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A322/01

GENERAL CERTIFICATE OF SECONDARY EDUCATION

TWENTY FIRST CENTURY SCIENCE

CHEMISTRY A

Unit 2 Modules C4 C5 C6 (Foundation Tier)

FRIDAY 25 JANUARY 2008

Morning

Time: 40 minutes

* C U P / T 4 3 5 4 4 *

Candidates answer on the question paper.

Additional materials (enclosed):

None

Calculators may be used.

Additional materials: Pencil
Ruler (cm/mm)

Candidate
Forename
Candidate
Surname
Centre
Number

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Candidate
Number

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INSTRUCTIONS TO CANDIDATES

- Write your name in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use blue or black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **all** the questions.
- Do **not** write in the bar codes.
- Do **not** write outside the box bordering each page.
- Write your answer to each question in the space provided.

FOR EXAMINER'S USE		
Qu.	Max.	Mark
1	6	
2	9	
3	5	
4	5	
5	5	
6	5	
7	7	
TOTAL	42	

INFORMATION FOR CANDIDATES

- The number of marks for each question is given in brackets [] at the end of each question or part question.
- The total number of marks for this paper is **42**.
- The Periodic Table is printed on the back page.

This document consists of **16** printed pages.

Answer **all** the questions.

1 Elements in Group 7 are called the halogens. The table gives some information about the physical properties of three of the halogens.

halogen	proton number	melting point in °C	boiling point in °C	state at 25 °C	colour
chlorine	17	-101	-35		pale green
bromine	35	-7	59	liquid	deep red
iodine	53	114	184	solid	dark grey

(a) The halogens show trends in physical properties with increasing proton number.

Use information from the table to help you answer these questions.

(i) Finish the sentence about the trend in melting point.

As the proton number the melting point [1]

(ii) What is the state of chlorine at 25 °C?

Put a **ring** around the correct answer.

gas liquid solid

[1]

(iii) **Astatine** is a halogen. The proton number of astatine is 85.

The halogens get darker in colour as the proton number increases.

Predict the colour of **astatine**.

Put a **ring** around the correct answer.

yellow orange black

[1]

(b) The halogens also show a trend in reactivity.

This can be shown by the displacement reactions when halogens are added to solutions of halides.

A student made the following observations.

- When chlorine is added to potassium bromide solution, red bromine appears.
- When bromine is added to potassium iodide solution, brown iodine appears.
- When bromine is added to potassium chloride solution, there is no displacement.

(i) What is the correct order of reactivity for these halogens?

Put a tick (✓) in the box next to the correct answer.

decreasing reactivity

→ bromine chlorine iodine

chlorine bromine iodine

iodine bromine chlorine

[1]

(ii) Fluorine is a halogen with proton number 9.

Which statement describes the displacement reactions of fluorine?

Put a tick (✓) in the box next to the correct answer.

Fluorine displaces chlorine, bromine and iodine.

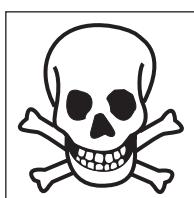
Fluorine displaces iodine but not chlorine or bromine.

Fluorine displaces chlorine and bromine but not iodine.

Fluorine displaces bromine and iodine but not chlorine.

[1]

(c) Hazard symbols are used to show the dangers involved in handling some chemicals.



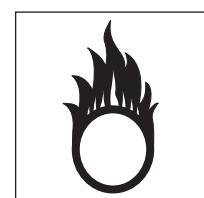
A



B



C



D

Iodine is **harmful**. Which hazard symbol, **A**, **B**, **C** or **D**, should be placed on a container of iodine?

answer [1]

[Total: 6]

2 This diagram shows part of the Periodic Table.

					He
Li	Be		C		Ne
Na	Mg				Cl
K	Ca				Br

(a) (i) Which three elements shown in the diagram are in the same **group**?

Put a tick (✓) in the box next to the correct answer.

