

**Chemistry**
Standard level
Paper 2

Friday 13 November 2015 (afternoon)

Candidate session number

1 hour 15 minutes

--	--	--	--	--	--	--	--	--	--

Instructions to candidates

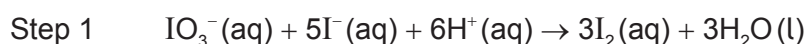
- Write your session number in the boxes above.
- Do not open this examination paper until instructed to do so.
- Section A: answer all questions.
- Section B: answer one question.
- Write your answers in the boxes provided.
- A calculator is required for this paper.
- A clean copy of the **chemistry data booklet** is required for this paper.
- The maximum mark for this examination paper is **[50 marks]**.



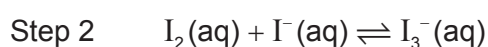
Section A

Answer **all** questions. Write your answers in the boxes provided.

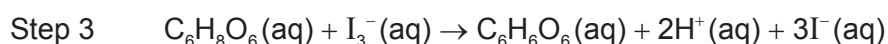
1. A student used the technique of titration to determine the concentration of ascorbic acid ($C_6H_8O_6$) in a sample of orange juice. Excess potassium iodide, $KI(aq)$, was added to acidified orange juice. The resulting solution was titrated with potassium iodate, $KIO_3(aq)$, in the presence of starch as an indicator. The end-point of the titration was shown by a blue-black colour.



Iodine is only slightly soluble in water; but in the presence of excess iodide ions, $I^-(aq)$, it forms the soluble tri-iodide ion, $I_3^-(aq)$.



Ascorbic acid reacts with tri-iodide ions as follows.



- (a) (i) Deduce the changes in oxidation number of iodine in step 1. [2]

<p>IO_3^- to I_2:</p> <p>.....</p> <p>I^- to I_2:</p> <p>.....</p>
--

- (ii) Identify the oxidizing **and** reducing agents in step 1. [1]

<p>Oxidizing agent:</p> <p>.....</p> <p>Reducing agent:</p> <p>.....</p>
--

(This question continues on the following page)



(Question 1 continued)

- (b) The concentration of KIO_3 used in the titration was $2.00 \times 10^{-3} \text{ mol dm}^{-3}$.
The titration produced the following results.

	Titration 1	Titration 2	Titration 3
Final volume of KIO_3 ($\pm 0.05 \text{ cm}^3$)	7.10	14.40	21.60
Initial volume of KIO_3 ($\pm 0.05 \text{ cm}^3$)	0.00	7.10	14.40
Volume added of KIO_3 ($\pm 0.10 \text{ cm}^3$)	7.10	7.30	7.20
Mean volume added of KIO_3 ($\pm 0.10 \text{ cm}^3$)	7.20		

- (i) Calculate the percentage uncertainty associated with the mean volume of $\text{KIO}_3(\text{aq})$.

[1]

.....

.....

.....

- (ii) The colour of orange juice interfered with the blue-black colour at the equivalence point. State the name of this type of error **and** suggest how this can be minimized.

[2]

.....

.....

.....

- (iii) Determine the amount, in mol, of $\text{KIO}_3(\text{aq})$, in the mean volume.

[1]

.....

.....

(This question continues on the following page)



(Question 1 continued)

- (c) Determine the amount, in mol, of ascorbic acid, $C_6H_8O_6(aq)$, in the sample of acidified orange juice. [2]

.....
.....
.....
.....

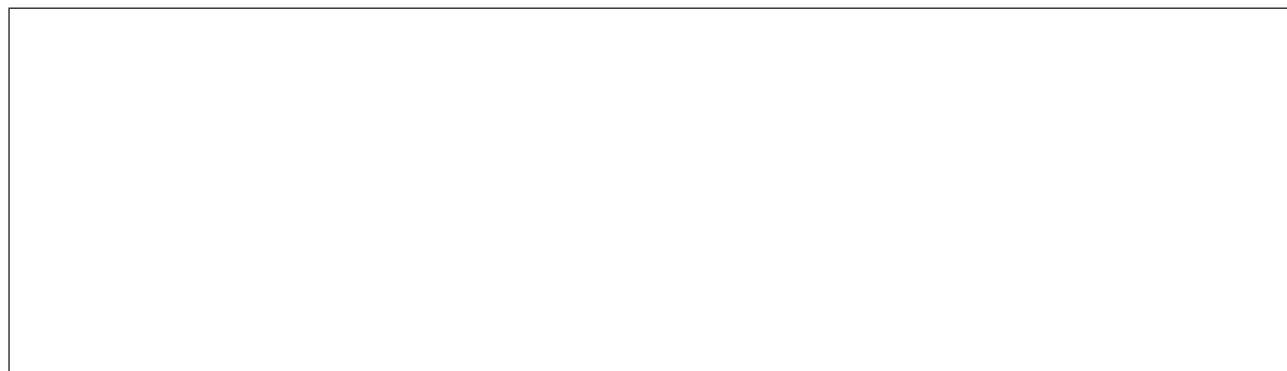
- (d) Calculate the mass, in g, of ascorbic acid, $C_6H_8O_6(aq)$, present in the sample of acidified orange juice. [1]

.....
.....



