



Mark Scheme (Results)

Summer 2018

Pearson Edexcel International GCSE
In Chemistry (4CH0) Paper 2CR

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Question number	Answer	Notes	Marks
1 (a)	carbon monoxide	ACCEPT CO	1
(b)	oxygen	ACCEPT O ₂ IGNORE O	1
(c)	chlorine	ACCEPT Cl ₂ IGNORE Cl	1
(d)	nitrogen	ACCEPT N ₂ IGNORE N	1
(e)	chlorine	ACCEPT Cl ₂ IGNORE Cl	1
(f)	nitrogen AND carbon monoxide	In either order ACCEPT N ₂ AND CO IGNORE N	1

Total for Question 1 = 6

Question number	Answer	Notes	Marks
2 (a) (i)	M1 (it/mass) decreases M2 (as) carbon dioxide/CO ₂ /gas escapes out of/leaves (the flask)	ALLOW carbon dioxide/CO ₂ /gas goes into air IGNORE carbon dioxide/CO ₂ /gas given off/produced /formed on its own	2
(ii)	<u>all</u> (hydrochloric) acid has reacted	ACCEPT the acid is used up ALLOW there was not enough acid to react with (all the) marble chips ALLOW acid was the limiting factor IGNORE marble chips was in excess	1
(iii)	mass (of flask and contents) stays the same/stops decreasing	ACCEPT stops bubbling/fizzing/effervescing IGNORE no more gas produced/given off	1
(b) (i)	A 1 minute (correct answer as the gradient of the curve is steeper than at 4, 6 or 8 minutes) B 4 minutes (incorrect - the gradient of the curve is less steep than at 1 minute) C 6 minutes (incorrect - the gradient of the curve is less steep than at 1 minute) D 8 minutes (incorrect - the gradient of the curve is less steep than at 1 minute)		1
(ii)	M1 starts from 0,0 and is steeper than original curve M2 levels off before original curve but at same height		2

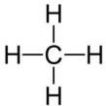
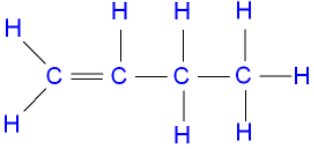
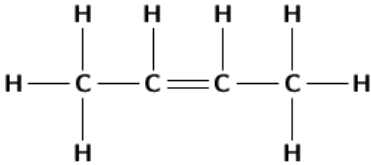
Question number	Answer	Notes	Marks
2 (c)	<p>M1 fewer (acid) particles in same volume</p> <p>M2 (so) fewer (successful) collisions per second</p> <p>M3 (so) lower rate of reaction</p>	<p>IGNORE references to wrong type of particles e.g. molecules</p> <p>ACCEPT less frequent (successful) collisions</p> <p>IGNORE references to less chance of collision</p> <p>Any reference to particles losing energy / moving more slowly can score M1 only</p>	3

Total for Question 2 = 10

Question number	Answer	Notes	Marks						
3 (a)	(i) pipette (ii) to see colour (change) more easily/clearly OWTTE (iii) M1 (before) pink M2 (after) colourless (iv) water does not affect number of moles/amount of NaOH (present) OWTTE	 ACCEPT magenta ALLOW red ACCEPT no colour ALLOW decolourised REJECT discoloured IGNORE clear ALLOW 1 mark for both correct but in reverse order ALLOW water will not affect the (titration) result/volume of acid needed REJECT water does not affect the concentration of NaOH IGNORE water does not affect the pH of NaOH (solution) IGNORE references to effect on indicator	 1 1 2 1						
(b)	<table border="1" data-bbox="371 1402 995 1659"> <tbody> <tr> <td data-bbox="371 1402 855 1503">Burette reading after adding the acid</td> <td data-bbox="855 1402 995 1503">(22.80)</td> </tr> <tr> <td data-bbox="371 1503 855 1603">Burette reading before adding the acid</td> <td data-bbox="855 1503 995 1603">1.45</td> </tr> <tr> <td data-bbox="371 1603 855 1659">Volume in cm³ of acid added</td> <td data-bbox="855 1603 995 1659">21.35</td> </tr> </tbody> </table> M1 for reading before adding the acid M2 for volume added	Burette reading after adding the acid	(22.80)	Burette reading before adding the acid	1.45	Volume in cm ³ of acid added	21.35	 Mark CQ	 2
Burette reading after adding the acid	(22.80)								
Burette reading before adding the acid	1.45								
Volume in cm ³ of acid added	21.35								

