

Candidate Forename						Candidate Surname				
Centre Number						Candidate Number				

**OXFORD CAMBRIDGE AND RSA EXAMINATIONS
GENERAL CERTIFICATE OF SECONDARY EDUCATION**

B642/02

**GATEWAY SCIENCE
CHEMISTRY B**

**Unit 2 Modules C4 C5 C6
(Higher Tier)**

**WEDNESDAY 17 JUNE 2009: Morning
DURATION: 1 hour**

SUITABLE FOR VISUALLY IMPAIRED CANDIDATES

**Candidates answer on the question paper
A calculator may be used for this paper**

OCR SUPPLIED MATERIALS:

None

OTHER MATERIALS REQUIRED:

**Pencil
Ruler (cm/mm)**

READ INSTRUCTIONS OVERLEAF

INSTRUCTIONS TO CANDIDATES

- **Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes on the first page.**
- **Use black ink. Pencil may be used for graphs and diagrams only.**
- **Read each question carefully and make sure that you know what you have to do before starting your answer.**
- **Answer ALL the questions.**
- **Write your answer to each question in the space provided, however additional paper may be used if necessary.**

INFORMATION FOR CANDIDATES

- **The number of marks is given in brackets [] at the end of each question or part question.**
- **The total number of marks for this paper is 60.**
- **The Periodic Table is printed on the back page.**

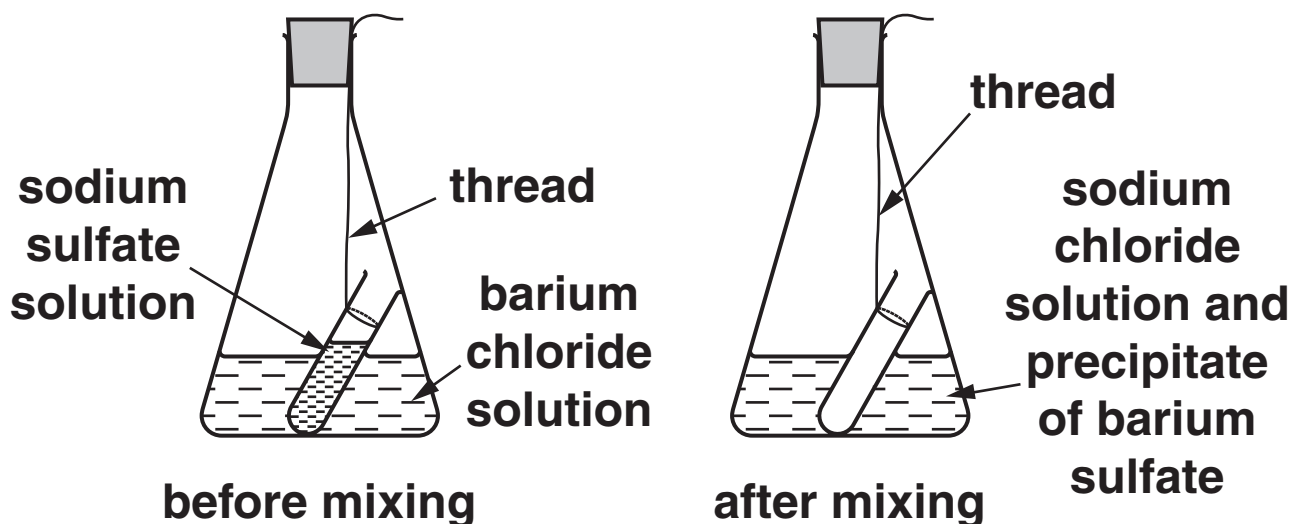
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Answer ALL the questions.

SECTION A – MODULE C4

1 Nick and Sloane investigate precipitation reactions.

Look at the apparatus they use.



They record the total mass of the flask, test-tube and contents.

They then take the flask off the balance and tip the flask upside down.

Nick and Sloane are careful not to let any liquid leak out of the flask.

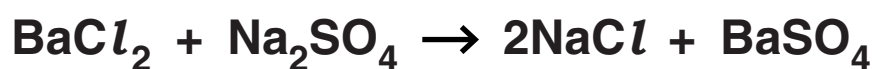
The solutions mix and react. A precipitate is made.

They put the flask back on the balance and record the mass again.

(a) Look at the results of this experiment.

	MASS IN GRAMS
mass of flask, test-tube and contents before mixing	142.6
mass of flask, test-tube and contents after mixing	142.6

Look at the balanced symbol equation for the reaction taking place.



Use the balanced symbol equation to help explain why the mass does not change during the reaction.

[1]

(b) Nick and Sloane do another experiment.

This time they use barium chloride solution and ammonium sulfate solution.

Complete the WORD equation opposite for this reaction.

[1]

barium
chloride + ammonium
sulfate →

+

(c) Look at the table.

It shows the formula of some of the compounds used by Nick and Sloane.

