

**GENERAL CERTIFICATE OF SECONDARY EDUCATION
TWENTY FIRST CENTURY SCIENCE
CHEMISTRY A**

A322/01

Unit 2: Modules C4 C5 C6 (Foundation Tier)

Candidates answer on the Question Paper
A calculator may be used for this paper

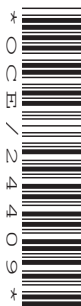
OCR Supplied Materials:
None

Other Materials Required:

- Pencil
- Ruler (cm/mm)

**Wednesday 27 January 2010
Afternoon**

Duration: 40 minutes



Candidate Forename		Candidate Surname	
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Centre Number						Candidate Number				
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MODIFIED LANGUAGE

INSTRUCTIONS TO CANDIDATES

- Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **all** the questions.
- Do **not** write in the bar codes.
- Write your answer to each question in the space provided, however additional paper may be used if necessary.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- The total number of marks for this paper is **42**.
- The Periodic Table is printed on the back page.
- This document consists of **16** pages. Any blank pages are indicated.

2						
3		Mg				
4	K	Ca	transition elements	Ga		
5						

Put a ring around the correct answer.

potassium

[1]

Put a ring around the correct answer.

potassium

[1]

[Total: 4]

- 2 Liz makes some notes about the properties of some elements in Group 1.

Group 1	lithium Li	<u>Lithium</u> Atomic number: 3 Melting point: 181°C Density: 0.53 g/cm³	<u>Rubidium</u> Atomic number: 37 Melting point: 39°C Density: 1.53 g/cm³
	sodium Na		
	potassium K		
	rubidium Rb	<u>Sodium</u> Atomic number: 11 Melting point: 98°C Density: 0.97 g/cm³	

- (a) Explain how Liz could use her notes to predict the properties of potassium.

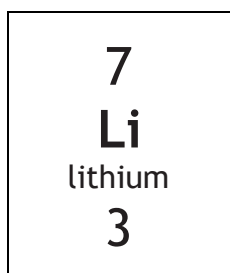
.....
.....
..... [2]

- (b) Describe **two** patterns in the properties of Group 1 elements shown by the information.

.....
.....
..... [2]

5

(c) This is the information for lithium on the Periodic Table.



Complete the sentences about the structure of a lithium atom.

Choose words from this list.

electrons

elements

ions

molecules

neutrons

protons

The shells around a lithium nucleus contain three

The central nucleus of the atom is made up of three

and four

[2]

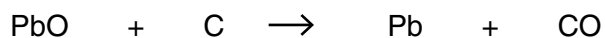
[Total: 6]

6

3 Some types of car batteries contain metals such as lead.

(a) Lead can be extracted by heating lead oxide with carbon.

The equation shows what happens when lead oxide is heated with carbon.



(i) Which statement about the reaction is true?

Put a tick (✓) in the box next to the correct answer.

The reaction involves only oxidation.

☐

The reaction involves only reduction.

☐

The reaction involves both oxidation and reduction.

☐

The reaction does not involve either oxidation or reduction.

☐

[1]

(ii) Which other metals can be extracted by heating their oxides with carbon?

Put a ring around each of the **two** correct answers.

aluminium

copper

potassium

sodium

zinc

[2]

Some car batteries also contain small amounts of other metals including lithium and calcium.

(b) Lithium cannot be extracted by heating lithium oxide with carbon.

Which of the statements gives the **best** reason for this?

Put a tick (✓) in the box next to the correct answer.

Lithium metal reacts with water.

☐

Lithium oxide is ionic.

☐

Lithium is very reactive.

☐

Lithium oxide has a very high melting point.

☐

[1]

7

(c) Calcium can be extracted using electrolysis.

Complete the passage about the extraction of calcium.

Choose words from this list.

electrodes	evaporates	ions	melts
molecules	negative	neutral	positive

Calcium oxide is heated until it

This allows the to move.

