



GCE

Chemistry B (Salters)

Unit **H033/02**: Chemistry in depth

Advanced Subsidiary GCE

Mark Scheme for June 2016

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All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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Annotations

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore

Abbreviations, annotations and conventions

Annotation	Meaning
/	alternative and acceptable answers for the same marking point
✓	Separates marking points
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

Question		Answer	Marks	Guidance
1	(a)	$C_4H_{10}(g/l) + 6\frac{1}{2}O_2(g) \rightarrow 4CO_2(g) + 5H_2O(l)$ ✓ for balanced equation ✓ for state symbols	2	NOT multiples
	(b)	FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = $\Delta_c H = -1419$ (kJ mol⁻¹) award 3 marks Calculates energy transferred to water $q = 50.00 \times 4.18 \times (74 - 19) = 11495$ (J) ✓ Calculates number of moles of butane burned $= 0.47 / 58.0$ or $0.008(1)$ mol ✓ $(\Delta_c H = -[11495 / 0.008(1)] = -1419136$ J mol ⁻¹) $\Delta_c H = -1419$ (kJ mol ⁻¹) ✓	3	Correct answer -1419 ±1 kJ mol⁻¹ scores 3 marks ALLOW ECF between steps ALLOW final answer to 2 or more sf (eg -1440kJ mol ⁻¹ if early rounding is evident) Final MP must include negative sign
	(c) (i)	-2850 - -2950 (kJ mol ⁻¹) ✓	1	Must have negative sign
	(ii)	Any two from: ✓ ✓ loss of fuel by evaporation / escape of unburned butane evaporation of water incomplete combustion / reaction non-standard conditions /states heat used to raise temp of calorimeter	2	Answers can be in any order Ignore 'not fully reacted' as this makes it unclear whether the candidate is talking about the vol of butane or the combustion reaction Ignore measurement errors
	(iii)	One from: ✓ use a (draught) shield because this will reduce heat lost (to the surroundings); burn the butane in oxygen / because this will ensure that the combustion is more complete; use bomb calorimeter ensures complete combustion / reduces heat loss; use cover over lighter during weighing to prevent evaporation; insulate can to reduce heat loss	1	Must have method plus explanation to score Ignore changes to vol of water / mass of fuel / length of time for combustion / move flame nearer to can / ALLOW 'put lid on can' NOT 'use a polystyrene cup'

Question	Answer	Marks	Guidance
(d)	<p>Any one from:</p> <p>reacting carbon and hydrogen doesn't (necessarily) make butane ✓</p> <p>Carbon and Hydrogen do not react together (under standard conditions) ✓</p>	1	<p>ALLOW reference to formation of a mixture of products or alternative product(s) / side reactions</p>
(e)	<p style="text-align: center;">  skeletal formula </p> <p>systematic name: (2-)methylpropane ✓</p>	1	<p>IGNORE dashes, commas and spaces in the name</p> <p>Needs BOTH skeletal formula AND name</p>
(f)	<p>FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = (+)486.6 / (+)487 (kJ mol⁻¹) award 3 marks</p> <p>energy absorbed in breaking bonds = 3(413) + (358) + x + 1½(498) = 2344 + x (kJ)</p> <p>AND</p> <p>energy evolved in making bonds = 2(805) + 4(x) = 1610 + 4x (kJ) ✓</p> <p>Overall energy change (Bonds broken – bonds made = ΔH) = [2344 + x] – [1610 + 4x] = -726 kJ mol⁻¹</p> <p>OR</p> <p>2344 - 1610 + 726 = 3x 1460 = 3x ✓</p> <p>x = (+)486.6 / (+)487 (kJ mol⁻¹) ✓</p>	3	<p>Correct answer +486.6 / +487 kJ mol⁻¹ scores 3 marks</p> <p>ALLOW ECF between steps</p> <p>2344 and 1610 in calculation scores 1 mark if no other mark scored</p> <p>ALLOW OH for x in calculation</p> <p>ALLOW –sign if evaluation of their expression for x is correct</p>
	Total	14	