Question number	Answer	Notes	Marks
3 c) (i)	ticks under last three results		1
(ii)	<p>M1 $(21.50 + 21.35 + 21.40) \div 3$</p> <p>M2 21.42</p>	<p>mark CQ on any combination of ticked results</p> <p>If no results are ticked then M1 can only be awarded if the last three results are averaged</p> <p>If only one result ticked then no marks can be scored in (ii)</p> <p>Mark CQ on M1</p> <p>All answers should be correctly rounded to 2 dp</p>	2
3(d)	<p>M1 $n[(\text{H}_3\text{PO}_4)] = (0.02 \times 21.30) \div 1000$ OR 0.000426</p> <p style="text-align: right;">OR</p> <p>4.26×10^{-4}</p> <p>M2 $n[\text{NaOH}] = 3 \times \text{M1}$ OR evaluated correctly and quoted to at least two significant figures</p> <p>M3 concⁿ of NaOH = M2 $\times 1000 \div 25$ evaluated correctly and quoted to at least two significant figures</p>	<p>ACCEPT 0.00043 OR 4.3×10^{-4}</p> <p>If M1 is correct this should be 0.001278 ACCEPT 0.00128 OR 0.0013</p> <p>If M1 and M2 are correct this should be 0.05112 ACCEPT 0.052 from 0.0013 ACCEPT 0.0512/0.051 from 0.00128</p> <p>correct answer with no working scores 3</p> <p>examples of possible answers using incorrect mole ratio 0.01704 scores 2 marks 0.00568 scores 2 marks</p>	3

Total for Question 3 = 13

Question number	Answer	Notes	Marks
4 (a) (i)	T / C ₂ H ₄ Br ₂	ACCEPT displayed formulae	1
(ii)	P / CH ₄	ACCEPT displayed formulae	1
(b)(i)		IGNORE bond angles	1
(ii)	<p data-bbox="379 674 427 703">M1</p>  <p data-bbox="379 972 427 1001">M2</p> 	<p data-bbox="1040 674 1369 703">IGNORE bond angles</p> <p data-bbox="1040 741 1294 808">ALLOW in either order</p> <p data-bbox="1040 1205 1342 1308">ALLOW cis- and trans- isomers for 2 marks</p>	2

Question number	Answer	Notes	Marks
4 (c) (i)	(bromine changes colour from red-brown to) colourless/decolourised	If initial colour given must be red-brown /orange/brown/yellow or any combination of orange/brown/yellow e.g. orange-brown REJECT discoloured IGNORE clear	1
	(ii) M1 (react with) bromine/Br ₂ M2 UV radiation/light	IGNORE (aq) IGNORE Br IGNORE heat IGNORE pressure mark(s) can be scored if reactants are shown in an equation/word equation (even if equation is incomplete or not fully correct)	2
(d) (i)	addition (polymers)		1
	(ii)		
	$\left(\begin{array}{cc} \text{H} & \text{H} \\ & \\ \text{---C} & \text{---C---} \\ & \\ \text{H} & \text{Cl} \end{array} \right)_n$		2
	M1 three Hs, one Cl and single bond between two Cs	ALLOW Cl in any of the four positions	
	M2 extension bonds and n	ACCEPT n anywhere after brackets but not before Extension bonds do not need to go out of brackets M2 DEP M1	

Question number	Answer	Notes	Marks
4 (e) (i)	condensation (polymerisation)		1
4 (e) (ii)	<p>Any one from:</p> <p>M1 in condensation polymerisation (small) molecule also formed</p> <p>M2 addition polymerisation involves identical/same monomers/starting molecules</p> <p>M3 condensation polymerisation involves two/ different monomers/starting molecules</p> <p>M4 addition polymerisation produces only one product</p> <p>M5 addition polymerisation involves the breaking of (carbon to carbon) double bonds</p> <p>M6 condensation polymerisation forms ester/amide links</p>	<p>ACCEPT water/H₂O/HCl also formed</p> <p>ALLOW two products formed</p> <p>ALLOW another compound/product/by-product formed</p> <p>ALLOW in addition polymerisation (carbon to carbon) double bonds become single</p>	1

