



GCSE

Chemistry A

General Certificate of Secondary Education

Unit A173/02: Module C7 (Higher Tier)

Mark Scheme for June 2013

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

OCR will not enter into any discussion or correspondence in connection with this mark scheme.

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Annotations

Used in the detailed Mark Scheme:

Annotation	Meaning
/	alternative and acceptable answers for the same marking point
(1)	separates marking points
not/reject	answers which are not worthy of credit
ignore	statements which are irrelevant - applies to neutral answers
allow/accept	answers that can be accepted
(words)	words which are not essential to gain credit
<u>words</u>	underlined words must be present in answer to score a mark
ecf	error carried forward
AW/owtte	credit alternative wording / or words to that effect
ORA	or reverse argument

Available in scoris to annotate scripts:

	correct response
	incorrect response
	benefit of doubt
	no benefit of doubt
	error carried forward
	indicate level awarded for a question marked by level of response
	information omitted
	contradiction
	reject

	indicate uncertainty or ambiguity
	draw attention to particular part of candidate's response

Subject-specific Marking Instructions

- a. Accept any clear, unambiguous response (including mis-spellings of scientific terms if they are *phonetically* correct, but always check the guidance column for exclusions).
- b. Crossed out answers should be considered only if no other response has been made. When marking crossed out responses, accept correct answers which are clear and unambiguous.

e.g. for a one-mark question where ticks in the third and fourth boxes are required for the mark:

<input type="checkbox"/>
<input type="checkbox"/>
<input checked="" type="checkbox"/>
<input type="checkbox"/>
<input type="checkbox"/>

<input type="checkbox"/>

<input checked="" type="checkbox"/>
<input checked="" type="checkbox"/>
<input type="checkbox"/>
<input type="checkbox"/>
<input type="checkbox"/>

*This would be worth
1 mark.*

*This would be worth
0 marks.*

*This would be worth
1 mark.*

- c. The list principle:
If a list of responses greater than the number requested is given, work through the list from the beginning. Award one mark for each correct response, ignore any neutral response, and deduct one mark for any incorrect response, e.g. one which has an error of science. If the number of incorrect responses is equal to or greater than the number of correct responses, no marks are awarded. A neutral response is correct but irrelevant to the question.

d. Marking method for tick-box questions:

If there is a set of boxes, some of which should be ticked and others left empty, then judge the entire set of boxes.

If there is at least one tick, ignore crosses and other markings. If there are no ticks, accept clear, unambiguous indications, e.g.

shading or crosses. Credit should be given according to the instructions given in the guidance column for the question. If more boxes are ticked than there are correct answers, then deduct one mark for each additional tick. Candidates cannot score less than zero marks.

e.g. *if a question requires candidates to identify cities in England:*

Edinburgh	<input type="checkbox"/>
Manchester	<input type="checkbox"/>
Paris	<input type="checkbox"/>
Southampton	<input type="checkbox"/>

the second and fourth boxes should have ticks (or other clear indication of choice) and the first and third should be blank (or have indication of choice crossed out).

Edinburgh		ü		ü	ü	ü	ü	
Manchester	ü	x	ü	ü				ü
Paris			ü	ü	ü	ü	ü	
Southampton	ü	x	ü	ü	ü	ü		
Score:	2	2	1	1	1	1	0	0
								NR

e. For answers marked by levels of response:

i. **Read through the whole answer from start to finish**

ii. **Decide the level that best fits** the answer – match the quality of the answer to the closest level descriptor

iii. **To determine the mark within the level**, consider the following:

Descriptor	Award mark
A good match to the level descriptor	The higher mark in the level
Just matches the level descriptor	The lower mark in the level

iv. Use the **L1, L2, L3** annotations in Scoris to show your decision; do not use ticks.

Quality of Written Communication skills assessed in 6-mark extended writing questions include:

- appropriate use of correct scientific terms
- spelling, punctuation and grammar
- developing a structured, persuasive argument
- selecting and using evidence to support an argument
- considering different sides of a debate in a balanced way
- logical sequencing.

Question		Answer	Marks	Guidance
1	(a)	catalyst / speed up reaction / increase rate of reaction	1	allow to lower activation energy / to provide an alternative route for the reaction ignore reference to getting rid of water
	(b) (i)	condenser	1	allow Liebig condenser / reflux condenser
	(ii)	reflux / refluxing	1	do not allow condensation / distillation
	(c) (i)	$C_3H_7OH + C_4H_9COOH \xrightarrow{a} C_4H_9COOC_3H_7 + H_2O$ or $CH_3(CH_2)_2OH + CH_3(CH_2)_3COOH \xrightarrow{a} CH_3(CH_2)_3COO(CH_2)_2CH_3 + H_2O$ or $C_3H_8O + C_5H_{10}O_2 \xrightarrow{a} C_8H_{16}O_2 + H_2O$	1	allow reactants in either order and products in either order allow O_2 for OO in formulae ignore state symbols even if incorrect allow multiples that balance
	(ii)	flavouring / add pineapple flavour	1	ignore vague answers eg add to food allow fragrances eg perfume/deodorant/air freshener
		Total	5	