COMPOUND	FORMULA
ammonium sulfate	$(\text{NH}_4)_2\text{SO}_4$
barium chloride	BaCl_2
iron(III) sulfate	$\text{Fe}_2(\text{SO}_4)_3$
sodium sulfate	Na_2SO_4

- (i) How many ATOMS are there in the formula for ammonium sulfate?

_____ [1]

- (ii) The relative formula mass (M_r) of barium chloride is 208.

What is the relative formula mass of iron(III) sulfate, $\text{Fe}_2(\text{SO}_4)_3$?

The relative atomic mass (A_r) of O is 16, of S is 32, of Cl is 35.5 of Fe is 56 and of Ba is 137.

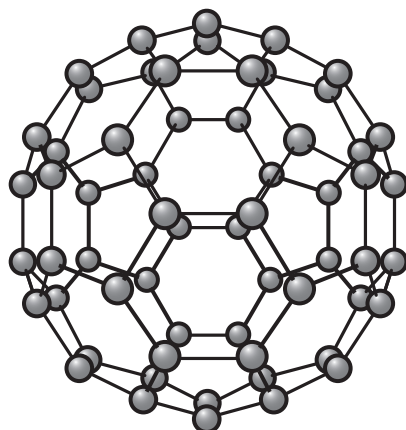
relative formula mass _____ [1]

[Total: 4]

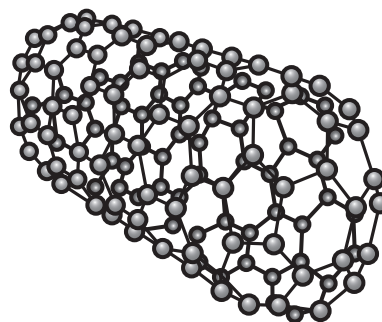
- 2 Buckminster fullerene and nanotubes are recently discovered substances.

Look at the diagrams.

They show the structures of Buckminster fullerene and of a nanotube.



BUCKMINSTER FULLERENE



A NANOTUBE

- (a) (i) What is the formula for a molecule of Buckminster fullerene?

_____ [1]

- (ii) Buckminster fullerene can 'cage' other molecules.

Describe ONE use of caged molecules.

_____ [1]

- (b) (i) Write down TWO reasons why nanotubes are used as catalysts.

1 _____

2 _____

_____ [2]

- (ii) Nanotubes are very strong and conduct electricity.

One use for nanotubes is as industrial catalysts.

Describe ONE OTHER use of nanotubes.

_____ [1]

[Total: 5]

3 Dirty clothes can be cleaned using a solvent.

(a) One way of using a solvent to clean clothes is called dry cleaning.

Suggest why it is called DRY cleaning.

_____ **[1]**

(b) Grease will dissolve in dry cleaning solvents.

Explain why.

Use ideas about

- **solvent molecules**
- **intermolecular forces**
- **molecules of grease.**

A labelled diagram may help you answer the question.

[2]

(c) Dry cleaning solvents are often harmful.

Eric decides to use solid carbon dioxide to dry clean some clothes.

Small lumps of solid carbon dioxide are fired at the clothing.

These dislodge the dirt.

Solid carbon dioxide quickly changes into a gas.

Suggest one ADVANTAGE of using solid carbon dioxide rather than a dry cleaning solvent.

[1]

[Total: 4]

4 A continuous process is used to make ammonia.

A batch process is used to make speciality chemicals such as medicines.

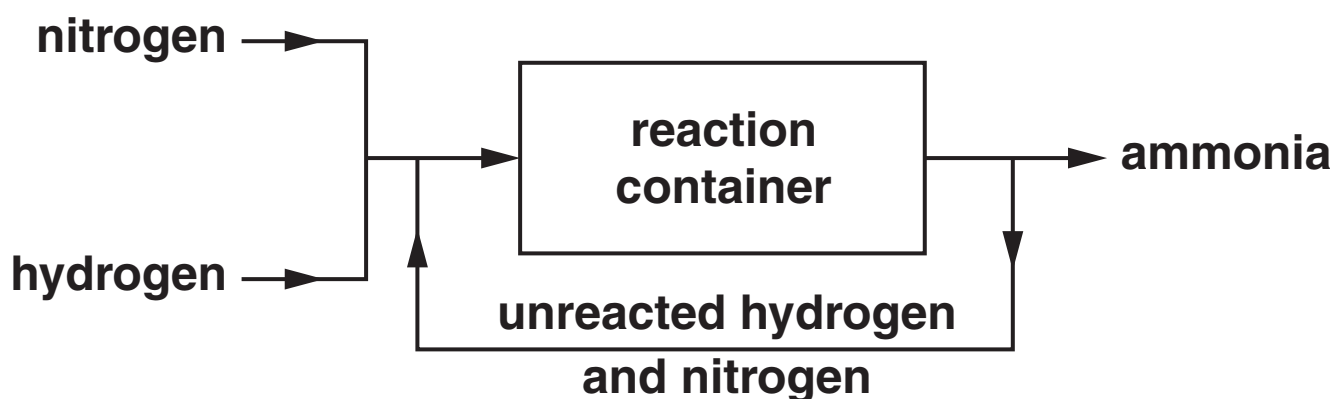
(a) Ammonia is made by a continuous process rather than a batch process. Suggest why.

