

Write your name here

Surname					Other names									
Pearson Edexcel Certificate					Centre Number					Candidate Number				
Pearson Edexcel International GCSE					<input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/>					<input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/> <input type="text"/>				
<h1>Chemistry</h1> <p>Unit: KCH0/4CH0 Paper: 2C</p>														
Friday 16 January 2015 – Morning										Paper Reference				
Time: 1 hour										KCH0/2C 4CH0/2C				
You must have: Calculator										Total Marks				

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Show all the steps in any calculations and state the units.
- Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

P44255A

©2015 Pearson Education Ltd.

1/1/1/1/



PEARSON

THE PERIODIC TABLE

Period 1 2 3 4 5 6 7 0

Group

4	He	Helium	2
---	----	--------	---

1	H	Hydrogen	1
---	---	----------	---

7	Li	Lithium	3	9	Be	Beryllium	4	20	Ne	Neon	10
23	Na	Sodium	11	24	Mg	Magnesium	12	32	O	Oxygen	8
39	K	Potassium	19	40	Ca	Calcium	20	31	N	Nitrogen	7
86	Rb	Rubidium	37	88	Sr	Strontium	38	73	Ge	Germanium	32
133	Cs	Caesium	55	137	Ba	Barium	56	115	In	Indium	49
223	Fr	Francium	87	226	Ra	Radium	88	204	Tl	Thallium	81
				227	Ac	Actinium	89				
								65	Zn	Zinc	30
								59	Ni	Nickel	28
								58	Co	Cobalt	27
								56	Fe	Iron	26
								55	Mn	Manganese	25
								52	Cr	Chromium	24
								51	V	Vanadium	23
								48	Ti	Titanium	22
								45	Sc	Scandium	21
								89	Y	Yttrium	39
								91	Zr	Zirconium	40
								93	Nb	Niobium	41
								96	Mo	Molybdenum	42
								101	Ru	Ruthenium	44
								103	Rh	Rhodium	45
								106	Pd	Palladium	46
								108	Ag	Silver	47
								112	Cd	Cadmium	48
								119	Sn	Tin	50
								122	Sb	Antimony	51
								127	I	Iodine	53
								128	Te	Tellurium	52
								131	Xe	Xenon	54
								135.5	Cl	Chlorine	17
								15	P	Phosphorus	15
								14	Si	Silicon	14
								13	Al	Aluminium	13
								79	Se	Selenium	34
								75	As	Arsenic	33
								80	Br	Bromine	35
								84	Kr	Krypton	36
								85	At	Astatine	85
								86	Rn	Radon	86

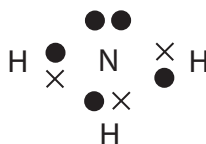
Key

Relative atomic mass
Symbol
Name
Atomic number



Answer ALL questions.

1 The diagram represents a particle of ammonia.



(a) This particle of ammonia is

(1)

- A an atom
- B an ion
- C a lattice
- D a molecule

(b) Which type of bonding is present in this particle of ammonia?

(1)

- A covalent
- B hydrogen
- C ionic
- D metallic

(c) What is the formula of ammonia?

(1)