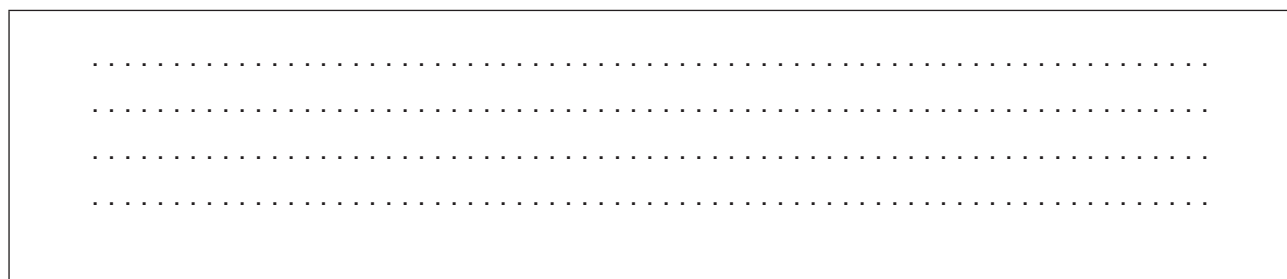
2. (a) Draw the Lewis (electron dot) structures of oxygen, O_2 , ozone, O_3 , and hydrogen peroxide, H_2O_2 .

[3]



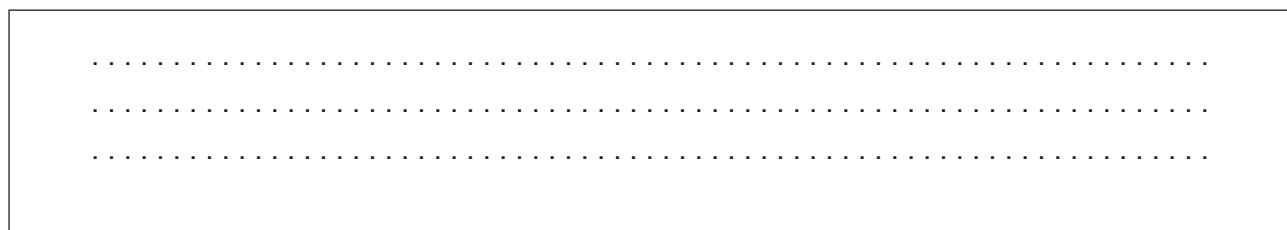
- (b) Deduce, giving a reason, the relative lengths of the oxygen to oxygen bonds in oxygen and hydrogen peroxide.

[1]



- (c) Predict, with a reason, the O–O–O bond angle in O_3 .

[2]



3. Propane, $C_3H_8(g)$, undergoes complete combustion to form carbon dioxide, $CO_2(g)$, and water, $H_2O(g)$.

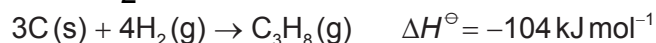
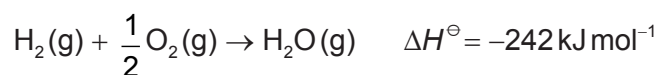
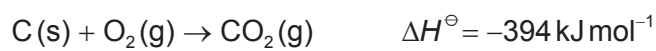
(a) State an equation for the complete combustion of propane, $C_3H_8(g)$. [1]

.....
.....

(b) Calculate the standard enthalpy change for the reaction in part (a) using bond enthalpy values given in table 10 of the data booklet. [3]

.....
.....
.....
.....
.....
.....
.....
.....

(c) Determine, using Hess's law, the enthalpy change, ΔH^\ominus , in kJ mol^{-1} , for the complete combustion of propane by using the following data. [4]



.....
.....
.....
.....
.....
.....

(This question continues on the following page)



(Question 3 continued)

(d) Suggest, with a reason, why the values obtained in parts (b) and (c) are different. [1]

.....
.....

4. (a) (i) Define the term electronegativity. [1]

.....
.....

(ii) Suggest why the noble gases are generally not assigned electronegativity values. [1]

.....
.....