_____ [1]

(b) Ammonia is made in the Haber process.

Look at the diagram.

It shows a flowchart of the Haber process.



Write about how ammonia is made in the Haber process.

Include in your answer

- the word equation for the reaction taking place in the Haber process
- the conditions used in the Haber process.

[3]

(c) Ammonium phosphate is made by reacting ammonia with an acid.

Which acid?

[1]

(d) Ammonia is used to make fertilisers.

Fertilisers increase crop yield.

Explain why fertilisers increase crop yield.

Use ideas about

- **essential elements**
- **plant protein.**

[2]

[Total: 7]

SECTION B – MODULE C5

5 Josh looks at this label on his packet of cornflakes.

It shows some information about 100g of cornflakes.

NUTRIENT	MASS OF INGREDIENT IN MILLIGRAMS	PERCENTAGE OF RECOMMENDED DAILY ALLOWANCE (RDA)
folic acid	0.2	85
iron	7.9	55
niacin	13.2	75
vitamin B1	0.7	50

(a) What mass of cornflakes will contain the RDA for vitamin B1?

mass of cornflakes = _____ g [1]

(b) Look at the table.

It gives information about the amount of sodium and of salt in 100 g of cornflakes.

	MASS IN GRAMS
sodium	0.7
salt	1.8

The RDA for salt is 6.0 g.

- (i) What is the percentage of the RDA of SALT in 100 g of cornflakes?

percentage = _____ [1]

- (ii) Show by calculation that all of the sodium in cornflakes comes from salt, NaCl.

The relative atomic mass of Na is 23 and of Cl is 35.5.

[2]

[Total: 4]

6 Insoluble compounds can be prepared by using a precipitation reaction.

(a) Emma wants to prepare a pure, dry sample of lead iodide.

She mixes potassium iodide solution and lead nitrate solution in a beaker.

An insoluble precipitate of lead iodide is made.

Describe the next steps Emma must do to get a PURE, DRY sample of lead iodide.

A labelled diagram may help you answer this question.

[3]

(b) Potassium iodide solution, KI, reacts with lead nitrate solution, $\text{Pb}(\text{NO}_3)_2$.

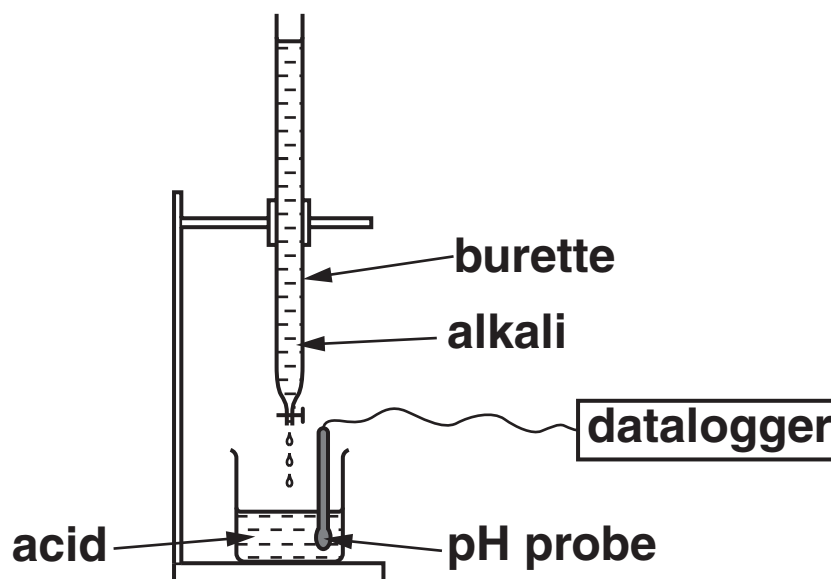
Lead iodide, PbI_2 , and potassium nitrate, KNO_3 , are made.

Write down the BALANCED SYMBOL equation for this reaction.

[2]

[Total: 5]