Question		Answer	Marks	Guidance
2	(a)	<p>Level 3 (5–6 marks) Answer uses idea of equilibrium position between phases to explain why some chemicals move quicker than others. Quality of written communication does not impede communication of the science at this level.</p> <p>Level 2 (3–4 marks) Answer partially explains why some chemicals move quicker than others. Quality of written communication partly impedes communication of the science at this level.</p> <p>Level 1 (1–2 marks) Answer includes idea that some chemicals move quicker than others with little or no explanation. Quality of written communication impedes communication of the science at this level.</p> <p>Level 0 (0 marks) Insufficient or irrelevant science. Answer not worthy of credit.</p>	6	<p>This question is targeted at grades up to A*</p> <p>Indicative scientific points may include:</p> <ul style="list-style-type: none"> solvent is mobile phase (water in) paper is stationary phase chemicals are attracted to (water in) paper chemicals are attracted to solvent an equilibrium is established for each chemical in solvent and in paper for each chemical there is distribution/partition between solvent and paper solvent moves up paper different chemicals move up paper at different rates the more the chemical is attracted to the solvent the quicker it moves up the paper / ora the more the chemical is towards the solvent side of the equilibrium the quicker it moves up the paper / ora. <p>accept dissolve in/spend more time in instead of attracted to</p> <p>do not accept ideas of chemicals reacting with solvent/paper/phases</p> <p>Use the L1, L2, L3 annotations in Scoris; do not use ticks.</p>

Question		Answer	Marks	Guidance
2	(b) (i)	$R_f = 4.2/5.6 \text{ (1)}$ $= 0.75 \text{ (1)}$ [for braille scripts only: $R_f = 4/6 \text{ (1)}$ $= 0.66 \text{ (1)}]$	2	both marks for correct answer without working allow (use of 4.0 / 4.4 or between in calculation giving) answer of 0.71 / 0.79 or between for one mark only allow answers with more than two figures after decimal point do not credit 5.6/4.2
	(ii)	‘method’ mark compare measured dye chemical R_f with published/reference R_f values of banned dye chemicals / run known samples of banned dye chemicals on a chromatogram and calculate R_f values for comparison (1) ‘use of results’ mark if measured dye chemical R_f is same as R_f of a banned dye chemical the soft drink contains a banned dye chemical / the same chemical will always have the same R_f value (1)	2	allow only one mark for idea of running banned dye chemicals and soft drink dye chemicals on same chromatogram and comparing distances travelled ignore idea that same chemical always travels same distance unless made clear that this is under identical conditions
		Total	10	

Question		Answer	Marks	Guidance
3	(a) (i)	<p>Level 3 (5–6 marks) Answer uses ideas of equilibrium and rate to explain the compromise choices of both temperature and pressure and why a catalyst is used. Quality of written communication does not impede communication of the science at this level.</p> <p>Level 2 (3–4 marks) Answer gives a good explanation for either temperature or pressure or a partial explanation of both. May write about yield rather than equilibrium. Quality of written communication partly impedes communication of the science at this level.</p> <p>Level 1 (1–2 marks) Answer discusses the data but gives little or no explanation of the compromise conditions. Quality of written communication impedes communication of the science at this level.</p> <p>Level 0 (0 marks) Insufficient or irrelevant science. Answer not worthy of credit.</p>	6	<p>This question is targeted at grades up to A</p> <p>Indicative scientific points may include:</p> <ul style="list-style-type: none"> to be economically viable need to make as much ammonia as possible as quickly as possible temperature ideas higher temperature shifts equilibrium to left the higher the temperature the lower the yield high temperature costs more to maintain because forward reaction is exothermic / releases heat energy at low temperature rate is too slow compromise medium temperature eg about 450°C <p>pressure ideas</p> <ul style="list-style-type: none"> higher pressure shifts equilibrium to right the higher the pressure the higher the yield because forward reaction decreases pressure / decreases number of molecules higher the pressure the tougher the reaction vessel has to be so that it does not burst higher pressure increases cost / risk of process compromise medium pressure of 250 atmospheres. <p>catalyst ideas</p> <ul style="list-style-type: none"> catalyst gives more product in same time catalyst allows use of lower temperature for same rate catalyst does not change yield/equilibrium position <p>Use the L1, L2, L3 annotations in Scoris; do not use ticks.</p>
	(ii)	tick in box 2 (1) tick in box 5 (1)	2	

Question		Answer	Marks	Guidance
3	(b)	(i) 17	1	
		(ii) $1.0 \times (2 \times 17) / 28 \text{ (1)}$ $= 1.2 \text{ (1)}$	2	allow 1.21 ignore additional figures beyond 3 sig fig if they would round down to 1.21 allow ecf from (i) allow both marks for correct answer without working
		(iii) $(1.2 \times 95/100 =) 1.14$	1	allow 1.1 allow ecf from (ii) (ie candidate's answer in (ii) x 0.95)
	(c)	fertiliser is washed into rivers causing pollution/eutrophication (1) production of ammonia has increased / use of fertilisers has increased (1) plus a link from more ammonia/fertilisers to more pollution (1)	3	allow description of eutrophication
		Total	15	

