

GENERAL CERTIFICATE OF SECONDARY EDUCATION

GATEWAY SCIENCE

ADDITIONAL SCIENCE B

Unit 2 Modules B4 C4 P4 (Higher Tier)

B624/02

Candidates answer on the Question Paper
A calculator may be used for this paper

OCR Supplied Materials:
None

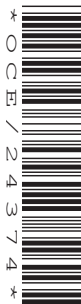
Other Materials Required:

- Pencil
- Ruler (cm/mm)

Monday 25 January 2010

Afternoon

Duration: 1 hour



Candidate Forename		Candidate Surname	
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Centre Number						Candidate Number				
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MODIFIED LANGUAGE

INSTRUCTIONS TO CANDIDATES

- Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **all** the questions.
- Do **not** write in the bar codes.
- Write your answer to each question in the space provided, however additional paper may be used if necessary.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- A list of physics equations is printed on page two.
- The Periodic Table is printed on the back page.
- The total number of marks for this paper is **60**.
- This document consists of **24** pages. Any blank pages are indicated.

2

EQUATIONS

$$\text{speed} = \frac{\text{distance}}{\text{time taken}}$$

$$\text{acceleration} = \frac{\text{change in speed}}{\text{time taken}}$$

$$\text{force} = \text{mass} \times \text{acceleration}$$

$$\text{work done} = \text{force} \times \text{distance}$$

$$\text{power} = \frac{\text{work done}}{\text{time}}$$

$$\text{kinetic energy} = \frac{1}{2}mv^2$$

$$\text{potential energy} = mgh$$

$$\text{weight} = \text{mass} \times \text{gravitational field strength}$$

$$\text{resistance} = \frac{\text{voltage}}{\text{current}}$$

3

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Question 1 begins on page 4.

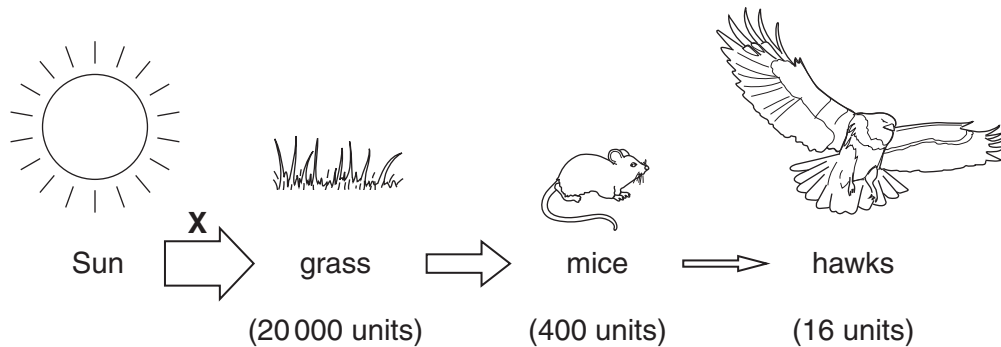
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Answer **all** the questions.

Section A – Module B4

- 1 Look at the diagram of a food chain.

The numbers show the amount of energy used for growth at each stage of the food chain.



- (a) Process **X** transfers energy from the Sun to the grass.

What is process **X**?

..... [1]

- (b) 2% of the energy in the grass is transferred to the mice.

This is a lower figure than the percentage of energy transferred from the mice to the hawks.

- (i) What percentage of the energy in the mice is transferred to the hawks?

answer% [2]

- (ii) **Not** all the energy at one stage of a food chain is transferred to the next stage.

Write down **one** reason why.

..... [1]

- (iii) The percentage of the energy transferred from the mice to the hawks is **more** than that transferred from the grass to the mice.

Suggest why.

.....
 [1]

5

(c) Look again at the diagram of a food chain.

A disease reduces the number of hawks.

What is likely to happen to the amount of grass?