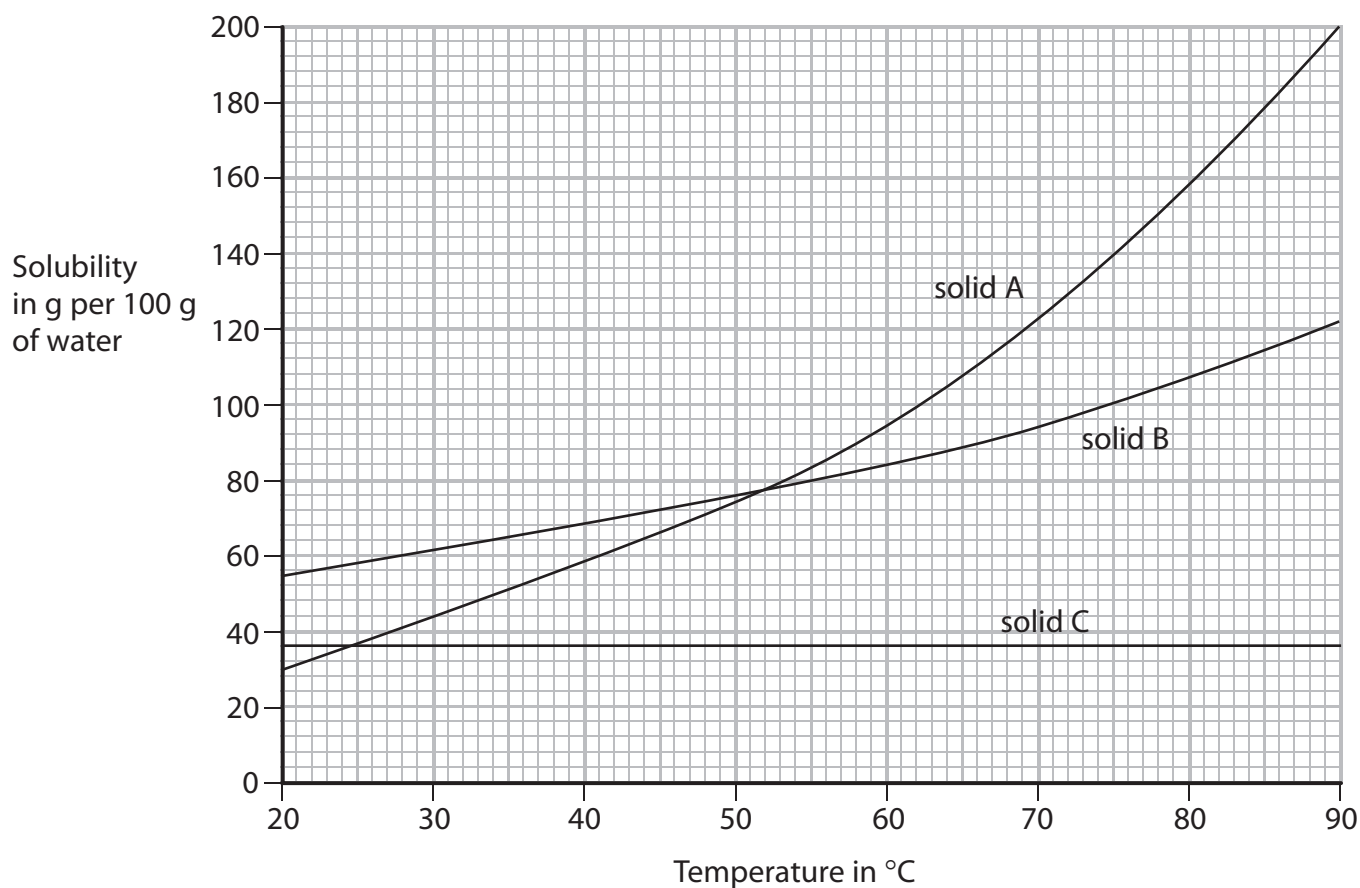
(Total for Question 1 = 3 marks)



- 2 The solubility of a solid in water is the maximum mass of the solid that can dissolve in 100 g of water at a given temperature.

An aqueous solution containing this maximum mass can be described as a saturated solution.

The graph shows the solubilities of three solids at different temperatures.



- (a) (i) What is the relationship between solubility and temperature for solid A?

(1)

- (ii) Which solid is the most soluble at 30°C?

(1)



(b) Explain what you would observe if a saturated solution of solid A were cooled from 90°C to 20°C.

(2)

.....

.....

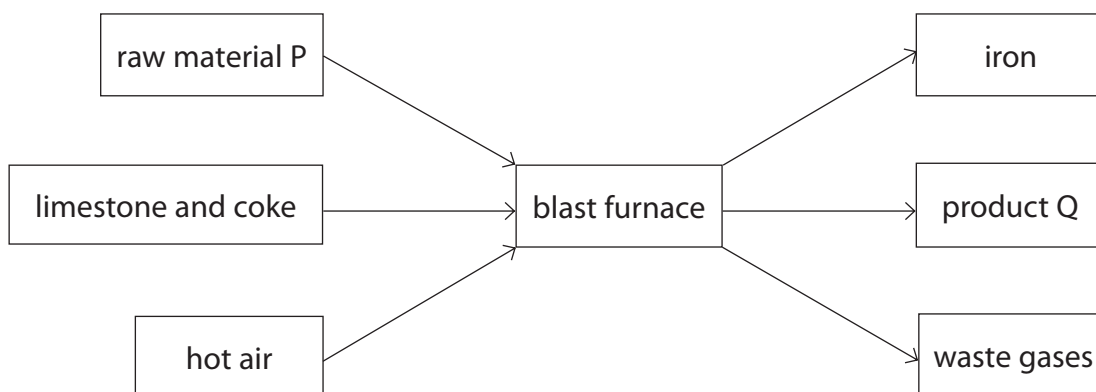
.....