Question		Answer	Marks	Guidance
4	(a)	(5 x 411) + 348 + 358 + 459 = (1) 3220 (1)	2	allow both marks for correct answer without working
	(b)	6 x 459 = (1) 2754 (1)	2	allow both marks for correct answer without working
	(c)	4714 in first box and 5950 in second box (1) in third box (–) 1236 (1)	2	allow ecf from (a) and (b) give first mark for correct addition and transfer of both figures second mark for correct subtraction of figures they have used ignore sign
	(d)	any two from as number of bonds (between two carbon atoms) increases bond length decreases (1) as number of bonds (between two carbon atoms) increases bond energy increases (1) as bond energy increases bond length decreases (1)	2	ora throughout ignore references to strength or weakness of bonds allow triple bond has more energy than double which has more energy than single
		Total	8	

Question		Answer	Marks	Guidance
5	(a) (i)	tick in box 2 (1) tick in box 4 (1)	2	
	(ii)	make sure correct/exact amount of acid is added (to react with magnesium hydroxide) (1) to show end point/when mixture neutral (by change of colour) (1)	2	
	(b) (i)	58	1	
	(ii)	0.0151×73.0 (1) = 1.1 (1)	2	allow both marks for correct answer without working ignore additional figures after sig fig if they would round down to 1.1
	(iii)	$58 \times 1.1/73$ (1) = 0.87 (1)	2	allow ecf from (i) and (ii) ie (i) x (ii) / 73 allow both marks for correct answer without working allow 0.9
	(c)	batch A is OK because each tablet is within the allowed range (1) batch B is unsatisfactory because the tablets contain less (than the minimum permitted amount) (1) batch C is unsatisfactory because the tablets contain more (than the maximum permitted amount) (1) batch D is unsatisfactory because some tablets contain more (than the maximum permitted amount) (1)	4	each answer must say indicate whether the batch is satisfactory or unsatisfactory and explain why do not credit calculation and use of average/mean values
		Total	13	

Question		Answer	Marks	Guidance
6	(a)	<p>Level 3 (5–6 marks) Answer deals fully with sustainability in terms of feedstock and another factor. Quality of written communication does not impede communication of the science at this level.</p> <p>Level 2 (3–4 marks) Answer deals fully with sustainability in terms feedstock or partially in terms of feedstock and another factor. Quality of written communication partly impedes communication of the science at this level.</p> <p>Level 1 (1–2 marks) Answer offers some explanation of sustainability. Quality of written communication impedes communication of the science at this level.</p> <p>Level 0 (0 marks) Insufficient or irrelevant science. Answer not worthy of credit.</p>	6	<p>This question is targeted at grades up to C</p> <p>Indicative scientific points may include: feedstock/raw materials</p> <ul style="list-style-type: none"> · method 1 uses ethene from crude oil · natural gas/crude oil is finite · method 2 uses sugar · more sugar beet/cane can be grown · in terms of feedstock fermentation is more sustainable fuel/energy input · method 1 uses high temperature which needs a lot of fuel/energy · burning fuel releases pollutant gases · method 2 uses low temperature with uses little fuel · in terms of fuel needed fermentation is more sustainable equipment · method 1 is carried out at high pressure and so needs reaction vessel to withstand this/has safety issues · method 2 is carried out at normal/atmospheric pressure · making reaction vessel for method 1 uses more resources than making reaction vessel for 2 · in terms of equipment fermentation more sustainable other ideas · method 1 has a better atom economy than method 2 and is therefore more sustainable in this aspect · method 2 releases carbon dioxide which is a greenhouse gas, but so does method 1 as fuel is burned <p>Use the L1, L2, L3 annotations in Scoris; do not use ticks.</p>

Question		Answer			Marks	Guidance																											
6	(b)		<table border="1"> <thead> <tr> <th></th> <th>ethanol</th> <th>ethane</th> <th>water</th> </tr> </thead> <tbody> <tr> <td>violent reaction</td> <td></td> <td></td> <td>ü</td> </tr> <tr> <td>slow reaction</td> <td>ü</td> <td></td> <td></td> </tr> <tr> <td>no reaction</td> <td></td> <td>ü</td> <td></td> </tr> <tr> <td>hydrogen made</td> <td>ü</td> <td></td> <td>ü</td> </tr> <tr> <td>sodium ethoxide made</td> <td>ü</td> <td></td> <td></td> </tr> <tr> <td>sodium hydroxide made</td> <td></td> <td></td> <td>ü</td> </tr> </tbody> </table>		ethanol	ethane	water	violent reaction			ü	slow reaction	ü			no reaction		ü		hydrogen made	ü		ü	sodium ethoxide made	ü			sodium hydroxide made			ü	3	one mark for each correct column
	ethanol	ethane	water																														
violent reaction			ü																														
slow reaction	ü																																
no reaction		ü																															
hydrogen made	ü		ü																														
sodium ethoxide made	ü																																
sodium hydroxide made			ü																														
			Total	9																													

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