

THIS IS A NEW SPECIFICATION

**H****Thursday 20 June 2013 – Afternoon****GCSE TWENTY FIRST CENTURY SCIENCE  
CHEMISTRY A****A173/02** Module C7 (Higher Tier)

Candidates answer on the Question Paper.  
A calculator may be used for this paper.

**OCR supplied materials:**  
None

**Other materials required:**

- Pencil
- Ruler (cm/mm)

**Duration:** 1 hour

|                       |  |                      |  |
|-----------------------|--|----------------------|--|
| Candidate<br>forename |  | Candidate<br>surname |  |
| Centre number         |  | Candidate number     |  |

**INSTRUCTIONS TO CANDIDATES**

- Write your name, centre number and candidate number in the boxes above. Please write clearly and in capital letters.
- Use black ink. HB pencil may be used for graphs and diagrams only.
- Answer **all** the questions.
- Read each question carefully. Make sure you know what you have to do before starting your answer.
- Write your answer to each question in the space provided. Additional paper may be used if necessary but you must clearly show your candidate number, centre number and question number(s).
- Do **not** write in the bar codes.

**INFORMATION FOR CANDIDATES**

- Your quality of written communication is assessed in questions marked with a pencil (✎).
- The number of marks is given in brackets [ ] at the end of each question or part question.
- The total number of marks for this paper is **60**.
- This document consists of **20** pages. Any blank pages are indicated.
- The Periodic Table is printed on the back page.

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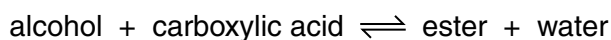
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## 3

Answer **all** the questions.

- 1 Esters are made by reacting an alcohol with a carboxylic acid.

A few drops of concentrated sulfuric acid are added to the mixture.



- (a) What is the job of the concentrated sulfuric acid?

..... [1]

- (b) The reaction mixture has to be heated for some time.

The reactants are volatile and would escape from an open flask as vapour.

- (i) What piece of apparatus is fitted to the flask to prevent this loss of vapour?

..... [1]

- (ii) What is the name of the technique used to prevent this loss of vapour?

..... [1]

- (c) Propyl pentanoate,  $\text{C}_4\text{H}_9\text{COOC}_3\text{H}_7$ , is an ester that tastes of pineapples.

It is made by reacting propanol,  $\text{C}_3\text{H}_7\text{OH}$  with pentanoic acid  $\text{C}_4\text{H}_9\text{COOH}$ .

- (i) Write a symbol equation for the reaction that makes propyl pentanoate.

.....  $\rightleftharpoons$  ..... [1]

- (ii) Suggest a use for propyl pentanoate.

..... [1]

[Total: 5]