Be	C	He	<input type="checkbox"/>
Na	Mg	Ar	<input type="checkbox"/>
He	Be	Ar	<input type="checkbox"/>
Li	Na	K	<input type="checkbox"/>

[1]

(ii) Which three elements shown in the diagram are in the same **period**?

Put a tick (✓) in the box next to the correct answer.

Na	Mg	Ar	<input type="checkbox"/>
Li	Na	Ca	<input type="checkbox"/>
He	Ne	Ar	<input type="checkbox"/>
Na	Ca	Br	<input type="checkbox"/>

[1]

(b) One element in the diagram has the symbol Be. This element is in Group 2.

Write down the following information about the element with the symbol **Be**.

Use the Periodic Table on the back page to help you.

name

proton number

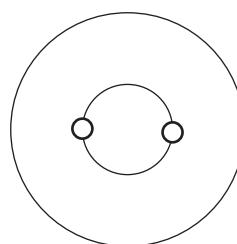
relative atomic mass

[3]

(c) Finish the diagram to show the arrangement of electrons in an atom of the element **carbon**.

Use a circle \circ to show the position of each electron.

The positions of two electrons have already been drawn to help you.



[1]

(d) The elements sodium and chlorine react to form the compound sodium chloride.

(i) Sodium chloride is made of sodium ions, Na^+ , and chloride ions, Cl^- .

Which statement describes evidence that sodium chloride is made of ions?

Put a tick (\checkmark) in the box next to the correct answer.

Sodium chloride is a solid.

Sodium chloride is made of crystals.

Sodium chloride has a high melting point.

Molten sodium chloride conducts electricity.

[1]

(ii) The table shows the arrangement of electrons in sodium atoms and chlorine atoms.

Complete the table to show the arrangement of electrons in sodium ions and chloride ions.

sodium atom Na	sodium ion Na^+	chlorine atom Cl	chloride ion Cl^-
2.8.1		2.8.7	

[2]

[Total: 9]

3 The table gives information about ions dissolved in sea water.

ion	symbol	percentage by mass of the total dissolved solids (%)
chloride	Cl^-	55
sodium	Na^+	30
sulfate	SO_4^{2-}	8
magnesium	Mg^{2+}	4
calcium	Ca^{2+}	1
potassium	K^+	1
carbonate	CO_3^{2-}	0.5
bromide	Br^-	0.2

These ions enter the sea water when crystals of ionic compounds in rocks dissolve.

Each of these ionic compounds is made up of one type of positive ion and one type of negative ion shown in the table.

(a) One compound that dissolved from the rocks into the water is magnesium sulfate.

Suggest the name of one **other** ionic compound that dissolved from the rocks into the water.

Use information from the table to help you.

..... [1]

(b) What holds together the ions in the crystals of ionic compounds?

Put a tick (✓) in the box next to the correct answer.

sharing of pairs of electrons

attraction between ions of opposite charge

attraction between ions of the same charge

[1]

(c) Sea water conducts electricity.

Which **two** statements give the best explanation for this?

Put ticks (✓) in the boxes next to the **two** correct answers.

Ions are able to move around in the sea water.

Electrons can pass from ion to ion in the sea water.

The sea water contains more ions with positive charges than ions with negative charges.

The sea water contains ions that have positive charges and ions that have negative charges.

[2]

(d) When a sample of sea water is evaporated to dryness, a white solid is left. This is a mixture of several ionic compounds.

Look at the **percentage by mass of the total dissolved solids** column in the table.

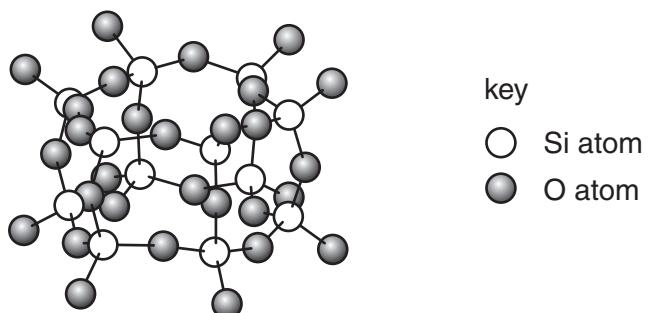
Use the information to name the ionic compound that makes up **most** of the white solid.