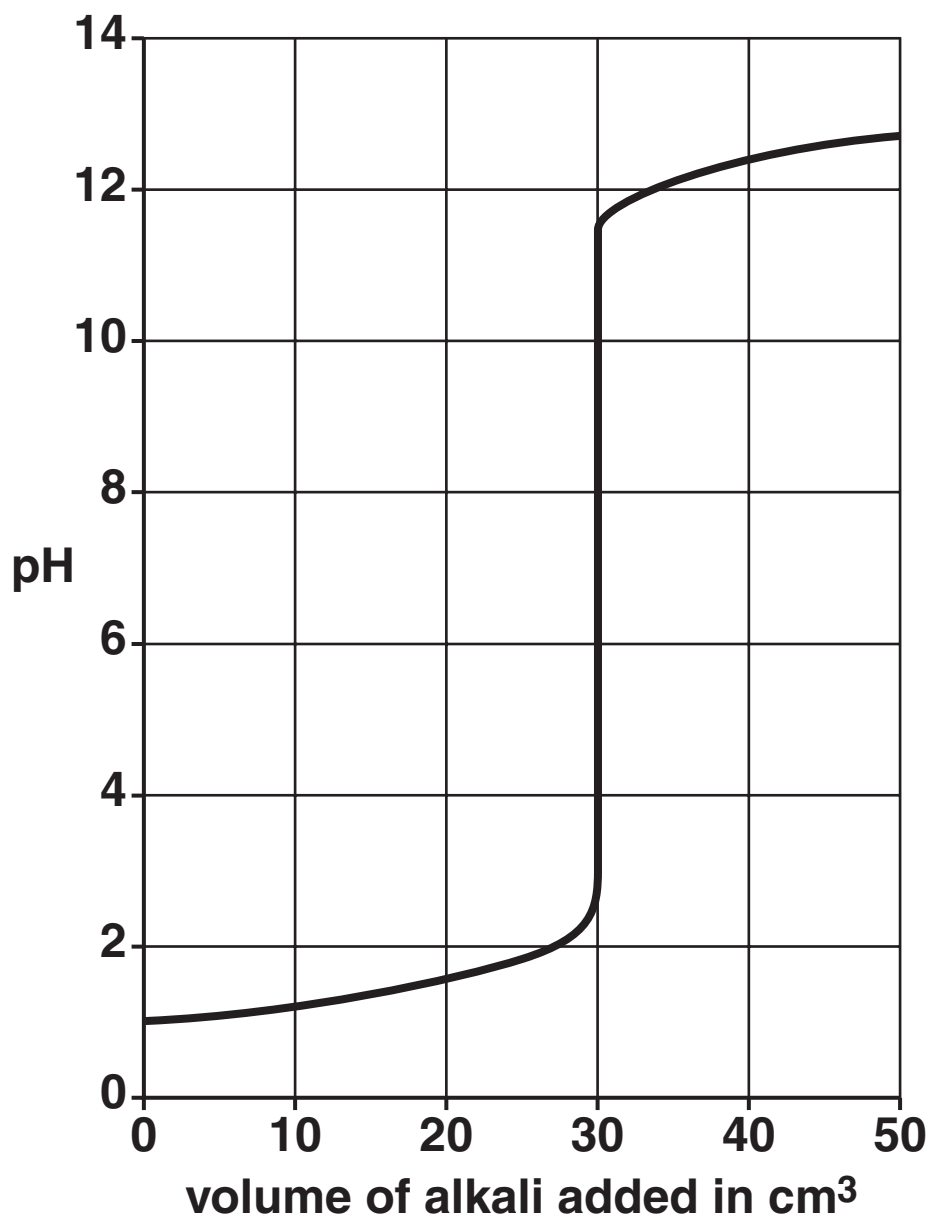
- 7 Tina investigates how the pH value changes as an acid is neutralised by an alkali.
Look at the apparatus she uses.



Sodium hydroxide solution is added slowly into the beaker of dilute sulfuric acid.

The pH probe is connected to a datalogger.

Look at the display opposite from the datalogger.



- (a) (i) What is the pH value when 15.0 cm^3 of sodium hydroxide has been added?

[1]

- (ii) What volume of alkali is needed to exactly neutralise the sulfuric acid?

_____ cm^3

[1]

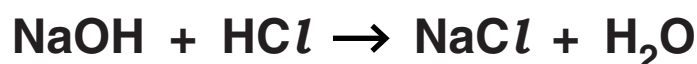
(b) Tina does another experiment.

This time she uses

- **25.0 cm³ of dilute hydrochloric acid in the beaker**
- **sodium hydroxide solution of concentration 0.100 mol/dm³ in the burette.**

The hydrochloric acid is exactly neutralised by 20.0 cm³ of this sodium hydroxide solution.

Look at the balanced symbol equation for the reaction.



Calculate

- **the number of moles of sodium hydroxide in 20.0 cm³ of a 0.100 mol/dm³ solution**
- **the number of moles of hydrochloric acid that reacted with this amount of sodium hydroxide**
- **the concentration, in mol/dm³, of the hydrochloric acid.**