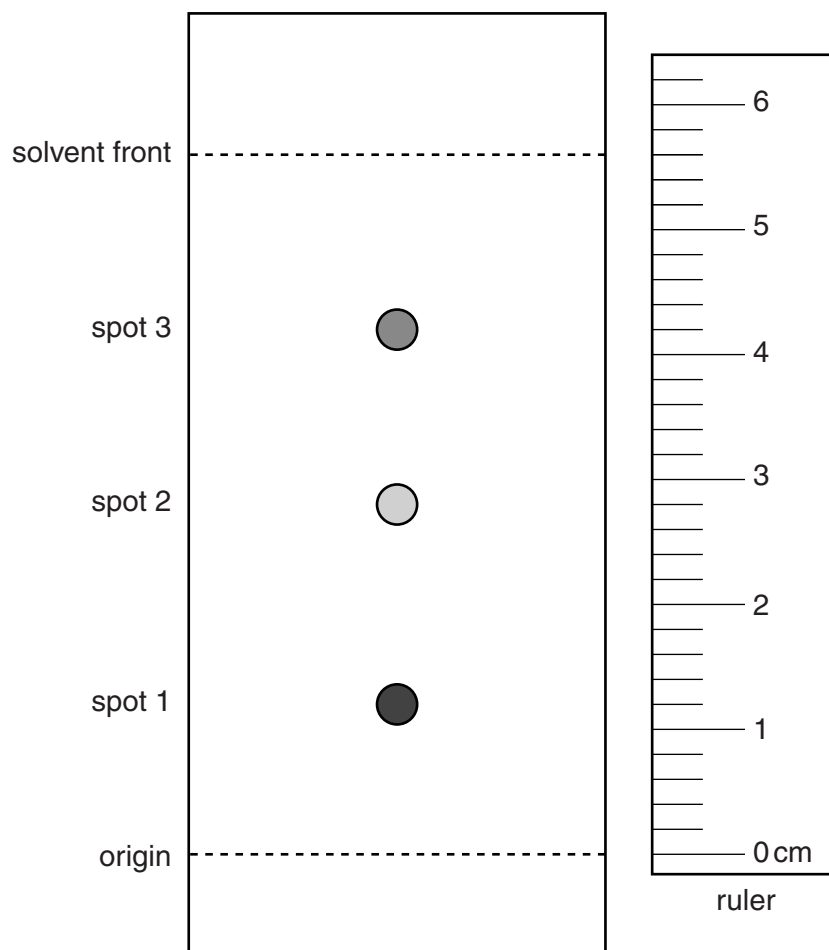
4

2 Peter is testing the food colourings in soft drinks made by different companies.

He wants to find out if any of these drinks contains a banned dye chemical.

Peter uses paper chromatography to separate and identify the dye chemicals in the soft drinks.

Here is Peter's chromatogram for one of the soft drinks.



5

- (a) Explain why the three spots travel different distances up the chromatogram.



*The quality of written communication will be assessed in your answer.*

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..... [6]

- (b) Peter uses R<sub>f</sub> values to identify the dye chemicals in a soft drink.

- (i) Work out the R<sub>f</sub> value of spot 3.

R<sub>f</sub> value = ..... [2]

- (ii) Peter finds the R<sub>f</sub> values for dye chemicals in several different soft drinks.

How can Peter check if any of the soft drinks contain a banned dye chemical?

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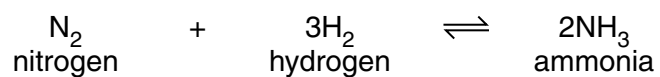
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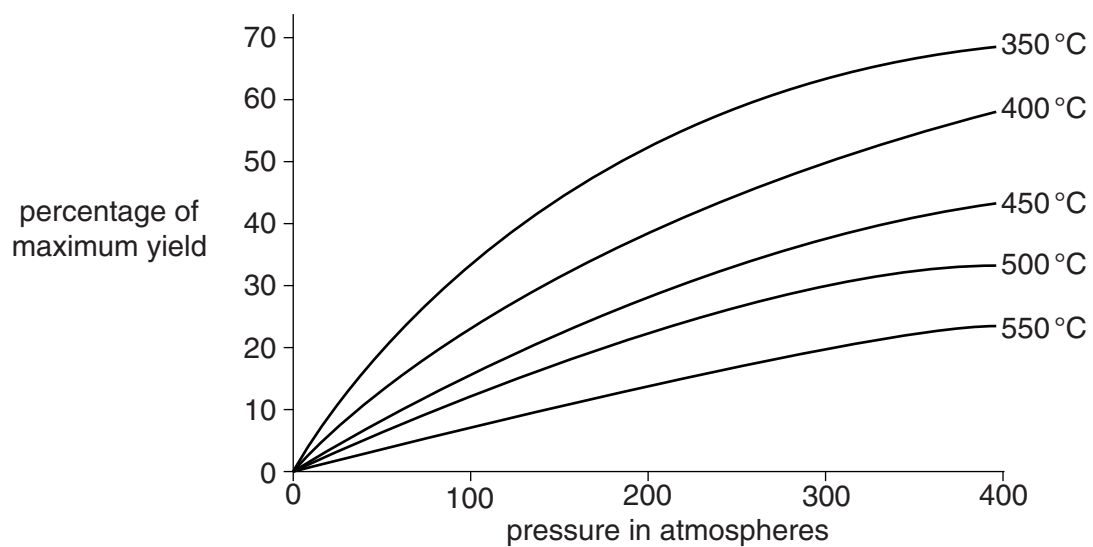
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6

- 3 Ammonia,  $\text{NH}_3$ , is made from nitrogen and hydrogen by the Haber process.



- (a) The graph shows how the yield of ammonia is related to both the temperature and the pressure used.