During electrolysis calcium metal collects at the electrode.

Oxygen gas is made at the electrode.

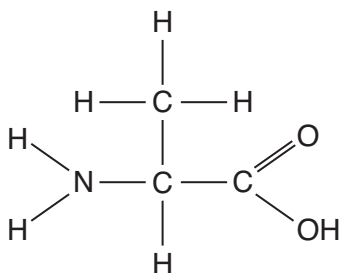
[3]

[Total: 7]

8

- 4 Proteins in the human body are formed from amino acids.

The diagram shows the structure of an amino acid.



The table below shows some information about the elements in the amino acid.

Complete the table by filling in the three empty boxes.

name of element	number of atoms in molecule	percentage (%) by mass
carbon	3	40
oxygen	2	36
.....	1	16
hydrogen		

[3]

[Total: 3]

5 Space probes have gathered data about the atmosphere on Mars.

The table compares the gases in the atmosphere on Mars and on Earth.

name of gas	percentage (%) in atmosphere on Mars	percentage (%) in atmosphere on Earth
carbon dioxide	95.3	less than 1.0
nitrogen	2.7	78.0
argon	1.6	0.9
oxygen	0.2	20.7

(a) Put ticks (✓) in the correct boxes to show whether each gas is an element or a compound.

name of gas	element	compound
carbon dioxide	☐	☐
nitrogen	☐	☐
argon	☐	☐
oxygen	☐	☐

[2]

(b) Look at the table.

Describe two ways that the atmospheres of Mars and Earth are **similar** and two ways that they are **different**.

.....

.....

.....

.....

..... [4]

(c) The percentages in the table for gases on Mars do not add up to 100%.

Suggest a reason why.

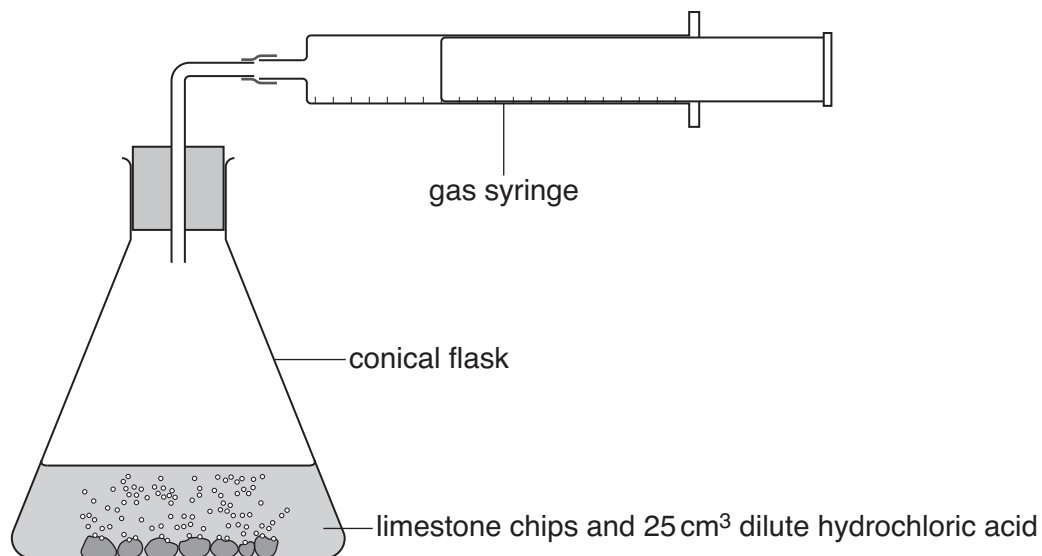
.....

..... [1]

[Total: 7]

Turn over

- 6 Eve carries out an experiment.
She adds 25 cm^3 of dilute hydrochloric acid to limestone chips (calcium carbonate).
Every 30 seconds she records the total volume of gas that has been given off.