Question		Answer	Marks	Guidance
2	(a)	8 - 11 ✓	1	Accept any value between 7.1 - 14
	(b)	<p>FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = 0.34 (mol) (2 sf) award 3 marks</p> <p>mass of Mg(OH)_2 in $250 \text{ cm}^3 = 8/100 \times 250 (= 20) \text{ g}$ ✓</p> <p>$M_r \text{ Mg(OH)}_2 = 58.3 \text{ g mol}^{-1}$ ✓</p> <p>Moles $\text{Mg(OH)}_2 = (20/58.3) = 0.34 \text{ (mol) (2 sf)}$ ✓</p>	3	<p>Allow ecf throughout</p> <p>Final answer MUST be to 2sf</p>
	(c)	(i)	1	<p>incorrect number is CON ALLOW 6/+6/6+ ALLOW oxidation/</p>
	(c)	(ii)	1	<p>ALLOW cloudiness of suspension obscures colour of indicator/ makes it difficult to identify end-point ALLOW medicine for Mg(OH)_2</p>
	(d)	(i)	1	<p>ALLOW only used titres agreeing to within 0.1 cm^3 / repeat 2 not included as it is an anomalous result(outlier)</p>
	(d)	(ii)	2	<p>Allow ecf from incorrect titre used in calculation</p> <p>Do not accept 0.0165, incorrect rounding of 0.01655 Final answer to 2sf or more.</p>

Question		Answer	Marks	Guidance
(d)	(iii)	<p>FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = 19.20 - 19.55 (g) award 3 marks amount of H₂SO₄ initially = (25.0/1000 x 2.00) = 0.05 mol ✓</p> <p>amount of H₂SO₄ used up = number of moles Mg(OH)₂ = (0.05 – 0.0166) = 0.0334 mol ✓</p> <p>mass of Mg(OH)₂ in 250 cm³ = (250/25 x 0.0334 x 58.3) = 19.47 / 19.5 (g) ✓</p>	3	<p>ALLOW ecf from incorrect value in d(ii)</p> <p>ALLOW ecf throughout</p> <p>ALLOW 3 or more sf throughout, but rounding must be correct</p>
(d)	(iv)	(0.06/25.0 x 100) = 0.2(%) ✓	1	Correct answer without working scores the mark
(e)	(i)	Mg → Mg ²⁺ + 2e ⁻ / Mg – 2e ⁻ → Mg ²⁺ ✓	1	Ignore state symbols
(e)	(ii)	(magnesium) loses (two) electrons ✓	1	<p>ACCEPT the oxidation state (of the magnesium) increases (from 0 to +2)</p> <p>ignore species losing electrons unless incorrectly named</p>
(e)	(iii)	H ⁺ /hydrogen (ion) ✓	1	<p>ACCEPT hydrochloric acid/HCl</p> <p>NOT H / H₂</p>
(f)	(i)	<p>Mg²⁺(aq) + 2OH⁻(aq) → Mg(OH)₂(s)</p> <p>✓ for balanced ionic equation</p> <p>✓ for state symbols</p>	2	<p>DO NOT ALLOW spectator ions</p> <p>ALLOW state symbol mark if 'magnesium hydroxide' given as solid and all other species as aq</p>
	(ii)	Ba(OH) ₂ is <u>more</u> soluble (in water) ✓	1	<p>ORA</p> <p>ALLOW Ba(OH)₂ will not precipitate</p>
Total			19	

