

GENERAL CERTIFICATE OF SECONDARY EDUCATION
GATEWAY SCIENCE
CHEMISTRY B

B641/02

Unit 1 Modules C1 C2 C3 (Higher Tier)

Candidates answer on the Question Paper
A calculator may be used for this paper

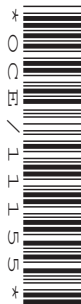
OCR Supplied Materials:
None

Other Materials Required:

- Pencil
- Ruler (cm/mm)

Monday 18 January 2010
Morning

Duration: 1 hour



Candidate Forename		Candidate Surname	
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Centre Number						Candidate Number				
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INSTRUCTIONS TO CANDIDATES

- Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **all** the questions.
- Do **not** write in the bar codes.
- Write your answer to each question in the space provided, however additional paper may be used if necessary.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- The Periodic Table is printed on the back page.
- The total number of marks for this paper is **60**.
- This document consists of **20** pages. Any blank pages are indicated.

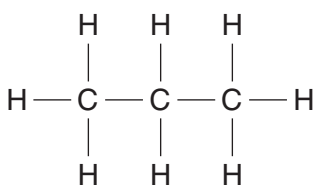
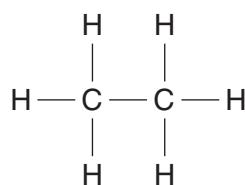
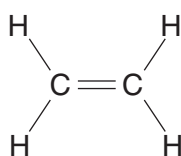
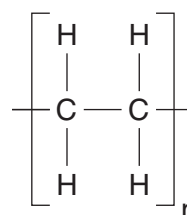
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PLEASE DO NOT WRITE ON THIS PAGE

3

Answer **all** the questions.**Section A – Module C1**

- 1 Look at the displayed formulas of some compounds of carbon.

compound **A**compound **B**compound **C**compound **D**

- (a) Which one of the compounds is a **polymer**?

Choose from **A**, **B**, **C** or **D**.

answer [1]

- (b) Which one of the compounds is an **alkene**?

Choose from **A**, **B**, **C** or **D**.

answer [1]

- (c) Compounds **A**, **B**, **C** and **D** are hydrocarbons.

The atoms of carbon and hydrogen are joined together by covalent bonds.

Describe how a covalent bond is formed between an atom of carbon and an atom of hydrogen.

.....

.....

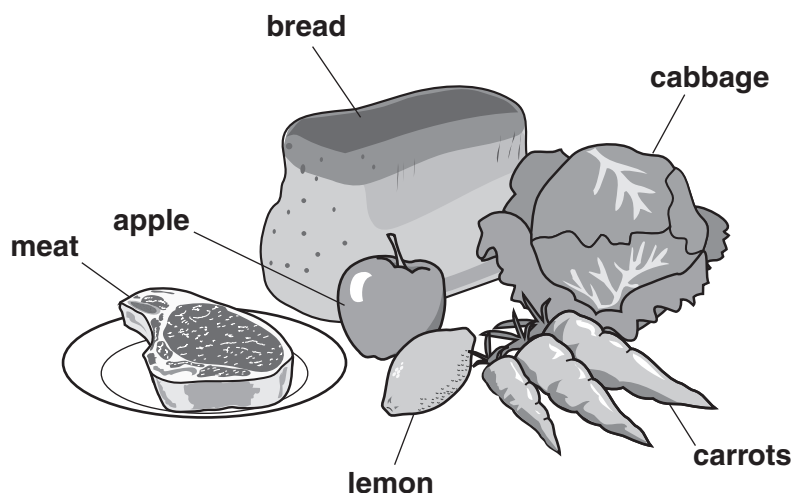
..... [2]

[Total: 4]

4

- 2 This question is about cooking and foods.

Look at the picture of some foods.



- (a) Write down the name of one food that contains a lot of **protein**.

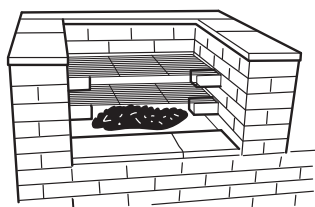
Choose from the foods in the picture.