The table shows her results.

time in s	total volume of gas in cm^3
0	0
30	80
60	120
90	140
120	150
150	150

- (a) (i) How long does it take for the reaction to finish?

answer s [1]

- (ii) Lumps of limestone are left in the flask after the reaction has stopped.
Why does the reaction stop?
Put a tick (✓) in the box next to the correct answer.

The temperature cools during the reaction.

☐

All the gas has been used up.

☐

All the acid has been used up.

☐

The limestone chips become unreactive.

☐

[1]

11

- (b) During the reaction, solid calcium carbonate reacts with dilute hydrochloric acid and the gas syringe fills with a gas.

At the end of the experiment the flask contains a solution of calcium chloride in water.

- (i) What is the name of the gas made during the reaction?

Put a ring around the correct answer.

carbon dioxide carbon monoxide hydrogen nitrogen oxygen

[1]

- (ii) Draw a line from each **chemical** to the correct **state symbol**.

chemical	state symbol
water	(s)
calcium carbonate	(g)
gas made in the reaction	(aq)
calcium chloride solution	(l)

[2]

- (iii) Draw a line from each **chemical** to the correct **formula**.

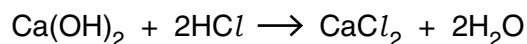
chemical	formula
water	CaCO_3
calcium carbonate	H_2O
hydrochloric acid	CaCl_2
calcium chloride	HCl

[2]

[Total: 7]

12

- 7 Joe carries out an experiment to make a salt.
He makes calcium chloride by reacting calcium hydroxide with dilute hydrochloric acid.



- (a) Joe works out what mass of calcium chloride he can make.

The box below shows some of Joe's working.

Complete Joe's working by filling in the gaps.

	relative atomic mass
Ca	
O	
H	
Cl	35.5

relative formula mass of $\text{Ca(OH)}_2 = 74$

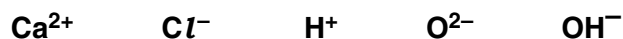
relative formula mass of $\text{CaCl}_2 = \dots\dots\dots$

[2]

- (b) The reaction between calcium hydroxide and hydrochloric acid is a neutralisation reaction.

Which ion is always present in a solution of an alkali?

Put a ring around the correct answer in this list.



[1]

- (c) Write the general equation for a neutralisation reaction by filling in the boxes.

Choose from the formulae in this list.

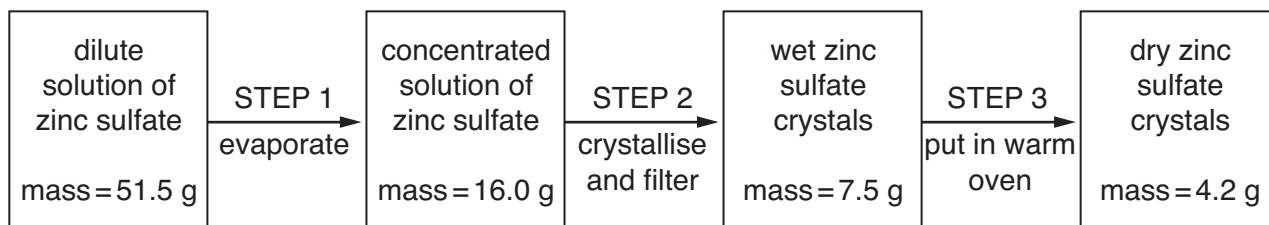


[1]

[Total: 4]

- 8 Sam works for a medicine company.
The company makes zinc sulfate for use in medicines.
She makes some zinc sulfate crystals from zinc sulfate solution.
She measures the mass after each step.

The flow chart shows what she does.



- (a) What happens to the mass of zinc sulfate solution during STEP 1?
Explain why.

.....
.....
..... [2]

- (b) Suggest why the crystals were put in a warm oven.

..... [1]

- (c) What is the actual yield from the experiment?

answer g [1]

[Total: 4]

END OF QUESTION PAPER

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* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.