(b) Explain why the melting points of the group 1 metals (Li → Cs) decrease down the group whereas the melting points of the group 7 elements (F → I) increase down the group. [3]

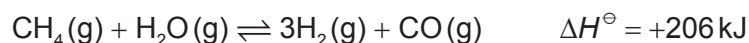
.....
.....
.....
.....
.....
.....
.....
.....



Section B

Answer **one** question. Write your answers in the boxes provided.

5. (a) The following reaction is used in industry to obtain hydrogen from natural gas by partial oxidation with steam.



- (i) Describe the effect, if any, of each of the following changes on the equilibrium amount of hydrogen, giving a reason in each case. [4]

Increasing the pressure, at constant temperature:

.....

Increasing the temperature, at constant pressure:

.....

- (ii) Identify which of the changes in part (a) (i) will affect the value of K_c and whether the value will increase or decrease. [1]

.....

- (iii) Discuss the effects of adding a solid catalyst to the mixture of methane and steam, at constant pressure and temperature. [3]

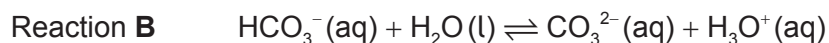
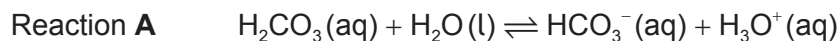
.....

(This question continues on the following page)



(Question 5 continued)

(b) The equations of two acid-base reactions are given below.



(i) Explain whether $\text{HCO}_3^-(\text{aq})$ behaves as an acid or a base in each of the reactions **A** and **B**. [2]

<p>Reaction A:</p> <p>.....</p> <p>.....</p> <p>Reaction B:</p> <p>.....</p> <p>.....</p>

(ii) Deduce **two** conjugate acid-base pairs from reactions **A** and **B**. [2]

	Acid	Base
Conjugate acid-base pair 1
Conjugate acid-base pair 2

(This question continues on the following page)



(Question 5 continued)

(c) Nitric acid, HNO_3 , and nitrous acid, HNO_2 , are described as strong and weak acids respectively.

(i) Distinguish between *strong* and *weak* acids.

[1]

.....
.....
.....

(ii) A 1.00 g sample of solid magnesium carbonate, MgCO_3 , is added to separate solutions of HNO_3 and HNO_2 of the same concentration and temperature. State **one** similarity and **one** difference in the observations made in these reactions.

[2]

Similarity:
.....
.....

Difference:
.....
.....

(iii) A solution of HNO_3 has a pH of 1, while a solution of HNO_2 has a pH of 5. Determine the ratio of the hydrogen ion concentration in HNO_3 : HNO_2 .

[1]

.....
.....

(This question continues on the following page)



(Question 5 continued)

- (d) (i) State the acid-base character of the oxides of the period 3 elements Na to Ar. [2]

.....
.....
.....
.....
.....
.....

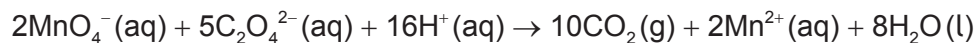
- (ii) State balanced equations to illustrate the acid-base character of sodium oxide and sulfur trioxide. [2]

Sodium oxide:
.....
.....

Sulfur trioxide:
.....
.....



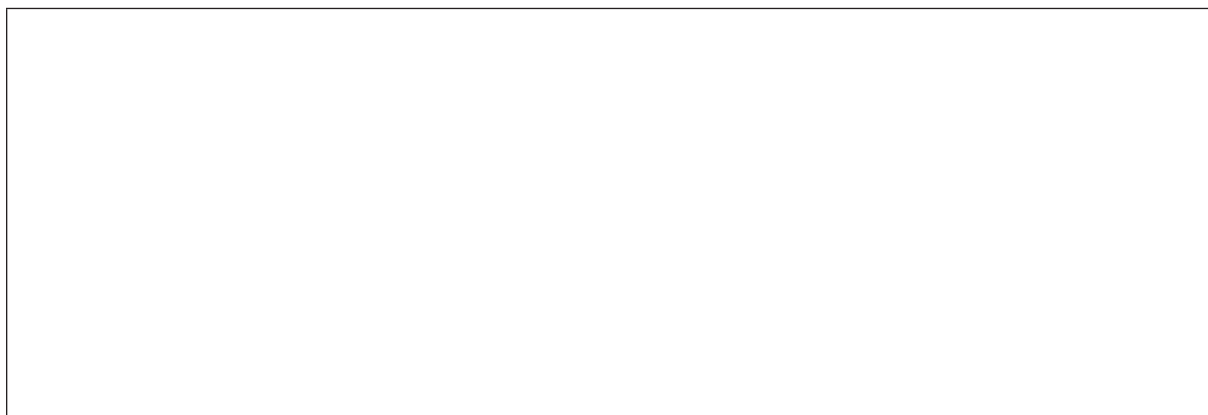
6. (a) A purple solution of potassium manganate(VII), KMnO_4 , reacts with ethanedioate ions according to the following equation.



