

OXFORD CAMBRIDGE AND RSA EXAMINATIONS
GCSE

A171/02

TWENTY FIRST CENTURY SCIENCE
CHEMISTRY A/SCIENCE A

Modules C1 C2 C3 (Higher Tier)

THURSDAY 18 MAY 2017: Morning

DURATION: 1 hour
plus your additional time allowance

MODIFIED ENLARGED

Candidate forename		Candidate surname	
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Centre number						Candidate number				
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Candidates answer on the Question Paper.
A calculator may be used for this paper.

OCR SUPPLIED MATERIALS:
A copy of The Periodic Table

OTHER MATERIALS REQUIRED:
Pencil
Ruler (cm/mm)

READ INSTRUCTIONS OVERLEAF



INSTRUCTIONS TO CANDIDATES

Write your name, centre number and candidate number in the boxes on the first page. Please write clearly and in capital letters.

Use black ink. HB pencil may be used for graphs and diagrams only.

Answer ALL the questions.

Read each question carefully. Make sure you know what you have to do before starting your answer.

Write your answer to each question in the space provided. If additional space is required, you should use the lined page(s) at the end of this booklet. The question number(s) must be clearly shown.

INFORMATION FOR CANDIDATES

The quality of written communication is assessed in questions marked with a pencil ().

The number of marks is given in brackets [] at the end of each question or part question.

The total number of marks for this paper is 60.

Any blank pages are indicated.

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- 1 The table shows some information about how the Earth's atmosphere has changed over time.

Gas	Approximate percentage composition of atmosphere in %		
	4 billion years ago	500 years ago	Today
Carbon dioxide	20	0.03	0.04
Water vapour	50	small	small
Nitrogen	3	78	78
Oxygen	0	21	21

- (a) The atmosphere 4 billion years ago contained other gases in addition to those named in the table. The other gases contained mainly methane.

Use the data in the table to estimate the percentage of methane gas in the atmosphere 4 billion years ago.

Explain your reasoning.

[2]

- (b) Describe and give reasons for the changes in the percentages of carbon dioxide, water vapour and oxygen from 4 billion years ago until today.**

Use data from the table to support your answer.



The quality of written communication will be assessed in your answer.

[6]

[TOTAL: 8]

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2 The exhaust gases of cars contain pollutants.

(a) The pollutants include nitrogen monoxide and carbon monoxide.

(i) Describe how nitrogen monoxide is formed in a car engine.

[2]

(ii) Describe how carbon monoxide is formed in the car engine.

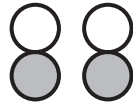
[1]

- (iii) Nitrogen monoxide and carbon monoxide are removed from exhaust gases in a catalytic converter.**

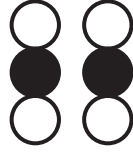
In the converter, nitrogen monoxide reacts with carbon monoxide to form carbon dioxide and nitrogen.

Complete the diagram to show the missing molecules. [3]

nitrogen monoxide + carbon monoxide → carbon dioxide + nitrogen

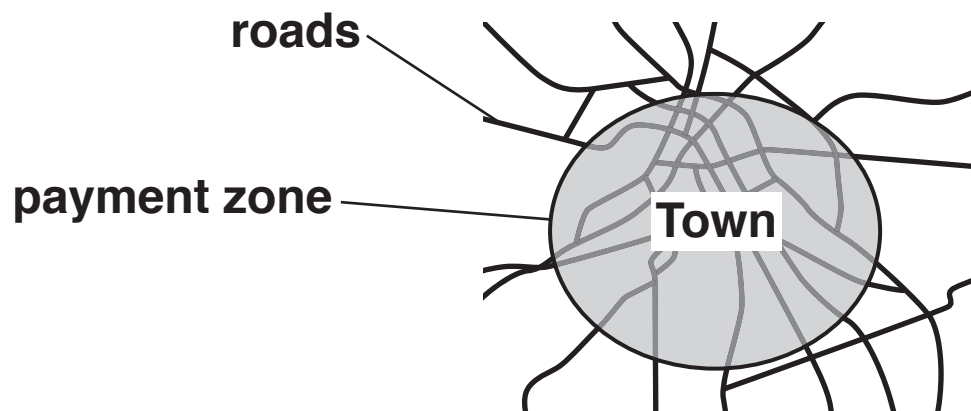


+



+

- (b) A town council wanted to reduce the amount of air pollutants in a town.
The council decided to introduce a payment zone for cars.**



Alex works for the town council.

