



GCSE

Chemistry B

Unit **B742/01**: Modules C4, C5, C6 (Foundation Tier)

General Certificate of Secondary Education

Mark Scheme for June 2014

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.


All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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These are the annotations, (including abbreviations), including those used in scoris, which are used when marking

Annotation	Meaning of annotation
	Blank Page – this annotation must be used on all blank pages within an answer booklet (structured or unstructured) and on each page of an additional object where there is no candidate response.

Abbreviations, annotations and conventions used in the detailed Mark Scheme.

- / = alternative and acceptable answers for the same marking point
- (1) = separates marking points
- allow = answers that can be accepted
- not = answers which are not worthy of credit
- reject = answers which are not worthy of credit
- ignore = statements which are irrelevant
- () = words which are not essential to gain credit
- = underlined words must be present in answer to score a mark (although not correctly spelt unless otherwise stated)
- ecf = error carried forward
- AW = alternative wording
- ora = or reverse argument

Question			Answer	Marks	Guidance
1	a		beryllium / calcium / strontium/ barium / radium (1)	1	allow Be / Ca / Sr / Ba / Ra (1)
	b		sodium / aluminium / silicon / phosphorus / sulfur / chlorine / argon (1)	1	allow Na / Al / Si / P / S / Cl / Ar (1)
	c		group 2 (1) 3 rd period (1)	2	
			Total	4	

Question		Answer	Marks	Guidance
2	a	colour of chlorine – (pale) green (1) state of bromine – liquid (1) use of iodine – sterilise wounds / antiseptic (1)	3	allow light green (1) ignore yellow / green allow disinfectant / pharmaceuticals / medicines / photographic chemicals / printing ink / dyes / making animal feeds / added to salt (to avoid thyroid disease) / making ethanoic acid / making polymers (1) allow sterilisation (1) allow testing for starch (1) ignore cleaning medical equipment
	b	sodium + bromine → sodium bromide (1)	1	allow correct formulae i.e. $\text{Na} + \text{Br}_2 \rightarrow \text{NaBr}$ or mix of words and correct formulae If formulae used balancing is not necessary
		Total	4	

Question	Answer	Marks	Guidance
3 a	any two from hard (1) high density (1) high tensile strength / strong (1) (good) conductors of electricity (1) (good) conductors of heat (1) malleable (1) ductile / can be made into wires (1) sonorous / when hit makes ringing sound (1)	2	allow good conductors (1) if no marks awarded for conductors of heat and electricity allow can be hammered into shape (1) ignore bendy / flexible ignore durable / tough / hardwearing / long lasting
b	strong (1) low density / lightweight (1) does not corrode (1)	3	mark first three points but ignore irrelevant reasons e.g. hardwearing or durable allow stays rigid (1) ignore so it doesn't fall apart ignore light allow does not rust (1) allow a property (1) and a linked explanation (1) for example strong (1) so the frame will not buckle (1) ignore references to cost
	Total	5	

Question	Answer	Marks	Guidance
4 a	<p>Level 3 Recall or deduce one piece of information about atomic number AND one about mass number AND deduces the number of protons, neutrons and electrons.</p> <p>Quality of written communication does not impede communication of the science at this level. (5 – 6 marks)</p> <p>Level 2 Recall or deduce one piece of information about atomic number AND one about mass number OR deduces the number of protons, neutrons and electrons.</p> <p>Quality of written communication partly impedes communication of the science at this level. (3 – 4 marks)</p> <p>Level 1 Recall or deduce one piece of information about atomic number OR recall or deduce one piece of information about mass number OR deduce the number of protons or neutrons or electrons.</p> <p>Quality of written communication impedes communication of the science at this level. (1 – 2 marks)</p> <p>Level 0 Insufficient or irrelevant science. Answer not worthy of credit. (0marks)</p>	6	<p>This question is targeted at grades up to C.</p> <p>Indicative scientific points may include:</p> <ul style="list-style-type: none"> • bottom number is atomic number or proton number • atomic number or proton number is 13 • atomic number or proton number is the number of protons in the nucleus • atomic number or proton number is the number of electrons • top number is mass number or atomic mass • mass number or atomic mass is 27 • mass number or atomic mass is the total number of protons and neutrons in the nucleus • number of protons is 13 • number of neutrons is 14 • number of electrons is 13 <p>allow higher level answers at level 3 such as:</p> <ul style="list-style-type: none"> • electronic structure is 2.8.3. • Al is in the 3rd period or row • Al is in Group 3 or column 3 • Al forms 3+ ions <p>allow 2 or more higher level answers with no other points - 5 marks</p> <p>Use the L1, L2, L3 annotations in Scoris; do not use ticks.</p>