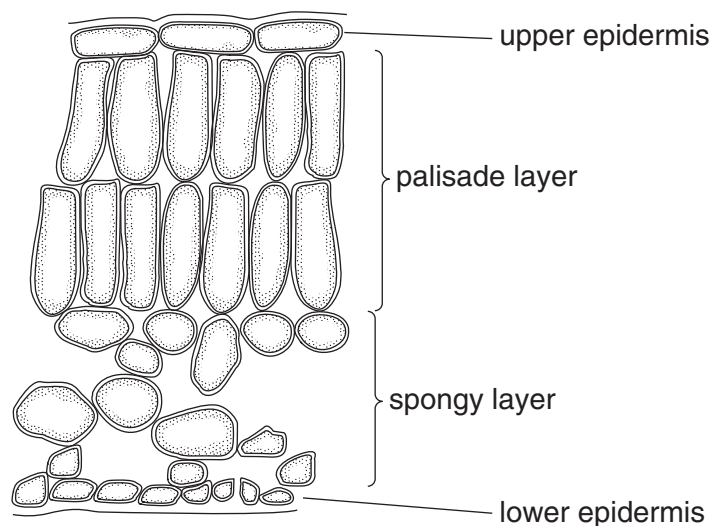
Explain your answer.

.....
..... [1]

[Total: 6]

6

2 The diagram shows a section through a leaf.



(a) The spongy layer contains air spaces.

Why are the air spaces important?

.....
 [1]

(b) (i) Which of the four layers labelled in the diagram contains most chloroplasts?

..... [1]

(ii) Chloroplasts contain chlorophyll.

What element found in minerals is present in chlorophyll?

Put a ring around the correct answer.

calcium **magnesium** **phosphorus** **potassium** [1]

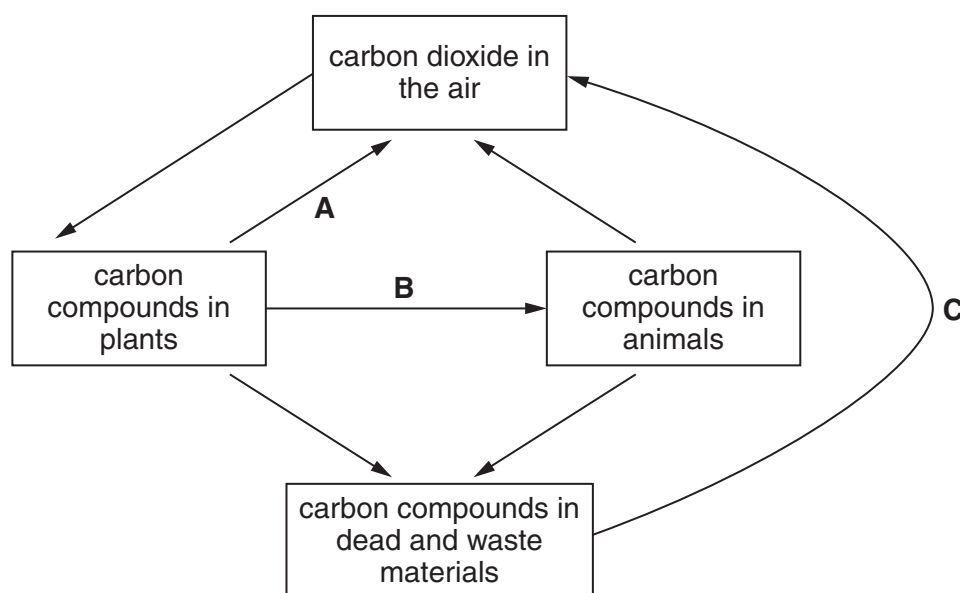
(iii) By what process do plants absorb minerals?

..... [1]

[Total: 4]

7

3 The diagram shows part of the carbon cycle.



(a) What process is shown by arrow **A**?

..... [1]

(b) What process is shown by arrow **B**?

..... [1]

(c) What type of organisms are responsible for the process shown by arrow **C**?

..... [1]

(d) There are other ways carbon dioxide is added to the air that are **not** shown on the diagram.

Write down **one other** way.

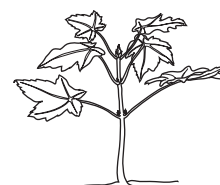
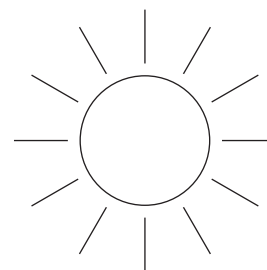
..... [1]

[Total: 4]

- 4 The diagram shows a plant on two different days.



rainy day



sunny day

- (a) Plants lose water from their leaves.