Alex measured the amount of pollutants in the air inside the payment zone and outside the payment zone.

He recorded data every day for a year before the payment was introduced and every day for a year afterwards.

The table shows Alex's data.

Site	Pollutant	Daily mean amount before the payment was introduced in $\mu\text{g}/\text{m}^3$	Daily mean amount after the payment was introduced in $\mu\text{g}/\text{m}^3$	Percentage change in %
Outside the payment zone	nitrogen oxides	560	476	-15
	carbon monoxide	25	22	-12
Inside the payment zone	nitrogen oxides	600	480	-20
	carbon monoxide	30	24	-20

Suzy and Martin talk about the data in the table.

Suzy says 'There is no need to introduce a payment zone. Air pollution is decreasing anyway.'

Martin says 'The table shows that the payment is helping to lower air pollution.'

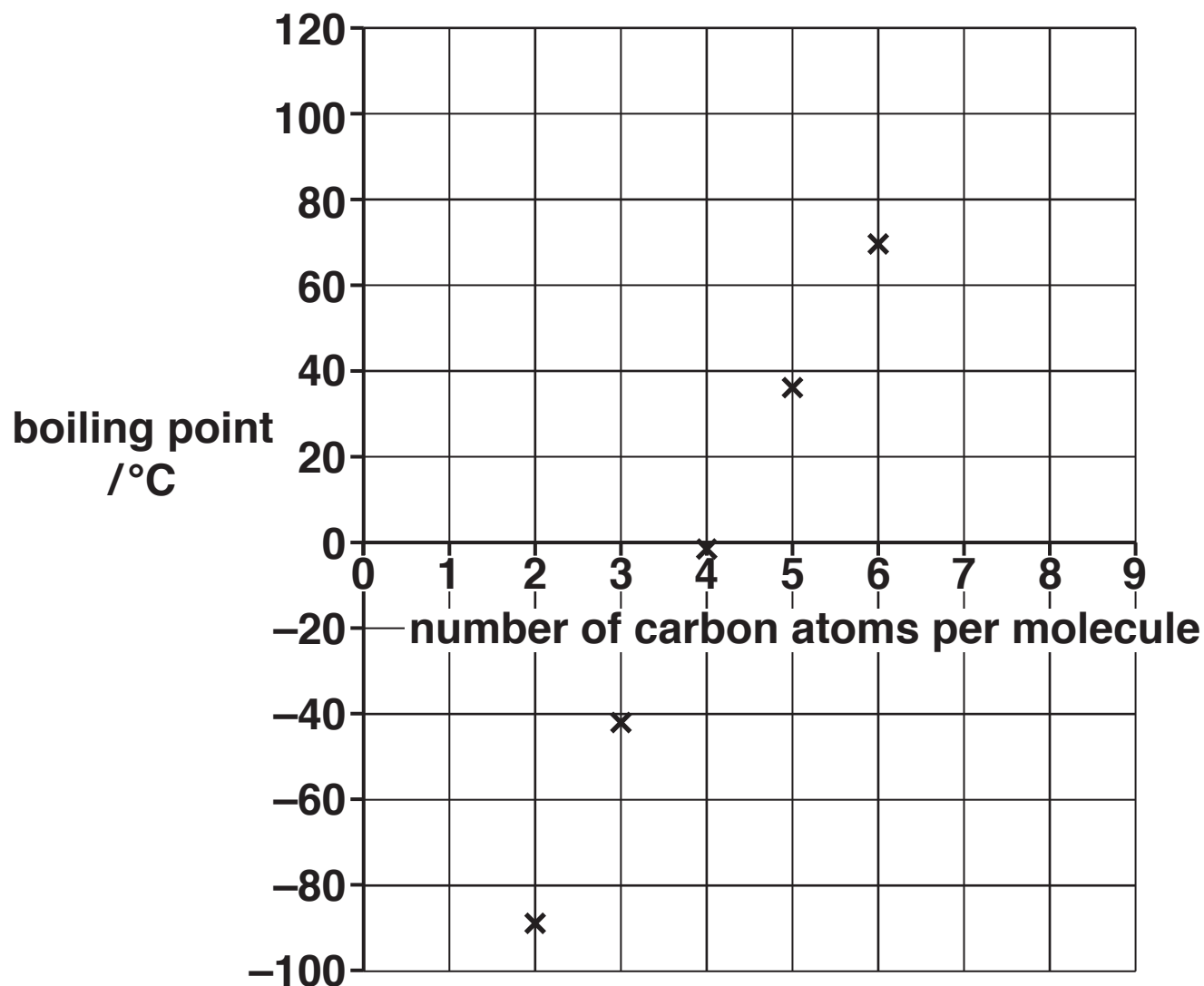
Explain how the data in the table supports the ideas of both Suzy and Martin.