b	(J.J. Thomson) discovered the electron (1) (Rutherford) suggested the nuclear atom (1)	2	allow references to discovery of electron and nuclear atom without assigning names (up to 2 marks) allow more detailed answers re J. J. Thomson or Rutherford for 2 marks
		8	

Question	Answer	Marks	Guidance
5 a	increased / got bigger (1) using more electricity / using more energy / more electricity generated (1)	2	allow in 2000 17000 to 19000 (megalitres) used and in 2006 24000 to 28000 (megalitres) used (1) ignore it has (developed) a wider range allow more power stations (1)
b	any two from nitrates (1) lead (compounds) (1) pesticides (1) herbicides (1)	2	allow fertilisers / named fertiliser (1) allow chlorinated hydrocarbons / aluminium (compounds) (1) ignore chlorine / fluoride / dust / allow other more obscure pollutants e.g. hormones (1)
	Total	4	

Question			Answer	Marks	Guidance																			
6	a		<table><tr><th rowspan="2">Indicator</th><th colspan="3">Colour in</th></tr><tr><th>Acid</th><th>Neutral</th><th>Alkali</th></tr><tr><td>litmus</td><td>red</td><td>purple</td><td>blue</td></tr><tr><td>phenolphthalein</td><td>colourless</td><td>colourless</td><td>pink</td></tr><tr><td>universal indicator</td><td>red, orange or yellow</td><td>green</td><td>blue or purple</td></tr></table> <p>phenolphthalein row correct (1)</p> <p>universal indicator row correct (1)</p>	Indicator	Colour in			Acid	Neutral	Alkali	litmus	red	purple	blue	phenolphthalein	colourless	colourless	pink	universal indicator	red, orange or yellow	green	blue or purple	2	allow purple / lilac / red (1)
Indicator	Colour in																							
	Acid	Neutral	Alkali																					
litmus	red	purple	blue																					
phenolphthalein	colourless	colourless	pink																					
universal indicator	red, orange or yellow	green	blue or purple																					
	b	i	burette (1)	1																				
		ii	<p>pipette filler (1) because the liquid is corrosive or harmful or an irritant (1)</p> <p>or</p> <p>(safety) goggles (1) because the liquid is corrosive or harmful or an irritant (1)</p>	2	<p>allow avoids getting liquid in mouth (1)</p> <p>ignore idea that protects eyes from chemicals</p> <p>allow wear gloves (1) because the liquid is corrosive or harmful or an irritant (1)</p> <p>allow protective clothing (1) because the liquid is corrosive or harmful or an irritant (1)</p>																			

		iii	idea of a colour change (1) but idea of a sudden colour change (2) or starts blue (1) then it changes to purple or red (1)	2	
			Total	7	

Question			Answer	Marks	Guidance
7	a		72 (1)	1	unit not needed
	b		C_4H_6 / H_6C_4 (1)	1	not if superscripts used for the numbers
	c		C_2H_2 and C_6H_6 (1)	1	both needed if no answer on answer line allow other ways of indicating the correct answer e.g. circling, ticking or underlining
	d		75% (1)	1	
			Total	4	

Question	Answer	Marks	Guidance										
8 a	copper hydroxide → copper oxide + water (1)	1	allow steam for water allow correct formulae and mix of formulae and names the equation does not need to be balanced Cu(OH) ₂ → CuO + H ₂ O (1) allow heat above arrow not copper hydroxide + heat → copper oxide + water										
b	<table border="1"><thead><tr><th>experiment number</th><th>Mass of water made in g</th></tr></thead><tbody><tr><td>1</td><td>0.09</td></tr><tr><td>2</td><td>0.18</td></tr><tr><td>3</td><td>0.28</td></tr><tr><td>4</td><td>0.37</td></tr></tbody></table> <p>mass of water calculated for missing values (1)</p> <p>prediction supported because as more copper hydroxide is used the mass of water increases (1)</p> <p>but</p> <p>prediction supported illustrated by examples showing the direct proportionality e.g. mass of Cu(OH)₂ doubles in expt 1 and expt 2 and so does the mass of water (2)</p>	experiment number	Mass of water made in g	1	0.09	2	0.18	3	0.28	4	0.37	3	
experiment number	Mass of water made in g												
1	0.09												
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3	0.28												
4	0.37												