What is this process called?

..... [1]

- (b) On the sunny day it is warmer, there is more light and it is less humid.

The plant loses water more quickly.

Explain why.

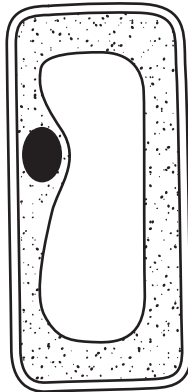
.....

 [3]

9

(c) The diagram shows a cell from the plant on the rainy day.

The cell is turgid.



On the sunny day the cell becomes plasmolysed.

Draw a diagram to show a plasmolysed cell.

Label your diagram to show the changes.

[2]

[Total: 6]

Section B – Module C4

5 This question is about fertilisers.

Fertilisers can be made by **neutralisation**.

(a) Complete the word equation to show what happens during neutralisation.

acid + base $\rightarrow$ + [2]

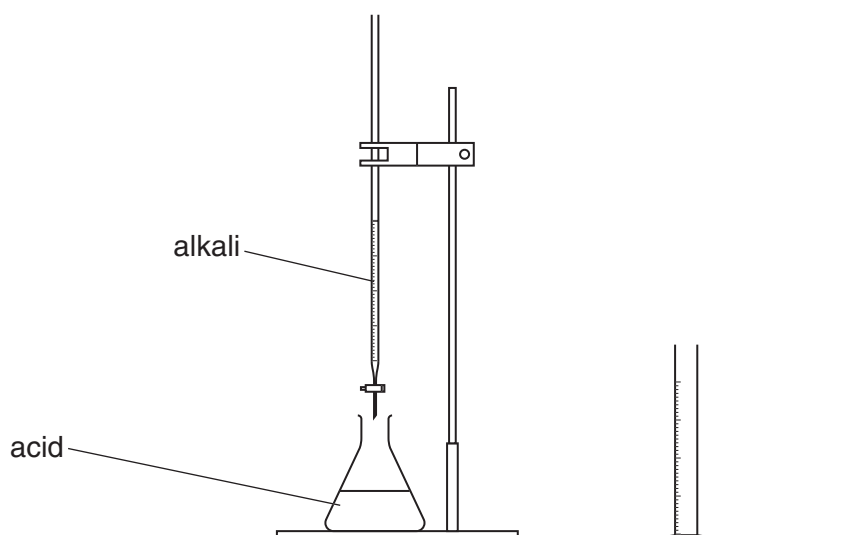
(b) Potassium hydroxide reacts with nitric acid.

What is the name of the fertiliser made?

..... [1]

(c) Acids react with alkalis when fertilisers are made.

Look at the diagram. It shows the equipment used.



Alkali from the burette is added to the acid in the flask.

Jo wants to make ammonium sulfate.

(i) Which acid and which alkali should she use?

acid used

alkali used [2]

(ii) Explain how Jo gets a **neutral** solution.

.....

.....

..... [2]

11

(iii) How does Jo get **solid** ammonium sulfate from the neutral solution?

..... [1]

(d) Acids contain hydrogen ions, H^+ . Alkalis contain hydroxide ions, OH^- .

Write an **ionic** equation for neutralisation.

..... [1]

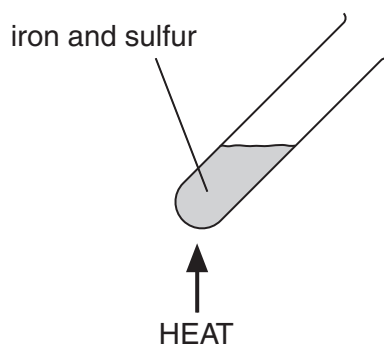
[Total: 9]

12

6 This question is about chemical calculations.

Jake and Monty make iron sulfide.

They heat a mixture of iron and sulfur.



(a) They make a sample of iron sulfide.

They predict that they will make 9.0 g.

They actually make 7.2 g.