[3]

[TOTAL: 9]

3 Crude oil contains hydrocarbons.

The graph shows the relationship between the number of carbon atoms in some hydrocarbons and their boiling points.



Describe the relationship shown by the graph and use ideas about forces between molecules to explain this relationship.



The quality of written communication will be assessed in your answer.

[6]

[TOTAL: 6]

4 Nanoparticles are very small particles.

(a) Put a ring around the correct range for the size of nanoparticles. [1]

0.1 to 1 nm

1 to 100 nm

100 to 200 nm

200 to 1000 nm

(b) Which statements about nanoparticles are TRUE and which are FALSE?

Put a tick (✓) in one box in each row. [2]

	True	False
Nanoparticles can be used to make sports equipment stronger.		
Nanoparticles can occur naturally.		
Nanoparticles have the same properties as larger particles.		
Nanoparticles are about the same size as some molecules.		

- (c) Doctors use stitches to hold together large cuts so that they can heal properly.**

A hospital is considering buying a new type of material to use for stitches.

They need to choose between a material that contains silver nanoparticles and a material that does not.

Which material should they choose?

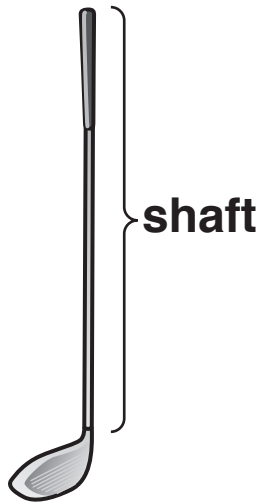
Justify your answer by explaining the risks and benefits of using each.

[3]

[TOTAL: 6]

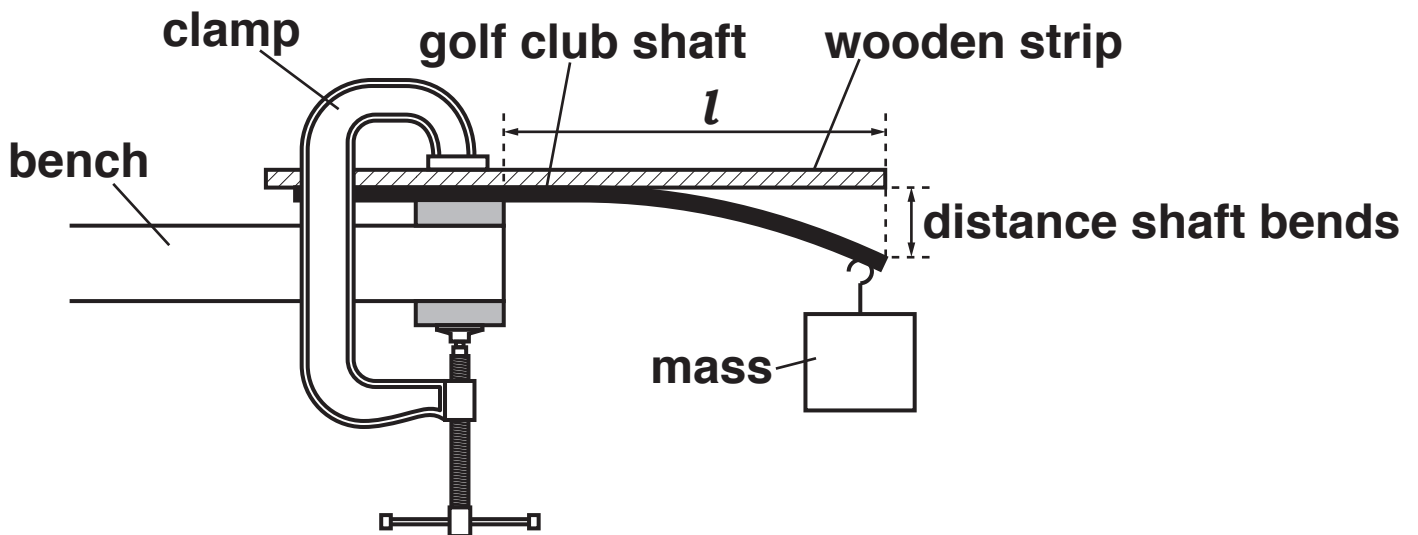
5 Chris works for a company that makes golf clubs.