.....

(Total for Question 2 = 4 marks)



3 The diagram shows how iron is produced in a blast furnace.



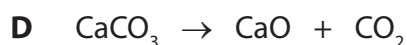
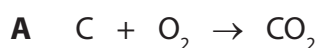
(a) Give the name of raw material P and of product Q.

(2)

raw material P

product Q

(b) The equations for some reactions in a blast furnace are



The table shows some types of reaction that occur in a blast furnace.

Complete the table by writing a letter, A, B, C, D, or E, to link each type of reaction to an appropriate reaction equation.

Each letter may be used once, more than once or not at all.

The first one has been done for you.

(3)

Type of reaction	Letter
one that gives out heat	A
one that is a thermal decomposition	
one that is a neutralisation	
one that forms a poisonous gas	



(c) The rusting of iron objects is a major problem.

Name the two substances needed for iron to rust.

(2)

1

2

(d) The order of reactivity of three metals is

most reactive

zinc

iron

tin

least reactive

Iron objects can be prevented from rusting by coating them with zinc or tin.

Some of these objects may be scratched when used, so the coating may come off.

Use the order of reactivity of the metals to suggest why coating these objects with zinc is more effective than coating them with tin.

(3)

.....

.....

.....

.....

.....

.....

.....

(Total for Question 3 = 10 marks)



4 (a) Wine can be made from grapes.

The grapes are crushed to produce an aqueous solution containing glucose.
Yeast is then added to this solution.

The solution is kept at a constant temperature for a period of time. The glucose is converted into ethanol.

(i) Name the process in which glucose is converted into ethanol.

(1)

(ii) What is the purpose of the yeast?

(1)

(b) Grape vines can be attacked by a fungus that ruins the grapes. The fungus can be killed using Bordeaux mixture, a solid containing copper(II) sulfate and calcium hydroxide.

(i) State a test to show that Bordeaux mixture contains calcium ions.

(2)

test for calcium ions

observation

(ii) A sample of Bordeaux mixture is dissolved in water.

Describe separate tests to show that this solution contains copper(II) ions and sulfate ions.

(5)

test for copper(II) ions

observation

test for sulfate ions

observation



- (c) Ethanol can be manufactured by passing a hot mixture of ethene and steam, at a high pressure, over a catalyst.

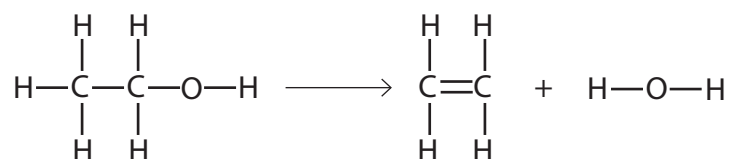
State the pressure used and name the catalyst.

(2)

pressure atm

catalyst

- (d) The equation for the conversion of ethanol into ethene can be written using displayed formulae.



The table gives some average bond energies.

Bond	Average bond energy in kJ/mol
C—C	348
C=C	612
C—H	412
C—O	360
O—H	463

Use information from the table to calculate the enthalpy change, in kJ/mol, for the conversion of ethanol into ethene.

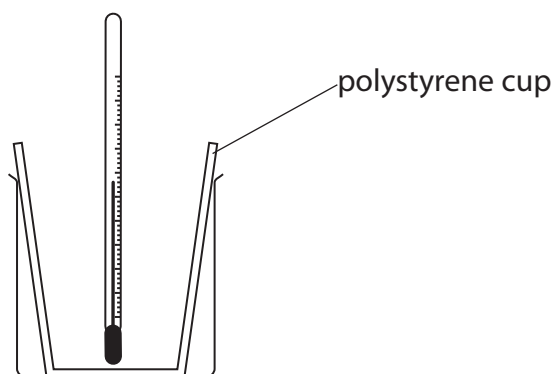
(4)

enthalpy change = kJ/mol

(Total for Question 4 = 15 marks)



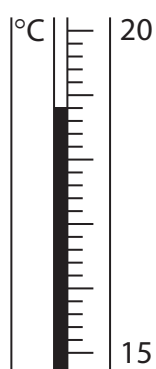
- 5 A student uses this apparatus to investigate the temperature change that occurs when potassium hydroxide is dissolved in water.