concentration of hydrochloric acid =

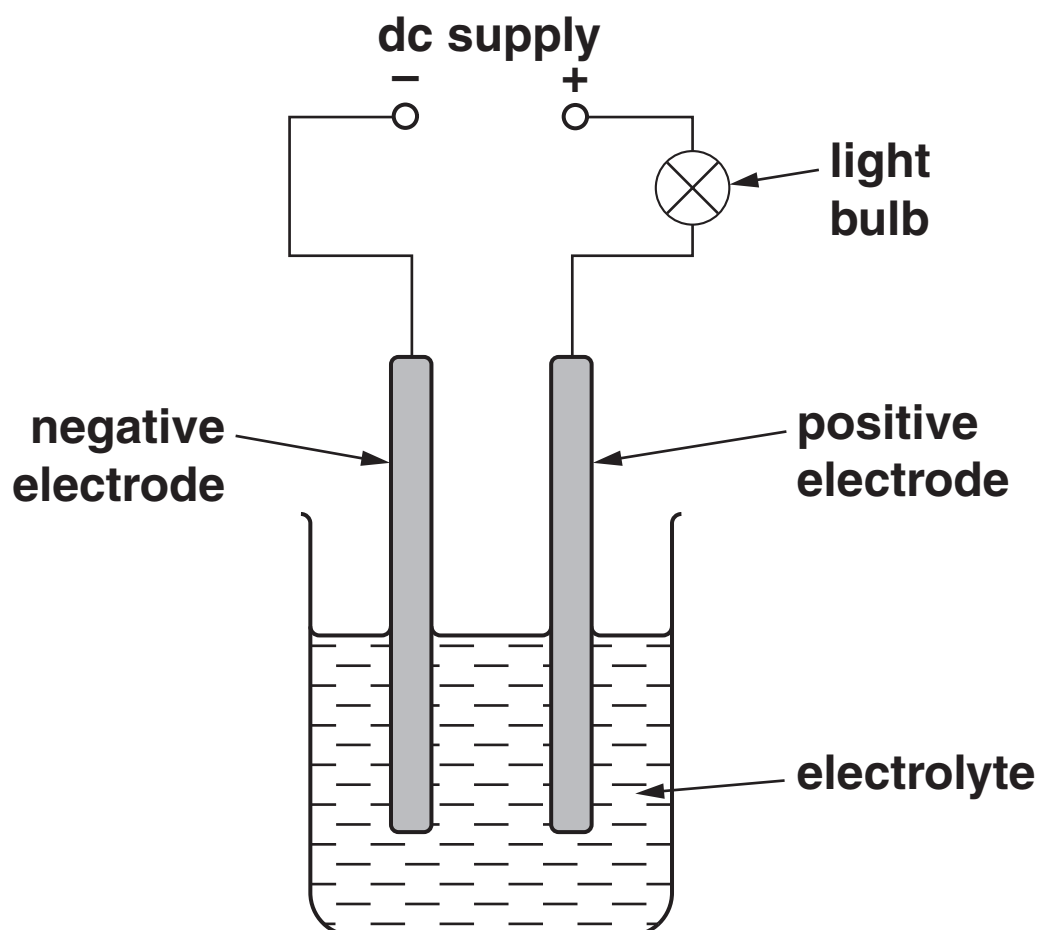
_____ **mol/dm³** **[3]**

[Total: 5]

8 This question is about electrolysis.

Look at the diagram.

It shows the apparatus that can be used to electrolyse solutions.



When a current flows the light bulb will glow.

The more a solution conducts the brighter the bulb.

(a) Luke uses potassium nitrate solution as the electrolyte.

A gas is made at the negative electrode.

What is the name of this gas?

Choose from the list.

HYDROGEN

NITROGEN

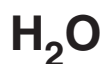
NITROGEN DIOXIDE

OXYGEN

answer _____ [1]

(b) Dilute ethanoic acid contains particles.

Look at the list of particles found in dilute ethanoic acid.



Luke uses ethanoic acid as the electrolyte.

Hydrogen is made at the negative electrode.

Explain why.

_____ [1]

(c) Luke uses hydrochloric acid as the electrolyte.

The light bulb glows much brighter than with dilute ethanoic acid of the same concentration.

Explain why.