Question			Answer	Marks	Guidance
3	(a)	(i)	CH ₃ COOH ✓	1	ALLOW any unambiguous structure
		(ii)	ester ✓	1	
		(iii)	FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = 75 (%) award 2 marks Mr of aspirin $= [(12.0 \times 9) + (16.0 \times 4) + (1.0 \times 8)] = 180.0$ ✓ % atom economy $= 180.0 / (138.0 + 102.0) \times 100$ OR $180 / (180 + 60) \times 100$ $= 75 (\%)$ ✓	2	ALLOW ecf from incorrect value for Mr of aspirin
	(b)		2-hydroxybenzoic acid ✓	1	ALLOW salicylic acid / phenol <u>group</u> (present) IGNORE dashes, commas and spaces NOT phenol on its own
	(c)		<u>dissolve</u> (crude) aspirin/solid in hot/warm ethanol/solvent ✓ use the minimum volume/amount of ethanol/solvent ✓ (allow to) <u>cool/crystallise</u> ✓ <u>filter</u> , <u>wash</u> (with cold ethanol) and (allow to) <u>dry</u> ✓	4	'Dissolve in a minimum amount of hot ethanol' scores 2 marks (MP1 and MP2) ALLOW crystals to form
	(d)		the range (it) will be wider ✓	1	ALLOW (it will be) lower ALLOW recrystallized product will be higher/narrower range
	(e)		FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = 45 - 46 (%) award 2 marks (138.0 g 2-hydroxybenzoic acid → 180.0 g aspirin) 1.15 g 2-hydroxybenzoic acid → $(1.15 / 138.0 \times 180.0)$ $= 1.50$ g aspirin ✓ % yield = $(0.68 / 1.50 \times 100) = 45(.3) (\%)$ ✓	2	Calculates number of moles as $1.15/138 = 0.00833$ mol and $0.68/180 = 0.00378$ mol (1) Allow ecf from incorrect value of Mr from 3a(iii) % yield = $0.00378/0.00833 \times 100 = 45.3\%$ ALLOW 2 or more sf

Question	Answer	Marks	Guidance
(f)	<p>LOR <i>Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question</i></p> <p>Level 3 (5 - 6 marks) Gives a detailed description (including some relevant fine detail) of ALL three phases</p> <p><i>The descriptions are well-developed, clear and logically structured.</i></p> <p>Level 2 (3 - 4 marks) Gives a basic description of all three phases of the process OR Describes two of the phases with one in some detail</p> <p><i>The method and analysis/further action is clear with some structure. The running is workable and in an acceptable order.</i></p> <p>Level 1 (1 - 2 marks) Gives a description of one of the phases</p> <p><i>Response shows some structure.</i></p> <p>Level 0 Insufficient or irrelevant science.</p>	6	<p>Indicative scientific points may include:</p> <p>Phase 1 Running the chromatography</p> <ul style="list-style-type: none"> • place plate in beaker with solvent • Allow solvent to rise through spots • remove plate • dry plate <p>Phase 2 Analysis of chromatogram</p> <ul style="list-style-type: none"> • crude product/CP contains (both aspirin and unreacted) 2-hydroxybenzoic acid • recrystallized/RP product contains some (unreacted) 2-hydroxybenzoic acid <p>Phase 3 Further action</p> <ul style="list-style-type: none"> • further purification is required <p>At Level 3 the fine detail may include</p> <ul style="list-style-type: none"> • solvent below line of dots • cover beaker with a lid • produces a saturated atmosphere • (remove plate) when solvent front near top • transfer to fume cupboard to evaporate solvent • difference in intensity of spots linked to quantity of unreacted 2-hydroxybenzoic acid present • further recrystallisation is required • Repeat chromatography after further recrystallization

Question		Answer	Marks	Guidance
	(g)	No (observable) reaction with paracetamol ✓ Effervescence/fizzing/bubbling/gas with aspirin ✓ Aspirin contains a carboxyl/carboxylic acid functional group (ORA) ✓	3	ALLOW 'nothing happens' for 'no reaction' ALLOW 'dissolve' for aspirin and 'does not dissolve' for paracetamol ALLOW CO ₂ gas but any other named gas is CON
		Total	21	

