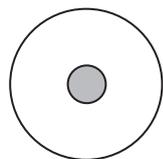
Suspect: _____

Explanation: _____

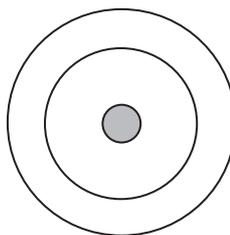
_____ [2]

Examiner Only	
Marks	Remark

- 9 (a) Complete the diagrams below to show how **all** of the electrons are arranged in an atom of hydrogen and an atom of oxygen.



hydrogen atom



oxygen atom

[2]

- (b) Explain fully how oxygen and hydrogen join to form a molecule of water. Explain this in terms of the atoms involved and their electrons. You should include the chemical formula for water in your answer.

In this question you will be assessed on your written communication skills including the use of specialist scientific terms.

[6]

Examiner Only	
Marks	Remark

10 Aluminium is extracted from aluminium oxide by electrolysis.

(a) Explain the meaning of the term **electrolysis**.

_____ [2]

(b) During the production of aluminium, cryolite is used. The chemical formula of cryolite is **Na₃AlF₆**.

(i) How many different elements does cryolite contain?

_____ [1]

(ii) Write down the name of the non-metal element in cryolite.

_____ [1]

(iii) What is the total number of atoms in Na₃AlF₆?

_____ [1]

The table below shows some information about aluminium and copper.

Metal	Relative electrical conductivity	Density g/cm ³	Relative strength	Cost per kg/£
aluminium	3.8	2.7	0.4	0.95
copper	5.7	8.9	0.6	3.10

(c) Aluminium is used in overhead electrical cables in the National Grid. The National Grid carries electricity throughout the country. Explain fully why aluminium is better than copper for this purpose. Use the information above to answer this question.

_____ [3]

Examiner Only

Marks Remark

11 Hydrocarbons are often used as fuels.

(a) Complete the table below about different hydrocarbons.

Hydrocarbon	Molecular formula	Structural formula
methane	CH_4	$\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \text{H} \\ \\ \text{H} \end{array}$
	C_2H_6	$\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H} - \text{C} - \text{C} - \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
propane	C_3H_8	
butane		$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{C} - \text{H} \\ \quad \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \end{array}$

[3]

(b) Propane is used as a fuel in camping stoves. Write a **balanced symbol** equation for the complete combustion of propane.

_____ [3]

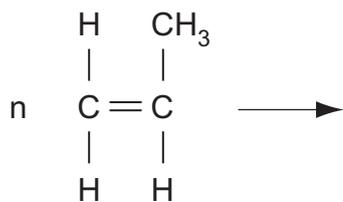
Examiner Only	
Marks	Remark

(c) Polypropene can be made by the polymerisation of propene.

(i) Explain the meaning of the term **polymerisation**.

_____ [2]

(ii) Complete the balanced symbol equation for the polymerisation of propene.



[3]

THIS IS THE END OF THE QUESTION PAPER

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Marks	Remark

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