- (i) Outline an experimental procedure which may be used to measure the rate of this reaction. [3]

.....
.....
.....
.....
.....
.....
.....
.....

- (ii) Sketch a graph to show the results of the experimental procedure outlined in (a) (i). [2]



- (iii) Outline how the rate of reaction at a particular time could be determined from the graph. [1]

.....
.....

(This question continues on the following page)



(Question 6 continued)

- (iv) Discuss, in terms of collision theory, the effect of increasing temperature on the rate of this reaction. [3]

.....

.....

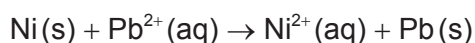
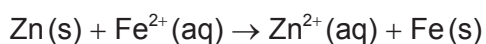
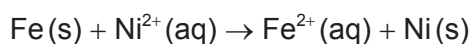
.....

.....

.....

.....

- (b) Consider the following spontaneous reactions.



- (i) Deduce the order of **increasing** reactivity of the metals based on the reactions above. [2]

.....

.....

.....

.....

- (ii) Identify the strongest oxidizing agent in the reactions above. [1]

.....

(This question continues on the following page)



(Question 6 continued)

- (c) Draw a diagram of an electrolytic cell for the electrolysis of molten nickel(II) bromide, $\text{NiBr}_2(\text{l})$. Include the direction of the electron flow, the polarity of electrodes and the half-equations at each electrode. [4]

Negative electrode (cathode):
.....

Positive electrode (anode):
.....

- (d) (i) In the operation of a mass spectrometer, the first stage is vaporization and the last is detection. State the names of the other three stages and outline what happens in each one. [3]

.....
.....
.....
.....
.....
.....
.....
.....

- (ii) Suggest why a very low pressure is maintained inside the mass spectrometer. [1]

.....
.....



7. (a) A 0.842g sample of a liquid halogenoalkane, RBr(l), was heated under reflux with 1.35×10^{-2} mol of aqueous sodium hydroxide, NaOH(aq). After cooling the mixture, the excess NaOH was titrated with hydrochloric acid, HCl(aq), and required 7.36×10^{-3} mol of the acid.

- (i) State the equation for the substitution reaction of the halogenoalkane with sodium hydroxide. [1]

.....
.....

- (ii) Calculate the amount, in mol, of sodium hydroxide that reacted with the halogenoalkane. [1]

.....
.....
.....

- (iii) Calculate the molar mass of the halogenoalkane. [1]

.....
.....
.....
.....

- (iv) Given that each molecule of the halogenoalkane contains one bromine atom, determine its molecular formula. [1]

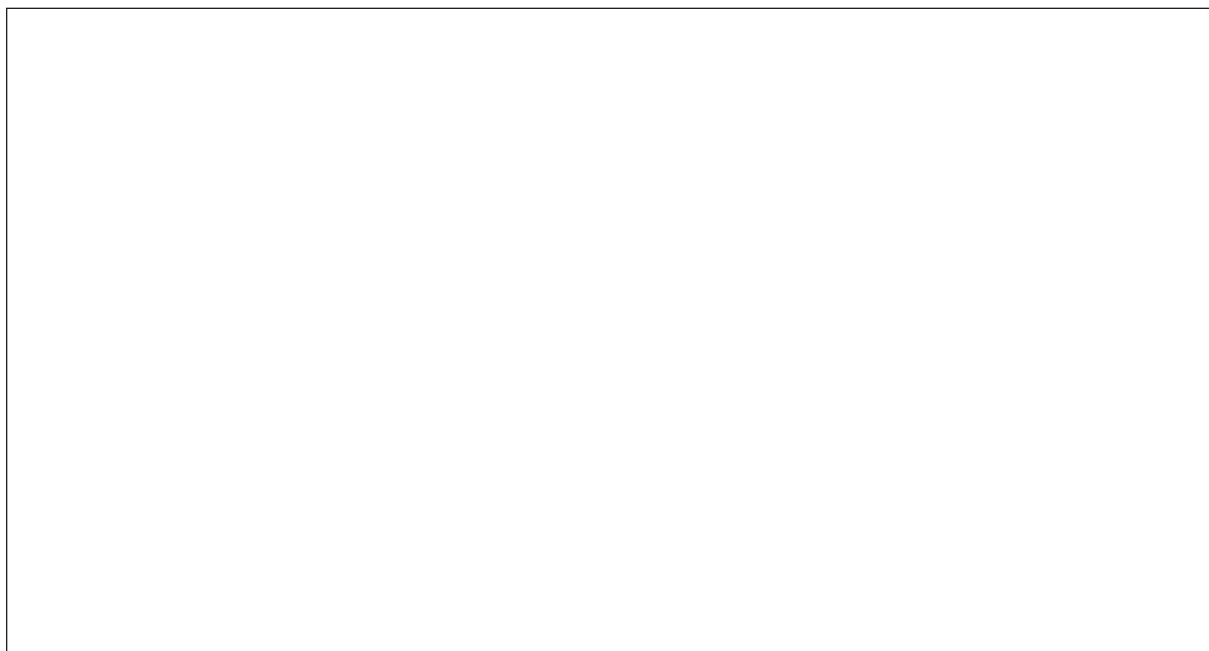
.....
.....
.....