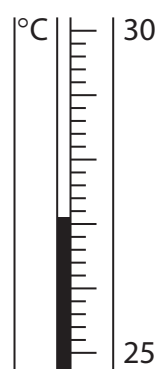
She uses this method.

- pour 50 cm^3 of water into the polystyrene cup and measure the temperature of the water
- add 3 g of potassium hydroxide and stir
- record the highest temperature of the solution

- (a) These diagrams show the thermometer readings before and after the student added the potassium hydroxide.



before



after

Use the readings to complete the table.

(3)

temperature in $^{\circ}\text{C}$ after adding potassium hydroxide	
temperature in $^{\circ}\text{C}$ before adding potassium hydroxide	
temperature change in $^{\circ}\text{C}$	



(b) The student uses her results to calculate the enthalpy change for dissolving potassium hydroxide in water.

She compares her value with a data book value.

Student's value = -32 kJ/mol .

Data book value = -55 kJ/mol .

There are no errors in the student's method or in the calculation.

Suggest two reasons why the student's value differs from the data book value.

(2)

1

.....

.....

2

.....

.....

(Total for Question 5 = 5 marks)



6 Potassium sulfide is an ionic compound.

(a) Complete the table to show the arrangement of electrons in the ions formed when potassium and sulfur react to form potassium sulfide.

Give the charge on each of the ions.

(3)

Element	Arrangement of electrons in atom	Arrangement of electrons in ion	Charge on ion
K	2.8.8.1		
S	2.8.6		

(b) (i) Explain why potassium sulfide conducts electricity when molten.

(1)

.....

.....

(ii) Explain why potassium sulfide has a high melting point.

(3)

.....

.....

.....

.....

.....

.....

.....

(Total for Question 6 = 7 marks)

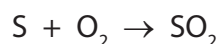


BLANK PAGE

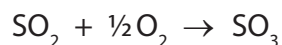


7 Sulfuric acid can be manufactured from sulfur in a four-stage process.

stage 1 sulfur is burned in air to form sulfur dioxide



stage 2 the sulfur dioxide is reacted with more oxygen to form sulfur trioxide



stage 3 the sulfur trioxide is absorbed in concentrated sulfuric acid to make oleum



stage 4 the oleum is carefully diluted with water to form sulfuric acid

(a) Write a chemical equation for the formation of sulfuric acid from oleum.

(1)

(b) A mass of 80 tonnes of sulfur is reacted with oxygen in stage 1.

Calculate the maximum mass, in tonnes, of sulfur trioxide that can be produced in stage 2.

[1 tonne = 1.0×10^6 g]

(3)

maximum mass =tonnes



(c) Calculate the minimum volume at rtp, in cubic decimetres (dm^3), of oxygen required to completely react with 64 tonnes of sulfur dioxide.

[1 mol of oxygen at rtp has a volume of 24 dm^3]

(2)

volume of oxygen = dm^3

(Total for Question 7 = 6 marks)



BLANK PAGE



9 Nitrogen dioxide (NO_2) is a brown gas.

Dinitrogen tetraoxide (N_2O_4) is a colourless gas.

The two gases can exist together in dynamic equilibrium according to the equation



A mixture of nitrogen dioxide gas and dinitrogen tetraoxide gas is allowed to reach equilibrium in a sealed container at 20°C . This equilibrium mixture is brown in colour.

(a) The sealed container is immersed in hot water at 60°C .

As the temperature of the gas mixture increases, the pressure of the gas mixture also increases.

(i) Predict the effect of the increase in temperature on the position of equilibrium.

(1)

(ii) Predict the effect of the increase in pressure on the position of equilibrium.

(1)

(iii) Suggest why it is difficult to predict which way the equilibrium will shift.

(1)



(b) Suggest why the equilibrium mixture is a darker shade of brown at 60°C than the equilibrium mixture at 20°C.

(2)

.....

.....

.....

.....

(Total for Question 9 = 5 marks)

TOTAL FOR PAPER = 60 MARKS



BLANK PAGE