Use ideas about

- **strong and weak acids**
- **ions.**

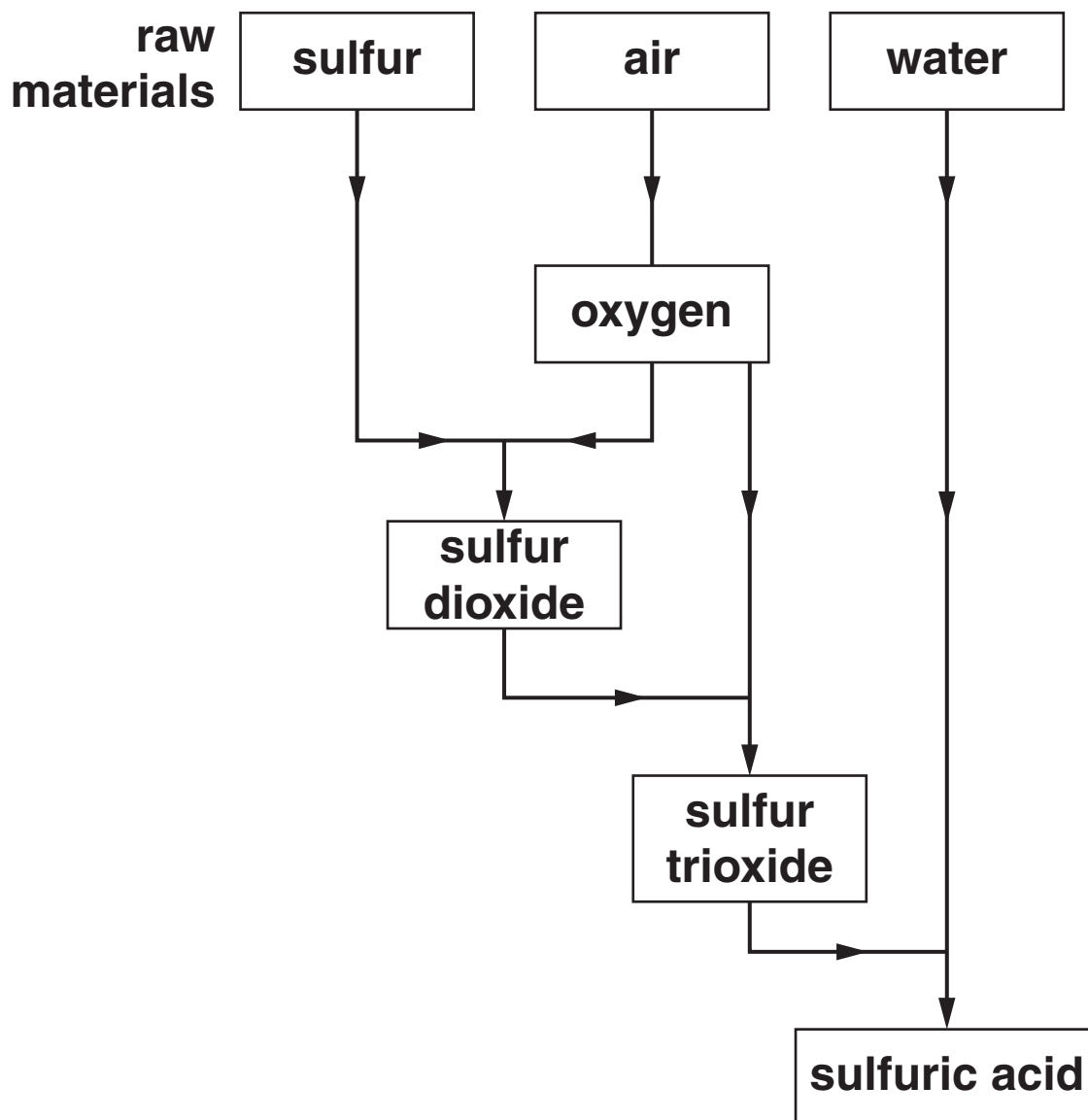
[2]

[Total: 4]

9 Sulfuric acid is made in the Contact Process.

Look at the flow chart.

It shows all the stages in the Contact Process.



There are three stages in the Contact Process.

- Stage 1 involves the reaction of sulfur and oxygen.
- Stage 2 involves the reaction of sulfur dioxide and oxygen.



- Stage 3 involves the reaction of sulfur trioxide and water.

Write the balanced symbol equations for stages 1 and 3.

stage 1 _____

stage 3 _____ [2]

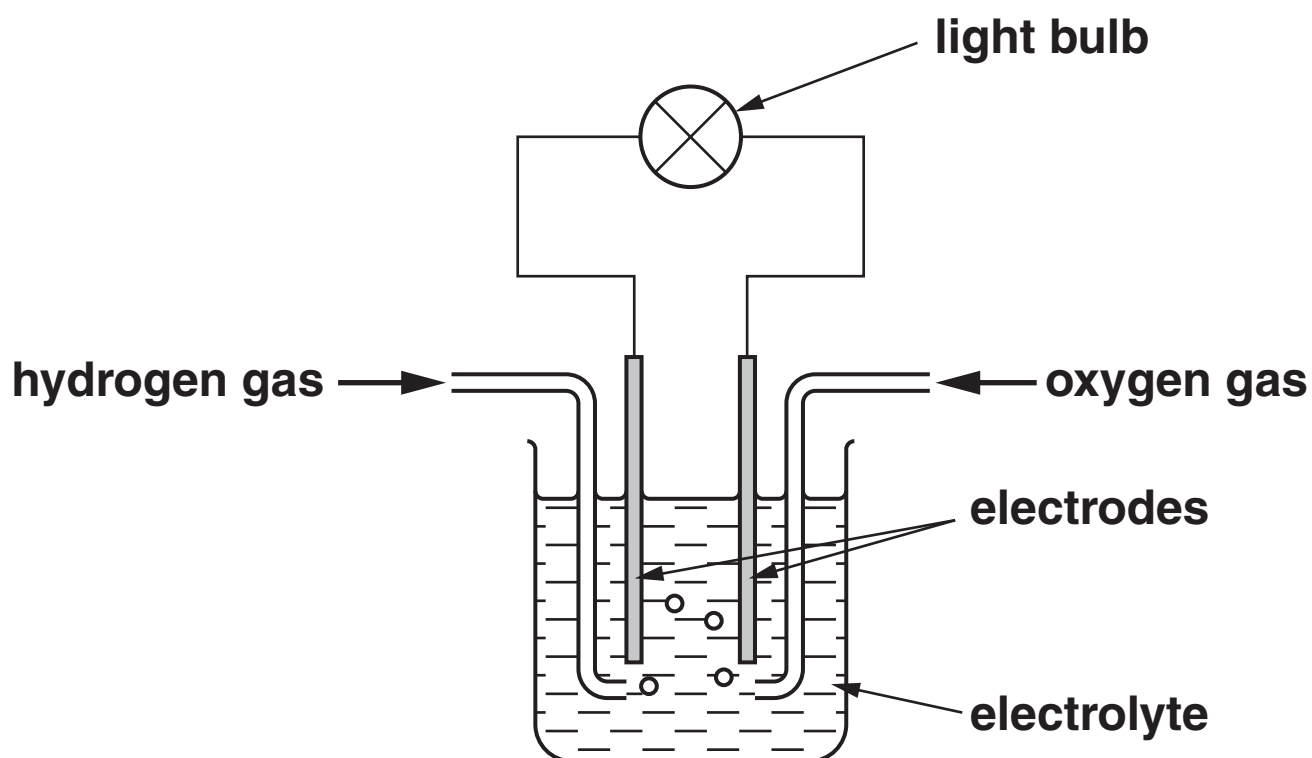
[Total: 2]