Total for Question 4 = 13

Question number	Answer	Notes	Marks
5 (a) (i)	electrolysis	ACCEPT decomposition	1
(ii)	zinc reacts with (sulfuric) acid	ACCEPT zinc is not inert (electrode) ALLOW zinc reacts with electrolyte/solution ALLOW zinc dissolves in the acid ALLOW zinc reacts with oxygen IGNORE zinc is (too/very) reactive IGNORE references to cost	1
(b) (i)	glowing splint/spill relights/rekindles/catches fire	REJECT lighted splint/spill	1
(ii)	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$	ALLOW $2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-$	1
(c)	M1 4 faraday give 1 mol O ₂ M2 0.010 faraday gives 0.0025 mol O ₂ M3 (so) volume of O ₂ = (0.0025 x 24000) = 60 (cm ³)	M2 subsumes M1 Mark CQ correct answer with no working scores 3 240/960/120 (cm ³) score 2 marks	3
(d)	very low concentration of hydroxide/OH ⁻ ions present in (sulfuric) acid	ALLOW (there are) no/few hydroxide/OH ⁻ ions in (sulfuric) acid	1

Total for Question 5 = 8

Question number	Answer	Notes	Marks
6 (a) (i)	<p>M1 volume of water</p> <p>M2 temperature of water before and after burning</p>	<p>ALLOW mass of water</p> <p>ALLOW temperature change</p>	2
6 (a) (ii)	<p>M1 $n[\text{heptanol}] = 0.75 \div M_r$ OR evaluated correctly and quoted to at least two significant figures</p> <p>M2 $19 \div M1$</p> <p>M3 evaluated correctly and quoted to at least two significant figures</p> <p>OR</p> <p>M1 (0.75g produces 19kJ)</p> <p>so 1g produces $19 \div 0.75 = 25.33\dots$ (kJ)</p> <p>M2 so 114g produces $114 \times 25.33\dots$</p> <p>M3 evaluated correctly and quoted to at least two significant figures</p>	<p>calculator answer from 114 is 0.0065789473684 (mol)</p> <p>calculator answer from 116 is 0.0064655172</p> <p>calculator answer from 114 is 2888 (kJ/mol)</p> <p>= 2900 (kJ/mol) to 2 sig fig</p> <p>calculator answer from 116 is 2938.66 (kJ/mol)</p> <p>= 2900 (kJ/mol) to 2 sig fig</p> <p>IGNORE sign in final answer</p> <p>from 114 calculator the answer is 2888 (kJ/mol)</p> <p>= 2900 (kJ/mol) to 2 sig fig</p> <p>IGNORE sign in final answer</p> <p>correct answer with no working scores 3</p>	3

Question number	Answer	Notes	Marks
6 (b) (i)	<p>M1 $\Sigma(\text{bonds broken}) = [(2 \times 436) + 498]$ OR 1370 (kJ)</p> <p>M2 $\Sigma(\text{bonds formed}) = 4 \times 464$ OR 1856 (kJ)</p> <p>M3 $\Delta H = - 486$ (kJ)</p> <p>OR</p> <p>if M1 and/or M2 incorrect:</p> <p>M3 numerical difference between M1 and M2</p> <p>If M2 greater than M1 answer must be negative</p> <p>If M2 less than M1 answer must be positive</p>	<p>IGNORE sign</p> <p>IGNORE sign</p> <p>negative sign is required</p> <p>-486 with or without working scores 3</p> <p>(+)486 with or without working scores 2</p> <p>unless a clear statement is made that reaction is exothermic then sign can be negative</p>	3

Question number	Answer	Notes	Marks
6 (b) (ii)	<p>M1 for $2\text{H}_2 + \text{O}_2$ and $2\text{H}_2\text{O}$ in correct positions</p> <p>M2 ΔH correctly placed and labelled</p>	<p>ALLOW hydrogen and oxygen and water ALLOW reactants and products</p> <p>Mark CQ on sign in (i)</p> <p>ACCEPT ΔH label on vertical line/double arrow/arrow pointing from reactants level to product level</p> <p>REJECT arrow pointing from products level to reactants level</p> <p>IGNORE activation energy attempts</p>	2

Total for Question 6 = 10