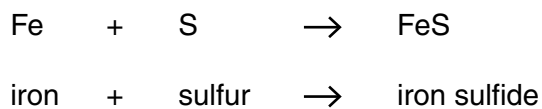
Calculate their percentage yield.

.....

answer%

[2]

(b) Look at the equation for the reaction.



Jake and Monty use 5.6 g of iron.

How much iron sulfide can they make?

The relative atomic mass of iron is 56 and of sulfur is 32.

.....

answer g

How much sulfur must Jake and Monty use?

.....

answer g

[2]

[Total: 4]

13

- 7 This question is about the manufacture of medicines.

Making and developing new medicines is very expensive.

One of the reasons for this is that less automation is possible.

- Write about **other** reasons why making and developing new medicines is very expensive.
- Explain your answers.

.....

.....

.....

.....

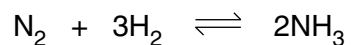
.....

..... [4]

[Total: 4]

- 8 Ammonia is made in the Haber process.

Look at the equation.



- (a) (i) One condition used in the Haber process is an iron catalyst.

Write down one **other** condition used.

..... [1]

- (ii) Why is a catalyst used?

..... [1]

- (b) Ammonia can be used to make urea, $\text{CO}(\text{NH}_2)_2$.

How many atoms are there in one molecule of urea, $\text{CO}(\text{NH}_2)_2$?

..... [1]

[Total: 3]

Section C – Module P4

- 9 Radioactive atoms can emit **three** different types of nuclear radiation.

One type is **alpha** (α) radiation.

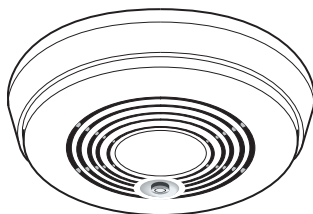
- (a) Write down the names of the **other** two types of nuclear radiation.

1

2 [2]

- (b) Americium-241 does not occur naturally.

It is a source of alpha radiation. It is used in smoke alarms.



- (i) Describe how a smoke detector containing americium-241 works.

Use ideas about **ionisation** to answer the question.

.....
.....
.....
.....
..... [2]

- (ii) Americium-241 is made when plutonium-241 decays.

Plutonium-241 is made in a nuclear reactor from plutonium-240.

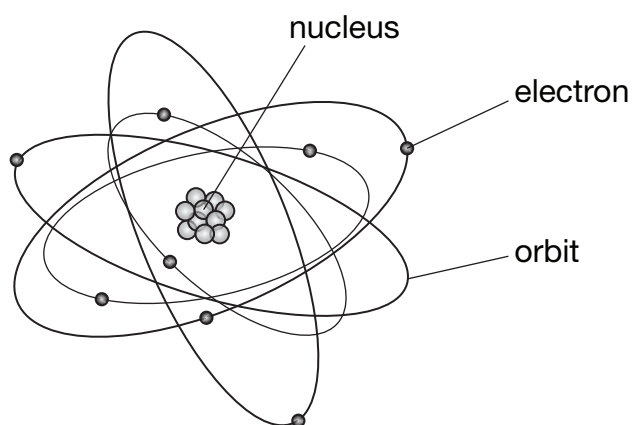
How does plutonium-240 change into plutonium-241 in a nuclear reactor?

..... [1]

[Total: 5]

16

10 (a) The diagram represents a radioactive atom.



Finish the sentences by choosing the **best** words from this list.

atom

electron

nucleus

orbit

stable

unstable

Radiation comes from the nucleus of the atom.

The radioactive atom is

The atom may emit an alpha particle.

The alpha particle is the same as a helium

[2]

(b) This table shows the atomic numbers of some elements.

element	symbol	atomic number
thallium	Tl	81
lead	Pb	82
bismuth	Bi	83
polonium	Po	84
astatine	At	85
radon	Rn	86

(i) Polonium-209 decays by emitting an alpha particle.

Which element is formed when polonium-209 decays?

Use the table to help you.

..... [1]

(ii) Finish and balance the equation to show what happens when polonium-209 decays.