Question	Answer	Marks	Guidance
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9	<p>[Level 3] Deduces how changing temperature and pressure affects the percentage yield AND explains the connection between percentage yield and position of equilibrium Quality of written communication does not impede communication of the science at this level (5 – 6 marks)</p> <p>[Level 2] Deduces how changing temperature and changing the pressure affects the percentage yield Quality of written communication partly impedes communication of the science at this level (3 – 4 marks)</p> <p>[Level 1] Deduces how changing temperature affects the percentage yield OR deduces how changing pressure affects the percentage yield Quality of written communication impedes communication of the science at this level (1 – 2 marks)</p> <p>[Level 0] Insufficient or irrelevant science. Answer not worthy of credit. (0 marks)</p>	6	<p>This question is targeted at grades up to C.</p> <p>Indicative scientific points at level 3 must include:</p> <ul style="list-style-type: none"> As percentage yield increases position of equilibrium shifts to the right / high percentage yield has position of equilibrium is on the right / ora <p>Relevant points at all levels could include explanations</p> <ul style="list-style-type: none"> as temperature increases percentage yield decreases / ora as pressure increases percentage yield increases / ora <p>Use the L1, L2, L3 annotations in scoris. Do not use ticks.</p>
	Total	6	

Question	Answer	Marks	Guidance
10 a	reading decreases (1)	2	2 nd mark is dependent on decreases

	(because) a gas is made / carbon dioxide is made (1)		not the name of an incorrect gas
b	reaction is slower / reading on balance decreases more slowly (1) ethanoic acid is a weak acid / fewer collisions / fewer hydrogen ions (1)	2	allow idea that reaction takes longer (1) allow has not decreased as much (in the same time) (1) ignore a longer rate of reaction allow ethanoic acid is not as strong as hydrochloric acid (1)
Total		4	

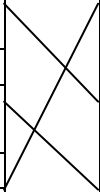
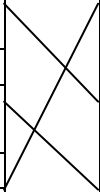
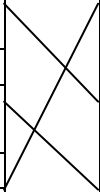
Question	Answer	Marks	Guidance
11 a	52 (%) (1)	1	
b i	carbon, hydrogen and fluorine (1)	1	all three required allow any order ignore symbols not fluoride
ii	8 (1)	1	
c	any two from: (increased risk of) sunburn (1) (accelerated) ageing of skin (1) (increased risk of) cataracts (1) (increased risk of) skin cancer (1)	2	allow sunstroke or damage to the skin (1) allow damage to eyes (1) ignore just 'cancer'
Total		5	

Question	Answer	Marks	Guidance
12 a i	5 (cm ³) (1)	1	allow any value between 5 and 6
ii	44°C (1)	2	allow 42-45 (°C) (1)

		idea of highest point on the curve / where most carbon dioxide is made (1)		second mark is dependent on the correct temperature allow optimum temperature (1)
	b	$C_6H_{12}O_6 \rightarrow 2CO_2 + 2C_2H_5OH$ formulae (1) balancing – dependent on correct formulae (1)	2	allow C_2H_6O as formula for ethanol allow any correct multiple e.g. $2C_6H_{12}O_6 \rightarrow 4CO_2 + 4C_2H_5OH$ allow = or \rightarrow for arrow not 'and' or & for + allow one mark for correct balanced equation with minor errors of case, subscript or superscript e.g. $C^6H^{12}O^6 \rightarrow 2Co_2 + 2C_2H_5OH$
Total			5	