..... [1]

[Total: 5]

4 All of the materials in the world are made up of elements.

(a) Much of the elements oxygen and silicon are in the compound silicon dioxide.



Here are some sentences about **silicon dioxide**.

Finish these sentences by putting a tick (✓) in the box next to the correct word in each pair.

Silicon dioxide is a giant

ionic	
covalent	

 structure.

It is very

soft	
hard	

 and has

low	
high	

 melting and boiling points.

Silicon dioxide is

soluble	
insoluble	

 in water and

does	
does not	

 conduct electricity.

[4]

(b) The compounds in living organisms are made **mainly** of four elements. Two of these elements are **carbon** and **hydrogen**.

Which are the other **two** elements?

Put a **ring** around each of the **two** correct answers.

calcium

nitrogen

oxygen

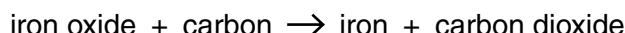
sodium

[1]

[Total: 5]

10

5 The ore haematite contains iron oxide. Iron is extracted from this ore by reaction with carbon.



(a) Oxidation **and** reduction take place in this reaction.

What is the name of the substance that has been **oxidised**?

..... [1]

(b) Which **two** of the metals below are also extracted by reaction with carbon?

Put a **(ring)** around each of the **two** correct answers.

copper

potassium

sodium

zinc

[2]

(c) The ore bauxite contains aluminium oxide.

Carbon is **not** used to extract aluminium from this ore.

(i) Why is carbon **not** used to extract aluminium from bauxite?

Put a tick (**✓**) in the box next to the correct answer.

The reaction is too fast.

Aluminium is more reactive than carbon.

Carbon is more reactive than aluminium.

The reaction would use too much carbon.

[1]

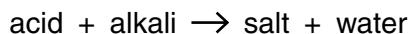
(ii) Suggest **one** method that can be used to extract large quantities of aluminium from bauxite.

..... [1]

[Total: 5]

11

6 An acid and an alkali react to form a salt and water.



(a) What type of reaction is this?

Put a (ring) around the correct answer.

decomposition

neutralisation

oxidation

polymerisation

[1]

(b) Draw straight lines to join up the boxes to show which **acid** reacts with which **alkali** to make each **salt**.

acid

alkali

salt

sulfuric acid

potassium hydroxide

sodium sulfate

hydrochloric acid

ammonium hydroxide

potassium chloride

nitric acid

sodium hydroxide

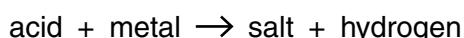
magnesium nitrate

phosphoric acid

magnesium hydroxide

[3]

(c) Acids also react with metals.



A piece of zinc is added to 20 cm³ of dilute hydrochloric acid.

Bubbles of hydrogen gas appear.

What is the formula of **hydrogen gas**?

Put a (ring) around the correct answer.

H

2H

H₂

[1]

[Total: 5]

12

7 Magnesium sulfate is one of the chemicals in detergent powder.

Mary makes some magnesium sulfate using this reaction.



She measures out 100cm^3 of dilute sulfuric acid and adds solid magnesium carbonate until no more bubbles appear.

(a) Some solid magnesium carbonate is left in the solution.

What technique can Mary use to remove the solid from the solution?

..... [1]

(b) Mary works out the theoretical yield to be 12.0g.

(i) To make this calculation Mary uses the relative formula mass of magnesium carbonate and magnesium sulfate.

She uses these relative atomic masses: C = 12; Mg = 24; O = 16; S = 32.

Use this information to work out these relative formula masses.

relative formula mass of magnesium carbonate, MgCO_3 =

relative formula mass of magnesium sulfate, MgSO_4 =

[2]

(ii) The theoretical yield for Mary's experiment is 12.0g.