..... [1]

- (b) Some of the foods in the picture need to be cooked before eating them.

Look at the picture of a barbecue.

This is used to cook food at a high temperature.



- (i) Write down **one** reason why some foods need to be cooked.

.....
 [1]

- (ii) Cooking food is an example of a chemical change.

Explain why.

.....
 [1]

[Total: 3]

5

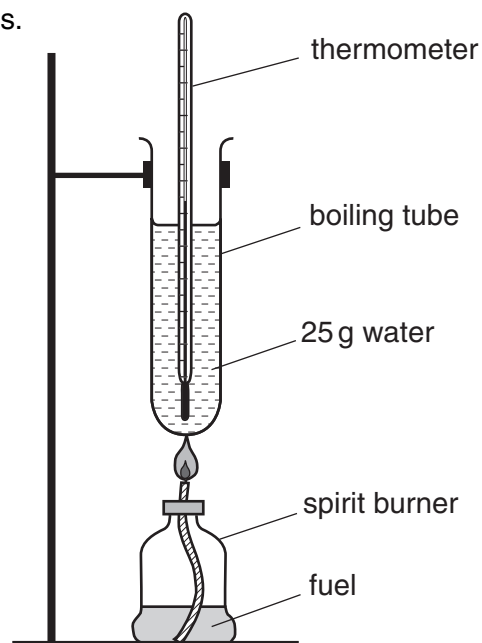
3 Luke and Sophie investigate the energy content of two fuels.

(a) (i) Look at the diagram.

It shows the apparatus they use.

They burn 1.0 g of fuel each time.

Look at their table of results.



fuel	starting temperature of water in °C	final temperature of water in °C
ethanol	20	35
paraffin	20	50

Calculate the amount of energy transferred when they burn 1.0 g of ethanol.

$$\text{energy} = \text{mass} \times \text{specific heat capacity} \times \text{temperature change}$$

The specific heat capacity of water is 4.2 J/g °C.

.....

answer J [3]

(ii) When 1.0 g of paraffin is burnt 3150 J is released.

Calculate the energy transferred in joules when burning 0.5 g of paraffin.

..... [1]

(b) Burning fuels is an **exothermic** reaction.

What is meant by an exothermic reaction?

..... [1]

[Total: 5]
 Turn over

6

4 This question is about polymers and plastics.

(a) Local councils find it difficult to dispose of plastics.

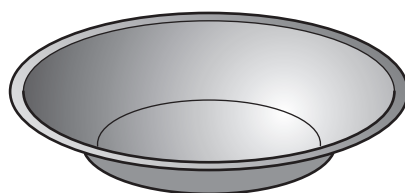
Explain why. Complete the table.

method of disposal	difficulties and problems
recycling	difficult to sort different types of plastics
burning	
landfill site	

[2]

7

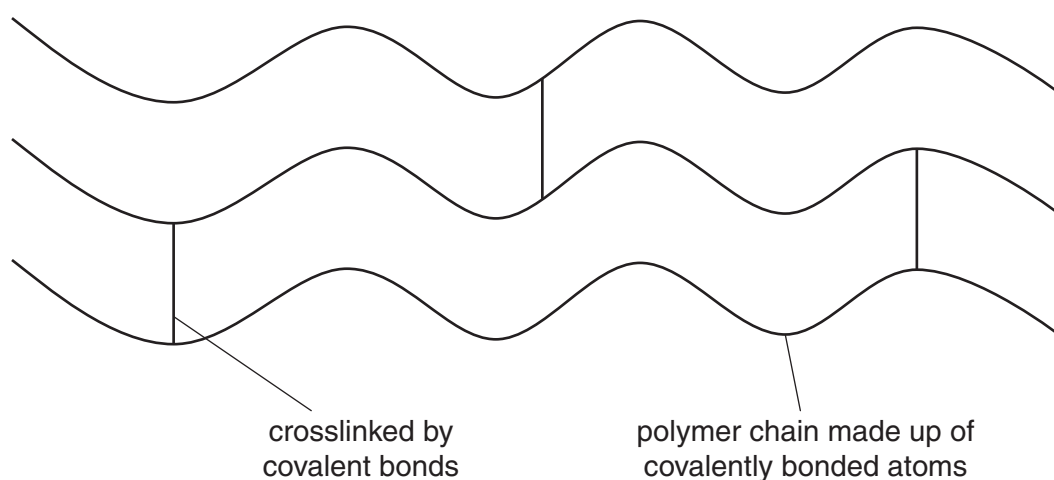
(b) Look at the picture.