Question	Answer	Marks	Guidance
13	<p>Level 3 Applies knowledge to correctly rank all of the samples in order of hardness AND identifies the type of hardness in two of the samples with a correct explanation for at least one of them. Quality of written communication does not impede communication of the science at this level. (5 – 6 marks)</p> <p>Level 2 Applies knowledge to correctly rank all of the samples in order of hardness AND identifies the type of hardness in one of the samples (no explanation needed). OR identifies the type of hardness in two of the samples (no explanation needed). OR Identifies the type of hardness in one of the samples with a correct explanation. Quality of written communication partly impedes communication of the science at this level. (3 – 4 marks)</p> <p>Level 1 Applies knowledge to either correctly rank two of the samples in order of hardness OR identifies the type of hardness in one of the samples (no explanation needed). Quality of written communication impedes communication of the science at this level. (1 – 2 marks)</p> <p>Level 0 Insufficient or irrelevant science. Answer not worthy of credit. (0 marks)</p>	6	<p>This question is targeted at grades up to C.</p> <p>Indicative scientific points may include: Order of hardness</p> <ul style="list-style-type: none"> • B>A>C <p>Types of hardness and explanation</p> <ul style="list-style-type: none"> • sample A contains permanent hardness • as not softened by boiling • sample B contains both temporary and permanent hardness • as some (but not all) of the hardness is removed by boiling • sample C contains only temporary hardness • as it completely softened by boiling <p>Use the L1, L2, L3 annotations in Scoris; do not use ticks.</p>
		6	

Question	Answer	Marks	Guidance
14 a	water (1) oxygen (1)	2	allow H ₂ O (1) allow O / O ₂ / air (1) allow moist air or damp air (2) ignore iron
b	any two from: oil or grease (1) paint (1) galvanising (1) idea of sacrificial protection (1) alloying (1) tin plating (1) then any one from: oil or grease or paint or tin plating – act as a barrier (so they stop water and oxygen reaching the iron) (1) galvanising – coated with zinc / so acts as a barrier / zinc corrodes first (1) sacrificial protection uses a more reactive metal / magnesium reacts instead of iron / zinc reacts instead of iron (1)	3	allow coat with zinc (1) allow put in contact with magnesium or zinc (1) allow coat in another metal (1) allow mix with another metal (1) allow use stainless steel (1) allow plastic coating (1) allow paint provides a protective layer (1)
	Total	5	

Question	Answer	Marks	Guidance														
15 a	<p>Yes (no mark)</p> <p>solvent A will remove the blue paint and solvent D will remove the yellow paint (1)</p> <p>solvents A and D will not damage the clothes (1)</p>	2	<p>allow solvent A removes blue paint and does not damage clothes (1)</p> <p>allow solvent D removes yellow paint and does not damage clothes (1)</p>														
b	<table><tr><td></td><td></td><td></td></tr><tr><td>active detergent</td><td rowspan="5"></td><td>removes food stains</td></tr><tr><td></td><td></td></tr><tr><td>optical brightener</td><td>does the cleaning</td></tr><tr><td></td><td></td></tr><tr><td>enzymes</td><td>gives a 'whiter than white' appearance</td></tr></table> <p>all three correct (2)</p> <p>but</p> <p>one or two correct (1)</p>				active detergent		removes food stains			optical brightener	does the cleaning			enzymes	gives a 'whiter than white' appearance	2	
active detergent		removes food stains															
optical brightener		does the cleaning															
enzymes		gives a 'whiter than white' appearance															
	Total	4															

Question			Answer	Marks	Guidance
16	a	i	electricity generation (1)	1	
		ii	any three from: decrease in total water usage (1) decrease in use for electricity generation (1) decrease in amount used for farming (1) other uses decreases (1)	3	allow (virtually) no change in use for farming (1)
		iii	$\frac{13000}{42000} \times 100$ (2) but if incorrect then $\text{percentage} = \frac{\text{volume for public watersupply}}{\text{total volume}} \times 100$ (1)	2	allow $\frac{13000}{42000} = 0.3095$ (1) allow $\frac{13}{42} \times 100$ (1) 0.3095 x 100 (1) No mark for 30.95%
		iv	increase (1)	1	allow went to 37.14 (%)
	b		water meters increases (so less water used) (1) population increases (so more water used) (1) idea of a balance between more and less water being used (1)	3	
Total				10	

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