Question		Answer	Marks	Guidance
4	(a) (i)	Reagents acidified (potassium) dichromate(VI) AND Conditions reflux ✓	1	ALLOW (potassium) dichromate in (sulfuric) acid (VI) in 'dichromate(VI)' is not required but must be correct if included ACCEPT $H^+/Cr_2O_7^{2-}$
	(a) (ii)	$C_2H_5OH + 2[O] \rightarrow CH_3COOH + H_2O$ ✓	1	ALLOW CH_3CH_2OH for ethanol / $2[O]$ over the arrow DO NOT ALLOW C_2H_6O for ethanol or $C_2H_4O_2$ for ethanoic acid as question asks for structural formulae Displayed or skeletal formulae are also both acceptable
	(b)	$C_2H_5OH + 2O_2 \rightarrow 2CO + 3H_2O$ ✓	1	ALLOW either CH_3CH_2OH or C_2H_6O for ethanol ALLOW $C_2H_5OH + O_2 \rightarrow 2C + 3H_2O$ OR $C_2H_5OH + 2.5O_2 \rightarrow CO_2 + CO + 3H_2O$ (or doubled) OR any other balanced equation that includes C and/or CO as a product Ignore state symbols
	(c)	ethanol and ethanoic acid – both hydrogen-bonds ✓ ethanoic acid has stronger/more H-bonds / id – id ✓ H-bonds stronger than pd-pd / H-bonds are the strongest and stronger bonds take more energy to break (ORA) ✓ Ethanal – permanent dipole - permanent dipole/pd - pd ✓	4	ALLOW pd : pd / Van der Waal forces as an alternative to id : id ALLOW Permanent dipole – dipole / dipole – permanent dipole / permanent – permanent dipole
	(d) (i)	FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = 750 (cm³) award 2 marks moles of $CH_3CHO = (0.55/44) = 0.0125$ mol 0.0125 mol CH_3CHO requires ($2\frac{1}{2} \times 0.0125$) = 0.03125 mol O_2 ✓ volume of $O_2 = (0.03125 \times 24000) = 750$ (cm ³) ✓	2	ALLOW 2 or more sf. Throughout Allow ecf ALLOW vol of $O_2 = 300$ (cm ³) from correct calculation of Moles $CH_3CHO \times 24000$ ie 0.0125×24000 for 1 mark
	(d) (ii)	$Mr\ CO_2 / (Mr\ CO_2 + Mr\ H_2O) \times 100$ $44 / (44 + 18) \times 100 = 71\%$ ✓	1	Correct answer = 71% without working scores Allow 2 or more sf

Question	Answer	Marks	Guidance
(e)	<p><i>Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question</i></p> <p>Level 3 (5 - 6 marks) Calculates both the empirical formula from the % composition and the molecular formula using MS. AND Uses IR spectrum to identify C=O bond and one other bond (or lack of bond) present in the structure. AND Draws correct detailed conclusion for ester formula/structure from above and MS fragment data</p> <p><i>The conclusion relates to the evidence and is clear and logically structured.</i></p> <p>Level 2 (3 - 4 marks) Concludes A is an ester supported by evidence from molecular formula/Mr and some IR data OR Concludes it is ethanal supported by empirical formula and appropriate evidence using IR / MS data</p> <p><i>The conclusion relates to the limited evidence and is clear and logically structured.</i></p> <p>Level 1 (1 - 2 marks) States empirical formula or Mr of Compound A using evidence from MS / % composition OR suggests it is an ester/aldehyde from IR evidence alone</p> <p><i>Pieces of evidence given are related in some way</i></p> <p>Level 0 Insufficient or irrelevant science</p>	6	<p>Indicative scientific points may include:</p> <p>Formula/Mass Spec evidence</p> <ul style="list-style-type: none"> empirical formula C_2H_4O with calculation from % data, $C = (54.5/12.0) = 4.54$, $H = (9.1/1.0) = 9.1$, $O = (36.4/16.0) = 2.275$, $C = 2(1.99)$, $H = 4$, $O = 1$ Mr = 88 identified from molecular ion peak in mass spectrum molecular formula = empirical formula (mass) x 2 = $C_4H_8O_2$ <p>Infra-red evidence</p> <ul style="list-style-type: none"> C=O bond in ester (aldehyde/ester) present, absorption is 1740 OR in range 1720-1740 (cm^{-1}) O-H bond in carboxylic acid not present, no (broad) absorption in range 2500-3300 (cm^{-1}) C-O bond present as absorption in range 1250 – 1300 (cm^{-1}) <p>Conclusion</p> <ul style="list-style-type: none"> ester $CH_3COOC_2H_5$ (structure or name, ethyl ethanoate) because of fragment(s) identified in mass spectrum
	Total	16	

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