The plastic container shown in the picture cannot be stretched.

This property is due to the way the polymer molecules are arranged.

The diagram shows how the polymer chains are arranged in the plastic.



Explain why the plastic cannot be stretched easily.

.....

.....

.....

.....

..... [3]

[Total: 5]

5 This question is about different types of chemical processes.

(a) Match the **process** with its correct **description**.

One has been done for you.

process	description
combustion	a reaction which converts large alkane molecules into smaller alkane and alkene molecules
cracking	a reaction which makes large molecules from many smaller molecules
fractional distillation	the separation of a mixture of hydrocarbons
polymerisation	a reaction in which carbon dioxide and water are made

[2]

(b) Alcohols react with acids to make an ester and water.

Write a **word** equation for this reaction.

..... [1]

[Total: 3]

Section B – Module C2

- 6 Limestone is a rock used to make buildings.

The chemical name for limestone is calcium carbonate, CaCO_3 .

When heated strongly calcium carbonate changes into calcium oxide.



This change is called **thermal decomposition**.

- (a) What is thermal decomposition?

.....
 [1]

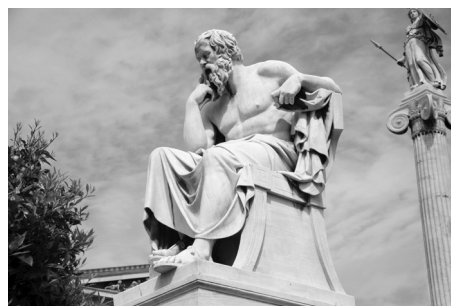
- (b) Write the balanced **symbol** equation for the thermal decomposition of calcium carbonate.

..... [1]

- (c) Look at the pictures. They show limestone and marble.



This building is made from limestone.
 Limestone is calcium carbonate.
 Limestone is a soft and crumbly rock.



This statue is made from marble.
 Marble is calcium carbonate.
 Marble is a hard rock.

Marble is harder than limestone.

Explain why. Use ideas about types of rock and how they are made.

.....

 [2]

[Total: 4]

Turn over

10

7 Iron reacts very slowly with dilute sulfuric acid.

The reaction makes iron sulfate and hydrogen.

(a) Write down the **word** equation for this reaction.

..... [1]

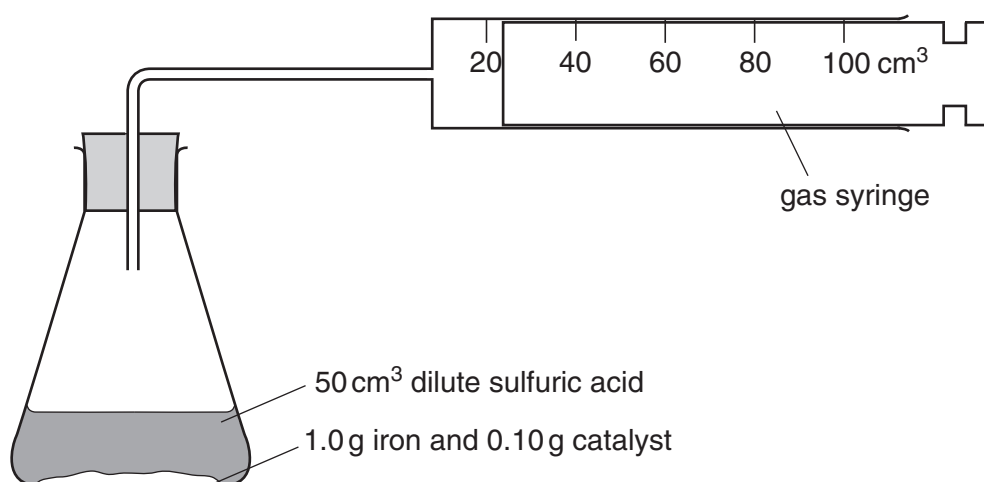
(b) Milly wants to make the reaction faster.