(This question continues on the following page)



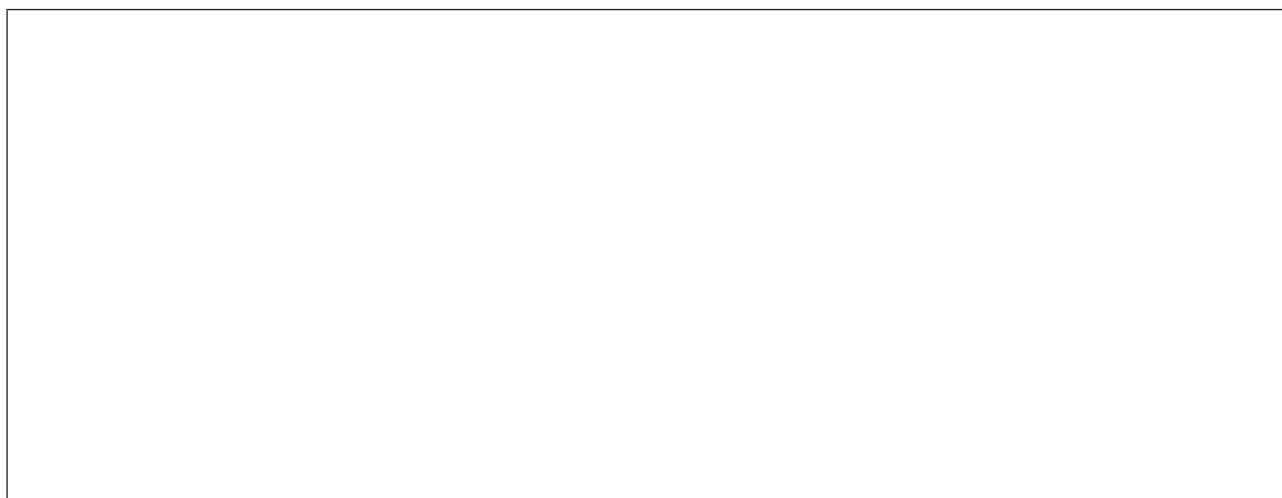
(Question 7 continued)

- (v) Deduce the structural formulas of **four** structural isomers of the halogenoalkane based on the molecular formula **and** label each isomer as primary, secondary or tertiary. (If you have not been able to determine the molecular formula in part (a) (iv), use $C_5H_{11}Br$ to deduce the four structural isomers.) [4]



- (b) The reaction between a primary halogenoalkane drawn in part (a) (v) and sodium hydroxide follows a S_N2 mechanism.

Explain the mechanism of the reaction using curly arrows to represent the movement of electron pairs. [4]



(This question continues on the following page)



(Question 7 continued)

- (c) (i) List the following compounds in order of **increasing** boiling point:
 CH_3CHO , $\text{CH}_3\text{CH}_2\text{CH}_3$, CH_3COOH , $\text{CH}_3\text{CH}_2\text{OH}$.

[2]

.....
.....

- (ii) Explain the order of boiling points in the compounds listed in part (c) (i), in terms of intermolecular forces.

[4]

.....
.....
.....
.....
.....
.....
.....
.....
.....
.....

- (iii) $\text{CH}_3\text{CH}_2\text{OH}$ can be oxidized in aqueous solution by acidified potassium dichromate(VI) by heating under reflux. Deduce the structural formula of the final product formed **and** state the colour change during the oxidation process.

[2]

.....
.....
.....
.....



Please **do not** write on this page.

Answers written on this page
will not be marked.



Please **do not** write on this page.

Answers written on this page
will not be marked.



Please **do not** write on this page.

Answers written on this page
will not be marked.