SECTION C – MODULE C6

10 This question is about fuel cells.

Look at the diagram.

It shows a fuel cell.



In a fuel cell, hydrogen and oxygen react to release energy.

(a) Hydrogen reacts with oxygen to make water.

Write a WORD equation for this reaction.

_____ [1]

- (b) Car makers may replace diesel or petrol engines with fuel cells.

Other than cost, write down TWO reasons for doing this.

1 _____

2 _____ [2]

- (c) Hydrogen gas, H_2 , reacts at one electrode in a fuel cell.

The reaction makes hydrogen ions, H^+ , and electrons.

Write an equation for this reaction. Use e^- to show an electron.

_____ [2]

- (d) Hydrogen can explode when it reacts with oxygen.

Heat energy is given out.

What type of reaction gives out heat?

Choose from the list.

DEHYDRATION

ELECTROLYSIS

ENDOTHERMIC

EXOTHERMIC

answer _____ [1]

[Total: 6]

11 This question is about reactions of metals.

(a) This question is about a large statue made from iron. It is going rusty.

Covering iron in oil or grease stops iron rusting.

(i) Explain why.

_____ **[1]**

(ii) The statue can also be coated with zinc.

This is called galvanising.

Explain why galvanising stops iron rusting.

_____ **[2]**

(b) Look at the equation.



The equation shows what happens when iron starts to rust.

What sort of reaction is this?

Choose from the list.

OXIDATION

REDUCTION

SAPONIFICATION

TINNING

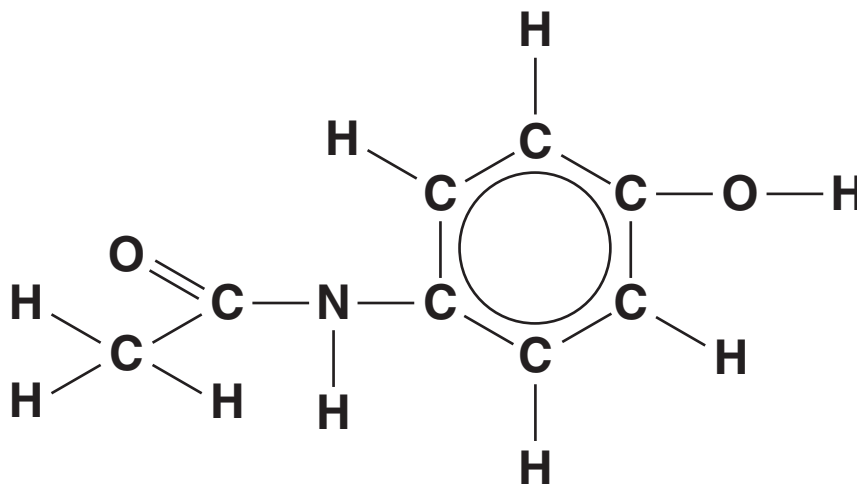
answer _____ [1]

[Total: 4]

12 This question is about drugs.

(a) Look at the diagram.

It shows the displayed formula of paracetamol.



Write down the MOLECULAR FORMULA of paracetamol.

_____ [2]

(b) Taking an overdose of aspirin can be dangerous.

Suggest why.

_____ [1]

(c) Aspirin is now available in a soluble form.

What is the advantage of using soluble aspirin?