She knows that catalysts make reactions faster.

She tries to find a catalyst for this reaction.

Look at the diagram.

It shows the apparatus she uses.



She measures the time it takes to collect 100 cm³ of hydrogen in the gas syringe.

In experiments **2** to **5** she uses 0.10 g of catalyst each time.

In experiment **1** no catalyst is used.

Look at the results table.

Experiment number	name of catalyst	colour of catalyst at start of reaction	colour of catalyst at end of reaction	mass of catalyst at the start of reaction in grams	mass of catalyst left at the end of reaction in grams	time to collect 100 cm ³ of hydrogen in seconds
1	no catalyst added					130
2	copper powder	pink	pink	0.10	0.10	20
3	copper sulfate powder	blue	pink	0.10	0.04	15
4	calcium sulfate powder	white	white	0.10	0.10	130
5	zinc powder	silver	silver	0.10	0.05	10

Milly thinks that copper powder is a catalyst for this reaction.

Explain how Milly made this conclusion from her results.

.....

 [2]

(c) Milly uses iron powder rather than a lump of iron.

This is because the reaction is faster with iron powder.

Explain why the reaction is faster.

Use ideas about collisions between particles.

.....

 [2]

[Total: 5]

12

- 8 Steel is an alloy that contains iron and carbon.

Iron rusts much more easily than steel.

- (a) Write one **other** way in which steel is more useful than iron.

..... [1]

- (b) Look at the word equation for rusting.

It is not finished.

Fill in the names of the missing substances in the **word** equation.

iron + water + → [2]

- (c) Solder is an alloy.

Which **two** metals are found in solder?

Choose from:

iron

lead

mercury

tin

zinc

answer and [1]

- (d) Fizzy drinks cans are made from metal.

Look at the diagram. It shows a can of fizzy drink.



The metal used to make the can must be malleable.

This is a property of the metal.

Write down two **other** properties that the metal used to make this fizzy drinks can must have.

1

2 [2]

[Total: 6]

- 9 This question is about paints.

Look at the table. It shows the ingredients of two different paints.

ingredient	percentage	
	in gloss paint	in emulsion paint
additives	4	1
binder	52	21
extender	11	21
pigment	23	19
solvent	10	38

- (a) Describe **two** differences between this gloss paint and this emulsion paint.

1

2 [2]

- (b) Paints are not solutions. They are examples of a colloid.

The solid particles in these colloids are mixed with particles of a liquid.

The particles do not settle at the bottom. Explain why.

..... [1]

- (c) Two processes are involved when an oil paint dries.

The first process is solvent evaporation.

What is the second process?

..... [1]

- (d) Years ago 'glow in the dark' watches used a radioactive substance.

Now a phosphorescent pigment is used.

Suggest why.

..... [1]

[Total: 5]

11 Iron and copper are transition elements.

(a) Brahim adds a small volume of sodium hydroxide solution to five different solutions.

An insoluble solid called a precipitate is made each time.

Look at the results table. It is not finished.

solution	formula	colour of precipitate made
copper chloride	CuCl_2	blue
copper nitrate	$\text{Cu}(\text{NO}_3)_2$	
iron(II) chloride	FeCl_2	green
iron(II) sulfate	FeSO_4	green
iron(III) nitrate	$\text{Fe}(\text{NO}_3)_3$	

(i) Finish the table.

[2]

(ii) Look at the formulas in the table.

Which formula contains **six** oxygen atoms?

Choose from the table.