(i) The Haber process uses:

- a catalyst
- a temperature of  $450^{\circ}\text{C}$
- 250 atmospheres pressure.

These conditions do not give the highest percentage yield of ammonia.

Explain why these conditions are a compromise to make the process economically viable.



*The quality of written communication will be assessed in your answer.*

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..... [6]

8

- (ii) The reaction between nitrogen and hydrogen to form ammonia is reversible.

The reaction mixture can reach a dynamic equilibrium.

Which of these statements describes what is happening at equilibrium?

Put ticks (✓) in the boxes next to the **two** correct statements.

The reaction between nitrogen and hydrogen has stopped.

☐

The forward and reverse reactions happen at the same rate.

☐

All of the nitrogen and hydrogen react to make ammonia.

☐

The concentration of ammonia is increasing.

☐

The concentrations of nitrogen, hydrogen and ammonia are constant.

☐

[2]

- (b) (i) Work out the relative formula mass (RFM) of ammonia,  $\text{NH}_3$ .

RFM of ammonia = ..... [1]

- (ii) What is the mass of ammonia that would be made if 1.0 tonne of nitrogen reacted completely with hydrogen?

Show your working.

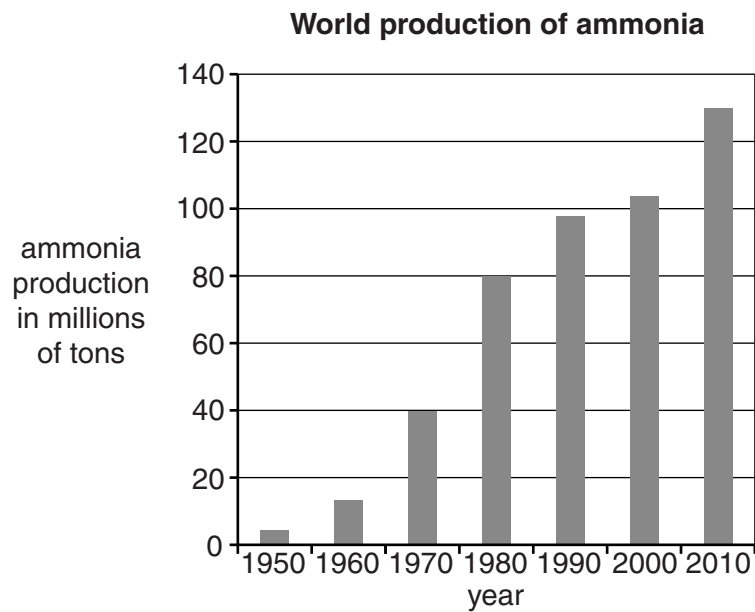
mass of ammonia = ..... tonne [2]

- (iii) A factory converts 95% of the nitrogen into ammonia.

What mass of ammonia does this factory make from each tonne of nitrogen?

mass of ammonia = ..... tonne [1]

(c) Look at the bar chart.



The main use of ammonia is to make fertilisers.

Large scale use of fertilisers made from ammonia causes environmental problems.

Write about these problems, and explain why they have got worse over the last 60 years.

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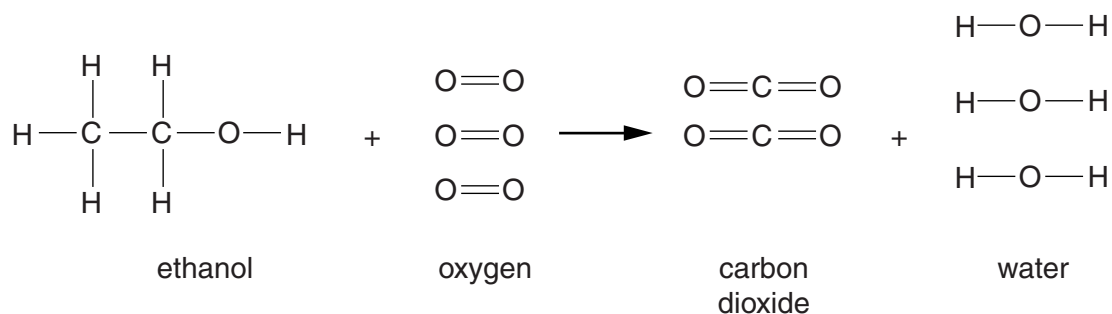
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[Total: 15]

- 4 The table shows the energy involved in the breaking of some bonds.

| Bond | Bond energy in kJ/mol |
|------|-----------------------|
| C—H  | 411                   |
| C—O  | 358                   |
| O=O  | 498                   |
| C=O  | 799                   |
| O—H  | 459                   |
| C—C  | 348                   |

Ethanol burns to make carbon dioxide and water.