[2]

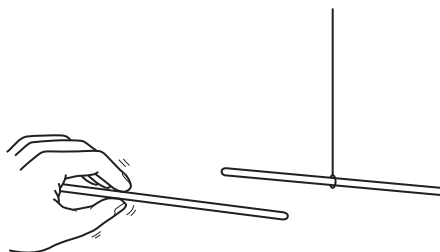
[Total: 5]

18

11 This question is about static electricity.

(a) Mel hangs up a charged plastic rod on a cotton thread.

She brings another charged rod towards it.



The rods move apart.

Why do the two rods move apart?

..... [1]

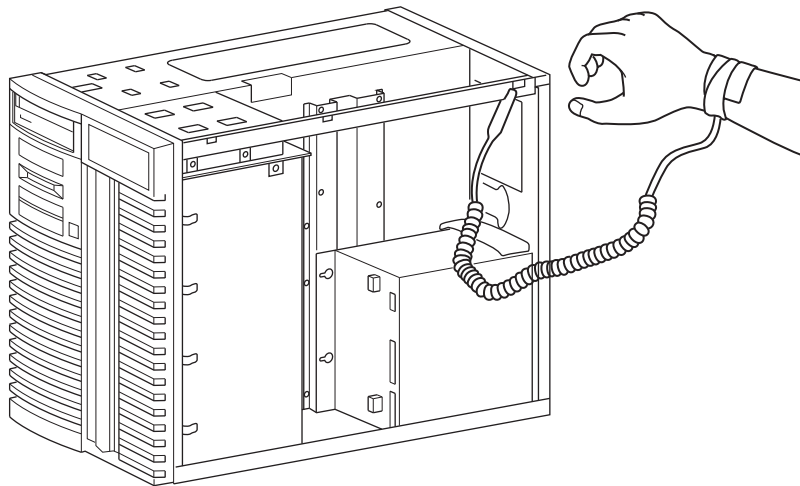
19

(b) Electronic components can be damaged by static electricity.

Look at the picture.

A technician builds a computer.

She is attached to the computer by an **antistatic strap**.



Write down two **other** examples where static electricity can be dangerous.

1.
.....
2.
..... [2]

(c) Technicians can sometimes get electrostatic shocks when using machines.

How do they avoid this?

..... [1]

[Total: 4]

12 The table shows the hearing ranges of some animals.

animal	frequency range in Hz
bullfrog	100 – 3000
canary	250 – 8000
chicken	125 – 2000
dog	67 – 45 000
elephant	16 – 12 000
goldfish	20 – 3000
horse	55 – 33 500
human	20 – 20 000
owl	200 – 12 000

(a) Which **two** animals can hear ultrasound?

..... and [1]

(b) Ultrasound is used to scan pregnant women to check on the growth of the unborn baby.

Write down **two** reasons why ultrasound is used instead of X-rays.

1

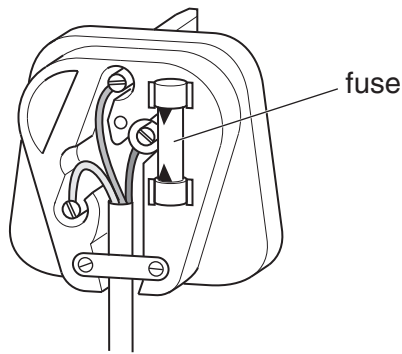
.....

2

..... [2]

[Total: 3]

13 A mains plug contains a fuse.



(a) This plug is connected to a table lamp.

The **fuse** is there for protection. It melts if the current is too large.

Explain **in detail** how this provides protection.

.....

 [1]

(b) The total resistance of the circuit containing the lamp is $460\ \Omega$.

A fuse is used in the 230V mains plug.

Calculate the current that passes through the fuse.

The equations on page 2 may help you.

.....

answer amps [2]

[Total: 3]

END OF QUESTION PAPER

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The Periodic Table of the Elements

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1	2	Key										3	4	5	6	7	0							
		relative atomic mass atomic symbol name atomic (proton) number																1 H hydrogen 1		4 He helium 2				
7 Li lithium 3	9 Be beryllium 4																	11 B boron 5		12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10
23 Na sodium 11	24 Mg magnesium 12																	27 Al aluminium 13		28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36							
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44	103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54							
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86							
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated													

Key

relative atomic mass
atomic symbol
name
atomic (proton) number

1	H	1
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* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.