..... [1]

(iii) Solutions of copper chloride and sodium hydroxide react.

Copper hydroxide, $\text{Cu}(\text{OH})_2$, is made.

In this reaction copper ions, Cu^{2+} , react with hydroxide ions, OH^- .

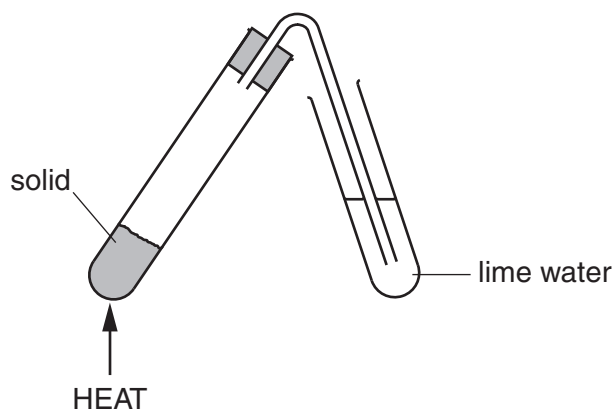
Write the **balanced ionic** equation for this reaction.

..... [2]

16

(b) Brahim investigates what happens when he heats some solids.

Look at the apparatus he uses.



Look at the results table.

solid	colour change of solid	effect on lime water
copper carbonate	green to black	goes milky
iron(II) sulfate	green to brown	stays colourless
potassium carbonate	stays white	stays colourless
zinc carbonate	white to yellow and back to white	goes milky

When heated, copper carbonate, CuCO_3 , makes two products.

What are the names of these two products?

..... and [1]

[Total: 6]

12 This question is about the elements in Group 7.

These elements are called the halogens.

(a) Look at the table. It shows information about some of the halogens.

element	atomic number	density in g/dm ³	melting point in °C	atomic radius in pm
chlorine	17	1.56	−101	99
bromine	35	2.93	−7	114
iodine	53	4.93	114	133

(i) Fluorine is a halogen with an atomic number of 9.

Predict the melting point of fluorine.

answer °C [1]

(ii) Astatine is a halogen with an atomic number of 85.

Predict the atomic radius for astatine.

answer pm [1]

(b) The reactivity of the halogens changes as the atomic number increases.

Describe how.

..... [1]

(c) Fluorine is bubbled through a solution of potassium iodide.

Predict the **names** of the **two** products of this reaction.

..... and [1]

(d) Look at the table. It shows information about two isotopes of chlorine.

	isotope 1	isotope 2
atomic number	17	17
mass number	35	37
number of protons	17	17
number of neutrons	18	20

What is an isotope? Use information from the table to help you.

.....

..... [1]

[Total: 5]

Turn over

18

13 This question is about alkali metals and their compounds.

(a) Hannah tests some metal compounds.

She uses the flame test.

Look at Hannah's results.

metal compound	colour of flame
potassium chloride	lilac
sodium chloride	yellow
lithium chloride	red

Describe how to do a flame test.

Include a labelled diagram of the apparatus she uses.

.....

.....

.....

..... **[3]**

19

- (b) Sodium chloride is an ionic compound.

Solid sodium chloride does not conduct electricity.

Explain why.

.....
..... [1]

- (c) Sodium has an electronic structure of 2.8.1.

Oxygen has an electronic structure of 2.6.

Sodium reacts with oxygen to make sodium oxide, Na_2O .

Sodium oxide is an ionic compound.

Draw 'dot and cross' diagrams to show the ions in sodium oxide.

Include the charges on the ions.

[2]

[Total: 6]

END OF QUESTION PAPER



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The Periodic Table of the Elements

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1	2	Key										3	4	5	6	7	0						
		relative atomic mass atomic symbol name atomic (proton) number																1 H hydrogen 1		4 He helium 2			
7 Li lithium 3	9 Be beryllium 4																	11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10
23 Na sodium 11	24 Mg magnesium 12																	27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36						
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44	103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54						
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86						
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated												

Key

relative atomic mass
atomic symbol
name
atomic (proton) number

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.