Mary dries and weighs the magnesium sulfate she makes. This is her actual yield.

Actual yield = 10.8g.

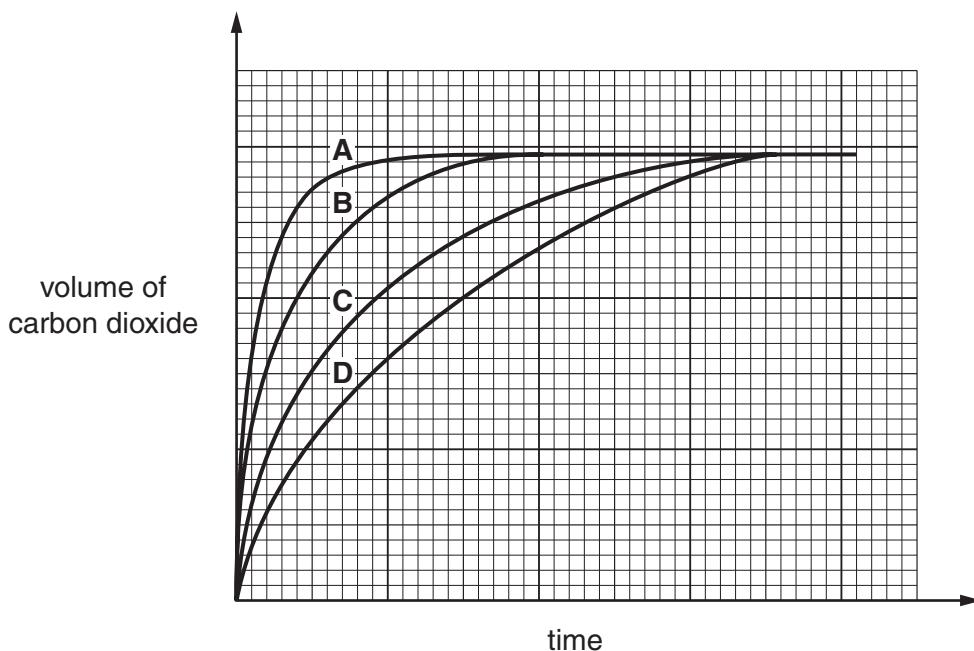
Work out the percentage yield for Mary's experiment.

percentage yield = [1]

13

(c) Mary investigates the rate of this reaction using the same sulfuric acid solution with different sized lumps of magnesium carbonate.

She measures the volume of carbon dioxide given off at time intervals and plots her results on a grid.



(i) How do these graphs show that Mary used the same mass of magnesium carbonate for each experiment?

Put a tick (✓) in the box next to the correct answer.

Each line is a curve.

Each line begins at the origin.

Each line finishes at the same time.

Each line finishes at the same volume.

[1]

(ii) Which line, **A**, **B**, **C** or **D**, shows results from:

the fastest rate of reaction?

answer

the largest lumps of magnesium carbonate?

answer [1]

14

(iii) What **other** change to Mary's experiment would make the reaction faster?

Put a tick (✓) in the box next to the correct answer.

decreasing the concentration of the acid

increasing the temperature of the acid

decreasing the mass of magnesium carbonate

[1]

[Total: 7]

END OF QUESTION PAPER

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The Periodic Table of the Elements

1	2	3	4	5	6	7	0
7 Li lithium 3	9 Be beryllium 4	11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10
23 Na sodium 11	24 Mg magnesium 12	27 Al aluminum 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Nb niobium 41	93 Zr zirconium 40	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[264] Sg seaborgium 106	[268] Mt meitnerium 107	[271] Ds darmstadtium 109
					[277] Hs hassium 108	[271] Rg roentgenium 110	[272] Rg roentgenium 111

Key

relative atomic mass
atomic symbol
name
atomic (proton) number

Elements with atomic numbers 112-116 have been reported but not fully authenticated

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number