The flexibility of the shaft of the golf club is important.



Golf clubs are given a Flex Rating as a measure of the flexibility of the shaft.

Chris measures the flexibility of a shaft using the following apparatus.



He measures the distance that the shaft bends when the mass is added.

- (a) Chris tests the flexibility of several different shafts.
He wants to make sure that his tests are 'fair tests'.

Give TWO factors that he needs to control when he tests each shaft and explain how these controls make his tests fair.

[3]

- (b) Chris repeats his test five times for the same shaft.

These are his results.

Distance shaft bends in mm				
Test 1	Test 2	Test 3	Test 4	Test 5
86	89	87	88	87

- (i) Calculate the mean value for the distance the shaft bends.

mean = _____ mm [2]

- (ii) The Flex Rating for a shaft can be worked out using the distance the shaft bends in metres. This is the formula:

$$\text{Flex Rating} = \frac{10}{3 \times \text{distance shaft bends in m}}$$

Ladies' golf club shafts must have a Flex Rating in the range 38–39.

Men's golf club shafts must have a Flex Rating in the range 45–46.

Is the shaft in (b)(i) suitable to be used for a men's or a ladies' golf club?

Use a calculation to support your answer.

[3]

[TOTAL: 8]

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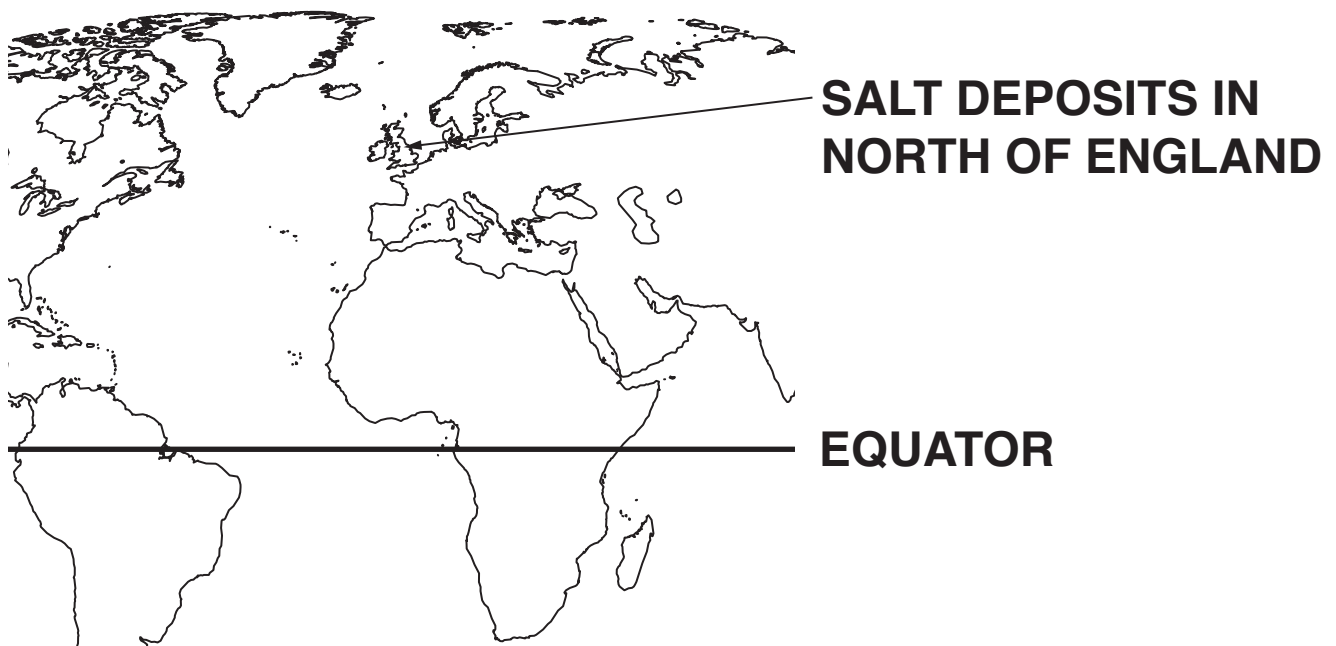
6 Large salt deposits are found deep underground in some parts of the World.