_____ [1]

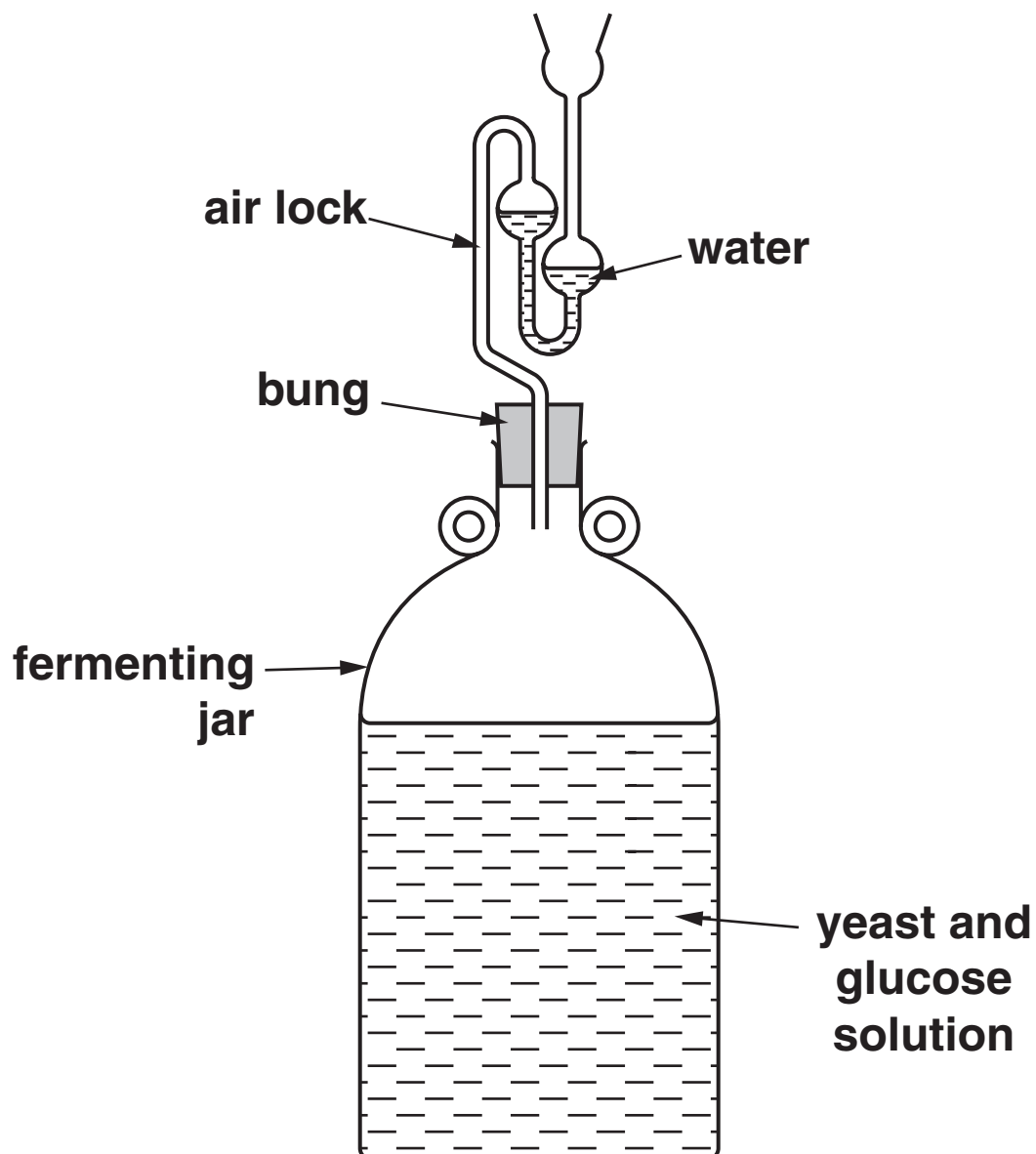
[Total: 4]

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13 This question is about ethanol.

Look at the diagram.

It shows how ethanol can be made by fermentation in a school laboratory.



(a) This reaction works best at 40 °C in the absence of air.

(i) The reaction does not work at 70 °C.

Explain why.

_____ [1]

(ii) The reaction is very slow at 10 °C.

Explain why.

_____ [1]

(iii) Air must be kept out of the fermenting jar.

Explain why.

_____ [1]

- (b) At the end of the process, the fermenting jar contains a mixture of ethanol and water.

How can the ethanol be separated from the water?

Choose from the list.

DISTILLATION

ELECTROLYSIS

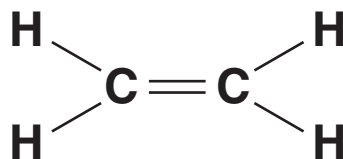
FILTRATION

SAPONIFICATION

answer _____ [1]

(c) The molecular formula of ethene is C_2H_4 .

Look at the DISPLAYED formula of ethene.



The molecular formula of ethanol is $\text{C}_2\text{H}_5\text{OH}$.

Draw the DISPLAYED formula of ethanol.

[1]

(d) The general formula for an alcohol is $C_nH_{2n+1}OH$.

Pentanol is an alcohol.

A molecule of pentanol contains five carbon atoms.

Write down the MOLECULAR FORMULA of pentanol.

_____ [1]

[Total: 6]

END OF QUESTION PAPER

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The Periodic Table of the Elements

1	2	Key					3	4	5	6	7	0						
1 H hydrogen 1		relative atomic mass atomic symbol name atomic (proton) number										11 B boron 5		12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10
7 Li lithium 3	9 Be beryllium 4											27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18	
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36	
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44	103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54	
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86	
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated							

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.