- (a) The energy needed to break all of the bonds in the oxygen,  $3\text{O}_2$ , is 1494 kJ.

Work out the energy needed to break all of the bonds in ethanol,  $\text{C}_2\text{H}_5\text{OH}$ .

energy = ..... kJ/mol [2]

- (b) The energy given out when new bonds in  $2\text{CO}_2$  are made is 3196 kJ.

Work out the energy given out when new bonds in the water,  $3\text{H}_2\text{O}$ , are made.

energy = ..... kJ/mol [2]

- (c) Use the information given, and your answers from (a) and (b), to complete the table.

|                                                                                 | Energy in kJ/mol |
|---------------------------------------------------------------------------------|------------------|
| energy <b>needed</b> to break all the bonds in ethanol and oxygen               |                  |
| energy <b>given out</b> when all the bonds in carbon dioxide and water are made |                  |
| energy change when ethanol burns                                                |                  |

[2]

11

(d) The table shows information about some bonds.

| Bond | Bond energy<br>in kJ/mol | Bond length<br>in pm |
|------|--------------------------|----------------------|
| C—C  | 348                      | 154                  |
| C=C  | 614                      | 134                  |
| C≡C  | 839                      | 120                  |

What conclusions can you make from this data?

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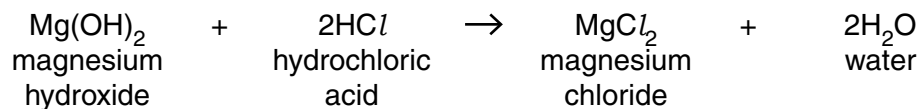
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[Total: 8]

- 5 Emma works for a company making indigestion tablets.

Her job is to find the mass of magnesium hydroxide in tablets from each batch.

Emma titrates the magnesium hydroxide in each tablet with hydrochloric acid.



- (a) (i) What steps should Emma take to ensure that her titration results are as accurate as possible?

Put ticks (✓) in the boxes next to the **two** correct statements.

Add the acid as quickly as possible.

☐

Use a burette to measure the volume of acid.

☐

Add acid 1 cm<sup>3</sup> at a time.

☐

Add acid drop by drop near the end point.

☐

Measure the mass of the flask every 30 seconds.

☐

Allow the product to crystallise.

☐

[2]

- (ii) Emma uses an indicator in her titration.

Explain why she uses an indicator.

.....

.....

..... [2]

(b) Emma analyses an indigestion tablet to find the mass of magnesium hydroxide.

(i) Work out the relative formula mass (RFM) of magnesium hydroxide,  $\text{Mg}(\text{OH})_2$ .

Show your working.

(Relative atomic masses:  $\text{H} = 1$ ;  $\text{Mg} = 24$ ;  $\text{O} = 16$ )

RFM of magnesium hydroxide = ..... [1]

(ii) Emma uses hydrochloric acid with 73.0 g of hydrogen chloride in each  $1.0\text{ dm}^3$  of the acid solution.

It takes  $15.1\text{ cm}^3$  of this hydrochloric acid to neutralise the tablet.

Work out the mass of hydrogen chloride in  $15.1\text{ cm}^3$  of the hydrochloric acid.

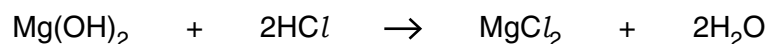
Give your answer to the nearest 0.1 g.

Show your working.

mass of hydrogen chloride = ..... g [2]

(iii) Work out the mass of magnesium hydroxide in the tablet.

Use your answers to (i) and (ii) and this equation to help you.



Show your working.

mass of magnesium hydroxide in the tablet = ..... g [2]

14

(c) Emma analyses six tablets from each batch.