(a) Describe how these salt deposits formed.

[2]

(b) There are large salt deposits in the North of England.

Scientists think these salt deposits formed much nearer to the equator.



Suggest how the salt deposits came to be in the North of England if they were formed nearer to the equator.

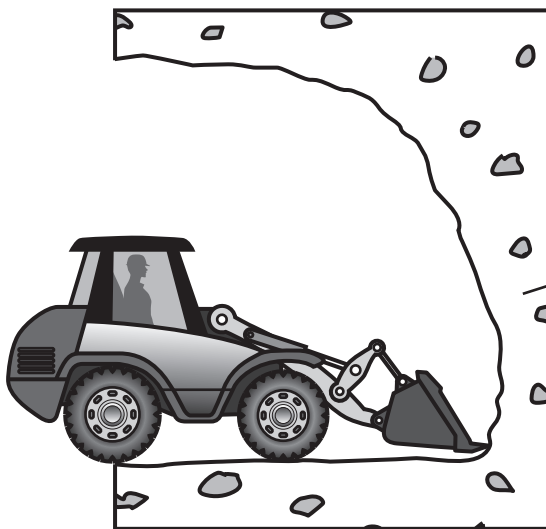
[2]

- (c) A company wants to extract the salt from underground and use it for making chemicals. Salt used for making chemicals needs to have a high purity.

The salt deposits are 200 m underground.
Salt can be extracted by two methods.

METHOD 1

Salt mixed with rocks is dug out from underground and brought up to the surface.

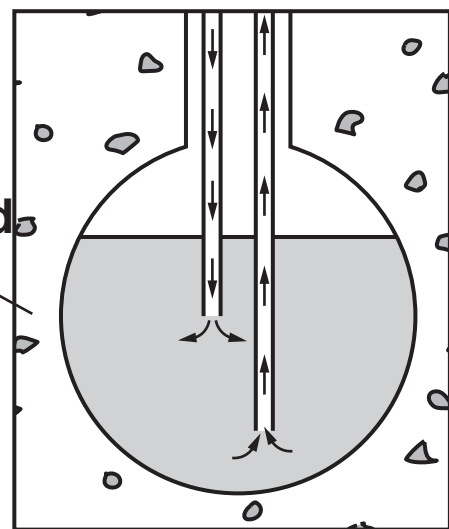


salt and
rock 200 m
underground

METHOD 2

Water is heated and pumped into the salt and rock. Salt dissolves and salt solution is pumped back to the surface.

hot water → salt solution



Compare the advantages and disadvantages of each method and explain which would be the best method to extract salt for making chemicals.



The quality of written communication will be assessed in your answer.

[illegible]

[TOTAL: 10]

7 Sodium carbonate was used as an alkali before the development of a modern chemical industry.

(a) (i) Give ONE example of how alkalis were used before the modern chemical industry developed.

_____ **[1]**

(ii) One traditional source of alkalis was from burnt wood.

Name another traditional source of alkalis.

_____ **[1]**

(b) In the 19th century sodium carbonate was made in a process which reacted sodium chloride (salt) and sulfuric acid with calcium carbonate (from limestone) and carbon (from coal).

The process had 2 stages

STAGE 1:

sodium chloride + sulfuric acid → sodium sulfate + hydrogen chloride

STAGE 2:

sodium sulfate + calcium carbonate + carbon → sodium carbonate + calcium sulfide + carbon dioxide

- (i) The process makes unwanted waste products that may cause harm to the environment.

One of these waste products is hydrogen chloride.

Name two OTHER waste products that are made.

1 _____

2 _____ [2]

- (ii) Waste hydrogen chloride from the process can be oxidised to make a useful chemical.

Give the name of this useful chemical and explain why it is useful.

_____ [2]

[TOTAL: 6]

- 8 PVC is a polymer that used to be used to make shoes and clothing.**



- (a) The chemical name for PVC is polychloroethene. Name the elements in PVC.**

_____ [2]

- (b) Plasticised PVC contains plasticisers to make it suitable for making clothing. Plasticisers change the properties of polymers.**

- (i) Explain how and why plasticisers change the properties of polymers.**

_____ [3]

- (ii) Some countries have banned the use of plasticised PVC for making containers for food or drinks.

Explain why polymers that contain plasticisers are not considered to be safe for making containers for food or drink.

[2]

[TOTAL: 7]

END OF QUESTION PAPER

ADDITIONAL ANSWER SPACE

If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).

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