The table shows Emma's results for four batches of tablets.

| Tablet number  | Mass of magnesium hydroxide in g |      |      |      |      |      |
|----------------|----------------------------------|------|------|------|------|------|
|                | 1st                              | 2nd  | 3rd  | 4th  | 5th  | 6th  |
| <b>Batch A</b> | 0.95                             | 0.93 | 0.95 | 0.96 | 0.94 | 0.93 |
| <b>Batch B</b> | 0.88                             | 0.86 | 0.89 | 0.87 | 0.89 | 0.87 |
| <b>Batch C</b> | 1.13                             | 1.16 | 1.14 | 1.15 | 1.13 | 1.16 |
| <b>Batch D</b> | 1.03                             | 1.13 | 1.05 | 1.04 | 1.15 | 1.03 |

The label on each pack of indigestion tablets says that each tablet contains 1.0 g of magnesium hydroxide.

The standard set by the company is that each tablet must be within 0.1 g of this figure.

For each batch decide whether it meets the standard and explain your answers.

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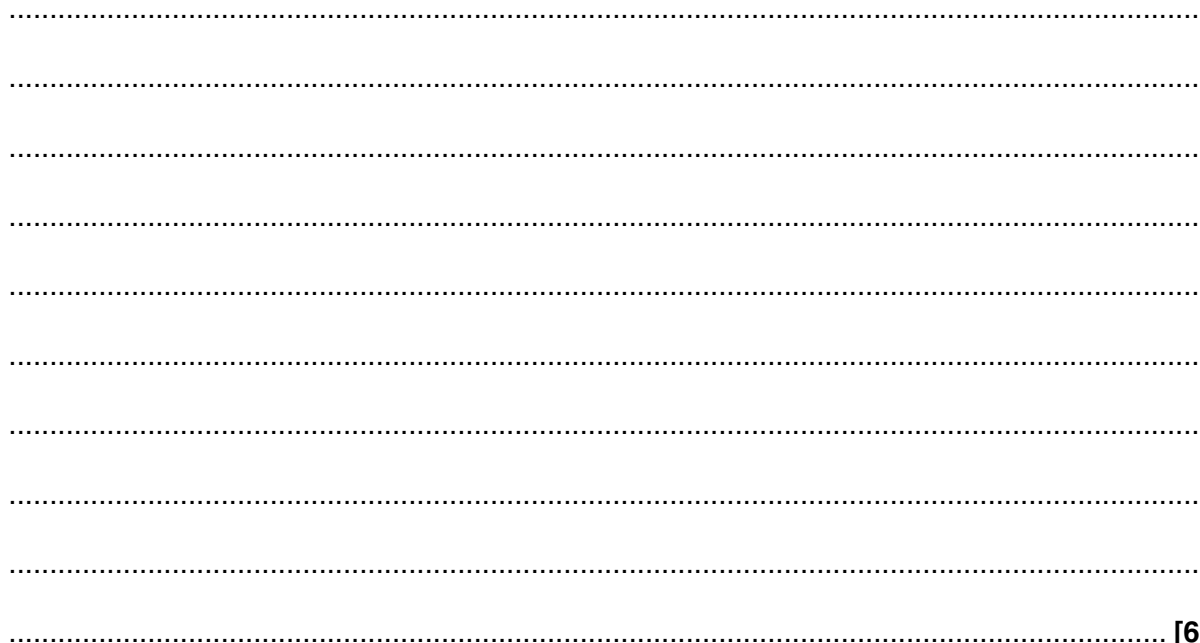
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..... [4]

[Total: 13]

**Method 1** – reaction of ethene with steam:



16

(b) How do the reactions of ethanol, ethane and water with sodium compare?

Put ticks (✓) in the correct boxes to show what happens in each reaction.

|                          | Reaction of<br>ethanol with<br>sodium | Reaction of<br>ethane with<br>sodium | Reaction of<br>water with<br>sodium |
|--------------------------|---------------------------------------|--------------------------------------|-------------------------------------|
| violent reaction         |                                       |                                      |                                     |
| steady reaction          |                                       |                                      |                                     |
| no reaction              |                                       |                                      |                                     |
| hydrogen made            |                                       |                                      |                                     |
| sodium ethoxide<br>made  |                                       |                                      |                                     |
| sodium hydroxide<br>made |                                       |                                      |                                     |

[3]

[Total: 9]

**END OF QUESTION PAPER**

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\* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.